Kf and Multiple Equilibria– Reading Guide

*sections 15.2 and 15.3 in OpenStax*

**Lewis Acids and Bases (section 15.2)**

Define **Lewis acid** and give one example:

Define **Lewis base** and give one example:

A **complex ion** is formed from the reaction of a Lewis acid and a Lewis base. Write the balanced equation below to show the formation reaction of [Cu(NH3)4]2+. Label the Lewis acid and Lewis base.

+ ⮀ [Cu(NH3)4]2+

Write the Kf equation for your balanced equation above.

Kf =

**Multiple Equilibria (section 15.3)**

Slightly soluble solids derived from \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ generally dissolve in strong acids.

Which of the following slightly soluble salts would be more soluble in acidic solution? **Explain** your answer.

PbSO4 Be(OH)2 KCl

**End of Chapter 15 Practice Problems**

#65, 75, 111

For detailed solutions to these problems, go to the [OpenStax website](https://openstaxcollege.org/textbooks/chemistry/resources) and download the “Student Answer and Solution Guide.”