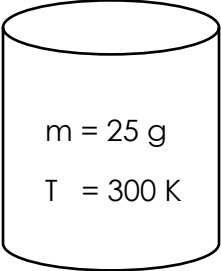
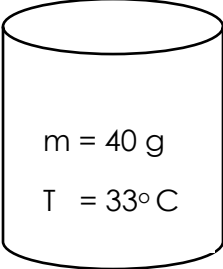
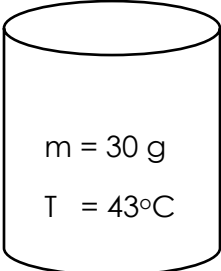
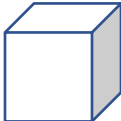
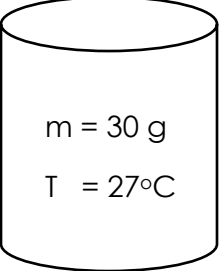
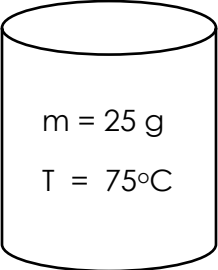
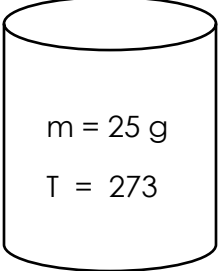


Calorimetry – Thermal Equilibrium (Final Temperature)

Below are six calorimeters with water of different mass and initial temperature. A sample of ice of 273 Kelvin is placed in each calorimeter. Rank the calorimeter according to the greatest to least final temperature of the system.

 2g	 5g	 7g
 m = 25 g T = 300 K	 m = 40 g T = 33°C	 m = 30 g T = 43°C
A	B	C
 5g	 2g	 10g
 m = 30 g T = 27°C	 m = 25 g T = 75°C	 m = 25 g T = 273
D	E	F

Greatest 1 _____ 2 _____ 3 _____ 4 _____ 5 _____ 6 _____ Least

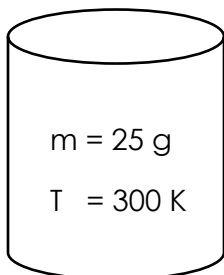
Are there any calorimeters that have the same final temperature? _____

Please carefully explain your reasoning.

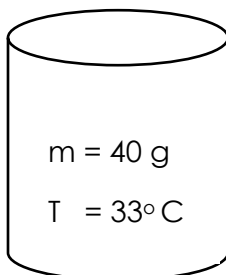
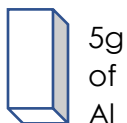
Calorimetry – Change in Temperature of the Metal

Evaluation: On a scale of 1-10, how sure are you of your ranking?

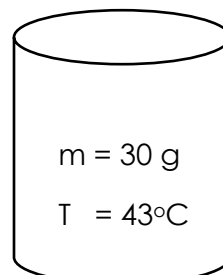
Below are six calorimeters with water of different mass and initial temperature. A sample of different metals with the same mass and each has an initial temperature of 75°C , are placed in each calorimeter. Rank the calorimeter according to the greatest to least change in temperature of the metal.



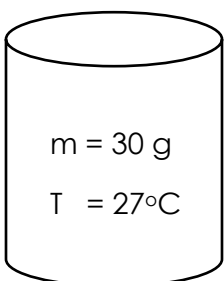
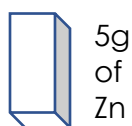
A



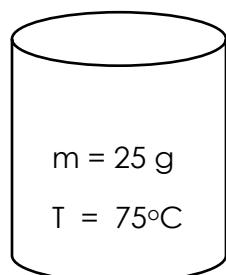
B



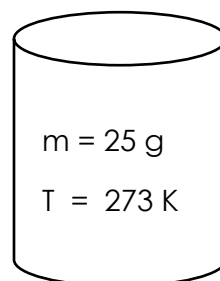
C



D



E



F

Greatest 1 _____ 2 _____ 3 _____ 4 _____ 5 _____ 6 _____ Least

Are there any calorimeters that have the same final temperature? _____

Please carefully explain your reasoning.

Evaluation: On a scale of 1-10, how sure are you of your ranking?