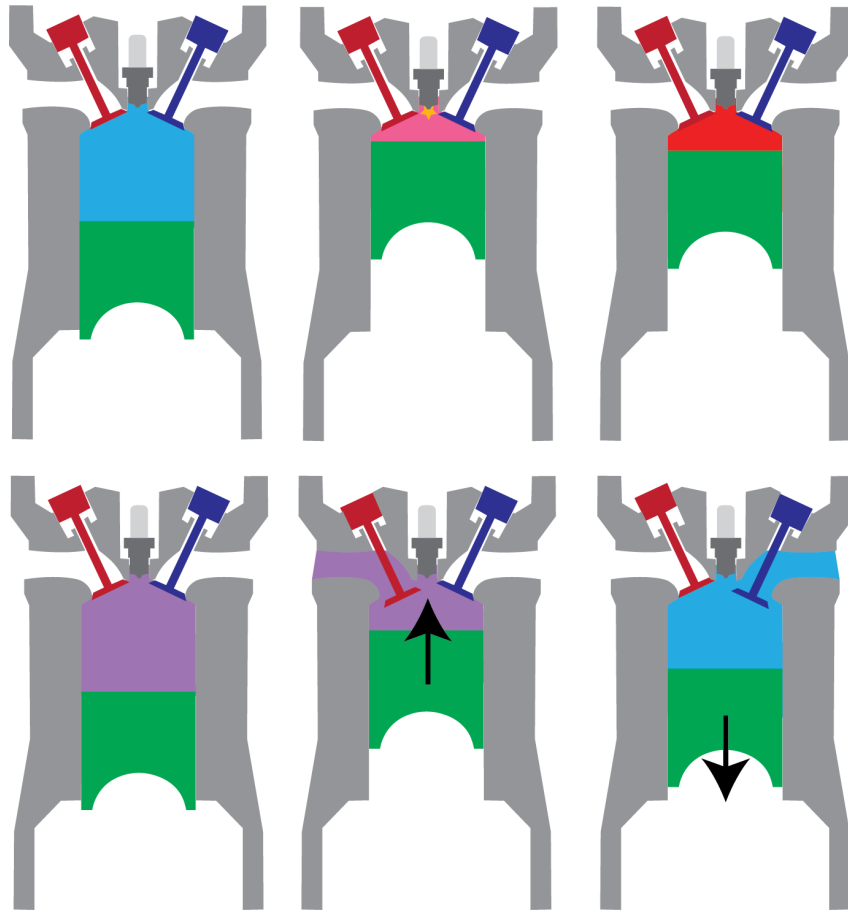


Thermodynamics for Engineers



Robert Brown

March 10, 2023



This work is licensed under the Creative Commons Attribution-NonCommercial-ShareAlike 3.0 United States ([CC BY-NC-SA 3.0 US](#)). Work has been modified from **Engineering Thermodynamics - A Graphical Approach** by Israel Urieli and **Thermodynamics** by Diana Bairaktarova.

Modifications include change of format from website to traditional book, reorganization of multiple chapters, revised and expanded explanations throughout the book, updated and revised examples, recreated and revised figures and plots, recreation of tables and plots in the Appendices, addition of multiple homework problems, and in some cases, omission of parts of the original text.

Contents

Preface	vii
1 Concepts, Dimensions, and Units	1
1.1 Thermodynamics and Energy	1
1.2 Force and Work	2
1.2.1 Nomenclature	2
1.2.2 Newton's Second Law	2
1.2.3 Work	2
1.3 Unit Conversions	3
1.4 Forms of Energy	5
1.4.1 Potential Energy	5
1.4.2 Kinetic Energy	5
1.4.3 Internal Energy	5
1.5 Basic Properties of Matter	7
1.5.1 Nomenclature	7
1.5.2 Density	7
1.5.3 Specific Volume	7
1.5.4 Pressure	7
1.5.5 Temperature	9
1.6 Types of Thermodynamic Systems	9
1.6.1 Closed Systems	9
1.6.2 Open Systems	10
1.7 Intensive and Extensive Properties	10
1.8 State and Equilibrium	11
1.9 Process and Cycle	11
1.10 Using Software: Google Colab (Python)	12
1.10.1 Setup and CoolProp Basics	12
1.10.2 Basic Math	13
1.10.3 Plotting in Python	14
2 Properties of Matter	19
2.1 Phase Change and Property Diagrams	20
2.2 Quality	22
2.3 The p - v Diagram for Water	26
2.4 Ideal Gas Equation of State	28
2.5 Non-Ideal Gas Behavior	30

2.6	Using Software: Google Colab (Python) with CoolProp	33
2.6.1	Setup and CoolProp Basics	33
2.6.2	A Simple Example	34
2.6.3	Using CoolProp	34
2.6.4	Exporting from Colab as a PDF	35
2.7	Summary	36
3	The First Law for Closed Systems	39
3.1	Introducing the First Law	39
3.1.1	Heat Transfer (Q)	40
3.1.2	Work (W)	41
3.1.3	Internal Energy (ΔU)	42
3.2	Enthalpy (h)- A New Property	44
3.3	Specific Heats	45
3.3.1	Constant Volume	45
3.3.2	Constant Pressure	46
3.3.3	Ideal Gases	47
3.4	Stirling Cycle Engine	49
3.4.1	Process 1→2 – Isothermal Compression	50
3.4.2	Process 2→3 – Constant Volume Heating	51
3.4.3	Process 3→4 – Isothermal Expansion	52
3.4.4	Process 4→1 – Constant Volume Cooling	53
3.4.5	Ideal Stirling Cycle Analysis	53
3.4.6	Stirling Engines without Regeneration	54
3.5	Stirling Cycle Cooling	57
3.6	Ideal Gas Adiabatic Processes	61
3.7	Air-standard Otto Cycle	63
3.7.1	Otto Cycle Analysis	64
3.8	Air-Standard Diesel Cycle	69
3.8.1	Diesel Cycle Analysis	70
3.9	Using Software: Plotting a Cycle in Python	73
3.9.1	Solving the Otto Cycle in Colab	73
3.9.2	Plotting the States	75
3.9.3	Revising the Plot	75
3.10	Summary	78
4	The First Law for Open Systems	83
4.1	Energy Conservation (First Law) for Open Systems	83
4.1.1	Mass Flow	84
4.1.2	Flow Work	84
4.1.3	Conservation of Energy for an Open System	85
4.2	The p - h (Pressure-Enthalpy) Diagram	86
4.3	Components of Thermodynamic Systems	87
4.3.1	Turbines	88
4.3.2	Compressors and Pumps	89
4.3.3	Boilers, Evaporators, and Steam Generators	90
4.3.4	Condensers	91
4.3.5	Mixers	92

4.3.6	Throttles	93
4.3.7	Heat Exchangers	94
4.3.8	De-aerators	94
4.4	Basic Rankine Cycle	95
4.5	Modifications to the Rankine Cycle	100
4.5.1	Rankine Cycle with Reheat	101
4.5.2	Rankine Cycle with Regeneration	105
4.6	Refrigeration Cycle	110
4.6.1	Types of Refrigerants	110
4.6.2	Refrigeration	111
4.7	Summary	116
5	The Second Law and Entropy	127
5.1	Entropy	127
5.1.1	Entropy as a Property	128
5.1.2	Entropy in Statistical Mechanics	128
5.2	Reversibility	129
5.2.1	Friction	129
5.2.2	Non-equilibrium Processes	129
5.2.3	Reversibility and Entropy	130
5.3	Entropy Transfer	131
5.3.1	Entropy Transfer in Statistical Mechanics	131
5.3.2	Defining Entropy for an Ideal Gas	132
5.3.3	Entropy for Liquids and Solids	132
5.3.4	Additional Thoughts on Pressure and Temperature	133
5.4	Defining the Second Law	135
5.4.1	The Clausius Statement of the Second Law	136
5.4.2	Kelvin-Planck Statement of the Second Law	136
5.4.3	Equivalence of Kelvin-Planck and Clausius Statements	137
5.5	Carnot Cycles and Ideal Efficiency	138
5.5.1	Carnot Cycle Analysis	139
6	Isentropic Efficiency and the Brayton Cycle	143
6.1	Isentropic Efficiency	143
6.1.1	Isentropic Efficiency of Turbines	144
6.1.2	Isentropic Efficiency of Compressors	146
6.2	Aircraft Engines and the Brayton Cycle	149
A	Steam Tables	165
A.1	Saturation Properties for Steam - Temperature	165
A.2	Saturation Properties for Steam - Pressure	167
A.3	Superheated Vapor Properties for Steam	168
A.4	Compressed Liquid Properties for Water	176
A.5	Steam Pressure-Enthalpy (p - h) Diagram	178
A.6	Steam Enthalpy-Entropy (h - s) Diagram	179

B	R134a (Tetrafluoroethane) Tables	181
B.1	Saturation Properties for R134a - Temperature	181
B.2	Saturation Properties for R134a - Pressure	183
B.3	Superheated Vapor Properties for Refrigerant R134a	184
B.4	R134a Pressure-Enthalpy (p - h) Diagram	188
B.5	R134a Enthalpy-Entropy (h - s) Diagram	189
C	Properties of Various Ideal Gases	191
C.1	Properties of Select Ideal Gases at 300 K	191
C.2	Ideal Gas Specific Heats of Air	191
C.3	Compressibility Charts	193

Preface

This work is an attempt to bring high-quality Open Educational Resources to Engineering Thermodynamics at Abilene Christian University, as supported by the Alternative Textbook Initiative Grant.

The goal of Open Educational Resources is to combat the increased cost of textbooks seen in recent years. Additionally, having resources that are completely free allows students to maintain a library of books, and encourages students to access the book, when they may not have otherwise purchased it.

Another benefit is the ability of the professor to curate exactly what is provided to the student. Many textbooks are padded with chapters that are skipped in the course, often leading to confusion with homeworks and textbook readings.

Finally, OER allows for collaboration and continuous improvement. As you work through this text, be on the lookout for things that are confusing, misleading, or simply poorly explained. Mistakes and errors can be fixed this semester, and any feedback can go into the next version of the book much faster than a traditional publisher.

This text is an adaptation of a web resource originally developed by Dr. Israel Urieli of Ohio University, titled [Engineering Thermodynamics - A Graphical Approach](#). The resource is unique for a number of reasons, the greatest of which being that it is licensed under a [Creative Commons license](#) which allows free sharing and adaptation by other parties. I should also mention Dr. Diana Bairaktarova, whose adaptation of Dr. Urieli's work was also invaluable in the process of building this version.

Chapter 1

Concepts, Dimensions, and Units

Thermodynamics, generally speaking, is the science of energy. It focuses on the transformation of energy from one form to another. This includes transforming heat into work, such as in an automobile engine or at a power plant, and transforming work into heat transfer, such as in refrigerators or heat pumps.

The application of Thermodynamics is almost everywhere in our daily life. Examples of some application areas of this subject are: propulsion, internal combustion engines, power plants, refrigeration and air conditioning, solar heating, the interaction of the human body with its surroundings, biomedical devices, the biology of the human body, animals, plants, ecological systems, etc. In this course on thermodynamics we will focus on the analysis of energy systems and the application of these systems to real world contexts.

There are two approaches of teaching Thermodynamics – microscopic and macroscopic. Classical thermodynamics (the macroscopic approach) does not require detailed knowledge of molecular motion to describe a system. Statistical Thermodynamics (the microscopic approach) considers quantum mechanical description of molecules. We will primarily be looking at the macroscopic view, while occasionally using the microscopic approach to improve our basic understanding.

This first chapter is dedicated to introductory concepts, dimensions, and units. The basic principles of thermodynamics are the conservation of mass, the conservation of energy, and the concept of entropy. Basic thermodynamics terminology covered in this chapter includes the concepts of a system, properties, states, equilibrium, and processes.

We will distinguish primary dimensions in Thermodynamics such as mass (m), length (L), time (t), temperature (T), electric current (I), the amount of a substance (n), and secondary (or derived) dimensions such as: velocity ($\frac{m}{s}$), density ($\frac{kg}{m^3}$), pressure ($\frac{N}{m^2}$), etc. In this course we will use the two most prevalent unit systems: the International System of Units (SI) and the United States Customary Units (US), focusing primarily on the first.

1.1 Thermodynamics and Energy

Thermodynamics is the science of energy, including energy storage and energy in transit. The Conservation of Energy Principle states that energy cannot be created or destroyed,

but can only change its form. The three forms of energy storage of greatest interest to us are Potential Energy (**PE**), Kinetic Energy (**KE**), and Internal Energy (**U**), which we introduce below. The two forms of energy in transit that we consider are Work (**W**) and Heat (**Q**), and the interactions between these various forms of energy are defined in terms of the First Law of Thermodynamics.

1.2 Force and Work

1.2.1 Nomenclature

Symbol	Meaning	SI Units	English Units
F	Force	Newton [N]	pounds-force [lbf]
m	Mass	kilogram [kg]	slugs OR pounds-mass [lbm]
a	Acceleration	meters/second ² [$\frac{m}{s^2}$]	feet/second ² [$\frac{ft}{s^2}$]
W	Work	Joules [J]	foot-pounds [ft-lb]

1.2.2 Newton's Second Law

You should have already seen Newton's Second Law in Engineering Physics.

Newton's Second Law states:

$$F = ma$$

Mass (m) and force (F) are two of the most common dimensions we will be using in this course. Mass is used in nearly everything, because we are typically using properties of matter given **per unit mass**. Force is used almost as often, but it tends to be hidden as energy ($F \cdot d = W[J]$) or pressure ($F/A = p [Pa]$)

You'll notice above that there are two possible units for mass in the US system. The US unit of slugs is not used from day-to-day, but it is the most direct analog of the SI unit kilograms [kg]. More commonly used is the pound-mass [lbm], which is defined by the amount of mass which creates a pound-force in Earth's standard gravity.

Since the acceleration due to gravity $g = 32.2 \frac{ft}{s^2}$ we have:

$$1 \text{ lbf} = 32.2 \frac{ft}{s^2} \cdot 1 \text{ lbm}$$

$$1 \text{ lbf} = 1 \frac{ft}{s^2} \cdot 1 \text{ slug}$$

$$1 \text{ slug} = 32.2 \text{ lbm}$$

1.2.3 Work

We now consider the work done (W), which is the energy transferred to an object in motion requiring both an applied force (F) and distance moved (d). If the force is constant over the distance moved then the work done is given by:

$$W = Fd,$$

In general, however, the force is not constant over the distance. For varying forces we need to sum all the incremental work processes taking into consideration the variation of the force. This leads to the integral form for determining work done as follows:

$$W = \int_0^d F dx$$

Notice that the integral form can be simplified to the original form when F is constant.

1.3 Unit Conversions

It is often necessary to convert data from one unit system to another. The lists below give conversions for the basic units, along with several shortcuts for commonly used unit combinations.

Base Units

1 kg = 2.2 lbm = 0.0685 slug	1 m = 3.281 ft
1 lbm = 0.031 slug = 0.4536 kg	1 ft = 0.3048 m
1 N = 0.225 lbf	1 cm = 0.3937 in.
1 lbf = 4.448 N	1 in. = 2.54 cm

Temperature

$$\begin{aligned} T[^\circ\text{C}] &= (T[^\circ\text{F}] - 32^\circ\text{F}) \cdot 5/9 & T[\text{K}] &= T[^\circ\text{C}] + 273.16\text{K} \\ T[^\circ\text{F}] &= T[^\circ\text{C}] \cdot 9/5 + 32^\circ\text{F} & T[^\circ\text{R}] &= T[^\circ\text{F}] + 459.67^\circ\text{R} \end{aligned}$$

Energy and Energy Transfer

$$\begin{aligned} 1 \text{ kJ} &= 737.56 \text{ ft} \cdot \text{lbf} = 0.9478 \text{ BTU} & 1 \text{ W} &= 1 \text{ J/s} = 3.41 \text{ BTU/hr} \\ 1 \text{ ft} \cdot \text{lbf} &= 1.3558 \text{ J} & 1 \text{ kW} &= 1.341 \text{ hp} = 737.56 \text{ ft} \cdot \text{lbf/s} \\ 1 \text{ BTU} &= 778.17 \text{ ft} \cdot \text{lbf} = 1.055 \text{ kJ} & 1 \text{ hp} &= 550 \text{ ft} \cdot \text{lbf/s} = 0.7457 \text{ kW} = 2545 \text{ BTU/hr} \end{aligned}$$

Common Combinations

$$\begin{aligned} 1 \text{ km/hr} &= 0.6214 \text{ mile/hr} = 0.2778 \text{ m/s} & 1 \text{ m}^3 &= 1000 \text{ L} = 35.315 \text{ ft}^3 = 264.2 \text{ gal} \\ 1 \text{ mile/hr} &= 1.609 \text{ km/hr} = 1.467 \text{ ft/s} & 1 \text{ cm}^3 &= 0.061 \text{ in}^3 \\ 1 \text{ kPa} &= 0.145 \text{ lbf/in}^2 = 0.021 \text{ lbf/ft}^2 & 1 \text{ in}^3 &= 16.39 \text{ cm}^3 \\ 1 \text{ atm} &= 101.325 \text{ kPa} = 14.7 \text{ lbf/in}^2 & 1 \text{ L} &= 10^{-3} \text{ m}^3 = 0.0353 \text{ ft}^3 = 0.264 \text{ gal} \\ 1 \text{ lbf/in}^2 &= 6.895 \text{ kPa} = 144 \text{ lbf/ft}^2 & 1 \text{ gal} &= 0.00378 \text{ m}^3 = 0.1337 \text{ ft}^3 = 3.78 \text{ L} \end{aligned}$$

Example 1.1: Unit Conversion

Figure 1.1 shows a speedometer. You can see that 50 mph is approximately the same as 80 km/h. How much error is present in that estimate?

1. Determine the approximate conversion created by the above information. Compare that to the value given in the Common Combinations section above.
2. Use the base unit conversions, along with 1 mile = 5280 feet and 1 hour = 3600 seconds to recreate the conversion value.



Figure 1.1: An old Ford speedometer. From [Wikimedia Commons](#).

1. The Common Combinations section gives the conversion “1 mile/hr = 1.609 km/hr”. We can create a similar equation from the speedometer, as shown below:

$$50 \frac{\text{mile}}{\text{hr}} = 80 \frac{\text{km}}{\text{hr}}$$

At this point, it’s simple to create a similar conversion to the one given:

$$\begin{aligned} \frac{1}{50} 50 \frac{\text{mile}}{\text{hr}} &= \frac{1}{50} 80 \frac{\text{km}}{\text{hr}} \\ 1 \frac{\text{mile}}{\text{hr}} &= 1.6 \frac{\text{km}}{\text{hr}} \end{aligned}$$

This is less than 1% off from the more accurate value given.

2. In order to build a conversion from base units, we need our starting point (1 mph), and fix one unit at a time.

$$\begin{aligned} 1 \frac{\text{mile}}{\text{hr}} \cdot \frac{5280 \text{ ft}}{1 \text{ mile}} &= 5280 \frac{\text{ft}}{\text{hr}} \\ 5280 \frac{\text{ft}}{\text{hr}} \cdot \frac{1 \text{ m}}{3.281 \text{ ft}} &= 1609.3 \frac{\text{m}}{\text{hr}} \\ 1609.3 \frac{\text{m}}{\text{hr}} \cdot \frac{1 \text{ km}}{1000 \text{ m}} &= 1.6093 \frac{\text{km}}{\text{hr}} \end{aligned}$$

We get a final result which is actually more accurate than the value given in the earlier list.

Outside of an exam or quiz, it is always good to double-check unit conversions with values found online. Typing “1 mph in kmph” into Google pulls up a very useful unit conversion tool.

1.4 Forms of Energy

The various forms of energy of interest to us are introduced in terms of a solid body having a mass m [kg]. These include potential, kinetic and internal energy.

1.4.1 Potential Energy

Potential energy (PE) is associated with the elevation of the body, and can be evaluated in terms of the work done to lift the body from one datum level to another under a constant acceleration due to gravity g [$\frac{m}{s^2}$], as follows:

$$W = \int_{h_1}^{h_2} F dx = \int_{h_1}^{h_2} m \cdot g dx = m \cdot g(h_2 - h_1) = m \cdot \Delta pe = \Delta PE \quad (1.1)$$

Typically, we simplify this to say that $PE = mgh$.

1.4.2 Kinetic Energy

Kinetic energy (KE) of a body is associated with its velocity V [$\frac{m}{s}$] and can be evaluated in terms of the work required to change the velocity of the body, as follows:

$$W = \int F dx = \int m \cdot a dx = \int m \cdot \frac{dV}{dt} dx \quad (1.2)$$

however, velocity $V = \frac{dx}{dt}$, thus integrating from V_1 to V_2 :

$$W = \int_{V_1}^{V_2} m \cdot V dV = m \cdot \left(\frac{V_2^2 - V_1^2}{2} \right) = m \cdot \Delta ke = \Delta KE \quad (1.3)$$

Typically, this is simplified to say that $KE = \frac{1}{2}mV^2$.

1.4.3 Internal Energy

Internal energy (U) of a body is that associated with the molecular activity of the body as indicated by its temperature T [$^{\circ}\text{C}$], and can be evaluated in terms of the heat required to change the temperature of the body having a specific heat capacity c [$\frac{\text{J}}{\text{kg} \cdot ^{\circ}\text{C}}$], as follows:

$$Q = m \cdot c \cdot \Delta T = m \cdot \Delta u = \Delta U \quad (1.4)$$

Unfortunately, the specific heat c changes with temperature. For this reason, it is typically necessary to look up the internal energy of a substance from a table.

Example 1.2: Cooking with Internal Energy

In order to gain an intuitive appreciation for the relative magnitudes of the different forms of energy we consider the (tongue-in-cheek) example of an attempt to cook a turkey by potential energy. The turkey is brought to the top of a 100 m building (about 30 stories) and then dropped from the ledge. The potential energy is thus converted into kinetic energy, and finally on impact the kinetic energy is converted into internal energy. The increase in internal energy is represented by an increase in temperature, and hopefully, if this experiment is repeated enough times the temperature increase will allow the turkey to cook. This remarkable experiment was first reported by R.C.Gimmi and Gloria J Browne – “Cooking with Potential Energy“, published in the Journal of Irreproducible Results (Vol. 33, 1987, pp 21-22).

Potential Energy

$$W = \int_0^h F \, dx = \int_0^h mg \, dx$$

$$W = m \cdot g \cdot h = \Delta PE$$

Internal Energy

$$Q = m \cdot C \cdot \Delta T = \Delta U$$

Kinetic Energy

$$\begin{aligned} W &= \int F \, dx = \int mg \, dx = \int m \frac{dV}{dt} \, dx \\ &= m \int \frac{dx}{dt} \, dV = m \int_0^V V \, dV \end{aligned}$$

$$W = \frac{m \cdot V^2}{2} = \Delta KE$$

Equating all three energy forms:

$$\Delta PE = \Delta KE = \Delta U \, [\text{J}]$$

$$m \cdot g \cdot h = \frac{m \cdot V^2}{2} = m \cdot C \cdot \Delta T$$

Since mass m is common, evaluate specific energy ($h=100 \, \text{m}$):

$$\Delta pe = \Delta ke = \Delta u \, \left[\frac{\text{J}}{\text{kg}} \right]$$

$$g \cdot h = \frac{V^2}{2} = C \cdot \Delta T \approx 1000 \, \left[\frac{\text{J}}{\text{kg}} \right]$$

$$\frac{V^2}{2} \left[\frac{\text{m}^2}{\text{s}^2} \right] = 1000 \left[\frac{\text{J}}{\text{kg}} \cdot \left(\frac{\text{N} \cdot \text{m}}{\text{J}} \right) \cdot \left(\frac{1}{\text{N}} \frac{\text{kg} \cdot \text{m}}{\text{s}^2} \right) \right]$$

$$V_{\text{impact}} = \sqrt{2000} = 44.7 \left[\frac{\text{m}}{\text{s}} \right] (\approx 100 \text{ mph!})$$

We estimate the specific heat of a turkey.

$$c = 3000 \text{ [J/kg}^\circ\text{C]} (\text{a little less than water})$$

Thus

$$c\Delta T = 3000 \Delta T = 1000 \text{ [J/kg]} \implies \Delta T = 0.33^\circ\text{C}$$

What a disappointment! At 0.33°C per fall it will require repeating the experiment 600 times just to reach the cooking temperature of 200°C .

1.5 Basic Properties of Matter

1.5.1 Nomenclature

Symbol	Meaning	SI Units	English Units
ρ	Density	kg/m^3	slug/ft^3 OR lbm/ft^3
p	Pressure	Pascal [Pa]	pounds/foot [psf]
v	Specific Volume	m^3/kg	ft^3/slug OR ft^3/lbm
T	Temperature	$^\circ\text{C}$ OR K	$^\circ\text{F}$ OR $^\circ\text{R}$

1.5.2 Density

Density is the amount of mass in a given volume. If the density is constant throughout the volume, we can say that $m = \rho \cdot \text{Vol}$. The units used for density are composed of a unit of mass, such as kg, and a unit of volume, such as m^3 .

The density of air under standard atmospheric conditions is around $1.2 \text{ kg}/\text{m}^3$. The density of water in normal conditions is about $1000 \text{ kg}/\text{m}^3$. Many solids have higher densities still.

1.5.3 Specific Volume

In Thermodynamics, we often prefer to use **mass-specific** properties. Mass-specific simply means that a property is given per unit mass. Density is unfortunately volume-specific, rather than mass-specific. In order to convert density to a mass-specific property, we simply need to invert it. Specific volume is the result: $v = 1/\rho$, represented in m^3/kg .

1.5.4 Pressure

Gases and liquids (together referred to as fluids) both push on all surfaces they touch. The force they apply is based on their pressure. If pressure is constant over the surface, we

can say that $F = p \cdot A$. A more in-depth study of this phenomenon is covered in Fluid Mechanics.

The basic unit of pressure is the Pascal [Pa], which is identical to a Newton per square meter $\left[\frac{\text{N}}{\text{m}^2}\right]$. However, practical units are kilopascal [kPa], bar [100 kPa] or atm (atmosphere) [101.32 kPa]. The **gauge** (or **vacuum**) pressure is related to the **absolute** pressure as shown in the diagram below:

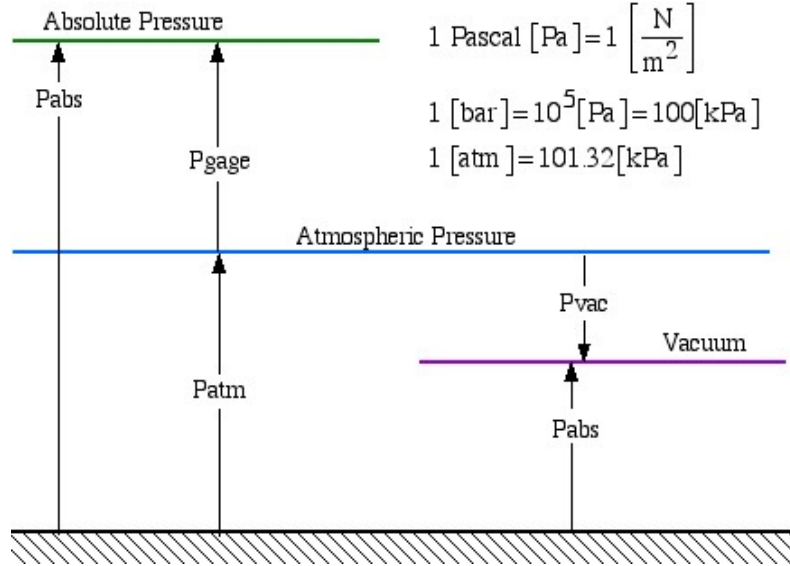


Figure 1.2: Various pressure measurements. Gauge and vacuum pressure are measured compared to atmosphere.

The basic method of measuring pressure is by means of a **manometer**, as shown below:

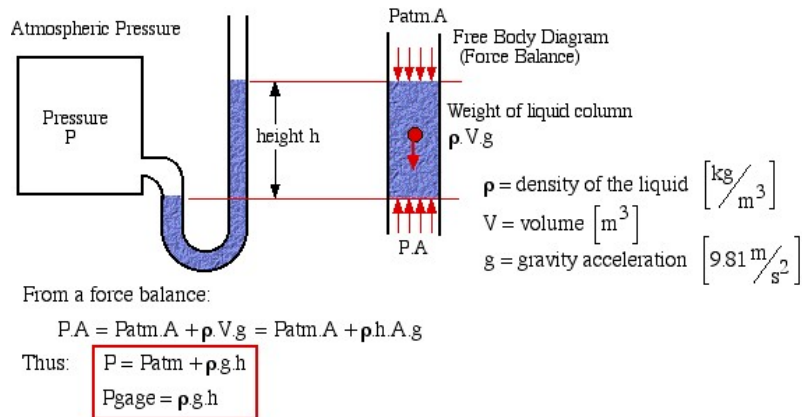


Figure 1.3: A diagram of a manometer, which is used to find the pressure difference between two fluids.

The basic principle here is that the force caused by excess pressure on one side is balanced by the extra weight of fluid on the other side. You can find the difference in pressure through the formula $\Delta p = \rho g \Delta h$. In this case, $\Delta p = p - p_{atm}$.

Manometers can only measure pressure differences, meaning that they are useful for gauge and vacuum pressures, but not for finding absolute or atmospheric pressure.

In order to find atmospheric pressure (and by extension, absolute pressure), a mercury **barometer** can be used as follows:

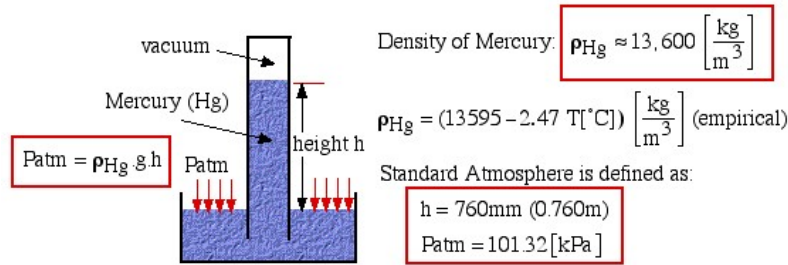


Figure 1.4: A diagram of a barometer, which is used to find the absolute pressure of a fluid.

The key here is the vacuum above the mercury. For a pure vacuum, absolute pressure is 0. With that knowledge, the barometer works the same as the manometer above, with $\Delta p = p_{atm} - 0$.

1.5.5 Temperature

Temperature is a measure of molecular activity, and a temperature difference between two bodies in contact (for example the immediate surroundings and the system) is the driving force leading to heat transfer between them.

Both the Fahrenheit and the Celsius scales are in common usage in the US, hence it is important to be able to convert between them. Furthermore we will find that in some cases we require the Absolute (Rankine and Kelvin) temperature scales (for example when using the Ideal Gas Equation of State).

To explain the difference between Fahrenheit and Celsius, remember that Fahrenheit describes how hot something is to a human (0 being very cold, and 100 being very hot). Celsius does the same for water (0 being freezing, and 100 being boiling). Rankine uses the Fahrenheit scale, but sets the zero point at absolute zero ($0^{\circ}\text{R} = -459.67^{\circ}\text{F}$). Kelvin does the same, but for Celsius ($0 \text{ K} = -273.15^{\circ}\text{C}$).

1.6 Types of Thermodynamic Systems

For purposes of analysis we consider two types of thermodynamic systems: closed systems and open systems.

1.6.1 Closed Systems

Closed systems are usually referred to as **control masses**. This type of system is separated from its surroundings by a physical boundary. Energy in the form of **work** or **heat** can

flow across the system boundary, however there can be no mass flow across the boundary. One typical example of a system is a piston/cylinder device in which the system is defined as the fixed mass of fluid contained within the cylinder.

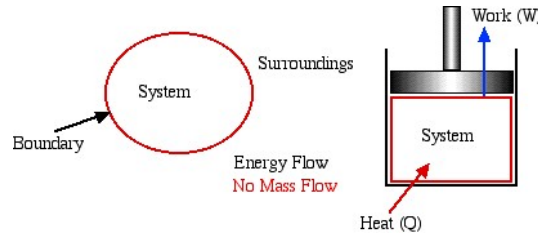


Figure 1.5: A closed system allows energy flow, but no mass flow.

A closed system that additionally restricts energy flow is known as an **isolated system**. An isolated system will always have the same amount of mass and energy that it started with.

1.6.2 Open Systems

Open Systems are usually referred to as **control volumes**. In this case, in addition to work or heat, we have mass flow of the working fluid across the system boundaries through inlet and outlet ports. In this course we will be exclusively concerned with **steady flow** control volumes, in that the net mass of working fluid within the system boundaries remains constant (i.e. mass flow in = mass flow out). The following sections refer mainly to systems – we will consider control volumes in more detail starting with Chapter 4.

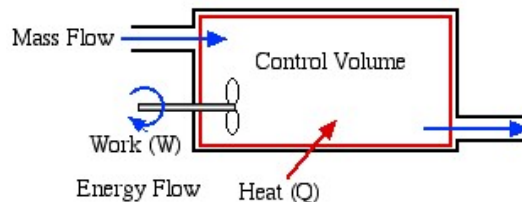


Figure 1.6: An open system allows both energy flow and mass flow.

1.7 Intensive and Extensive Properties

Systems are defined by the properties of the matter that make them up. For instance, you can define a piston by the total amount of mass inside, the temperature, the pressure, etc. A large part of Thermodynamics is defining and calculating the properties of matter.

Properties can be either **extensive** or **intensive**. Extensive properties depend on the “extent” of the system, meaning that if you have more mass in the system, you will have more of the property. Key examples of this are volume and mass, though you can define any of the energies as extensive. Intensive properties are independent of the size of the system. Pressure and temperature are key examples, though most properties can be redefined as

intensive. For instance, density (ρ) and specific volume (v) are both intensive properties defined from extensive properties.

Another example is the specific internal energy (u), which is simply the total internal energy (U) divided by mass:

$$u \left[\frac{\text{kJ}}{\text{kg}} \right] = \frac{U [\text{kJ}]}{m [\text{kg}]}$$

In general, a **specific** property is an intensive property which has been obtained by dividing the extensive property by the mass of the system.

1.8 State and Equilibrium

The **state** of a system is defined by the values of the various intensive properties of the system.

The **state postulate** states that if two independent intensive property values are defined, then all the other intensive property values (and thus the state of the system) are also defined. This can significantly simplify the graphical representation of a system, since only two-dimensional plots are required. Note that pressure and temperature are not necessarily independent properties, thus a boiling liquid will change its state from liquid to vapor at a constant temperature and pressure.

We assume that throughout the system **equilibrium** conditions prevail, which means that there are no temperature or pressure gradients or transient effects. At any instant the entire system is under chemical and phase equilibrium, meaning that all chemical reactions and phase transitions (such as boiling) happen instantly, with no time delay. An alternative viewpoint is that any processes (see Section 1.9) happen slowly enough that any chemical reactions or phase transitions have a chance to complete as we inch along.

Note that this is an assumption, and that in future classes, such as Heat and Mass Transfer, you will be interested in those gradients.

1.9 Process and Cycle

A **process** is a change of state of a system from an initial to a final state due to an energy interaction (work or heat) with its surroundings. For example in the following diagram the system has undergone a compression process in the piston-cylinder device.

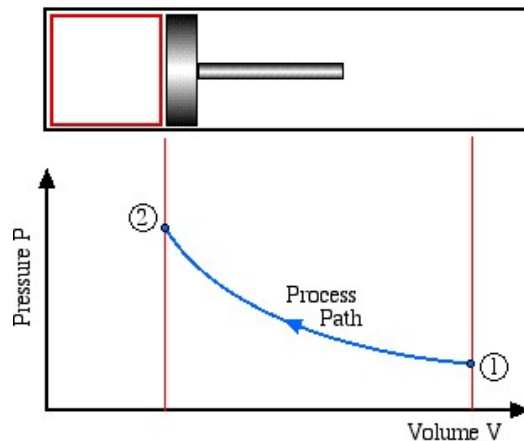


Figure 1.7: A compressing piston increases pressure and reduces volume in a system.

The **process path** defines the type of process undergone. Typical process paths are:

- Isothermal (constant temperature process)
- Isochoric or Isometric (constant volume process)
- Isobaric (constant pressure process)
- Adiabatic (no heat flow to or from the system during the process)

We assume that all processes are **quasi-static** in that equilibrium is attained after each incremental step of the process.

A system undergoes a **cycle** when it goes through a sequence of processes that leads the system back to its original state.

1.10 Using Software: Google Colab (Python)

Google Colab is a free IDE (integrated development environment) for Python. Through Colab, you can write Python code, link to countless existing libraries of code and data, and run the code you write.

1.10.1 Setup and CoolProp Basics

To start off, click the link above (or search Google for “Google Colab”). You want to create a New Notebook, so either select the option from the pop-up menu, or select “File” from the menu bar and click on “New notebook”.

Your empty notebook should look something like Figure 1.8.

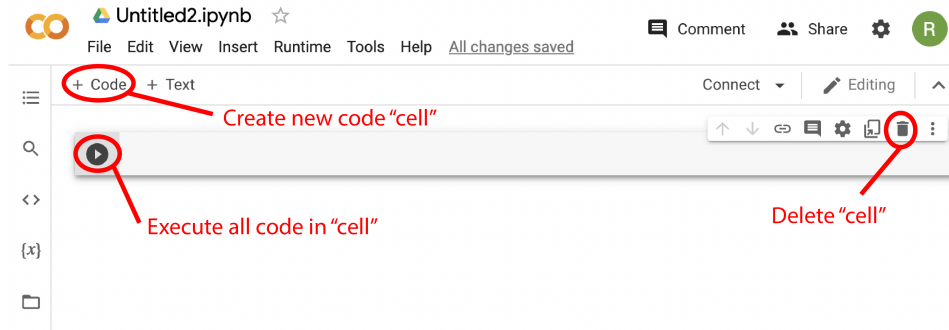


Figure 1.8: An empty notebook in Google Colab, with several important buttons labeled.

Once you have your notebook created, we need to start off with some commands to link to some common libraries:

```
# Clear all variable definitions
%reset -f
from numpy import *           # Import common numerical functions (like sqrt)
from matplotlib.pyplot import * # Import plotting functions (like plot)
```

Everything to the right of the `#` symbol is a comment (text that is ignored by Python). Comments are not part of the code, but are helpful in describing the code in plain language.

The `reset` command removes variable definitions, and essentially makes sure we're working from a clean slate. Some of the most frustrating errors to debug stem from the fact that something is defined in a way that you aren't expecting, and this helps avoid some of those errors.

The `from ... import *` command grabs all of the functions from a library and gives us access to them. We are getting function from the `numpy` library and the `matplotlib` library. The `numpy` library contains functions like `sqrt`, `sin`, and `log`, and generally allows you to do everything you could do on a scientific calculator (plus a lot more we won't talk about here). The `matplotlib` library gives you access to `plot` and all the functions you need to make plots look good.

1.10.2 Basic Math

Let's say that we know the area of a square (in acres) and would like to find out the length of one of the sides. We can make this happen through the following lines of code:

```
areaInAcres = 2.5 # 2.5 acres
areaInSqFt = areaInAcres*43560 # conversion to square feet
sideLength = sqrt(areaInSqFt)  # Area = length^2
```

This will convert from acres to square feet, then take the square root in order to find the length of one of the sides. Running this code will not produce any errors, but at the same time, it won't seem to do anything.

In order to get information out of the code, we need to use the `print` command. The simplest form is shown below:

```
print(sideLength)
```

This will output the numerical value of `sideLength`, but nothing else. The problem with this arises when you are printing several different values over the course of your code. It becomes very easy to forget which number is which, or at the very least waste time trying to figure it out.

What I prefer is a slightly more complicated version that includes a label and units:

```
print('length = ', sideLength, 'm')
```

The little bit of up-front effort means that you are less likely to forget what the number means in the future.

1.10.3 Plotting in Python

In order to make a plot, we first need to accumulate some data. The easiest way of doing this is to manually build a **list**. Once you have two lists, you can plot! You can do this with the following three lines of code:

```
X = [0, 1, 2, 3, 4]
Y = [4, 1, 0, 1, 4]
plot(X,Y)
```

After running, you should end up with a screen that looks like Figure 1.9

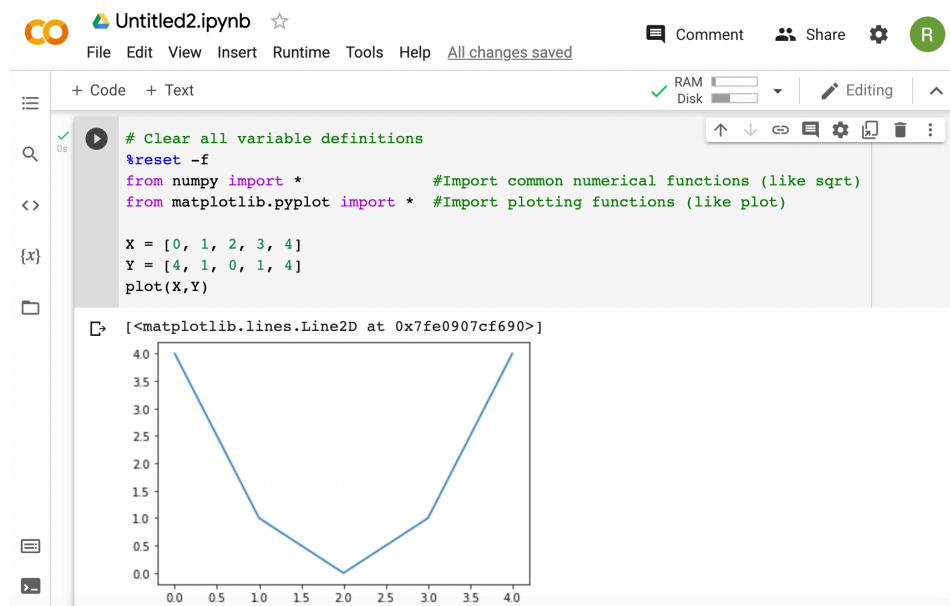


Figure 1.9: Google Colab, displaying a simple plot.

We could give labels, titles, etc., but there's not much point when our data doesn't actually mean anything. Instead, let's return to our original problem of find the side length of a plot of land.

In order to attack this problem, we will build a list and calculate our side length values all at once. To start, we need to define all of the areas we want to work with. Instead of manually putting those in a list, we will use the `linspace` command:

```
areaInAcres = linspace(0, 25, 100) # create 100 points between 0 and 25
```

`linspace` has three **arguments**, or input values. The first and second define the beginning and ending points of the list, respectively. The last argument is the number of points. If you want to see all the values in the list, you can use the command `print(areaInAcres)`.

Python can actually work with lists for most simple functions (like multiplication, `sqrt`, trig functions, etc.). This means we can re-use the code we wrote for Section 1.10.2:

```
areaInAcres = linspace(0, 25, 100) 2.5 # 2.5 acres
areaInSqFt = areaInAcres*43560      # conversion to square feet
sideLength = sqrt(areaInSqFt)       # Area = length^2
```

Then, we can plot `areaInAcres` as the x-axis, and `sideLength` as the y-axis.

```
plot(areaInAcres, sideLength)
```

This will give us a plot, but since our line actually has meaning, we should give the plot labels and a title. While we're at it, we'll also modify the axes and add grid lines.

The label and title commands are very straightforward. The important thing to note is that the arguments must be **strings**, which are collections of characters. You denote a string using quotation marks.

```
xlabel('Plot Area [ac]')
ylabel('Side Length [m]')
title('Length of Side of Land vs. Acreage')
```

Next, we'll use the `axis` command. `axis` takes a list of values as an argument, which should be in the order [`<xmin>`, `<xmax>`, `<ymin>`, `<ymax>`]. We'll implement this as follows:

```
axis([0, 25, 0, 1200])
```

Finally, let's add some grid lines to make data a little easier to extract.

```
grid(visible=True)
```

When you've put everything together, Colab should look like Figure 1.10.

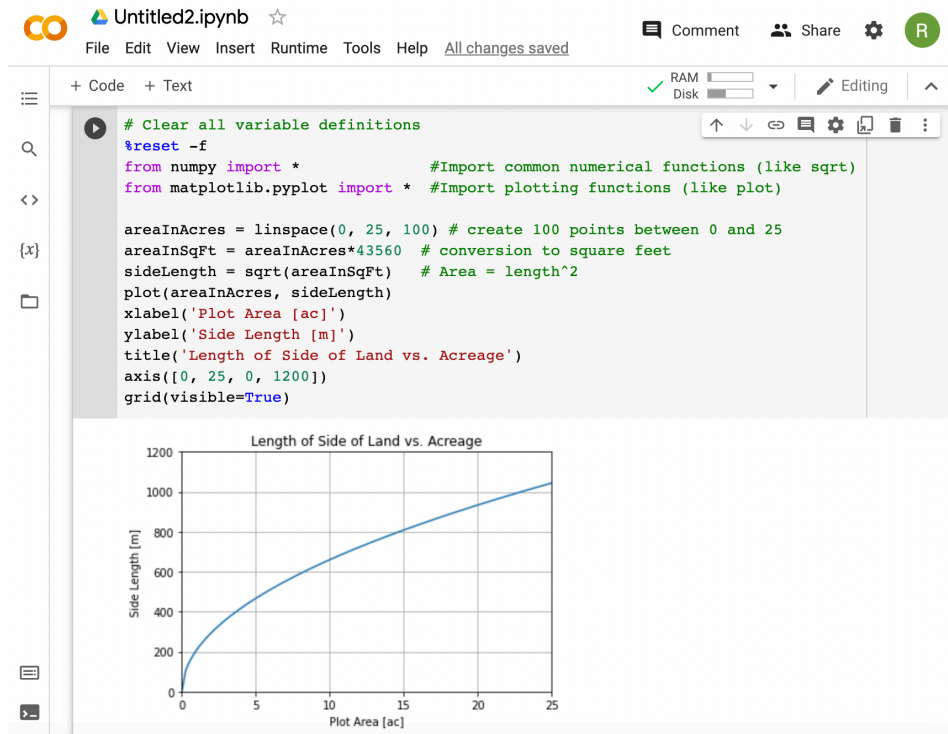


Figure 1.10: Google Colab, displaying a fully labeled plot.

Chapter 1 Homework:**Question 1.1:** How are states, processes, and cycles related?**Question 1.2:** Describe the difference between an intensive and extensive property.**Question 1.3:** For each of the following properties, state whether it is intrinsic or extrinsic.

- Pressure
- Volume
- Kinetic energy
- Specific internal energy

Question 1.4: For each of the following systems, state whether it is open, closed, and/or isolated.

- The ice and water inside of a well-insulated cooler
- An airtight piston, which is heated from the outside
- A water spigot

Question 1.5: Describe the difference between an isothermal process and an adiabatic process.**Question 1.6:** A soda can has a diameter of 2.13 inches and a height of 4.83 inches. Additionally, it has a mass of 384 g. Assuming the can is perfectly cylindrical, and the mass of aluminum is 14.7 g, determine:

- The volume of soda in the can (in ft^3). [0.0100 ft^3]
- The mass of soda in the can (in lbf). [0.812 lbf]
- The density of soda in the can (in $\frac{\text{kg}}{\text{m}^3}$) [1304 $\frac{\text{kg}}{\text{m}^3}$]

Question 1.7: The temperature in Chicago in winter can be as low as 14°F. What is the temperature in °C, K, and °R?**Question 1.8:** A piston with a diameter of 3" compresses air, which is at an **absolute** pressure of 2 atm. What is the net force required to hold the piston still? Provide your answer in lbf. [103.9 lbf]Hint: don't forget atmosphere ($p_{\text{atm}} = 1 \text{ atm}$) pushing on the other side.**Question 1.9:** Rework the previous question in Google Colab. Plot the force required to hold the piston still, while varying the diameter of the piston between 1" and 10". Plot your result, including a title and labels on both the x- and y- axes.**Question 1.10:** A barometer measures 690 mmHg. What is the atmospheric pressure in kPa? [92 kPa]**Question 1.11:** A manometer connects a tank with an unknown pressure to atmosphere. If the manometer is filled with water at 20°C ($\rho = 1000 \text{ kg/m}^3$), and a height difference of 20 cm is measured, what is the gauge pressure in Pa? [1962 Pa] If the atmospheric pressure is 101 kPa, what is the absolute pressure in Pa? [102.96 kPa]

Chapter 2

Properties of Matter

In this chapter we consider the properties and relationships between properties of a pure substance (such as water) which can exist in three phases – solid, liquid and gas. We will not consider the solid phase in this course.

As you read through this chapter, you will notice that we always define a state through two properties. Based on the state postulate, this is sufficient to fully define the state. In other words, all properties of a state can be found, provided we know at least two.

2.1 Phase Change and Property Diagrams

In order to introduce the rather complex phase change interactions that occur in pure substances we consider an experiment in which we have liquid water in a piston-cylinder device at 20°C and 100kPa pressure. Heat is added to the cylinder while the pressure is maintained constant until the temperature reaches 300°C, as shown the T - v diagram (temperature vs. specific volume) in Figure 2.1.

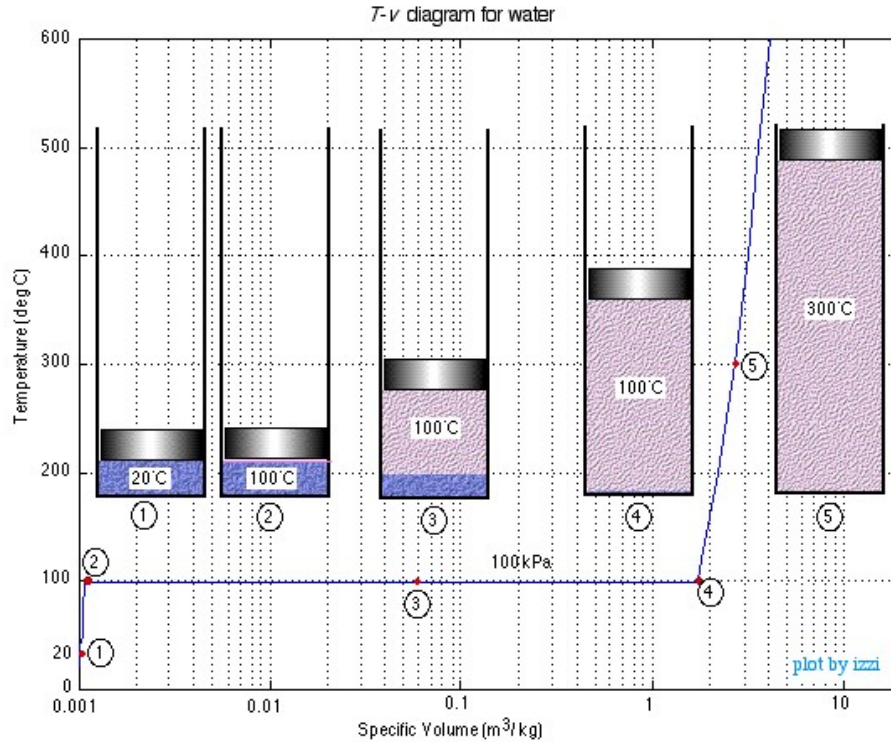


Figure 2.1: Heat is added to water at constant pressure, leading to boiling and an increase of volume.

From State (1) to State (2) the water maintains its liquid phase and the specific volume increases very slightly until the temperature reaches close to 100°C (State (2) – **saturated liquid**). As more heat is added the water progressively changes phase from liquid to water vapor (steam) while maintaining the temperature at 100°C (**saturation temperature** – T_{sat}) until there is no liquid remaining in the cylinder (State (4) – **saturated vapor**). If heating continues then the water vapor temperature increases ($T > T_{sat}$) and is said to be **superheated** (State (5)).

Notice that during this entire process the specific volume of the water increased by more than three orders of magnitude, which made it necessary to use a logarithmic scale for the specific volume axis.

This first experiment was performed at constant pressure. If we increased the force on the piston (perhaps by increasing the mass of the piston), we could increase the pressure in

the system. Then, we could cool the system until it was a liquid, and repeat the experiment again. Figure 2.2 contains data from four such experiments, at 100 kPa, 1 MPa, 10 MPa, and 22 MPa.

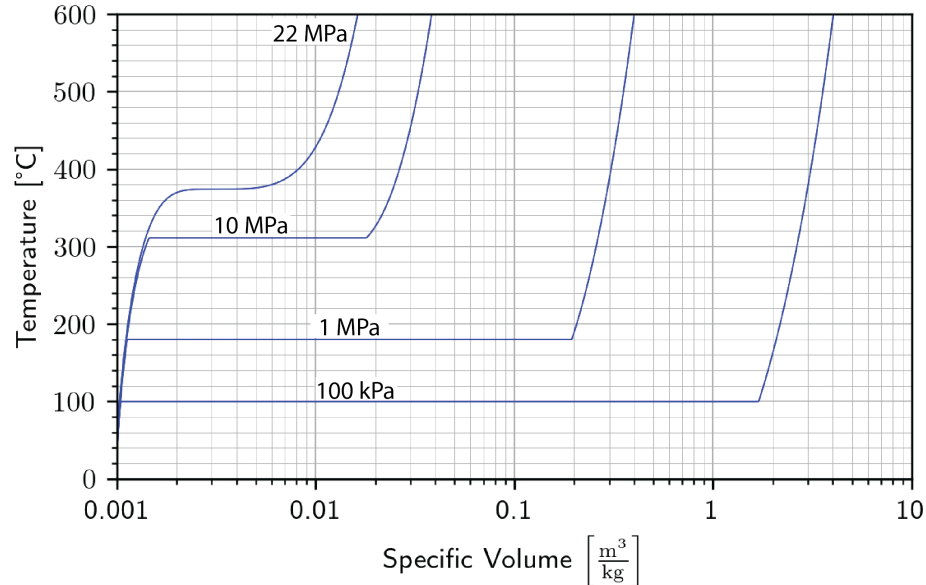


Figure 2.2: Heat is added to water at multiple pressures.

As the pressure increases, the constant temperature region between saturated liquid and saturated vapor becomes smaller and smaller until it is eliminated completely at the critical point. Above the critical point, there is no clear distinction between the liquid and vapor states.

Fluids with temperatures above the **critical temperature** are known as **supercritical fluids** or **superheated** fluids. Increasing the pressure will cause the fluid to be more liquid-like, and decreasing the pressure will cause the fluid to be more gas-like.

To help distinguish between the regions of the T - v diagram, we separate the liquid, vapor, and liquid-vapor mixtures with **saturation lines**. These lines are drawn by noting each saturated liquid and saturated vapor points, and connecting them. The end result is shown in Figure 2.3.

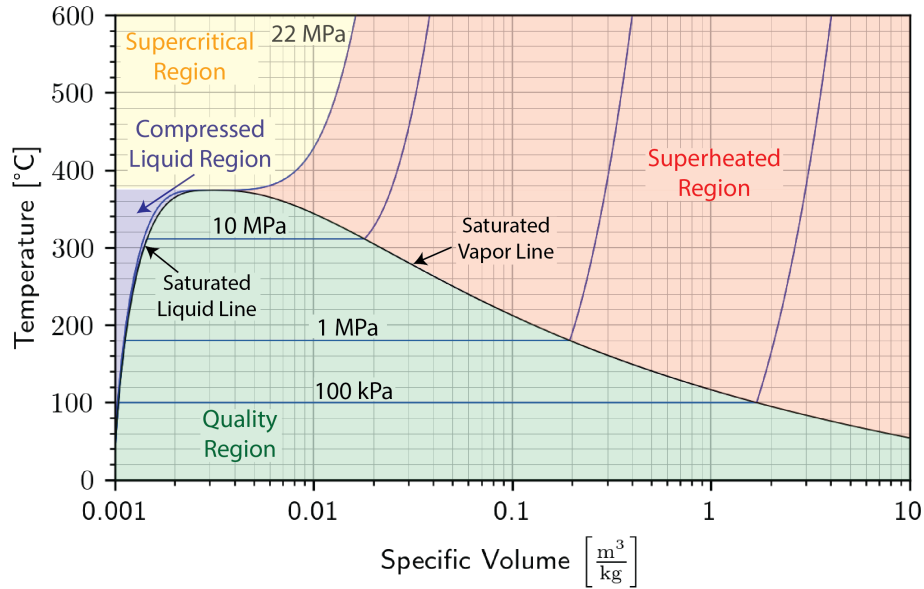


Figure 2.3: T - v diagram with saturation curves shown.

The saturation lines define the regions of interest as shown in the diagram, being the **compressed liquid** region to the left (blue), the **quality** region or **saturated** region enclosed by the saturation lines (green), and the **superheated** region to the right of the saturated vapor line and critical pressure line (red), and the **supercritical** region at temperatures and pressures higher than the critical point (yellow).

We will use **property tables** associated with the regions in order to evaluate the various properties. Notice that we have provided property tables of steam and Refrigerant R134a in Appendix A and B.

2.2 Quality

The **quality region**, also referred to as the **saturated liquid-vapor mixture region**, is the area enclosed between the saturated liquid line and the saturated vapor line. At any point within this region the quality of the mixture (represented by the symbol x) is defined as the mass of vapor divided by the total mass of the fluid, as defined in Equation 2.1.

$$x = \frac{m_{\text{gas}}}{m_{\text{total}}} \quad (2.1)$$

We can also use the **rule of mixtures** to define mass-specific properties. As an example,

this is done for specific volume as follows:

$$\begin{aligned}
 V &= V_f + V_g \\
 mv &= m_f v_f + m_g v_g \\
 mv &= (m - m_g) v_f + m_g v_g \\
 v &= \left(1 - \frac{m_g}{m}\right) v_f + \frac{m_g}{m} v_g \\
 v &= (1 - x) v_f + x v_g
 \end{aligned}$$

Notice that properties relating to the saturated liquid have the subscript f , and those relating to the saturated vapor have the subscript g . If we consider a volume V containing a mass m of a saturated liquid-vapor mixture, we can calculate the quality x as follows:

$$\begin{aligned}
 v &= (1 - x) v_f + x v_g \\
 v &= v_f + x (v_g - v_f) \\
 x &= \frac{v - v_f}{v_g - v_f}
 \end{aligned} \tag{2.2}$$

This is demonstrated graphically in Figure 2.4.

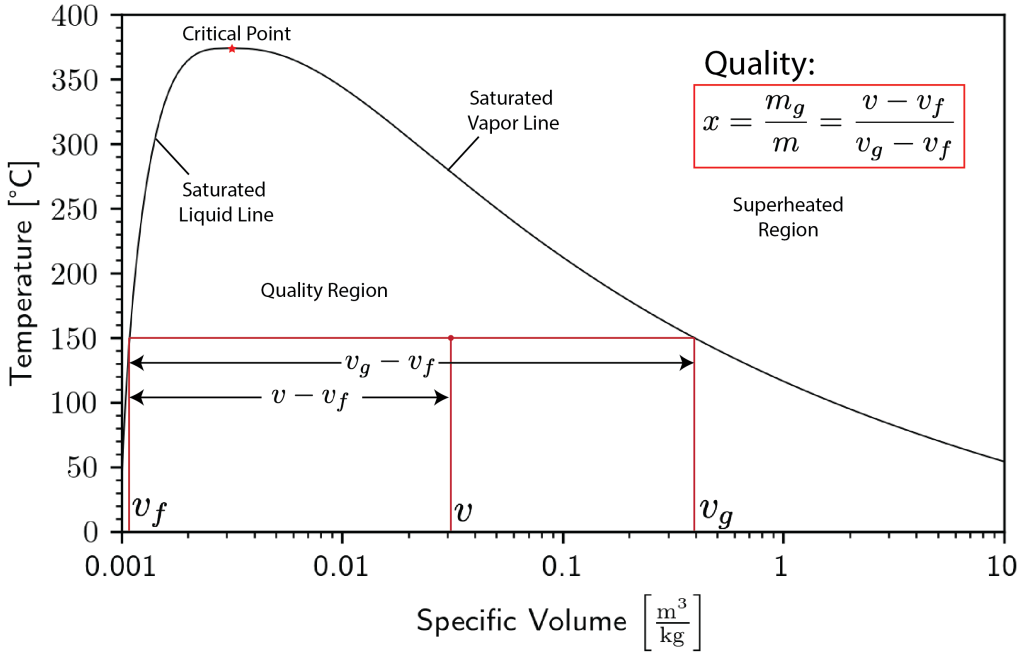


Figure 2.4: Quality is the ratio of the mass of gas to the total mass within the system.

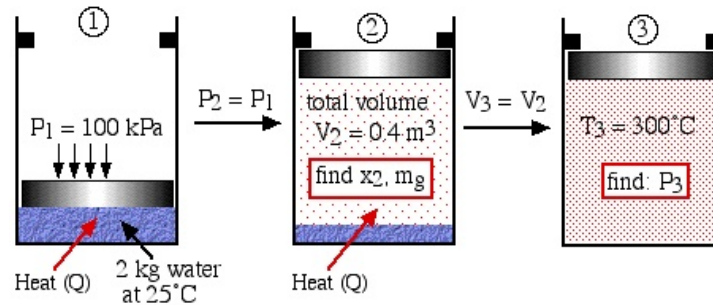
Typically, because of the extremely large range of specific volume values of interest, the T - v diagram can only be done on a semi-log plot (logarithmic scaling on the x -axis). This is

extremely inconvenient, so instead the T - v diagram is normally not drawn to scale. Instead, it is only sketched in order to help define the problem, which is then solved in terms of the steam tables. This approach is illustrated in an example below.

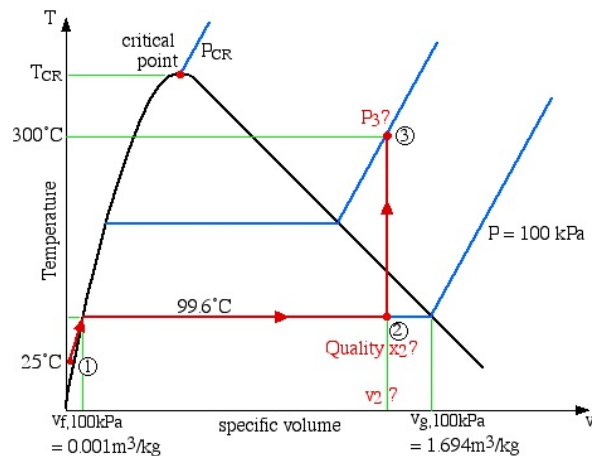
Example 2.1: Boiling Water

Two kilograms of water at 25°C are placed in a piston cylinder device under 100 kPa pressure (state (1)). Heat is added to the water at constant pressure until the piston reaches the stops at a total volume of 0.4 m^3 (state (2)). More heat is then added at constant volume until the temperature of the water reaches 300°C (state (3)). Determine (a) the quality of the fluid and the mass of the vapor at state (2), and (b) the pressure of the fluid at state (3).

Step 1: Diagram Always draw a complete diagram of the states and processes of the problem and include all the relevant information on the diagram. In this case there are three states and two processes (constant pressure and constant volume).



Step 2: p - v or T - v In the case of a closed system with a phase change fluid, always sketch a T - v or p - v diagram indicating all the relevant states and processes on the diagram. As mentioned above this diagram will not be drawn to scale, however it will help to define the problem and the approach to solution. In the case of steam, as we determine various values from the steam tables we add these values to the diagram, typically as shown below:



Step 3: Address Prompt Notice that the T - v diagram is based exclusively on intensive properties, hence mass is not indicated on the diagram. Thus we indicate on the diagram that in order to determine the quality at state (2) we need to first evaluate the specific volume v_2 , which can then be compared to the saturation values v_f and v_g at the pressure of 100 kPa.

Thus:

$$v_2 = \frac{V}{m} = \frac{0.4 \text{ m}^3}{2 \text{ kg}} = 0.2 \frac{\text{m}^3}{\text{kg}}$$

Quality can be found via Equation 2.2:

$$x_2 = \frac{v_2 - v_f}{v_g - v_f} = \frac{0.2 - 0.001}{1.694 - 0.001} \rightarrow x_2 = 0.118$$

Finally, the mass of water vapor:

$$x = \frac{m_g}{m_{tot}} \rightarrow m_{g2} = x_2 \cdot m_{tot} = 0.118 \cdot 2 \text{ kg} \rightarrow m_{g2} = 0.235 \text{ kg}$$

Concerning state (3), the problem statement did not specify that it is in the superheated region. We needed to first determine the saturated vapor specific volume v_g at 300°C. This value is 0.0216 m³/kg, which is much less than the specific volume v_3 of 0.2 m³/kg, thus placing state (3) well into the superheated region. Thus the two intensive properties which we use to determine the pressure at state (3) are $T_3=300^\circ\text{C}$, and $v_3=0.2 \text{ m}^3 / \text{kg}$. On scanning the superheated tables we find that the closest values lie somewhere between 1.2 MPa and 1.4 MPa, thus we use linear interpolation techniques to determine the actual pressure p_3 as shown below:

Superheat Vapor Tables at 300°C

v	0.2139	$v_3 = 0.2$	0.1823	m^3/kg
P	1.2	$P_3 ?$	1.4	MPa

Linear interpolation has a good description from [Wikipedia](#), but we will use it as follows:

$$\frac{\Delta p_A}{\Delta p_b} = \frac{\Delta v_A}{\Delta v_b} \rightarrow \frac{p_3 - 1.2}{1.4 - 1.2} = \frac{0.2 - 0.2139}{0.1823 - 0.2139} = 0.440$$

Therefore, $p_3 = 1.29 \text{ MPa}$.

If you check the steam property tables, you will see that we have also included three new properties: internal energy u [kJ/kg], enthalpy h [kJ/kg], and entropy s [kJ/kgK]. All of these will be defined as needed in future sections. At this stage we note that the 3 equations relating quality and specific volume can also be evaluated in terms of these three additional properties.

$$\begin{aligned}
 u &= (1 - x) u_f + x u_g \\
 h &= (1 - x) h_f + x h_g \\
 s &= (1 - x) s_f + x s_g
 \end{aligned}$$

2.3 The p - v Diagram for Water

The above discussion was done in terms of the temperature (T) and specific volume (v). You may recall from Chapter 1 when we defined the State Postulate however, that any two independent intensive properties can be used to completely define all other intensive state properties. This means we can also evaluate a substance in terms of pressure (p) and specific volume (v) as shown below:

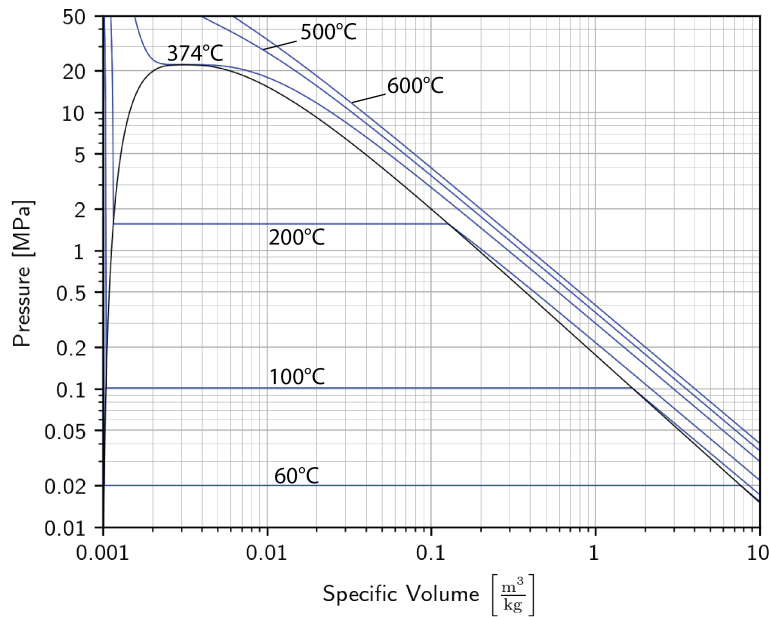


Figure 2.5: Variation of pressure and specific volume for water at various temperatures. Note that both the x- and y- axes are logarithmic.

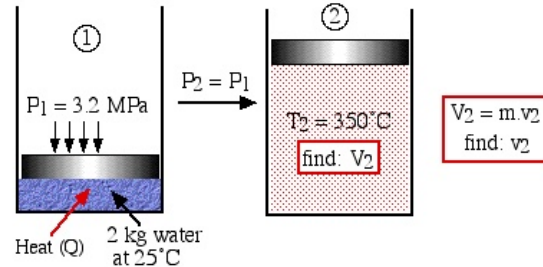
Similar to the T - v diagram, we typically only sketch the p - v diagram to solve problems, as seeing all relevant data requires a log-log plot. Again, the example below illustrates this process.

Example 2.2: Constant Pressure Expansion

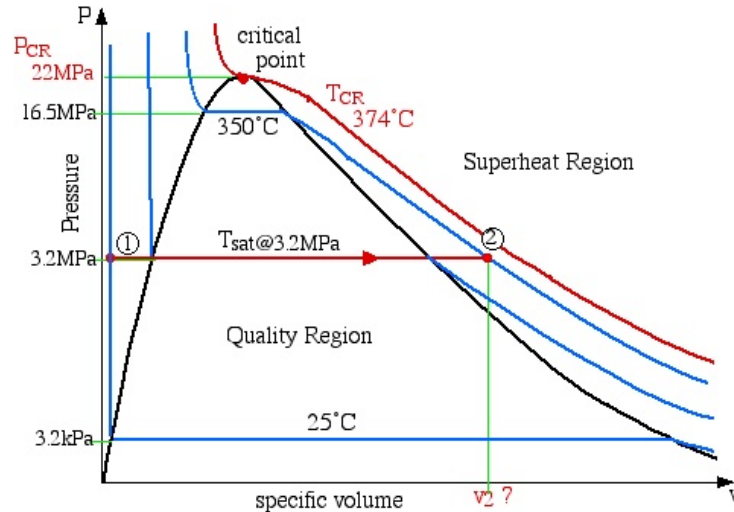
Two kilograms of water at 25°C are placed in a piston cylinder device under 3.2 MPa pressure as shown in the diagram (State (1)). Heat is added to the water at constant pressure until the temperature of the fluid reaches 350°C (State (2)). Determine the

final volume of the fluid at state (2).

Step 1: Diagram In this case, there are only two states.



Step 2: p - v or T - v In this example since the pressure is known (3.2 MPa) and remains constant throughout the process, we find it convenient to draw a p - v diagram indicating the process (1) - (2) as follows.



Step 3: Address Prompt As in the previous example, on scanning the superheated tables we find that we need to interpolate between pressure $p=3.0 \text{ MPa}$ and $p=3.5 \text{ MPa}$ in order to determine the specific volume at the required pressure of 3.2 MPa as follows:

Superheat Vapor Tables				
P	3.0	3.2	3.5	MPa
T_{sat}	233.9°C		242.6°C	
$v @ 350^\circ\text{C}$	0.0906		0.0768	m^3/kg

$$\frac{\Delta p_A}{\Delta p_b} = \frac{\Delta v_A}{\Delta v_b} \rightarrow \frac{3.2 - 3.0}{3.5 - 3.0} = \frac{v_2 - 0.0906}{0.0768 - 0.0906} = 0.4$$

Therefore, $v_2 = 0.085 \frac{\text{m}^3}{\text{kg}}$ and $V_2 = 0.17 \text{ m}^3$.

2.4 Ideal Gas Equation of State

We find that for a pure substance in the superheated region, at specific volumes much higher than that at the critical point, the p - v - T relation can be conveniently expressed by the **ideal gas equation of state** to a high degree of accuracy, as follows:

$$pv = RT \quad (2.3)$$

where: R is constant for a particular substance and is called the **gas constant**.

Note that for the ideal gas equation both the pressure p and the temperature T must be expressed in absolute quantities.

The gas constant R can be expressed as follows:

$$R = \frac{R_u}{M} \left[\frac{\text{kJ}}{\text{kg} \cdot \text{K}} \right] \quad (2.4)$$

where $R_u = 8.314 \frac{\text{kJ}}{\text{kmol} \cdot \text{K}}$, and is known as the **universal gas constant**. M is the molar mass of the gas, measured in $\left[\frac{\text{kg}}{\text{kmol}} \right]$ or $\left[\frac{\text{lbm}}{\text{lbmol}} \right]$ (which yield the same numerical value). For air,

$$M = 28.97 \frac{\text{kg}}{\text{kmol}} \rightarrow R_{air} = 0.287 \frac{\text{kJ}}{\text{kg} \cdot \text{K}}$$

For water/steam,

$$M = 18.02 \frac{\text{kg}}{\text{kmol}} \rightarrow R_{H_2O} = 0.4615 \frac{\text{kJ}}{\text{kg} \cdot \text{K}}$$

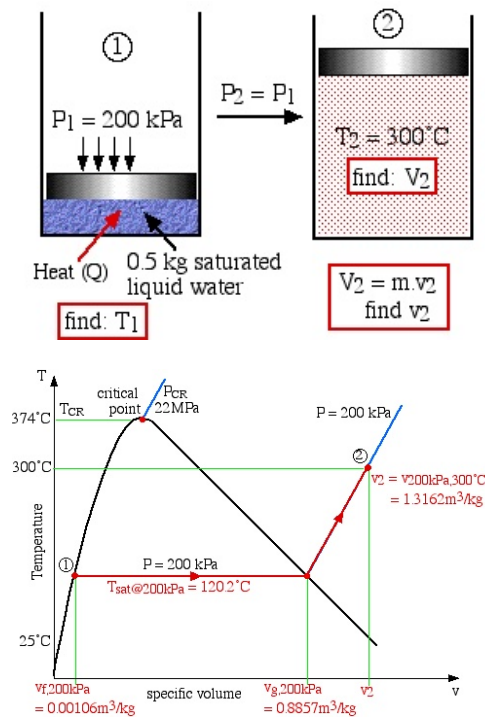
Example 2.3: Checking the Ideal Gas Law

A piston-cylinder device contains 0.5 kg saturated liquid water at a pressure of 200 kPa. Heat is added and the steam expands at constant pressure until it reaches 300°C.

- Draw a diagram representing the process showing the initial and final states of the system.
- Sketch this process on a T - v (temperature-specific volume) diagram with respect to the saturation lines, critical point, and relevant constant pressure lines, clearly indicating the initial and final states.
- Using steam tables determine the initial temperature of the steam prior to heating.
- Using steam tables determine the final volume of the steam after heating.
- Using the ideal gas equation of state determine the final volume of the steam after heating. Determine the percentage error of using this method compared to that of using the steam tables.

Solution:

Even if questions a) and b) were not required, this should always be the first priority item in solving a thermodynamic problem.



c) Since state (1) is specified as saturated liquid at 200 kPa, we use the saturated pressure steam tables to determine that $T_1 = T_{sat}(200\text{kPa}) = 120.2^\circ\text{C}$.

d) From the $T-v$ diagram we determine that state (2) is in the superheated region, thus we use the superheated steam tables to determine that $v_2 = v(200\text{kPa}, 300^\circ\text{C}) = 1.3162 \text{ m}^3/\text{kg}$. Thus $V_2 = m \cdot v_2 = (0.5\text{kg}) \cdot (1.3162 \text{ m}^3/\text{kg}) = 0.658 \text{ m}^3 = 658 \text{ L}$.

e) We use $pv = RT$ as our base equation. Note that pressure and temperature must be absolute!

$$v = \frac{RT}{p} \rightarrow v_2 = \frac{0.4615 \text{ kJ}/(\text{kg} \cdot \text{K}) \cdot (300 + 273) \text{ K}}{200 \text{ kPa}} = 1.322 \frac{\text{m}^3}{\text{kg}}$$

Therefore, $V_2 = m \cdot v_2 = (0.5\text{kg}) \cdot (1.322 \text{ m}^3/\text{kg}) = 0.661 \text{ m}^3 = 661 \text{ L}$.

Finally, we calculate error:

$$\text{error} = \left| \frac{\text{true value} - \text{estimated value}}{\text{true value}} \right| = \left| \frac{0.658 - 0.661}{0.658} \right| \approx 0.005 = 0.5\%$$

As a side-note, we see the unit ratio kJ/kPa quite often. Expanding these out results in:

$$\frac{\text{kJ}}{\text{kPa}} = \frac{\text{kN} \cdot \text{m}}{\text{kN}/\text{m}^2} = \text{m}^3$$

2.5 Non-Ideal Gas Behavior

The T - v diagram for water is shown below, with the addition of error that arises from calculating the pressure through the ideal gas law.

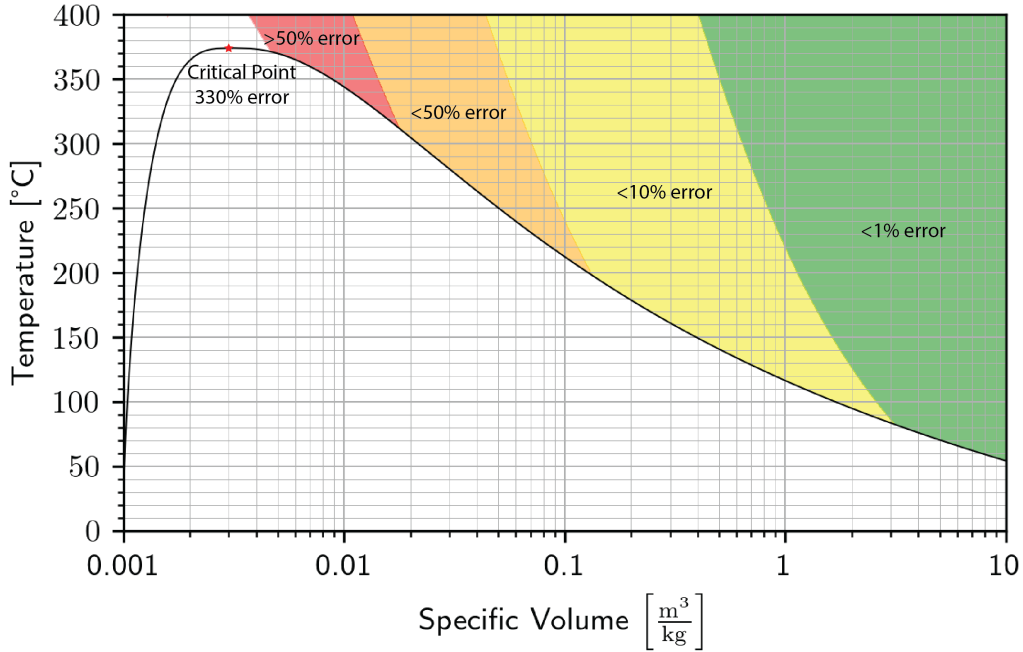


Figure 2.6: T - v diagram, including errors from ideal gas equation of state.

Note that error tends to decrease somewhat as temperature increases, but is much more affected by specific volume (and therefore pressure). Note that at the critical point the error is 330%. This non-ideal behavior can be accounted for by a correction factor called the **compressibility factor** Z defined as follows:

$$pv = ZRT \rightarrow Z = \frac{pv}{RT} \quad (2.5)$$

When the compressibility factor Z approaches 1, the gas behaves as an ideal gas. By specifying both temperature and pressure, the compressibility factor can be expressed as:

$$Z = \frac{v_{actual}}{v_{ideal}} \quad (2.6)$$

All fluids normalized in this manner exhibit similar non-ideal gas behavior within a few percent, thus they can all be plotted on a Generalized Compressibility Chart. A number of these charts are available, however we prefer to use the Lee-Kesler (logarithmic) Compressibility Chart, shown below.

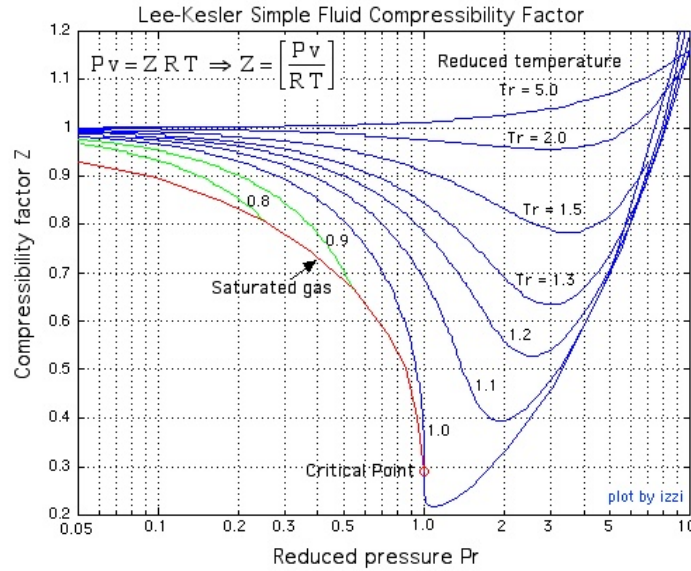


Figure 2.7: Compressibility factor for non-ideal gases.

Notice that the compressibility factor is dependent on the reduced pressure and temperature. These are defined as follows:

$$p_R = \frac{p}{p_{crit}} \qquad T_R = \frac{T}{T_{crit}} \qquad (2.7)$$

Different fluids have different values of critical point pressure and temperature data p_{crit} and T_{crit} , and these can be determined from the Table of Critical Point Data of Various Substances in Appendix C.1.

Example 2.4: Correcting the Ideal Gas Law

Carbon Dioxide gas is stored in a 100 liter tank at 6 MPa and 30°C. Determine the mass of CO₂ in the tank based on (a) the ideal gas equation of state, (b) the generalized compressibility chart, and (c) values obtained from the CO₂ tables of data. Determine the percentage error in each case.

Solution Approach:

There is only one state here, so there is little to gain by drawing the state diagrams. Likewise, the $T-v$ diagram is useful when using tables, but we want to use the ideal gas law for our calculations.

For CO₂, the molar mass is $M = 44.01 \frac{\text{kg}}{\text{kmol}}$, which leads to a gas constant of $R_{CO_2} = 0.1889 \frac{\text{kJ}}{\text{kg} \cdot \text{K}}$. We can use this in the ideal gas equation of state:

$$v_{ideal} = \frac{R \cdot T}{p} = \frac{0.1889 \frac{\text{kJ}}{\text{kg} \cdot \text{K}} \cdot 303 \text{ K}}{6000 \text{ kPa}} = 0.00954 \frac{\text{m}^3}{\text{kg}}$$

Mass can then be found from the equation $V = m \cdot v$:

$$m_{ideal} = \frac{V}{v} = \frac{100 \text{ L}}{0.00954 \frac{\text{m}^3}{\text{kg}}} \cdot \frac{1 \text{ m}^3}{1000 \text{ L}} \rightarrow m_{ideal} = 10.5 \text{ kg}$$

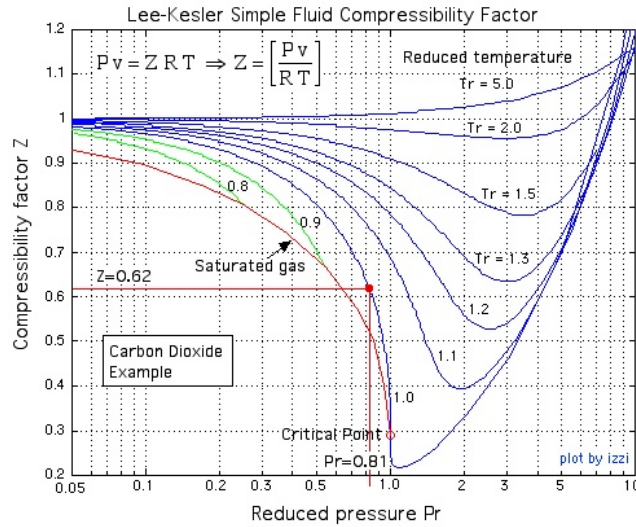
To use the compressibility chart, we need to find the critical temperature and pressure for CO_2 . This comes from the Table of Critical Point Data of Various Substances (Appendix C.1): $p_{crit} = 7.39 \text{ MPa}$, $T_{crit} = 304.2 \text{ K}$.

Next, we calculate our reduced temperature and pressure:

$$p_R = \frac{p}{p_{crit}} = \frac{6 \text{ MPa}}{7.39 \text{ MPa}} = 0.81$$

$$T_R = \frac{T}{T_{crit}} = \frac{30 + 273.15 \text{ K}}{304.2 \text{ K}} \approx 1.0$$

With our reduced pressure and temperature in hand, we go to the compressibility factor chart. We start at our pressure on the x-axis, shoot a line up until we hit our desired reduced temperature line, then the y-axis position of that point will be our Z value.



Since $Z = 0.62$, we can use Equation 2.6:

$$Z = \frac{v_{comp}}{v_{ideal}} \rightarrow v_{comp} = 0.62 \cdot 0.00954 \frac{\text{m}^3}{\text{kg}} = 0.00591 \frac{\text{m}^3}{\text{kg}}$$

Mass is found the same way as before:

$$m_{comp} = \frac{V}{v} = \frac{100 \text{ L}}{0.00591 \frac{\text{m}^3}{\text{kg}}} \cdot \frac{1 \text{ m}^3}{1000 \text{ L}} \rightarrow m_{comp} = 16.9 \text{ kg}$$

Finally, we'll get the most accurate answer from the CO₂ tables:

$$v_{actual} = 0.005833 \frac{\text{m}^3}{\text{kg}} \rightarrow m_{actual} = 17.14 \text{kg}$$

This means that the compressibility chart had an error of about 1% (mostly due to the inaccuracy of reading charts). The ideal gas law without a compressibility correction had an error of around 39%!

2.6 Using Software: Google Colab (Python) with CoolProp

As we saw in Section 1.10, [Google Colab](#) is a free IDE (integrated development environment) for Python. Through Colab, you can write Python code, link to countless existing libraries of code and data, and run the code you write.

Hopefully, you already have a basic understanding of how to use Python for your homework. In this section, we will be using a library called [CoolProp](#) to determine the properties of materials, instead of looking up values from tables.

2.6.1 Setup and CoolProp Basics

First off, in order to access the CoolProp library, we need to input the line

```
!pip install CoolProp
```

This tells Google Colab that we plan on using the library, and it should expect to link to it later.

In a separate cell (which is created using the “+ Code” button at the top left), you should include the following lines:

```
# Clear all variable definitions
%reset -f
from numpy import *           # Import common numerical functions (like sqrt)
from matplotlib.pyplot import * # Import plotting functions (like plot)
import CoolProp.CoolProp as CP # Import CoolProp library
```

The only new line in these commands is the “import CoolProp” line, which simply links to the CoolProp library. The phrasing we used means that we need to add “CP.” in front of any CoolProp commands we end up using. If you want to avoid this, you can type “from CoolProp.CoolProp import *” instead. In order to match the rest of the example, I would suggest leaving it as written.

You can now run the two cells by clicking the “play” button that appears at the top left of each cell. You can also run all of the cells in the script by clicking “Runtime” from the top menu, then selecting “Run all”.

2.6.2 A Simple Example

At this point, we are ready to build our first bit of code. Let's start by reworking Example 2.4.

We can calculate the ideal specific volume with the following lines of code:

```
R = 8.31446          # Universal Gas Constant (in kJ/kgK)
M = 44.01           # Molar Mass of CO2
RCO2 = R/M          # CO2 gas constant
p = 6000            # pressure in kPa
T = 30 + 273.16     # Conversion of temperature from Celsius to Kelvin
vIdeal = RCO2 * T / p # Ideal Gas Law solved for v
print('Ideal v = ', vIdeal, 'm^3/kg')
```

Note that the `print` statement at the end is needed to see the results:

```
Ideal v = 0.009545602111641295 m^3/kg
```

You can use the simpler `print(vIdeal)` as a shortcut, which will only print the number. However, with multiple print statements, you can lose track of the meanings of the numbers, which is I prefer the slightly longer version used above.

2.6.3 Using CoolProp

The general form of the CoolProp function is as follows:

```
<outputVariable> = CP.PropsSI(<outputType>, <input1Type>, <input1Value>,
<input2Type>, <input2Value>, <materialName>)
```

You can see the full list of Types [online](#), or you can use the short list below:

Table 2.1: List of CoolProp Types and Meanings

Type	Meaning
'D'	Density [kg/m ³]
'H'	Enthalpy [J/kg]
'P'	Pressure [Pa]
'Q'	Quality [-]
'S'	Entropy [J/kgK]
'T'	Temperature [K]
'U'	Internal Energy [J/kg]

Unfortunately, there is no option for specific volume. Instead, we must use the identity that $v = 1/\rho$ to switch between density and specific volume. Using this identity, the second half of Example 2.4 can be recreated using the CoolProp library:

```
p = 6000*1000      # Conversion of pressure from kPa to Pa
T = 30 + 273.16    # Conversion of temperature from Celsius to Kelvin
dActual = CP.PropsSI('D', 'P', p, 'T', T, 'CO2')
vActual = 1/dActual
print('Actual v = ', vActual, 'm^3/kg')
```

Note the conversion necessary for pressure (from kPa to Pa), as well as the conversion for temperature (from Celsius to Kelvin). This code has the following output:

Actual $v = 0.005833989613407337 \text{ m}^3/\text{kg}$

The difference between this result and Example 2.4 stems from the inaccuracy of reading from a plot.

2.6.4 Exporting from Colab as a PDF

You can create a PDF of your code, including the output, by clicking the “File” menu option and navigating to “Print” at the bottom.

In the dialog that pops up, choose “Save as PDF” in the Destination drop-down menu, then click save. The resulting PDF for the work above is shown in Figure 2.8.

```
1/4/22, 11:20 AM                               Untitled0.ipynb - Colaboratory

!pip install CoolProp

Collecting CoolProp
  Downloading CoolProp-6.4.1-cp37-cp37m-manylinux1_x86_64.whl (4.2 MB)
    |████████████████████████████████████████| 4.2 MB 4.2 MB/s
Installing collected packages: CoolProp
Successfully installed CoolProp-6.4.1

# Clear all variable definitions
%reset -f
from numpy import *           # Import common numerical functions (like sqrt)
from matplotlib.pyplot import * # Import plotting functions (like plot)
import CoolProp.CoolProp as CP # Import CoolProp library

R = 8.31446                    # Universal Gas Constant (in kJ/kgK)
M = 44.01                      # Molar Mass of CO2
RCO2 = R/M                     # CO2 gas constant
p = 6000                       # pressure in kPa
T = 30 + 273.16                # Conversion of temperature from Celsius to Kelvin
vIdeal = RCO2 * T / p          # Ideal Gas Law solved for v
print('Ideal v = ', vIdeal, 'm^3/kg')

↩ Ideal v = 0.009545073998333714 m^3/kg

# Clear all variable definitions
%reset -f
from numpy import *           # Import common numerical functions (like sqrt)
from matplotlib.pyplot import * # Import plotting functions (like plot)
import CoolProp.CoolProp as CP # Import CoolProp library

p = 6000*1000                  # Conversion of pressure from kPa to Pa
T = 30 + 273.16                # Conversion of temperature from Celsius to Kelvin
dActual = CP.PropsSI('D', 'P', p, 'T', T, 'CO2')
vActual = 1/dActual
print('Actual v = ', vActual, 'm^3/kg')

Actual v = 0.005833989613407337 m^3/kg
```

Figure 2.8: Output from Google Colab saved as a PDF.

2.7 Summary

In this chapter, we determined the state of matter, which was determined by two independent properties, and then used that state to determine any missing properties. As part of this process, we introduced linear interpolation to find values from the table that are not explicitly listed. We also considered a number of processes, which are simply the change between two states.

Finally, we looked at two alternatives to the steam tables: the ideal gas law and software. The ideal gas law requires a lot more calculation, but can be used when the steam tables are not available (which is typically the case). Most importantly, a compressibility correction is necessary for low temperatures and/or high pressures. Software can be used in most circumstances, but experience in coding is necessary to obtain accurate solutions.

Chapter 2 Homework:

Question 2.1: How many independent quantities are necessary to fully describe a state of matter?

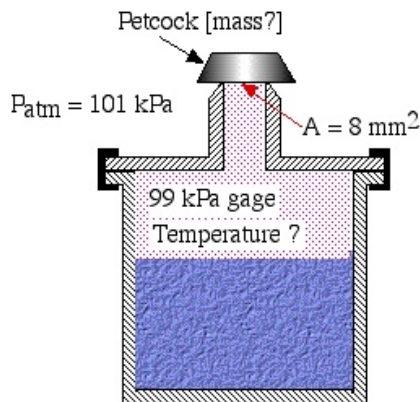
Question 2.2: Define the following: phase, state, property, process

Question 2.3: Two kilograms of water at 25°C are placed in a piston cylinder device under 3.2 MPa pressure. Heat is added to the water at constant pressure until the temperature of the fluid reaches 350°C (State (2)). Determine the final volume of the fluid at state (2). $[0.08508 \text{ m}^3/\text{kg}]$

Question 2.4: A piston-cylinder device contains a saturated mixture of steam and water having a total mass of 0.5 kg at a pressure of 160 kPa and an initial volume of 100 liters. Heat is then added and the fluid expands at constant pressure until it reaches a saturated vapor state.

- Draw a diagram representing the process showing the initial and final states of the system.
- Sketch this process on a P-v diagram with respect to the saturation lines, critical point, and relevant constant temperature lines, clearly indicating the initial and final states.
- Determine the initial quality and temperature of the fluid mixture prior to heating.
 $[x_1 = 0.182, T_1 = 113.3^\circ\text{C}]$
- Determine the final volume of the steam after heating. $[0.546 \text{ m}^3 \text{ (546 liters)}]$

Question 2.5: A pressure cooker allows much faster (and more tender) cooking by maintaining a higher boiling temperature of the water inside. It is well sealed, and steam can only escape through an opening on the lid, on which sits a metal petcock. When the pressure overcomes the weight of the petcock, the steam escapes, maintaining a constant high pressure while the water boils.



Assuming that the opening under the petcock has an area of 8 mm^2 , determine:

- the mass of the petcock required in order to maintain an operating pressure of 99 kPa gauge. $[80.7 \text{ g}]$

- b) the corresponding temperature of the boiling water. [120.2°C]

Note: Assume that the atmospheric pressure is 101 kPa. Draw a free body diagram of the petcock.

Question 2.6: Consider a rigid container having a volume of 100 liters, filled with steam at an initial state of 400 kPa and 300°C. The steam is then cooled until it reaches a temperature of 90°C.

- Draw a diagram representing the process showing the initial and final states of the system.
- Using steam tables determine the mass of steam in the container. [0.153 kg]
- Using the ideal gas equation of state determine the mass of steam in the container. [0.151 kg] Determine the percentage error of using this method compared to that of using the steam tables. [1%]
- Sketch this process on a T - v (temperature-specific volume) diagram with respect to the saturation lines, critical point, and relevant constant pressure lines, clearly indicating the initial and final states.
- Using steam tables determine the final pressure and quality of the fluid mixture after cooling. [70.2 kPa, $x = 0.277$]

Note: The critical point data and ideal gas constant for steam can be found on the first page of the steam tables.

Question 2.7: An automobile tire with a volume of 100 liters is inflated to a gauge pressure of 210 kPa. Determine:

- the mass of air in the tire if the temperature is 20°C [$m = 0.369$ kg]
- the increase in gauge pressure if the temperature in the tire reaches 50°C [$p_{2,gage} = 242$ kPa]

Assume that atmospheric pressure is 100 kPa.

Question 2.8: Compressed air is commonly used to power a large variety of power tools. Lowe's sells an air compressor that can fill an 8-gallon tank to 160 psi. At a temperature of 70°F, determine the mass of the air inside a full 8-gallon tank. Let $p_{atm} = 14.7$ psi.

- Use the ideal gas law (you will need to do a lot of unit conversions for this). [0.429 kg]
- Find the compressibility factor. How far off is your analysis above? [0.99]

Chapter 3

The First Law for Closed Systems

3.1 Introducing the First Law

The First Law of Thermodynamics states that energy is conserved (i.e. neither created nor destroyed). As we discussed in Section 1.4, we can consider many types of energy (potential, kinetic, chemical, nuclear, and internal), but most of those are either small compared to internal (such as potential and kinetic) or outside the scope of this class (nuclear and chemical).

Energy can be transferred between the system and the surroundings in the form of heat and work, resulting in a change of internal energy of the system. Internal energy change can be considered as a measure of molecular activity associated with change of phase or temperature of the system.

The statement of the first law for closed systems reflects that we are primarily interested in the internal energy (U), heat transfer (Q), and work (W).

$$Q - W = \Delta U \quad (3.1)$$

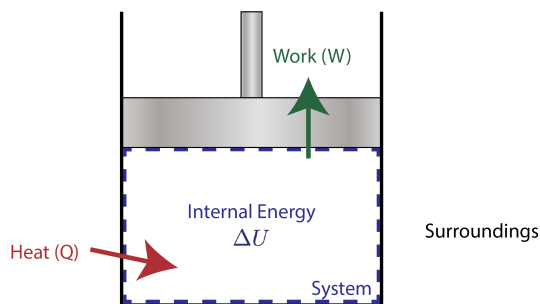


Figure 3.1: A closed system with heat transfer and work.

We can also write the first law by reducing each value to their **mass-specific** counterparts. Thus, Q becomes mq , where q is the mass-specific heat transfer. W becomes mw ,

where w is the mass-specific work. U becomes mu , where u is the mass-specific internal energy. As a general rule, a capital variable refers to the extensive value, while the lower-case variable refers to an intensive value (mass is the biggest exception).

After we divide through by mass (m), the first law therefore becomes:

$$q - w = \Delta u$$

3.1.1 Heat Transfer (Q)

Energy transferred across the boundary of a system in the form of heat always results from a difference in temperature between the system and its immediate surroundings. Heat transfer can occur through conduction, convection, or radiation.

Conduction is the transfer of heat through direct contact. When you touch an object hotter than your skin (a stove top, food heated in the oven, etc.), energy moves from the hot object to your colder skin. The same occurs in reverse when you touch an object colder than your skin. You typically perceive the movement of heat (not the temperature of the object itself).

Convection is the transfer of heat through contact with a moving fluid, such as air or water. Still air typically feels warmer than moving air (even when both are the same temperature) because the rate of heat transfer is higher. This leads to the concept of “wind chill” in North America, the “apparent temperature” in Australia, among others. These models attempt to show how windy days feel colder than a thermometer reading would indicate. The **wind chill** is calibrated so that the rate of heat transfer is the same between a person walking in the windy weather and on a still day at the wind chill temperature. The increased heat transfer from convection is also why fans can make a room feel cooler, even if they don’t decrease the temperature.

Radiation is the transfer of heat through the absorption of light. Photons are emitted from all substances above absolute zero, and these photons carry energy. You feel the heat of a fire not through contact with the superheated air, but because the air is hot enough to emit a significant amount of light in the infrared and visible spectrums, which is then absorbed by your skin. Likewise, you can feel the heat of the sun because the photons are able to travel through the vacuum of space.

The calculation of heat transfer due to temperature differences is left for a later class. In this book, the quantity of heat transferred during any process will either be specified as part of the problem or evaluated as the unknown of the energy equation.

By convention, **heat is that transferred into the system from the surroundings is considered positive**. Positive heat transfer results in an increase in internal energy of the system.

Heat transfer is **path-dependent** and not a property. It is dependent on the process path between the initial and final states.

Recall in Section 1.9 that we introduced some commonly used processes:

- Isothermal (constant temperature process)
- Isochoric or isometric (constant volume process)
- Isobaric (constant pressure process)
- Adiabatic (no heat flow to or from the system during the process)

3.1.2 Work (W)

In this course we are primarily concerned with **boundary work** due to compression or expansion of a system in a piston-cylinder device as shown above. In all cases we assume a perfect seal (no mass flow in or out of the system), no loss due to friction, and quasi-equilibrium processes in that for each incremental movement of the piston equilibrium conditions are maintained.

By convention positive work is that done by the system on the surroundings, and negative work is that done by the surroundings on the system. Since negative work results in an increase in internal energy of the system, this explains the negative sign in Equation 3.1.

Boundary work is evaluated by integrating the force F multiplied by the incremental distance moved dx between an initial state (1) to a final state (2). We normally deal with a piston-cylinder device, thus the force can be replaced by the piston area A multiplied by the pressure p , allowing us to replace $A dx$ by the change in volume dV , as follows:

$$W_{1-2} = \int_1^2 F dx = \int_1^2 pA dx = \int_1^2 p dV = m \int_1^2 p dv \quad (3.2)$$

Figure 3.2 provides an example of a piston being compressed. However, until the process path is specified, it is impossible to know how much work is associated with the compression. This is seen in the figure as each process has a different area under the curve, which is directly related to the integral in Equation 3.2. In other words, work done is **path-dependent** and not a property, exactly like heat transfer.

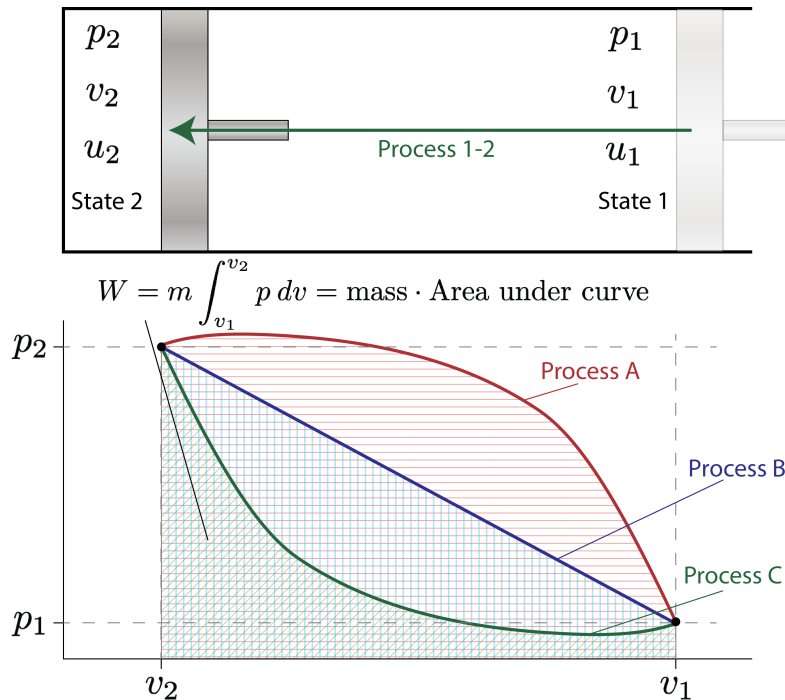


Figure 3.2: The state of a piston plotted as air is compressed. Multiple paths can be taken to connect the same two states. Processes A, B, and C will have different amounts of work and heat transfer associated with them.

We note that work done by the system on the surroundings (expansion process) is positive, and that done on the system by the surroundings (compression process) is negative.

Finally, **shaft work** (due to a paddle wheel) and **electrical work** (due to a voltage applied to an electrical resistor or motor driving a paddle wheel) will always be negative (work done on the system). Positive forms of shaft work, such as that due to a turbine, will be considered in Chapter 4 when we discuss open systems.

3.1.3 Internal Energy (ΔU)

The third component of Equation 3.1 is the change of internal energy. Recall from Section 1.4.3 that internal energy is closely tied to the temperature of a substance. In fact, specific internal energy is a property of the system, like temperature and pressure. From the State Postulate, we can find the internal energy of a state as long as we know at least two other properties.

Often, we prefer to use the mass-specific form of internal energy, which is related to the internal energy in our system as follows:

$$u = \frac{U}{m} \rightarrow \Delta u = \frac{\Delta U}{m}$$

Internal energy can be determined from the steam tables in Appendix A or the R134a tables in Appendix B. Heat transfer and work cannot be found from tables because they are not a property of a state, but rather a description of how a process moves between states.

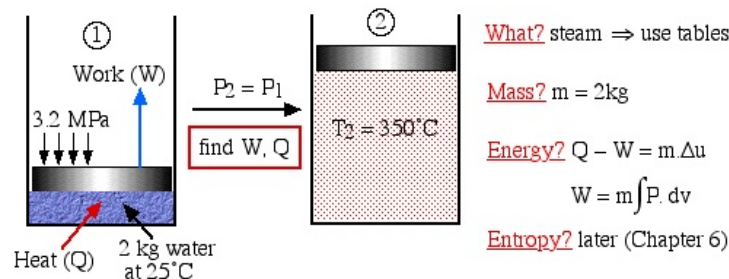
Example 3.1 makes use of the First Law to calculate the heat transfer for a constant pressure process.

Example 3.1: Constant Pressure Expansion Revisited

Recall Example 2.2 in which we presented a constant pressure process. We wish to extend the problem to include the energy interactions of the process, hence we restate it as follows:

Two kilograms of water at 25°C are placed in a piston cylinder device under 3.2 MPa pressure as shown in the diagram (State (1)). Heat is added to the water at constant pressure until the temperature of the steam reaches 350°C (State (2)). Determine the work done by the fluid (W) and heat transferred to the fluid (Q) during this process.

Solution Approach We first draw the diagram of the process including all the relevant data as follows:



Notice the four questions to the right of the diagram, which we should always ask

before attempting to solve any thermodynamic problem.

- **What?**

What are we dealing with - liquid? pure fluid, such as steam or refrigerant? ideal gas? In this case it is steam, thus we will use the steam tables to determine the various properties at the various states.

- **Mass?**

Is the mass or volume given? If so we will specify and evaluate the energy equation in kiloJoules rather than specific quantities (kJ/kg).

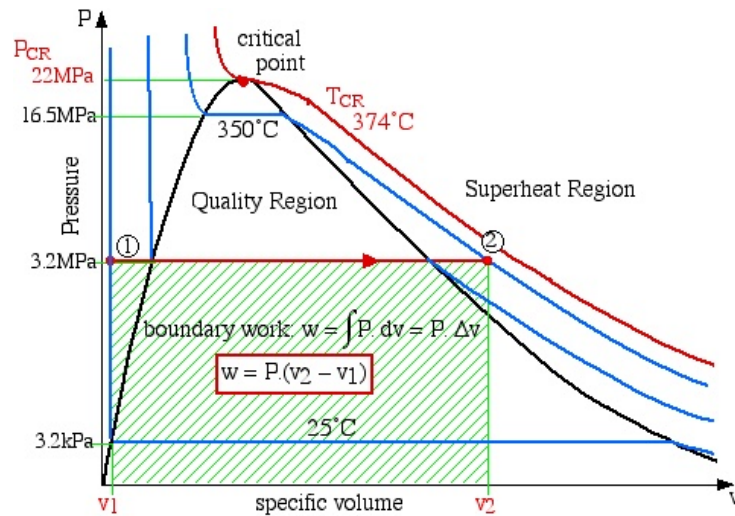
- **Energy?**

The energy equation will almost always be our starting point for working through these problems.

- **Entropy?**

What about entropy? Not so fast - wait until Chapter 6.

Since work involves the integral of $p \cdot dv$ we find it convenient to sketch the p - v diagram of the problem as follows:



Notice on the p - v diagram how we determine the specific work done as the area under the process curve. We also notice that in the compressed liquid region the constant temperature line is essentially vertical. Thus all the property values at State (1) (compressed liquid at 25°C) can be determined from the saturated liquid table values at 25°C .

$$u_1 = u_f, 25^\circ\text{C} = 104.8 \left[\frac{\text{kJ}}{\text{kg}} \right]$$

$$v_1 = v_f, 25^\circ\text{C} = 0.001 \left[\frac{\text{m}^3}{\text{kg}} \right]$$

Superheat Vapor Tables @ 350°C

P	3.0	3.2	3.5	MPa
v	0.0906		0.0768	m ³ /kg
u	2844.4		2836.0	kJ/kg

$$\frac{\Delta p_A}{\Delta p_b} = \frac{\Delta v_A}{\Delta v_b} \rightarrow \frac{3.2 - 3.0}{3.5 - 3.0} = \frac{v_2 - 0.0906}{0.0768 - 0.0906} = 0.4$$

$$v_2 = 0.085 \frac{\text{m}^3}{\text{kg}}$$

$$\frac{\Delta p_A}{\Delta p_b} = \frac{\Delta u_A}{\Delta u_b} \rightarrow \frac{3.2 - 3.0}{3.5 - 3.0} = \frac{u_2 - 2844.4}{2836.0 - 2844.4} = 0.4$$

$$u_2 = 2841 \frac{\text{kJ}}{\text{kg}}$$

Now we use Δv to define our work:

$$W = m \int p dv = m \cdot p \cdot \Delta v = 2.0 \text{ kg} \cdot 3200 \text{ kPa} \cdot (0.085 - 0.001) \frac{\text{m}^3}{\text{kg}}$$

$$W = 538 \text{ kJ}$$

Then, with work and ΔU , we can find our heat transfer:

$$\Delta U = Q - W \rightarrow Q = m \cdot (u_2 - u_1) + W$$

$$Q = 2.0 \text{ kg} \cdot (2841 - 104.8) \frac{\text{kJ}}{\text{kg}} + 538 \text{ kJ}$$

$$Q = 6010 \text{ kJ}$$

3.2 Enthalpy (h)- A New Property

In the case studies that follow we find that one of the major applications of the closed system energy equation is in heat engine processes in which the system is approximated by an ideal gas, thus we will develop relations to determine the internal energy for an ideal gas. We will find also that a new property called **enthalpy** will be useful both for closed systems and in particular for open systems, such as the components of steam power plants or refrigeration systems. Enthalpy is not a fundamental property. Instead, it is a combination of properties and is defined as follows:

$$h = u + p \cdot v \quad \text{OR} \quad H = U + p \cdot V \quad (3.3)$$

As with other properties, the lowercase h refers to the mass specific property and is measured in $\frac{\text{kJ}}{\text{kg}}$, whereas the capital H is an extensive property and is measured in kJ.

Enthalpy shows its use whenever applied to a constant pressure process, such as Example 3.1. For these processes, work can be defined as $w = p \cdot \Delta v$, meaning that the energy equation can be written as:

$$\begin{aligned}\Delta u &= q - p\Delta v \\ u_2 - u_1 &= q - p(v_2 - v_1) \\ (u_2 + p v_2) - (u_1 + p v_1) &= q \\ h_2 - h_1 &= q \rightarrow \boxed{q = \Delta h}\end{aligned}$$

This means that we can look up enthalpy directly to find heat transfer **for constant pressure processes**.

3.3 Specific Heats

For solids and liquids, we can define a relatively simple equation:

$$Q = mc\Delta T \quad \text{OR} \quad q = c\Delta T \quad (3.4)$$

The variable c is called the **specific heat**, and is a relationship between the amount of heat added to a material and the temperature increase we expect. This is often discussed in chemistry alongside devices known as calorimeters.

3.3.1 Constant Volume

In this section, we will expand the concept of specific heat to gases. To start, let's consider u as a function of T and v :

$$u = u(T, v)$$

Using the concept of the **total differential** from Calculus, we can write:

$$du = \frac{\partial u}{\partial T} dT + \frac{\partial u}{\partial v} dv$$

Since a partial derivative only considers change from a single variable, and we assumed that u was a function only of T and v , $\partial u / \partial T$ is the same as saying “the change of u with respect to T when v is constant”.

However, we could have written u as a function of any two independent properties, so to specify exactly which one, we'll add the property that stays constant as a subscript:

$$\frac{\partial u}{\partial T} = \left[\frac{du}{dT} \right]_v$$

Writing du again, we get:

$$du = \left[\frac{du}{dT} \right]_v dT + \left[\frac{du}{dv} \right]_T dv \quad (3.5)$$

Now, let's write the energy equation in differential form. Recall that $w = p\Delta v$, which can be written in differential form as $\delta w = p dv$.

$$\delta q - \delta w = du \quad \rightarrow \quad \delta q - p dv = du \quad (3.6)$$

We used δq and δw because q and w are not properties, and therefore not integrable in the same way as v and u .

Combining Equations 3.5 and 3.6, we get:

$$\delta q = \left[\frac{du}{dT} \right]_v dT + \left(\left[\frac{du}{dv} \right]_T + p \right) dv$$

All of this leads to a huge simplification for constant-volume processes. If $dv = 0$, we get the following:

$$\delta q = \left[\frac{du}{dT} \right]_v dT \quad (3.7)$$

Equation 3.7 is simply the differential form of Equation 3.4! In fact, $\left[\frac{du}{dT} \right]_v$ is simply the **constant volume specific heat**, c_v .

3.3.2 Constant Pressure

Now we could have started this discussion by defining h as a function of T and p .

$$h = h(T, p)$$

That would result in:

$$dh = \left[\frac{dh}{dT} \right]_p dT + \left[\frac{dh}{dp} \right]_T dp \quad (3.8)$$

We know that the second half will eventually fall off when we think about constant pressure systems.

Let's expand dh , using our definition of $h = u + pv$. In order to do so, we need to remember the Product Rule from Calculus: $(fg)' = f'g + g'f$

$$\begin{aligned} dh &= d(u + pv) = du + d(pv) \\ dh &= du + p dv + v dp \end{aligned} \quad (3.9)$$

Now, we can replace du with $\delta q - p dv$, based on Equation 3.6:

$$\begin{aligned} dh &= (\delta q - p dv) + p dv + v dp \\ dh &= \delta q + v dp \end{aligned}$$

Next, we plug our result into Equation 3.8:

$$\begin{aligned} \delta q + v dp &= \left[\frac{dh}{dT} \right]_p dT + \left[\frac{dh}{dp} \right]_T dp \\ \delta q &= \left[\frac{dh}{dT} \right]_p dT + \left(\left[\frac{dh}{dp} \right]_T - v \right) dp \end{aligned}$$

Finally, we implement the constant pressure assumption, which sets $dp = 0$:

$$\delta q = \left[\frac{dh}{dT} \right]_p dT \quad (3.10)$$

Again, with Equation 3.10, we can state that $\left[\frac{dh}{dT} \right]_p$ is the **constant pressure specific heat**, c_p .

3.3.3 Ideal Gases

Now, consider u and h for ideal gases. We state without proof that for an ideal gas, internal energy depends on temperature alone. In other words, u and c_v are purely functions of temperature.

With $u = u(T)$ and $pv = RT$, we can write:

$$h = u + pv = u(T) + RT$$

This means that h is also only a function of temperature! In fact, by evaluating dh/dT , we can say:

$$\frac{dh}{dT} = \frac{du}{dT} + R \rightarrow c_p = c_v + R \quad (3.11)$$

Because both c_p and c_v are only functions of temperature, we can drastically reduce the amount of information needed for the tables. A single table (Appendix C.2) is sufficient for all ideal gas properties. In fact, we can even assume a constant specific heat and retain good accuracy, as long as the specific heat is evaluate at the average temperature, T_{avg} .

$$\Delta u = c_v(T_{avg}) \cdot \Delta T \quad \Delta h = c_p(T_{avg}) \cdot \Delta T \quad (3.12)$$

Example 3.2: Ideal Gas u and h

Air at 300 K undergoes a process in which heat is added at constant pressure. Consider processes that end at 900 K, 1200 K, and 1500 K. Determine the heat added using a) $c_p(T_{avg})$ and b) the enthalpies given in the table.

Solution Approach Since the process occurs at constant pressure, we know that the heat added is simply the difference in enthalpy:

$$Q = h_2 - h_1$$

If we assume constant specific heat, we can also say that:

$$Q = c_p (T_2 - T_1)$$

a) We first need to find the c_p at the average process temperatures.

T_2	T_{avg}	$c_p(T_{avg})$
900 K	600 K	1.0509 kJ/kgK
1200 K	750 K	1.0866 kJ/kgK
1500 K	900 K	1.1208 kJ/kgK

Note that we needed to use linear interpolation in order to find the c_p at 750 K.

With the c_p values in hand, we can find Q :

T_2	ΔT	$c_p(T_{avg})$	Q
900 K	600 K	1.0509 kJ/kgK	630.54 kJ/kg
1200 K	900 K	1.0866 kJ/kgK	977.94 kJ/kg
1500 K	1200 K	1.1208 kJ/kgK	1344.96 kJ/kg

The rightmost column above is our desired result.

- b) The second method requires us to find the enthalpies directly. The enthalpy of air at 300 K is 426.53 kJ/kg. The enthalpies at the end of the process, along with the resulting heat transfers, are given below:

T_2	h_2	Q
900 K	1059.36 kJ/kg	632.83 kJ/kg
1200 K	1404.14 kJ/kg	977.61 kJ/kg
1500 K	1762.29 kJ/kg	1335.76 kJ/kg

Again, the rightmost column is our desired result.

Comparing the two methods, we can find the percent error through the formula:

$$err = \frac{Q_a - Q_b}{Q_b} \times 100\%$$

Finally, we see the errors below:

T_2	Error
900 K	-0.36%
1200 K	0.034%
1500 K	0.69%

In all three cases, we see less than 1% error, meaning that the average c_p method is very accurate for air.

3.4 Stirling Cycle Engine

Conceptually, the Stirling engine, invented by Robert Stirling, is the simplest of all heat engines. It has no valves, and includes an externally heated space and an externally cooled space.

In its original form the working gas (typically air or helium) is sealed within a single cylinder by the piston and shuttled between the hot and cold spaces by a displacer, as seen in Figure 3.3. The linkage driving the piston and displacer will move them such that the gas will compress while it is mainly in the cool compression space and expand while in the hot expansion space. This is clearly illustrated in [this animation](#) which was produced by Richard Wheeler (Zephyris) of Wikipedia.

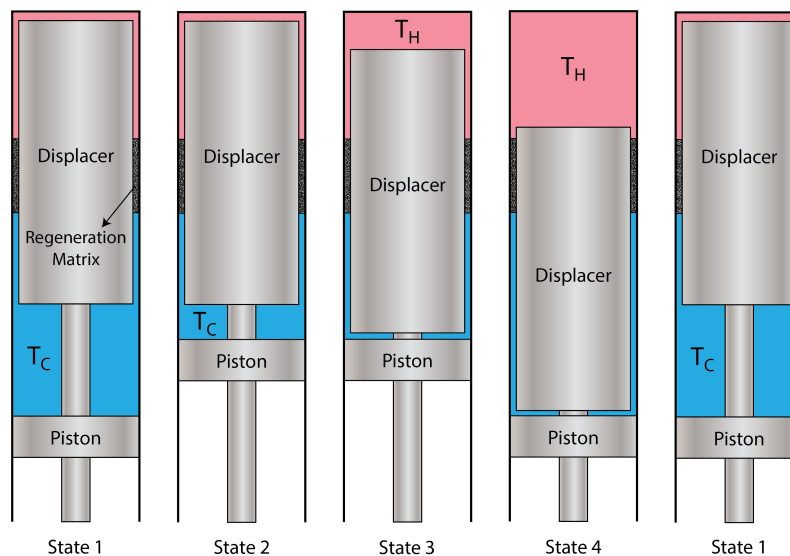


Figure 3.3: The states of the Stirling engine cycle.

Refer also to the [animation produced by Matt Keveney](#) in his Stirling engine animation website. Since the gas is at a higher temperature, and therefore pressure, during its expansion than during its compression, more power is produced during expansion than is reabsorbed during compression, and this net excess power is the useful output of the engine. Note that there are no valves or intermittent combustion, which is the major source of noise in an internal combustion engine. The same working gas is used over and over again, making the Stirling engine a sealed, closed cycle system. All that is added to the system is steady high temperature heat, and all that is removed from the system is low temperature (waste) heat and mechanical power.

Please note that the following analysis of Stirling cycle engines is ideal, and is intended only as an example of First Law Analysis of closed systems. In the real world we cannot expect actual machines to perform any better than 40 - 50% of the ideal machine.

The four states of the Stirling engine are summarized in Figure 3.3 and Figure 3.4. Note specifically the motion of both the piston and the displacer. The movement of the piston correlates to the compression or expansion of the fluid, while the action of the displacer is to move the fluid between the hot and cold regions, thereby heating or cooling the fluid.

In essence, the ideal Stirling engine undergoes four distinct processes, each of which can be analyzed separately: isothermal compression at T_C , constant volume heating at v_{min} , isothermal expansion at T_H , and constant volume cooling (or heat rejection) at v_{max} .

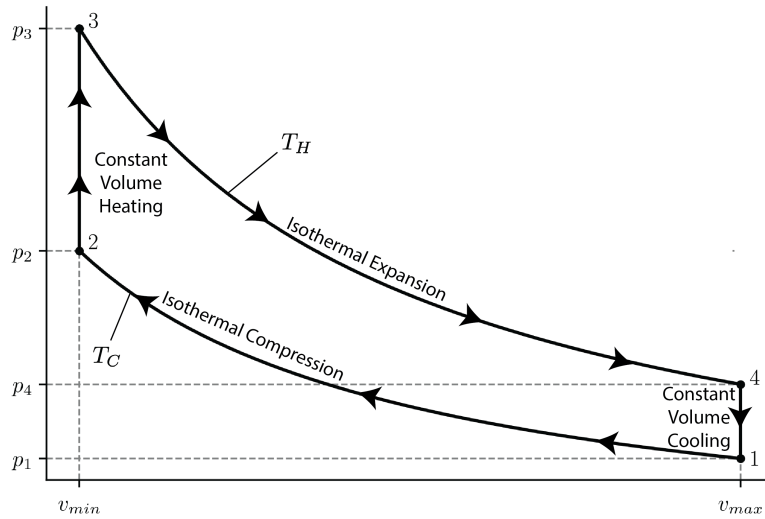


Figure 3.4: The states of the Stirling engine cycle.

3.4.1 Process 1→2 – Isothermal Compression

We start our analysis at State 1, which is the coolest and least compressed point of the cycle.

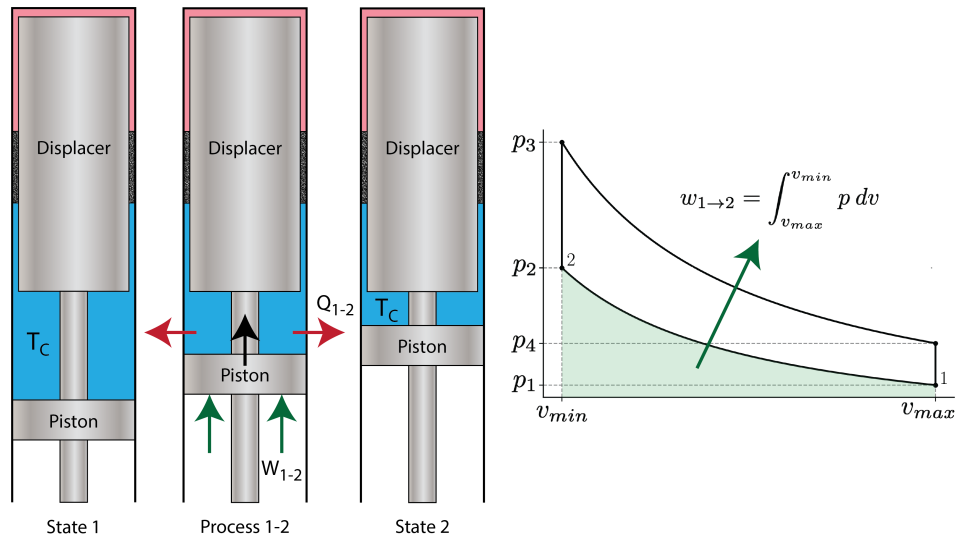


Figure 3.5: The compression process of the Stirling Engine.

Figure 3.5 shows the compression from State 1 to State 2. In process 1-2, the gas is compressed by the piston while the displacer is at the top of the cylinder. Thus during this process the gas is cooled in order to maintain a constant temperature T_C . Note that without cooling, the work added by compression would increase the internal energy, and by extension temperature.

The work W_{1-2} required to compress the gas is shown as the area under the p - V curve, and is evaluated as follows, using the ideal gas law with the constant temperature during compression to say $p = \frac{RT_C}{v}$.

$$w_{1-2} = \int_{v_1}^{v_2} p \, dv = RT_C \int_{v_1}^{v_2} \frac{dv}{v} = RT_C \cdot \ln \left(\frac{v_2}{v_1} \right) \quad (3.13)$$

Another consequence of the constant temperature is that the internal energy of the gas will also stay unchanged.

$$q_{1-2} - w_{1-2} = \Delta u_{1-2} \xrightarrow{0} q_{1-2} = w_{1-2} \quad (3.14)$$

3.4.2 Process 2→3 – Constant Volume Heating

After compression, the displacer piston pushes the air from the cold region to the hot region, warming it up as it moves through the regeneration matrix.

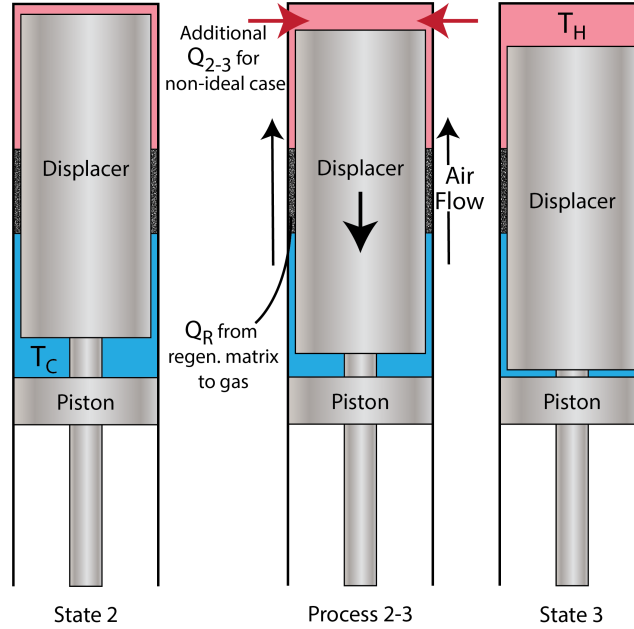


Figure 3.6: The heating process of the Stirling Engine.

Figure 3.6 shows the movement of the displacer down into the cold region. This in turn drives the air through the regenerator into the hot region. The end result of this is that the air is significantly heated and ends up at T_H .

Importantly, there is no compression or expansion in Process 2-3. This means no work is done between the two states, greatly simplifying our analysis:

$$\Delta u_{2-3} = q_{2-3} - \cancel{w_{2-3}} \xrightarrow{0} \quad q_{2-3} = c_v (T_H - T_C) \quad (3.15)$$

3.4.3 Process 3→4 – Isothermal Expansion

After the air is displaced, the hot air expands, forcing back the piston. Once again, this process is isothermal, but this time at a temperature T_H .

Figure 3.7 shows Process 3-4, the expansion of the air at constant temperature. The analysis mirrors that of Process 1-2.

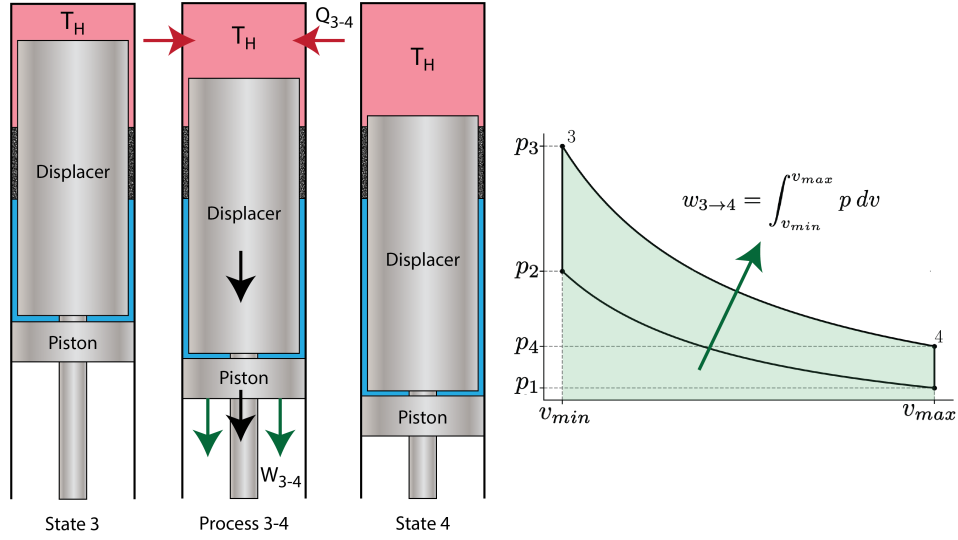


Figure 3.7: The expansion process of the Stirling Engine.

$$w_{3-4} = \int_{v_3}^{v_4} p dv = RT_H \int_{v_3}^{v_4} \frac{dv}{v} = RT_H \cdot \ln \left(\frac{v_4}{v_3} \right) \quad (3.16)$$

Fortunately, because Processes 2-3 and 4-1 are isometric (constant volume), we can make the replacements $v_3 \rightarrow v_2$ and $v_4 \rightarrow v_1$ in Equation 3.16:

$$w_{3-4} = RT_H \cdot \ln \left(\frac{v_1}{v_2} \right) = -\frac{T_H}{T_C} w_{1-2} \quad (3.17)$$

3.4.4 Process 4→1 – Constant Volume Cooling

Finally, Process 4-1 (shown in Figure 3.8) is essentially Process 2-3 in reverse (though with a larger volume). The gas loses its heat to the regenerator matrix and is cooled back to T_C .

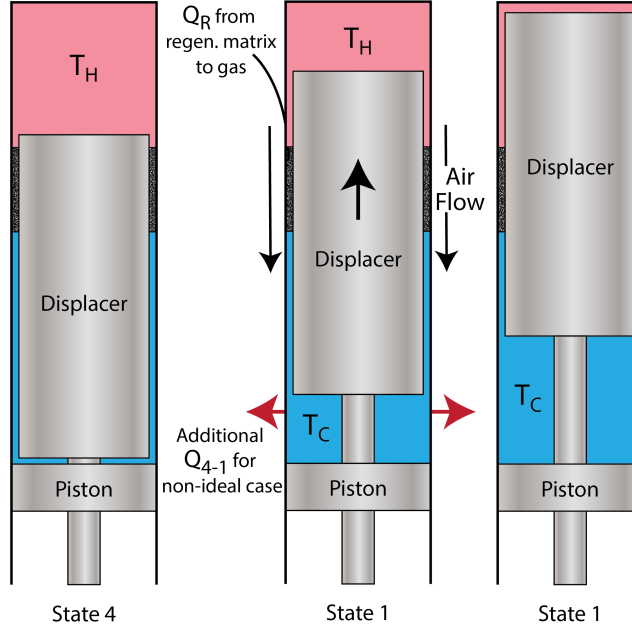


Figure 3.8: The cooling process of the Stirling Engine.

Again, the heat transfer is purely based on the change in temperature, as there is no work involved: $q_{4-1} = -c_v(T_H - T_C)$. The sign is opposite of that for q_{2-3} since we are cooling, rather than heating.

3.4.5 Ideal Stirling Cycle Analysis

For cycles, it is often helpful to write out the energy equation in tabular format, as shown in Table 3.1.

Table 3.1: Tabular Representation of Energy Equation for Stirling Cycle

	Δu	q	w
Process 1-2	0	$-RT_C \cdot \ln\left(\frac{v_1}{v_2}\right)$	$-RT_C \cdot \ln\left(\frac{v_1}{v_2}\right)$
Process 2-3	$c_v(T_H - T_C)$	$c_v(T_H - T_C)$	0
Process 3-4	0	$RT_H \cdot \ln\left(\frac{v_1}{v_2}\right)$	$RT_H \cdot \ln\left(\frac{v_1}{v_2}\right)$
Process 4-1	$-c_v(T_H - T_C)$	$-c_v(T_H - T_C)$	0
Total (net)	0	$R(T_H - T_C) \cdot \ln\left(\frac{v_1}{v_2}\right)$	$R(T_H - T_C) \cdot \ln\left(\frac{v_1}{v_2}\right)$

We expect that Δu for a cycle is always zero, as the cycle starts and ends at the same point. This implies in turn that $q_{net} = w_{net}$, which is exactly what we see in Table 3.1. In fact, the purpose of the Stirling engine is to turn the net heat transfer into work. The heat transfer to and from the regenerator doesn't come into play, as the heat transfer in Process 2-3 perfectly cancels with the heat transfer in Process 4-1. The total work done (and the total heat added to the system) can be simplified to:

$$q_{net} = w_{net} = R\Delta T \cdot \ln(r) \quad (3.18)$$

where ΔT is the difference between the hot and cold temperatures and r is the **compression ratio**, defined as:

$$r = \frac{v_{\max}}{v_{\min}} = \frac{V_{\max}}{V_{\min}} \quad (3.19)$$

We can also define the **thermal efficiency** as the ratio between the input heat and the net work. This efficiency is a picture of how much work we can get out of an engine based on the amount of fuel we burn.

$$\begin{aligned} \eta_{th} &= \frac{w_{net}}{q_{in}} = \frac{w_{net}}{q_{3-4}} = \frac{R\Delta T \cdot \ln(r)}{RT_H \cdot \ln(r)} \\ \eta_{th} &= \frac{\Delta T}{T_H} = 1 - \frac{T_C}{T_H} \end{aligned} \quad (3.20)$$

Equation 3.20 is the maximum theoretical efficiency achievable from a heat engine, and is typically referred to as the **Carnot efficiency**.

3.4.6 Stirling Engines without Regeneration

Note that the ideal case assumes that the regenerator is able to fully heat/cool the air in Processes 2-3 and 4-1. If this is not the case (or if there is no regenerator present), then the extra heat Q_{2-3} must be provided by the heater, which will reduce the efficiency. In the case of no regenerator, the efficiency becomes:

$$\eta_{th} = \frac{w_{net}}{q_{in} + q_{2-3}} = \frac{R\Delta T \cdot \ln(r)}{(RT_H \cdot \ln(r) + c_v \Delta T)}$$

This is a little difficult to draw conclusions from, but let's look at a sample case with air as the working fluid ($c_v \approx 2.5R$), $T_C = 300$ K, $T_H = 600$ K, and $r = 1.3$. That gives us an ideal thermal efficiency of $\eta_{th, \text{regen}} = 0.5$ with the regenerator, and an ideal thermal efficiency of $\eta_{th, \text{no regen}} = 0.097$! Having a regenerator is hugely important!

Another important note is that even with a regenerator, the practical Stirling cycle has many losses associated with it and does not really involve isothermal processes, nor ideal regeneration. Furthermore, Stirling cycle machines do not typically start and stop the pistons perfectly, meaning that the p - V diagram has a more oval shape, rather than the sharp edges defined in the above diagrams. Nevertheless we use the ideal Stirling cycle to get an initial understanding and appreciation of the cycle performance.

Example 3.3: Stirling Cycle Engine/Generator

This exercise concerns the ideal performance of the **Microgen 1050 Watt engine** which is gas fired and was designed to generate electricity (1kWe) as well as to provide hot water for a private home. This engine has a similar configuration to the Stirling engine discussed in Section 3.4. (Note that the values presented here are not actual values of the Microgen Engine, however were devised by your instructor for purposes of this exercise only).

The working fluid is helium, which has an gas constant $R = 2.077 \text{ kJ/kg K}$, and constant volume specific heat $c_v = 3.116 \text{ kJ/kg K}$. (refer: Ideal Gas Properties). State (1) is defined at a maximum volume of 650 cm^3 and a pressure of 10 bar, and State (2) is defined at a minimum volume of 550 cm^3 . Process 1-2 is isothermal compression at $T_C = 50^\circ\text{C}$, Process 2-3 is isometric heating, using a regenerator to source the heat. Process 3-4 is isothermal expansion at $T_H = 500^\circ\text{C}$, and Process 4-1 is isometric cooling, again using the regenerator as a heat sink.

- From the given conditions at state 1 determine the mass of working gas (helium) used in the cycle.
- Determine the net work done per cycle (kiloJoules): $W_{exp} + W_{comp}$ (Note that the compression work W_{comp} is always negative). At a frequency of 50 Hertz (cycles per second) determine the power output produced by the engine.
- Determine the heat absorbed in the expansion space Q_{in} during the expansion process (3)-(4).
- Evaluate the Thermal Efficiency η_{th} , defined as: $\eta_{th} = (W_{exp} + W_{comp})/Q_{in}$. (Net mechanical work done divided by the heat supplied externally by the gas burner).
- Determine the amount of thermal power rejected to the cooling water. Note that at a temperature of 50°C this is suitable for providing hot water for the home, as well as providing home space heating capability.
- Determine the amount of heat transferred to the working fluid Q_R as it passes through the regenerator during process (2)-(3). If this heat were to be supplied externally by the gas burner, (i.e. no regenerator) what would be the new value of thermal efficiency η_{th} ?

Solution Approach

- We can find the mass through the ideal gas law, letting $1 \text{ cm}^3 = 1 \times 10^{-6} \text{ m}^3$:

$$\begin{aligned}
 p_1 v_1 &= RT_1 \quad \rightarrow \quad v_1 = \frac{RT_1}{p_1} \\
 v_1 &= \frac{2.077 \text{ kJ/kgK} \cdot 323.15 \text{ K}}{1000 \text{ kPa}} \\
 v_1 &= 0.6712 \text{ m}^3/\text{kg} \\
 v_1 &= \frac{V_1}{m} \quad \rightarrow \quad m = \frac{6.5 \times 10^{-4} \text{ m}^3}{0.6712 \text{ m}^3/\text{kg}} = 9.68 \times 10^{-4} \text{ kg}
 \end{aligned}$$

- b) The net work comes from Equation 3.18:

$$W_{net} = mR(T_H - T_C) \ln \left(\frac{V_1}{V_2} \right)$$

$$W_{net} = (0.968 \text{ g}) \left(2.077 \frac{\text{kJ}}{\text{kgK}} \right) (450 \text{ K}) \ln \left(\frac{650 \text{ cm}^3}{550 \text{ cm}^3} \right) \rightarrow \boxed{W_{net} = 0.151 \text{ kJ}}$$

We can convert from work to power by multiplying by the frequency (units of Hz are the same as 1/s):

$$\dot{W}_{net} = W_{net} \cdot f = 0.151 \text{ kJ} \cdot 50 \frac{1}{\text{s}} \rightarrow \boxed{\dot{W}_{net} = 7.55 \text{ kW}}$$

- c) The heat absorbed in the expansion space can be calculated from Equation 3.16, recognizing that the constant-temperature process results in no change in internal energy:

$$W_{exp} = Q_{exp} = mRT_H \ln \left(\frac{V_4}{V_3} \right)$$

$$Q_{exp} = (0.968 \text{ g}) \left(2.077 \frac{\text{kJ}}{\text{kgK}} \right) (773.15 \text{ K}) \ln \left(\frac{650 \text{ cm}^3}{550 \text{ cm}^3} \right) \rightarrow \boxed{Q_{exp} = 0.260 \text{ kJ}}$$

This is equivalent to $\boxed{\dot{Q}_{exp} = 13.0 \text{ kJ}}$.

- d) We can find the thermal efficiency from the results of the previous two parts:

$$\eta_{th} = \frac{W_{net}}{Q_{in}} = \frac{0.151 \text{ kJ}}{0.260 \text{ kJ}} \rightarrow \boxed{\eta = 58\%}$$

- e) The heat rejected to the cooling water can be calculated directly as in part c), or by recognizing that $Q_{in} - Q_{out} = W_{net}$.

$$Q_{out} = Q_{in} - W_{net} = 0.260 \text{ kJ} - 0.151 \text{ kJ} \rightarrow \boxed{Q_{out} = 0.109 \text{ kJ}}$$

- f) The regenerator transfers enough heat to change the temperature of the helium from 50°C to 500°C. This heat can be calculated through the definition of the c_v , as in Equation 3.15:

$$Q_R = mc_v \Delta T = (0.968 \text{ g}) \left(3.116 \frac{\text{kJ}}{\text{kgK}} \right) (450 \text{ K}) \rightarrow \boxed{Q_R = 1.36 \text{ kJ}}$$

Adding the regenerator heat transfer to the heat supplied to the gas burner allows us to find the new thermal efficiency:

$$\eta_{th} = \frac{W_{net}}{Q_{in} + Q_R} = \frac{0.151 \text{ kJ}}{0.260 \text{ kJ} + 1.35 \text{ kJ}} \rightarrow \boxed{\eta_{th} = 9.3\%}$$

This value is in line with the efficiency calculated in the previous section.

3.5 Stirling Cycle Cooling

One important aspect of Stirling cycle machines that we need to consider is that the cycle can be reversed - if we put net work into the cycle then it can be used to pump heat from a low temperature source to a high temperature sink. This results in a very compact refrigerator, which can reach very low temperatures. This technology has been used in a few specific instances, primarily in **ultra-low-temperature freezers for biomedical uses**. The cycle used in standard air conditioning units is discussed in Section 4.6.

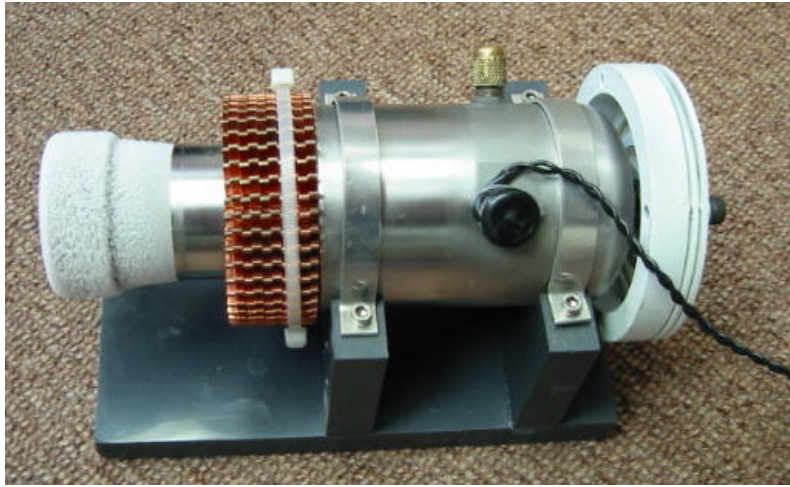


Figure 3.9: The M100B Stirling Cooler from Global Cooling (now Stirling Ultracold).

Figure 3.9 shows an example of a Stirling Cooler. Power is provided through the small electrical cord, while ice crystals can be seen forming on the left-hand side of the cooler. Refer to the schematic in Figure 3.10 to identify the various parts of the cooler.

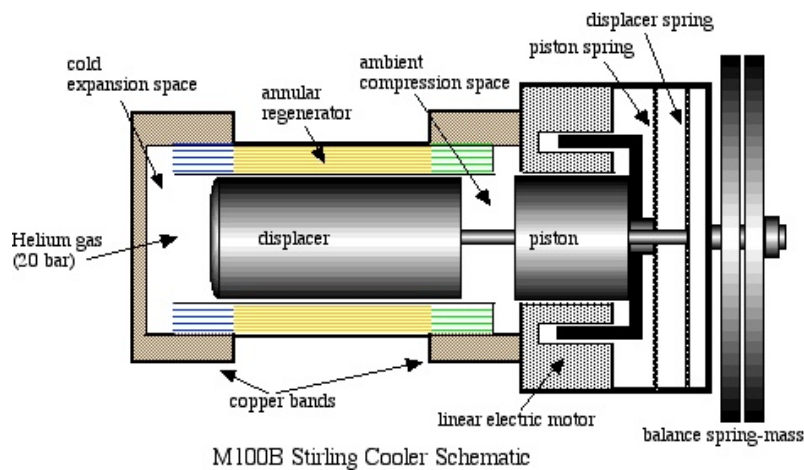


Figure 3.10: Schematic of M100B Stirling Cooler from Global Cooling.

Figure 3.11 shows the cycle used by the Stirling Cooler. Heat is rejected to the environment (Q_H) in Process 1-2, rejected to the regenerator (Q_R) in Process 2-3, absorbed from the refrigerated space (Q_C) in Process 3-4, and finally absorbed from the regenerator (Q_R) in Process 4-1. More work is required to drive Process 1-2 than is regained in Process 3-4, meaning that the net work for this cycle is negative. In other words, work is required to move heat from the cold side to the hot side.

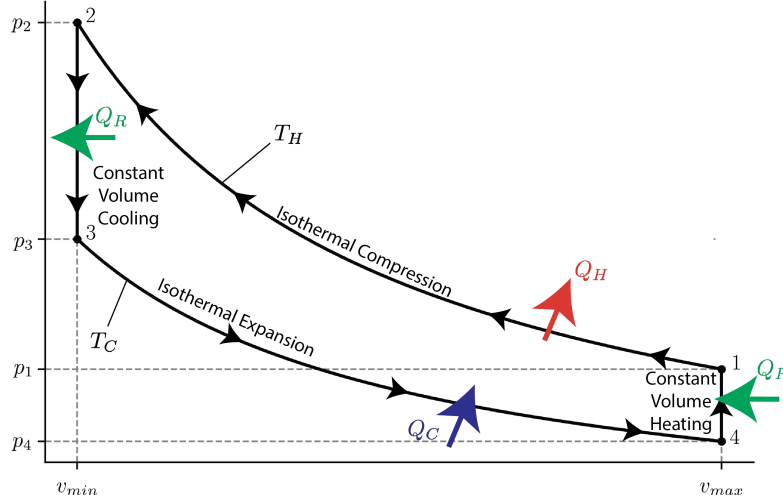


Figure 3.11: p - v diagram of the Stirling Cooling Cycle.

Cooling cycles do not use the thermal efficiency as an indicator of performance. Following the general idea that “efficiency” is “what we get out”/“what we put in”, the efficiency of a cooling cycle is defined at the Coefficient of Performance for a refrigeration cycle, COP_R :

$$COP_R = \frac{q_{in}}{w_{net}}$$

The heat absorbed, q_{in} , is the benefit we get from the cooling cycle, as we are seeking to remove heat from a refrigerated space. The net work, w_{net} , is what we have to pay to get that benefit, as we are doing work on the cycle.

Conversely, if we chose to think of this cycle as a way to add heat to a heated space, we would arrive at the Coefficient of Performance for a heat pump cycle, COP_{HP} :

$$COP_{HP} = \frac{q_{out}}{w_{net}}$$

In this case, the heat rejected to the heated space is the benefit from the cycle, which leads to a change in our “efficiency” calculations.

Example 3.4 investigates the Stirling cooling cycle in more detail.

Example 3.4: Stirling Cycle Cooling

A Stirling cooling cycle is defined as follows:

- Compression occurs at a constant temperature of $T_H = 30^\circ\text{C}$.
 - Expansion occurs at a constant temperature of $T_C = -20^\circ\text{C}$.
 - The maximum volume is 35 cm^3
 - The minimum volume is 30 cm^3
 - The pressure immediately before compression is 1.9 MPa .
 - The working fluid is helium, which has an gas constant $R = 2.077\text{ kJ/kg K}$, and constant volume specific heat $C_v = 3.116\text{ kJ/kg K}$.
- a) Determine the heat absorbed in the expansion stroke (Joules). Determine also the heat power absorbed (Watts). Note that the cycle frequency is the line frequency ($f = 60\text{ Hz}$).
 - b) Determine the net work done per cycle (Joules): $W_{net} = W_{comp} + W_{exp}$ (Note that the compression work W_{comp} is always negative). Determine also the power supplied to the linear electric motor (Watts).
 - c) Evaluate the Coefficient of Performance of the refrigerator defined as: $COP_R = Q_C/W_{net}$. (heat absorbed in the expansion space divided by the net work done).
 - d) Determine the amount of heat rejected by the working fluid Q_R as it passes through the regenerator matrix during constant volume cooling.
 - e) If there were no regenerator present then this heat would need to be removed from the gas by the expansion process in order to reduce the temperature to the cold temperature of the freezer. How would this affect the performance of the cooler? Discuss the importance of an effective regenerator in the Stirling cycle cooler.

Solution Approach

First off, let's find the mass of the working fluid.

For the first state, pressure and temperature are both defined, as is the total volume. This is sufficient for us to find the mass of the working fluid. As part of the process, we need to convert 35 cm^3 to $3.5 \times 10^{-5}\text{ m}^3$.

$$p_1 v_1 = RT_1 \quad \rightarrow \quad v_1 = \frac{RT_1}{p_1}$$

$$v_1 = \frac{2.077\text{ kJ/kgK} \cdot 303.15\text{ K}}{1900\text{ kPa}}$$

$$v_1 = 0.3314\text{ m}^3/\text{kg}$$

$$v_1 = \frac{V_1}{m} \quad \rightarrow \quad m = \frac{3.5 \times 10^{-5}\text{ m}^3}{0.3314\text{ m}^3/\text{kg}} = 1.056 \times 10^{-4}\text{ kg}$$

- a) Once we have the mass, we can use Table 3.1 for the analysis, except we need to replace T_C with T_H and vice-versa. The heat transfer is therefore:

$$Q_{exp} = mRT_C \ln\left(\frac{v_4}{v_3}\right) = mRT_C \ln\left(\frac{V_4}{V_3}\right)$$

$$Q_{exp} = (0.1056 \text{ g}) \left(2.077 \frac{\text{kJ}}{\text{kgK}}\right) (253.15 \text{ K}) \ln\left(\frac{35 \text{ cm}^3}{30 \text{ cm}^3}\right) \rightarrow Q = 8.56 \text{ J}$$

We can get to heat power absorbed simply by multiplying by the frequency of 60 Hz.

$$\dot{Q}_{exp} = Q \cdot f = 8.56 \text{ J} \cdot 60 \frac{1}{\text{s}} \rightarrow \dot{Q}_{exp} = 513.6 \text{ W}$$

- b) The net work can be obtained from Equation 3.18, again with T_C and T_H switched.

$$W_{net} = mR(T_C - T_H) \ln\left(\frac{V_1}{V_2}\right)$$

$$W_{net} = (0.1056 \text{ g}) \left(2.077 \frac{\text{kJ}}{\text{kgK}}\right) (50 \text{ K}) \ln\left(\frac{35 \text{ cm}^3}{30 \text{ cm}^3}\right) \rightarrow W_{net} = -1.69 \text{ J}$$

Note that the work done is negative, indicating that work is performed on the cycle, rather than provided by it.

We can again find the power by multiplying by f . This becomes

$$\dot{W}_{net} = -101 \text{ W}.$$

- c) The Coefficient of Performance is easily determined from the two previous results.

$$COP_R = \frac{Q_C}{W_{net}} = \frac{8.56 \text{ J}}{1.69 \text{ J}} \rightarrow COP_R = 5.07$$

- d) The heat absorbed/rejected by the regeneration matrix can be found from Equation 3.15.

$$Q_R = mc_v \Delta T = (0.1056 \text{ g}) \left(3.116 \frac{\text{kJ}}{\text{kgK}}\right) (50 \text{ K}) \rightarrow Q_R = 16.45 \text{ J}$$

This equates to $\dot{Q}_R = 988 \text{ W}$.

- e) Without regeneration, we would expect a drastically lower COP_R value. It's difficult to see exactly how this comes about, as there is no easy mechanism through which we can compensate for the lack of a regenerator (like we simply added more heat in the Stirling engine).

3.6 Ideal Gas Adiabatic Processes

The next two cycles both rely heavily on **adiabatic** processes. In order to discuss those cycles, we will look closer at adiabatic processes for an ideal gas.

Adiabatic processes are simply processes without heat transfer. This means that the change in internal energy is purely due to the work done.

$$\overset{0}{\delta q} - \delta w = du$$

Substituting $\delta w = p dV$ and $du = c_v dT$ gives us:

$$-p dv = c_v dT$$

With $pv = RT$, we calculate:

$$-\frac{RT}{v} dv = c_v dT \quad \rightarrow \quad c_v dT + \frac{RT}{v} dv = 0 \quad \rightarrow \quad c_v \frac{dT}{T} + R \frac{dv}{v} = 0$$

Now we integrate:

$$\begin{aligned} \int_1^2 \left(c_v \frac{dT}{T} + R \frac{dv}{v} \right) &= 0 \\ \int_{T_1}^{T_2} c_v \frac{dT}{T} + \int_{v_1}^{v_2} R \frac{dv}{v} &= 0 \\ c_v \ln \frac{T_2}{T_1} + R \ln \frac{v_2}{v_1} &= 0 \end{aligned}$$

Rearranging slightly gives us:

$$\ln \frac{T_2}{T_1} = -\frac{R}{c_v} \ln \frac{v_2}{v_1} \quad (3.21)$$

Remembering that $e^{c \ln x} = x^c$, we can arrive at:

$$\frac{T_2}{T_1} = \left(\frac{v_2}{v_1} \right)^{-\frac{R}{c_v}}$$

Now, we can define the **ratio of specific heats**, γ :

$$\gamma = \frac{c_p}{c_v} \quad (3.22)$$

With the definition of γ , we can use Equation 3.11 to define both c_p and c_v with respect to γ and R . This is done as follows:

$$\begin{aligned} c_p &= c_v + R \\ \gamma c_v &= c_v + R \\ (\gamma - 1) c_v &= R \end{aligned}$$

$$c_v = \frac{1}{\gamma - 1} R \quad (3.23)$$

$$c_p = \frac{\gamma}{\gamma - 1} R \quad (3.24)$$

Using Equation 3.23 with Equation 3.21 allows us a relationship between the temperature ratios and the volume ratios based only on γ :

$$\frac{T_2}{T_1} = \left(\frac{v_2}{v_1} \right)^{-(\gamma-1)} \quad (3.25)$$

Going back to the ideal gas law:

$$\frac{p_2 v_2}{p_1 v_1} = \frac{R T_2}{R T_1}$$

Substituting the temperature ratio with Equation 3.25 gives us:

$$\frac{p_2}{p_1} \frac{v_2}{v_1} = \left(\frac{v_2}{v_1} \right)^{-(\gamma-1)} = \left(\frac{v_1}{v_2} \right)^{(\gamma-1)}$$

$$\frac{p_2}{p_1} = \left(\frac{v_2}{v_1} \right)^{-\gamma} = \left(\frac{v_1}{v_2} \right)^{\gamma} \quad (3.26)$$

Substituting the volume ratio instead gives us:

$$\frac{p_2}{p_1} \left(\frac{T_2}{T_1} \right)^{-\frac{1}{\gamma-1}} = \frac{T_2}{T_1}$$

$$\frac{p_2}{p_1} = \left(\frac{T_2}{T_1} \right)^{\frac{\gamma}{\gamma-1}} \quad (3.27)$$

Equations 3.25, 3.26, and 3.27 are collectively known as the **isentropic relations** for an ideal gas. Isentropic means constant **entropy**, which will be discussed further in Chapter 6.

3.7 Air-standard Otto Cycle

The Air Standard Otto cycle is the ideal cycle for Spark-Ignition (SI) internal combustion engines, first proposed by Nikolaus Otto over 130 years ago, and which is currently used most motor vehicles. The following link by the Kruse Technology Partnership presents a description of the **four-stroke Otto cycle** operation including a short history of Nikolaus Otto. Once again we have excellent animations produced by Matt Keveney presenting the operation for both the **four-stroke** and the **two-stroke** spark-ignition internal combustion engines.

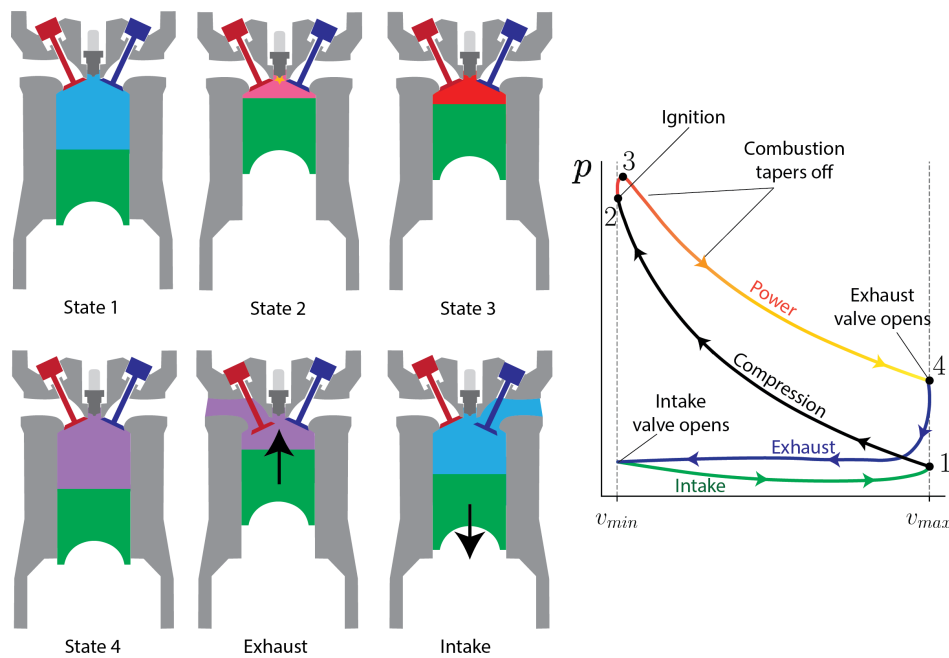


Figure 3.12: States of the four-stroke Otto Cycle with $p-v$ diagram.

Figure 3.12 gives an approximate picture of an actual four-stroke spark-ignition cycle. It is extremely complex, both due to the gradual release of heat after ignition, as well as the mass transfer that occurs as the exhaust and intake valves open and close. A perfect representation also needs to deal with the chemistry of combustion, the dynamics of fuel mixing, and the timing of various parts of the cycle.

Because the actual spark-ignition cycle is so complex, we make a number of simplifying assumptions. First off, we use an ideal "air-standard" assumption, in which the working fluid is a fixed mass of air undergoing the complete cycle. This simplification removes both the need for the exhaust/intake strokes and the need to consider the chemistry of combustion. Instead of ignition, we use instantaneous heat addition between states 2 and 3. Finally, the exhaust and intake strokes are replaced by heat rejection between states 4 and 1.

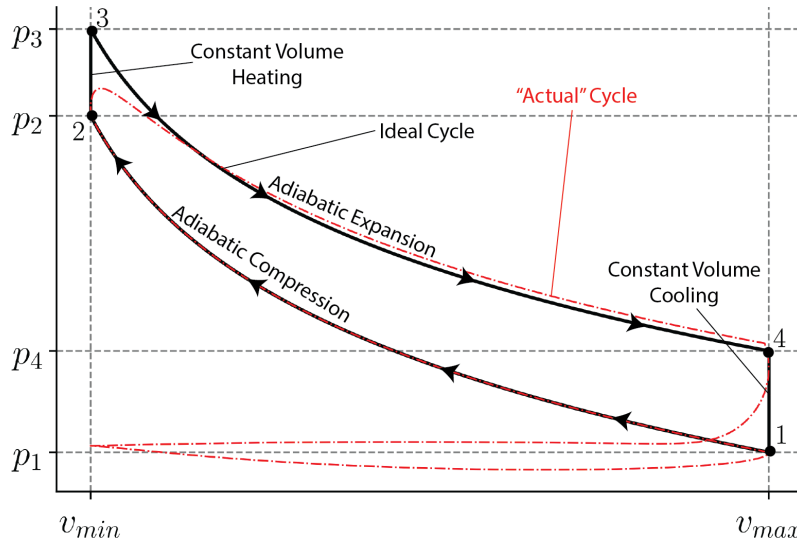


Figure 3.13: Ideal Otto cycle p - v diagram compared to a simulated Otto cycle.

The end result is the cycle shown in Figure 3.13. This looks very similar to the Stirling Cycle discussed in Section 3.4. The major difference between the two is that the compression between states 1 and 2 and the expansion between states 3 and 4 are both **adiabatic**, rather than isothermal. Analysis of these processes depends on the analysis done in Section .

Note that the ideal cycle in Figure 3.13 reaches significantly higher pressures than the actual cycle. Additionally, the pressure at state 4 is slightly lower in the ideal cycle. Both of these occur because of the assumption of instantaneous heat addition in the ideal cycle. The actual cycle never reaches the high temperatures because expansion happens at the same time as heating, and the temperature is higher at state 4 because heat is added at a lower temperature, which means the air has a lower c_v value.

The omission of the exhaust and intake strokes results in better efficiency for the ideal cycle, as work is required to both push out extra combustion products and to suck in fresh air.

3.7.1 Otto Cycle Analysis

The Otto Cycle looks very similar to the Stirling Cycle discussed in Section 3.4, with the key difference that the expansion and compression are adiabatic (no heat transfer) instead of isothermal.

The results of this can be seen in Table 3.2. Δu is always $c_v \Delta T$, though the definition of ΔT changes for each process. For processes 1-2 and 3-4, $q = 0$ because the processes are adiabatic. For processes 2-3 and 4-1, $w = 0$ because no work can be developed in a constant volume process. The First Law ($\Delta u = q - w$) is used to fill in the remaining blanks.

Table 3.2: Tabular Representation of Energy Equation for the Otto Cycle

	Δu	q	w
Process 1-2	$c_v(T_2 - T_1)$	0	$-c_v(T_2 - T_1)$
Process 2-3	$c_v(T_3 - T_2)$	$c_v(T_3 - T_2)$	0
Process 3-4	$c_v(T_4 - T_3)$	0	$-c_v(T_4 - T_3)$
Process 4-1	$c_v(T_1 - T_4)$	$c_v(T_1 - T_4)$	0
Total (net)	0	$c_v(T_1 - T_2 + T_3 - T_4)$	$c_v(T_1 - T_2 + T_3 - T_4)$

We are often interested in the **thermal efficiency** of an engine, which is defined as $\eta_{th} = q_{in}/w_{net}$ in Equation 3.20. For the Otto engine, this becomes:

$$\eta_{th} = \frac{\cancel{c_v}(T_1 - T_2 + T_3 - T_4)}{\cancel{c_v}(T_3 - T_2)} = 1 + \frac{T_1 - T_4}{T_3 - T_2}$$

Now we define the **compression ratio**, r :

$$r = \frac{v_{max}}{v_{min}} \quad (3.28)$$

We can recognize that $v_{max} = v_1 = v_4$ and $v_{min} = v_2 = v_3$ to say that $r = v_1/v_2 = v_4/v_3$.

We can then use the adiabatic relation for T and v from Equation 3.21 to make the following definitions:

$$\begin{aligned} T_2 &= T_1(v_2/v_1)^{-(\gamma-1)} = T_1 r^{\gamma-1} \\ T_3 &= T_4(v_3/v_4)^{-(\gamma-1)} = T_4 r^{\gamma-1} \end{aligned}$$

Using this, we can simplify the efficiency as follows:

$$\begin{aligned} \eta_{th} &= 1 + \frac{T_1 - T_4}{T_4 r^{\gamma-1} - T_1 r^{\gamma-1}} \\ \eta_{th} &= 1 - \frac{\cancel{(T_4 - T_1)} 1}{\cancel{(T_4 - T_1)} r^{\gamma-1}} \\ \eta_{th} &= 1 - \frac{1}{r^{\gamma-1}} \end{aligned} \quad (3.29)$$

Equation 3.29 shows that the thermal efficiency is a function only of the compression ratio, r , and the ratio of specific heats, γ . This simplified analysis requires the assumption that γ is a constant over the entire cycle. As long as γ is chosen at or near $T_{avg} = \frac{1}{2}(T_{max} + T_{min})$, this is a very reasonable assumption.

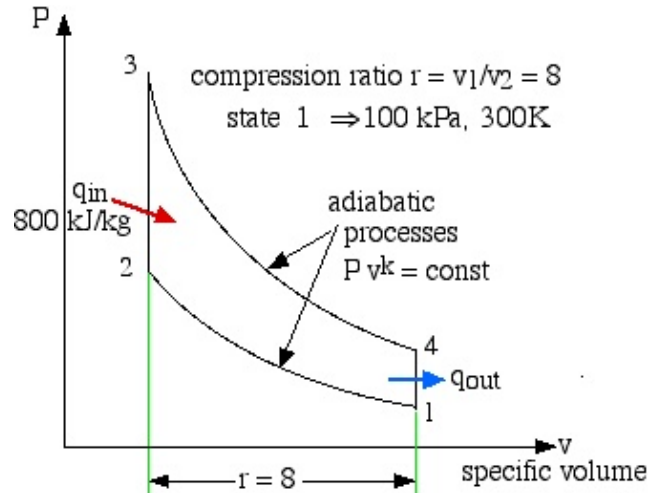
Example 3.5 provides a closer look at an ideal Otto cycle.

Example 3.5: Otto Cycle

An ideal air-standard Otto cycle engine has a compression ratio of 8. At the beginning of the compression process, the working fluid is at 100 kPa, 27°C (300 K), and 800 kJ/kg heat is supplied during the constant volume heat addition process. Neatly sketch the pressure-volume [p - v] diagram for this cycle, and using the specific heat values for air at a typical average cycle temperature of 900 K determine:

- the temperature and pressure of the air at the end of each process
- the net work output/cycle [kJ/kg], and
- the thermal efficiency [η_{th}] of this engine cycle.

The first step is to draw the p - v diagram of the complete cycle, including all the relevant information. We notice that neither volume nor mass have been provided, hence the diagram and solution will be in terms of specific quantities.



We assume that the fuel-air mixture is represented by pure air. The relevant equations of state, internal energy and adiabatic process for air follow:

$$pv = RT \qquad R = 0.287 \frac{\text{kJ}}{\text{kg K}}$$

$$\Delta u = c_v \Delta T$$

$$\frac{p_2}{p_1} = \left(\frac{v_1}{v_2} \right)^\gamma \qquad \frac{T_2}{T_1} = \left(\frac{v_1}{v_2} \right)^{\gamma-1}$$

Note that values for c_v and γ should come from the average value of temperature. We won't know what the actual average temperature is until we solve the problem, but we are asked to use the "typical average cycle temperature" of 900 K. We get this from the table of Specific Heat Capacities of Air.

$$c_{v,900 \text{ K}} = 0.834 \frac{\text{kJ}}{\text{kg K}} \qquad \gamma_{900 \text{ K}} = 1.344$$

Compare these values to the nominal specific heat capacity values used for air at 300K: $C_v = 0.717 \text{ kJ/kg K}$ and $\gamma = 1.4$, and you will see the importance of using the Ideal Gas Tables.

We now go through all four processes in order to determine the temperature and pressure at the end of each process, as well as the work done and heat transferred during each process.

Process 1-2 - Adiabatic Compression:

$$p_1 = 100 \text{ kPa}, \quad T_1 = 300 \text{ K}$$

Starting at state 1, we compress the fluid adiabatically with a compression ratio of 8. Keep in mind that the specific volume of state 1 will be **larger** than that of state 2.

Adiabatic	$\rightarrow q_{12} = 0 \frac{\text{kJ}}{\text{kg}}$
$r = \frac{v_1}{v_2} = 8$	
$\frac{T_2}{T_1} = \left(\frac{v_1}{v_2}\right)^{\gamma-1} = 8^{0.344} = 2.045$	$\rightarrow T_2 = 613 \text{ K}$
$\frac{p_2}{p_1} = \left(\frac{v_1}{v_2}\right)^{\gamma} = 8^{1.344} = 16.359$	$\rightarrow p_2 = 1635 \text{ kPa}$
$\Delta u_{12} = c_v \Delta T_{12} = \cancel{q_{12}}^0 w_{12}$	$\rightarrow w_{12} = -261 \frac{\text{kJ}}{\text{kg}}$

Note that we could also find the work through $\int p dv$. The above approach is much simpler.

Process 2-3 - Constant Volume Combustion: From state 2, we add heat at constant volume ($v_3 = v_2$).

Constant Volume	$\rightarrow w_{23} = 0 \frac{\text{kJ}}{\text{kg}}$
Given	$\rightarrow q_{23} = 800 \frac{\text{kJ}}{\text{kg}}$
$q = c_v \Delta T \quad \rightarrow \quad T_3 - T_2 = \frac{q}{c_v}$	
$T_3 - T_2 = \frac{800 \text{ kJ/kg}}{0.834 \text{ kJ/kgK}}$	$\rightarrow T_3 = 1572 \text{ K}$
$\frac{p_3 \cancel{v_3}}{p_2 \cancel{v_2}} = \frac{\cancel{R} T_3}{\cancel{R} T_2}$	$\rightarrow p_3 = 4193 \text{ kPa}$

Process 3-4 - Adiabatic Expansion: From state 3 to 4, we again use the compression ratio, but recognize that $v_4 > v_3$.

Adiabatic

$$\rightarrow q_{34} = 0 \frac{\text{kJ}}{\text{kg}}$$

$$r = \frac{v_3}{v_4} = \frac{1}{8}$$

$$\frac{T_4}{T_3} = \left(\frac{v_3}{v_4}\right)^{\gamma-1} = \frac{1}{8}^{0.344} = 0.489$$

$$\rightarrow T_3 = 769 \text{ K}$$

$$\frac{p_4}{p_3} = \left(\frac{v_3}{v_4}\right)^{\gamma} = \frac{1}{8}^{1.344} = 0.0611$$

$$\rightarrow p_4 = 256 \text{ kPa}$$

$$\Delta u_{12} = c_v \Delta T_{12} = q_{12} - w_{12}$$

$$\rightarrow w_{34} = 670 \frac{\text{kJ}}{\text{kg}}$$

Process 4-1 - Constant Volume Heat Rejection: From state 4, we reject heat at constant volume ($v_1 = v_4$).

Constant Volume

$$\rightarrow w_{23} = 0 \frac{\text{kJ}}{\text{kg}}$$

$$\Delta u_{41} = c_v \Delta T_{41} = q_{41} - w_{41}$$

$$\rightarrow q_{41} = -391 \frac{\text{kJ}}{\text{kg}}$$

Cycle Summary: Now we can build a summary of the processes using Table 3.2.

	ΔU	Q	W
Process 1-2	-261 kJ/kg	0	-261 kJ/kg
Process 2-3	800 kJ/kg	800 kJ/kg	0
Process 3-4	670 kJ/kg	0	670 kJ/kg
Process 4-1	-391 kJ/kg	-391 kJ/kg	0
Total (net)	0	409 kJ/kg	409 kJ/kg

The net work created by a single cycle was 409 kJ/kg, and the heat added per cycle was 800 kJ/kg. We can find the thermal efficiency as follows:

$$\eta_{th} = \frac{w_{net}}{q_{in}} = \frac{409 \text{ kJ/kg}}{800 \text{ kJ/kg}} \rightarrow \eta_{th} = 51\%$$

Note that using constant specific heat values over the cycle we can determine the thermal efficiency directly from the ratio of specific heat capacities γ and the compression ratio r using Equation 3.29:

$$\eta_{th} = 1 - \frac{1}{r^{\gamma-1}} = 1 - \frac{1}{8^{0.344}} = 0.511$$

This matches closely with the analysis above.

3.8 Air-Standard Diesel Cycle

The Air Standard Diesel cycle is the ideal cycle for Spark-Ignition (CI) reciprocating engines, first proposed by Rudolph Diesel over 100 years ago. The following link by the Kruse Technology Partnership describes the [four-stroke diesel cycle](#) operation including a short history of Rudolph Diesel. The four-stroke diesel engine is usually used in motor vehicle systems, whereas larger marine systems usually use the [two-stroke diesel cycle](#). Once again we have an excellent animation produced by Matt Keveney presenting the operation of the [four-stroke diesel cycle](#).

The analysis of the Diesel cycle is very similar to that of the Otto cycle which we analyzed in the previous section. We will use the ideal "air-standard" assumption in our analysis. Thus the working fluid is a fixed mass of air undergoing the complete cycle which is treated throughout as an ideal gas. All processes are ideal, combustion is replaced by heat addition to the air, and exhaust is replaced by a heat rejection process which restores the air to the initial state.

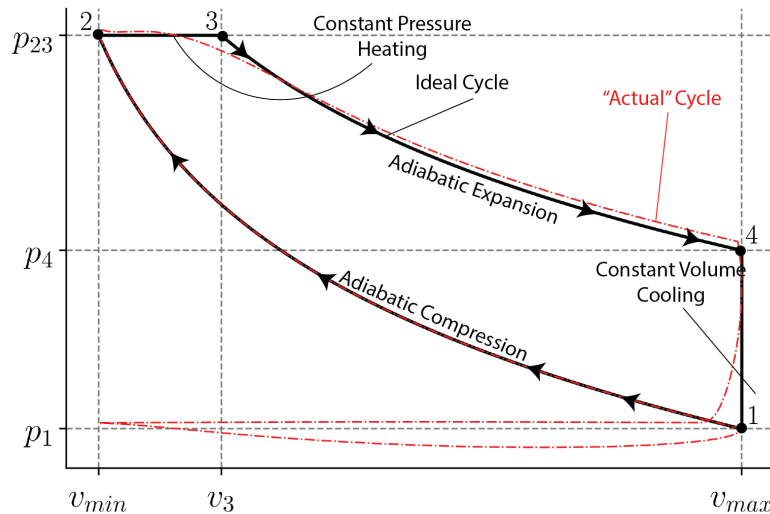


Figure 3.14: Ideal Diesel Cycle p - v diagram compared to a simulated Diesel Cycle.

The most significant difference between the ideal Otto cycle and the ideal Diesel cycle is the method of igniting the fuel-air mixture. For the Otto cycle, spark-ignition allows us to maintain a relatively low compression ratio (around 8:1), and time the combustion such that it results in a sudden jump in pressure while the volume remains essentially constant.

In contrast, in order to self-ignite, the ideal Diesel cycle needs an extremely high compression ratio (around 18:1) to allow the air to reach the ignition temperature of the fuel. Only then is the fuel injected. The timing of ignition is such that combustion adds heat as the piston is in the process of expanding, which leads to combustion at nearly a constant pressure, as seen in Figure 3.14.

After combustion, the rest of the cycle (expansion and exhaust) is essentially identical to that of the ideal Otto cycle.

3.8.1 Diesel Cycle Analysis

The Diesel cycle is very similar to the Otto cycle, with the major change that heating occurs at constant pressure, rather than constant volume. The major complication is that work occurs at the same time as heat transfer, unlike in the Otto cycle, where work and heat transfer are perfectly decoupled.

Table 3.3: Tabular Representation of Energy Equation for the Diesel Cycle

	Δu	q	w
Process 1-2	$c_v(T_2 - T_1)$	0	$-c_v(T_2 - T_1)$
Process 2-3	$c_v(T_3 - T_2)$	$c_p(T_3 - T_2)$	$p_{23}(v_3 - v_2) = R(T_3 - T_2)$
Process 3-4	$c_v(T_4 - T_3)$	0	$-c_v(T_4 - T_3)$
Process 4-1	$c_v(T_1 - T_4)$	$c_v(T_1 - T_4)$	0
Total (net)	0	$c_v(T_1 - T_4) + c_p(T_3 - T_2)$	$c_v(T_1 - T_4) + c_p(T_3 - T_2)$

Let's again find the thermal efficiency:

$$\eta_{th} = \frac{c_v(T_1 - T_4) + c_p(T_3 - T_2)}{c_p(T_3 - T_2)} = 1 - \frac{(T_4 - T_1)}{\gamma(T_3 - T_2)}$$

Often, rather than indicating the amount of heat from combustion, the heat addition process is defined by a **cutoff ratio**. This is another volume ratio, defined as:

$$r_c = \frac{v_3}{v_2} \quad (3.30)$$

The cutoff ratio will always be less than the compression ratio (Equation 3.19). Additionally, this ratio allows us to link each of our temperatures using r , r_c , and γ :

$$\begin{aligned} T_2 &= T_1 r^{\gamma-1} && \text{from adiabatic relations} \\ T_3 &= T_2 r_c && \text{from the ideal gas law} \\ T_4 &= T_3 \left(\frac{r_c}{r}\right)^{\gamma-1} && \text{from adiabatic relations} \end{aligned}$$

We can simplify T_3 and T_4 as follows:

$$\begin{aligned} T_3 &= T_1 r^{\gamma-1} r_c \\ T_4 &= T_1 r^{\gamma-1} r_c \left(\frac{r_c}{r}\right)^{\gamma-1} = T_1 r_c r_c^{\gamma-1} = T_1 r_c^\gamma \end{aligned}$$

Substituting the above relationships into the thermal efficiency gives us the following:

$$\begin{aligned} \eta_{th} &= 1 - \frac{T_1 r_c^\gamma - T_1}{\gamma(T_1 r^{\gamma-1} r_c - T_1 r^{\gamma-1})} \\ \eta_{th} &= 1 - \frac{1}{\gamma} \frac{r_c^\gamma - 1}{r^{\gamma-1} r_c - 1} \end{aligned} \quad (3.31)$$

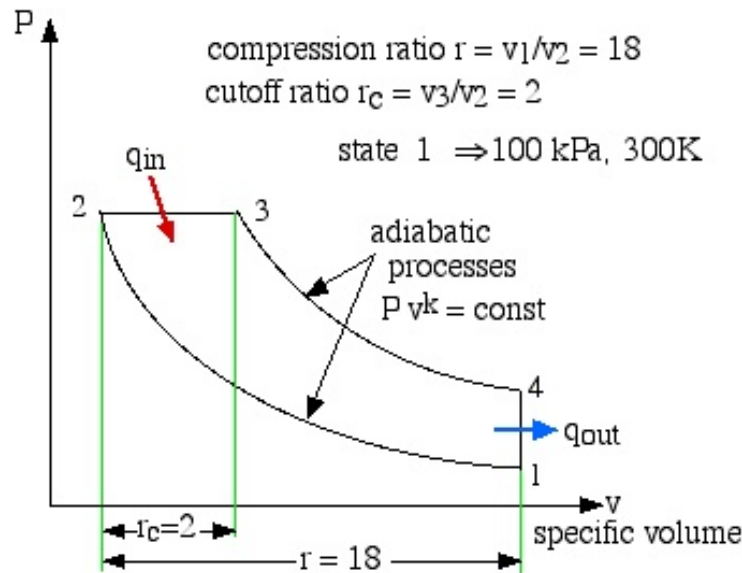
As $r_c \rightarrow 1$, Equation 3.31 approaches Equation 3.29, the Otto efficiency. Example 3.6 provides a closer look at an ideal Diesel cycle.

Example 3.6: Diesel Cycle

An ideal air-standard Diesel cycle engine has a compression ratio of 18 and a cutoff ratio of 2. At the beginning of the compression process, the working fluid is at 100 kPa, 27°C (300 K).

- Determine the temperature and pressure of the air at the end of each process.
- Determine the net work output per cycle [kJ/kg].
- Determine the thermal efficiency.

The first step is to draw a diagram representing the problem, including all the relevant information. We notice that neither volume nor mass is given, hence the diagram and solution will be in terms of specific quantities. The most useful diagram for a heat engine is the p - v diagram of the complete cycle:



In order to use constant specific heats, we must guess the average temperature of the cycle. For this analysis, we will use 900 K to calculate the specific heats. We will then check that value at the end of the analysis.

$$c_{p,900\text{ K}} = 1.121 \frac{\text{kJ}}{\text{kg K}} \quad \gamma_{900\text{ K}} = 1.344$$

We now go through all four processes in order to determine the temperature and pressure at the end of each process, as well as the work done and heat transferred during each process.

Process 1-2 - Adiabatic Compression:

$$p_1 = 100 \text{ kPa}, \quad T_1 = 300 \text{ K}$$

Starting at state 1, we compress the fluid adiabatically with a compression ratio of 18. Keep in mind that the specific volume of state 1 will be **larger** than that of state 2.

Adiabatic

$$\rightarrow q_{12} = 0 \frac{\text{kJ}}{\text{kg}}$$

$$r = \frac{v_1}{v_2} = 18$$

$$\frac{T_2}{T_1} = \left(\frac{v_1}{v_2} \right)^{\gamma-1} = 18^{0.344} = 2.703$$

$$\rightarrow T_2 = 811 \text{ K}$$

$$\frac{p_2}{p_1} = \left(\frac{v_1}{v_2} \right)^{\gamma} = 18^{1.344} = 48.650$$

$$\rightarrow p_2 = 4865 \text{ kPa}$$

Process 2-3 - Constant Pressure Combustion: From state 2, we add heat at constant pressure ($p_3 = p_2$), with a cutoff ratio of 2.

Constant Pressure

$$\rightarrow p_3 = 4865 \text{ kPa}$$

$$\frac{p_3 v_3}{p_2 v_2} = \frac{R T_3}{R T_2}$$

$$\rightarrow T_3 = 1622 \text{ K}$$

For this process, both work and heat transfer are present. However, we can find the heat transfer by using the constant pressure specific heat, c_p :

$$q_{23} = c_p(T_3 - T_2) = 1.121 \frac{\text{kJ}}{\text{kg K}} \cdot (1622 \text{ K} - 811 \text{ K}) \rightarrow q_{23} = 909 \frac{\text{kJ}}{\text{kg}}$$

Process 3-4 - Adiabatic Expansion: To calculate the volume ratio between states 3 and 4, we simply divide the compression ratio by the cutoff ratio:

$$\frac{v_4}{v_3} = \frac{v_1}{v_3} = \frac{v_2}{v_3} \frac{v_1}{v_2} = \frac{r}{r_c}$$

This allows us to follow a similar pattern to Process 1-2.

Adiabatic

$$\rightarrow q_{34} = 0 \frac{\text{kJ}}{\text{kg}}$$

$$\frac{v_3}{v_4} = \frac{r_c}{r} = \frac{1}{9}$$

$$\frac{T_4}{T_3} = \left(\frac{v_3}{v_4} \right)^{\gamma-1} = \frac{1}{9}^{0.344} = 0.470$$

$$\rightarrow T_4 = 762 \text{ K}$$

$$\frac{p_4}{p_3} = \left(\frac{v_3}{v_4} \right)^{\gamma} = \frac{1}{9}^{1.344} = 0.0522$$

$$\rightarrow p_4 = 254 \text{ kPa}$$

Process 4-1 - Constant Volume Heat Rejection: From state 4, we reject heat at constant volume ($v_1 = v_4$).

$$\Delta u_{41} = c_v \Delta T_{41} = q_{41} - w_{41} \rightarrow 0$$

$$\rightarrow q_{41} = -385 \frac{\text{kJ}}{\text{kg}}$$

Cycle Summary: Notice that we didn't calculate the work for any of the processes. The reason for this is we can more easily obtain the net work, w_{net} , by finding the net heat transfer, q_{net} . For a cycle, the net work and the net heat transfer will be equal.

$$w_{net} = q_{net} = 909 \frac{\text{kJ}}{\text{kg}} - 385 \frac{\text{kJ}}{\text{kg}} \rightarrow w_{net} = 524 \frac{\text{kJ}}{\text{kg}}$$

We can find the thermal efficiency as follows:

$$\eta_{th} = \frac{w_{net}}{q_{in}} = \frac{524 \text{ kJ/kg}}{909 \text{ kJ/kg}} \rightarrow \eta_{th} = 58\%$$

As a side note, our maximum temperature was about 1600 K, so an average temperature of 900 K for finding c_p and γ is reasonable.

You may wonder at the unrealistically high thermal efficiency obtained in this example. In this idealized analysis we have ignored many loss effects that exist in practical heat engines. We will begin to understand some of these loss mechanisms when we study the Second Law in Chapter 5.

3.9 Using Software: Plotting a Cycle in Python

In Section 1.10, we introduced the various plotting options available in Python through [Google Colab](#). In Section 2.6, we used Colab in order to implement the ideal gas law.

In this section, we will combine the two in order to plot the Otto cycle in Python.

3.9.1 Solving the Otto Cycle in Colab

To start off, let's solve Example 3.5 in Colab. We still need to start off by including our library imports:

```
# Clear all variable definitions
%reset -f
from numpy import *           # Import common numerical functions (like sqrt)
from matplotlib.pyplot import * # Import plotting functions (like plot)
import CoolProp.CoolProp as CP # Import CoolProp library
```

With those in place, we define the known values for this Otto cycle:

```
# Given Values
cv = 0.834      # Constant volume specific heat for air at 900 K (kJ/kgK)
gamma = 1.344    # Ratio of specific heats for air at 900 K (-)
Rgas = 8.314/28.97 # Gas constant for air (kJ/kgK)
r = 8           # Compression ratio of cycle (-)
qin = 800       # Combustion heat transfer (kJ/kg)
T1 = 300        # Initial temperature (K)
p1 = 100        # Initial pressure (kPa)
```

The compression ratio is our path to state 2, so we need to start off by finding the specific volume of state 1:

```
# Find v2
v1 = Rgas * T1 / p1 # Ideal Gas Law
v2 = v1 / r         # compression ratio r = v1/v2
```

The pressure and temperature of state 2 can be found using the adiabatic ideal gas equations (Equations 3.25 and 3.26):

```
# Find p2 and T2 with adiabatic ideal gas equations
T2 = T1 * (v1/v2)**(gamma-1)
p2 = p1 * (v1/v2)**gamma
```

The volume at state 3 is the same as at state 2. The temperature at state 3 is found using the equation for constant-volume heat addition. The pressure can be found through the ideal gas law.

```
# Find values of state 3 (constant volume heat addition)
v3 = v2 # constant volume process
T3 = T2 + qin/cv # q = cv*deltaT
p3 = T3*Rgas/v3 # Ideal gas law
```

The volume at state 4 is the same as state 1. The temperature and pressure can be found through the adiabatic ideal gas law equations.

```
# Find p4 and T4 with adiabatic ideal gas equations
v4 = v1
T4 = T3 * (v3/v4)**(gamma-1)
p4 = p3 * (v3/v4)**gamma
```

At this point, we have each of our four states fully defined. If we want to know the values, we can use the print command:

```
# Output state values
print('State      Pressure      Temperature      Spec. Volume')
print('  1      %9.3f      %9.3f      %9.3f' % (p1, T1, v1))
print('  2      %9.3f      %9.3f      %9.3f' % (p2, T2, v2))
print('  3      %9.3f      %9.3f      %9.3f' % (p3, T3, v3))
print('  4      %9.3f      %9.3f      %9.3f' % (p4, T4, v4))
```

This set of commands gives us the following table:

State	Pressure	Temperature	Spec. Volume
1	100.000	300.000	0.861
2	1635.886	613.457	0.108
3	4193.839	1572.690	0.108
4	256.365	769.095	0.861

This print command looks a lot different than what we've used in the past. Instead of manually putting in each of the values we wanted to print, we used **placeholders**, which were then filled after we completed the format. This allows us to create a small table without having to spend a ton of time fidgeting with spacing.

The placeholders have three components:

- The first value (9) indicates how much total space to give the number. A smaller number will provide less white space. If the number requires more space than given, Python will automatically give as much as needed (which would mess up our table).
- The second piece (.3) indicates the precision, or how many numbers after the decimal point will be included.
- The final piece (f) is an argument or conversion specifier. You may want to use `%9.3e` instead for very large or small numbers to use an exponential format (e.g. `4.194e+03` instead of `4193.839`).

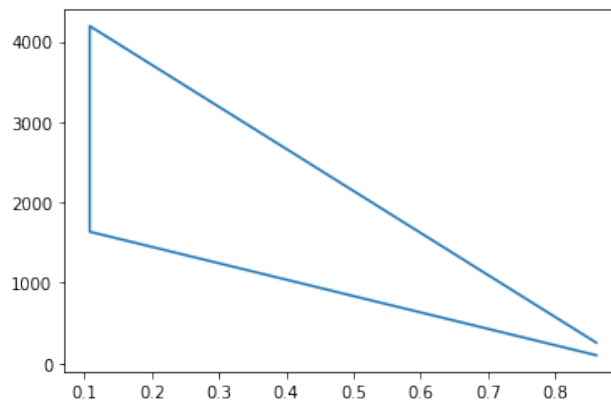
3.9.2 Plotting the States

With our states fully known, let's plot the properties on a p - v diagram.

We only have four values, so we can create our list manually:

```
plot([v1, v2, v3, v4], [p1, p2, p3, p4])
```

This results in the following plot:



This plot has a number of issues. We need labels and better limits for the axes, the cycle doesn't make a full circuit, and the adiabatic compression and expansion are straight lines, rather than following the path they should. Additionally, it would be nice to label each of the states with a number. Let's tackle each of these.

3.9.3 Revising the Plot

Axis Labels and Limits

We've done this before in Section 1.10, but to re-cap, we'll use the `xlabel`, `ylabel`, and `axis` commands.

```
xlabel('Spec. Volume [m^3/kg]')
ylabel('Pressure [kPa]')
axis([0, 1, 0, 4500])
```

Adiabatic Compression/Expansion Lines

This one is a bit more difficult. In order to create curved lines, we need to create an array to work with. We know the limits of specific volume, so we'll use that as our base.

```
vArray = linspace(v2, v1, 100)
```

Once we have that in hand, we want define our pressure array by substituting the volume array for v_2 in Equation 3.26.

```
pArrayComp = p1*(vArray/v1)**-gamma
```

We can do the exact same thing for the expansion stroke by substituting p_1 for p_3 and v_1 for v_3 :

```
pArrayExp = p3*(vArray/v3)**-gamma
```

Let's add two plot commands to plot these two lines, making them both blue.

```
plot(vArray, pArrayComp, 'b')
plot(vArray, pArrayExp, 'b')
```

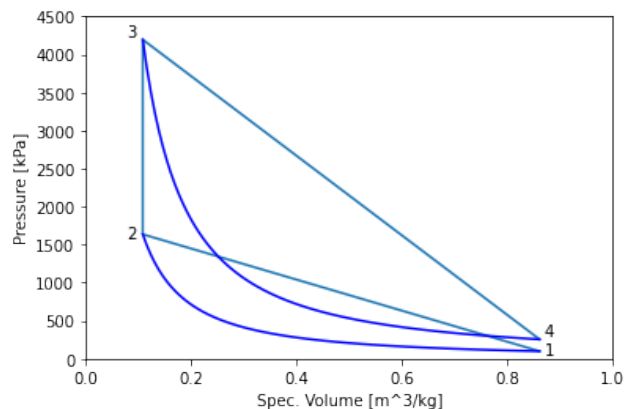
Adding State Labels

To label the states, we use matplotlib's `text` command:

```
text(v1+.01, p1-50, '1') # Place the text '1' next to state 1
text(v2-.03, p2-50, '2')
text(v3-.03, p3+50, '3')
text(v4+.01, p4+50, '4')
```

The first two arguments tell Colab where to place the text, and the third argument is what text to type. This can be finicky, and the easiest way to determine correct placement is just to alter the values added or subtracted from the specific volume and pressure.

After all of the changes above, we end up with the following:



Finishing Up

To clean up the plot, we want to get rid of our original lines and manually put in the lines for constant volume heating and cooling.

For the first bit, we'll change our lines to simple circles by adding 'ob' to the end of the plot command. The o piece makes small circles instead of the lines, while b keeps the circles blue.

The command becomes:

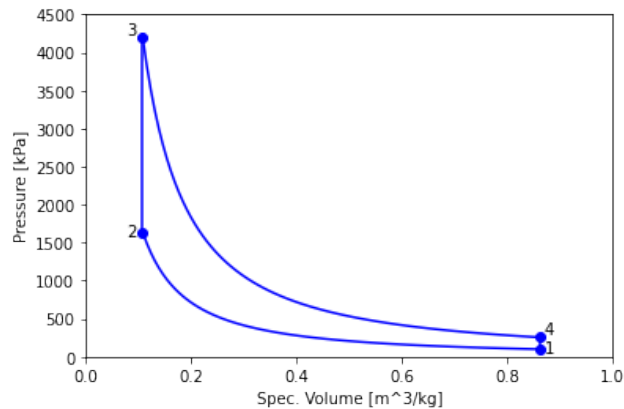
```
plot([v1, v2, v3, v4], [p1, p2, p3, p4], 'ob')
```

In order to create the constant volume lines, we just connect the two points:

```
plot([v2, v3], [p2, p3], 'b')
```

```
plot([v4, v1], [p4, p1], 'b')
```

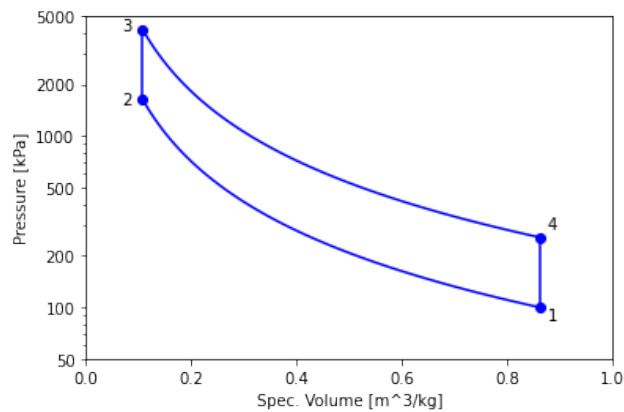
After all this, we end up with the final plot:



Semi-log Plots

There's an argument for using semi-log or log-log plots, depending on differences in pressure.

Replacing all of the `plot` commands with `semilogy` and changing the y-axis from a $0 < p < 4500$ to $50 < p < 5000$ yields the following plot:



In order for the plot to come out well, it is also necessary to fiddle with the `text` commands. In addition, to fix the tick labels on the y axis, it is also necessary to add a `yticks` command:

```
yticks(ticks=[50, 100, 200, 500, 1000, 2000, 5000],
      labels=[50, 100, 200, 500, 1000, 2000, 5000])
```

3.10 Summary

The main result from this chapter was the explanation of three engine cycles: Stirling, Otto, and Diesel. Each of these cycles can convert excess heat, typically from combustion, to work. The Stirling cycle can also use work in order to transfer heat from a cold environment to a hot environment, thus acting as either a heater or a refrigerator. All of these cycles typically use ideal gases as the working fluid, so the ideal gas law was used in abundance.

In order to analyze these cycles, we introduced the First Law of Thermodynamics:

$$\Delta U = Q - W,$$

which is a statement of the conservation of energy. The change in internal energy can be determined from the temperature change between states:

$$\Delta U = U_2 - U_1 = mc_v(T_2 - T_1) \quad \Delta u = u_2 - u_1 = c_v(T_2 - T_1)$$

The work is typically found through integration, which results in the following for special cases:

Isobaric	\rightarrow	$W = mp(v_2 - v_1) = p(V_2 - V_1)$
Isothermal	\rightarrow	$W = mRT \ln \left(\frac{v_2}{v_1} \right) = mRT \ln \left(\frac{V_2}{V_1} \right)$
Adiabatic	\rightarrow	$W = m \frac{p_1 v_1 - p_2 v_2}{\gamma - 1} = mc_v(T_2 - T_1)$
Isometric	\rightarrow	$W = 0$

For the special case of constant pressure, the heat transfer for a process can be found using the constant pressure specific heat, c_p :

$$Q = mc_p(T_2 - T_1)$$

In all other circumstances, the heat transfer can only be determined through the First Law.

Finally, because two of the cycles use adiabatic compression/expansion, it was necessary to develop equations for the properties. These are summarized below:

$$\begin{aligned} \frac{p_2}{p_1} &= \left(\frac{T_2}{T_1} \right)^{\frac{\gamma}{\gamma-1}} = \left(\frac{v_1}{v_2} \right)^{\gamma} \\ \frac{T_2}{T_1} &= \left(\frac{p_2}{p_1} \right)^{\frac{\gamma-1}{\gamma}} = \left(\frac{v_1}{v_2} \right)^{(\gamma-1)} \\ \frac{v_2}{v_1} &= \left(\frac{p_1}{p_2} \right)^{\frac{1}{\gamma}} = \left(\frac{T_1}{T_2} \right)^{\frac{1}{(\gamma-1)}} \end{aligned}$$

Chapter 3 Homework:

Question 3.1: Look at Figure 3.2. Which process would require the most work to complete? What does this say about the heat transfer during that process?

Question 3.2: Why is work negative in the First Law of Thermodynamics?

Question 3.3: In what circumstances is enthalpy useful? In other words, why would we use enthalpy instead of internal energy?

Question 3.4: Which is larger, internal energy or enthalpy? Is this always true? Why?

Question 3.5: In what ways is specific heat different between gases and solids?

Question 3.6: In words, what is c_p ?

Question 3.7: What is the function of the regenerator? Without a regenerator, what would be the effect on the engine?

Question 3.8: Is there heating that occurs during constant temperature expansion/-compression? If so, how much?

Question 3.9: Draw the cycles in a p - v diagram for the Stirling Cycle Engine and Stirling Cycle Cooler. What is different between the two? Label all heat flow and work in and out of the the system. Use Q_H , Q_C , Q_R , W_{in} , and W_{out} for your labels.

Question 3.10: What is γ ? How does it change with temperature (see Table C.2)?

Question 3.11: What assumptions are necessary to use Equations 3.25, 3.26, and 3.27?

Question 3.12: How does the Otto cycle differ from the Stirling cycle?

Question 3.13: What parameters determine the efficiency of the Otto cycle?

Question 3.14: How does the Diesel cycle differ from the Stirling and Otto cycles?

Question 3.15: Why does the Diesel cycle typically use a higher compression ratio than the Otto cycle?

Question 3.16: Four kilograms of water at 50°C are placed in a piston cylinder device under 6 MPa of pressure. Heat is added to the water at constant pressure until the temperature of the steam reaches 400°C.

- Determine the work done by the fluid (W) [1.1 kJ] and the heat transferred to the fluid (Q) [11.9 kJ] during this process.
- Plot the process on a p - v diagram.

Question 3.17: For each of the following states, use the steam tables (or CoolProp) to check the relationship $h = u + pv$:

- a) Steam at 300°C and a pressure of 200 kPa
[3072.1 kJ/kg = 2808.8 kJ/kg + 200 kPa · 1.3162 m³/kg]
- b) R-134a at a specific volume of 1 m³/kg and an internal energy of 200 kJ/kg [206.5 kJ/kg]
- c) Carbon dioxide at a temperature of -50°C and a pressure of 100 kPa (use CoolProp) [444.7 kJ/kg]
- d) Helium at a temperature of -150°C and a pressure of 5 MPa (use CoolProp) [658.4 kJ/kg]

Question 3.18: How much heat is required to raise the temperature of Helium by 50°C at constant pressure? [259.7 kJ/kg] What about at constant volume? [155.8 kJ/kg]

Question 3.19: Analyze a Stirling cycle engine which uses air as a working fluid. Let the minimum pressure be 400 kPa, the maximum volume be 1000 cm³, and the minimum volume be 400 cm³. The temperature of the hot side is 100°C, and the temperature of the cold side is 20°C.

- a) Sketch each process and state on a p - v diagram.
- b) Find the pressure at each state. [$p_1 = 400\text{kPa}$, $p_2 = 1000\text{kPa}$, $p_3 = 1273\text{kPa}$, $p_4 = 509\text{kPa}$]
- c) Create a table with the pressure, temperature, and specific volume at each state.
- d) Find the change of internal energy, work, and heat transfer for each process. Put this information in a table.
- e) Find the net work of a cycle [100 J], the input heat transfer [466 J], and the thermal efficiency [21%].

Question 3.20: Analyze a Stirling cycle cooler which uses helium as a working fluid. Let the minimum pressure be 800 kPa, the maximum volume be 200 cm³, and the minimum volume be 150 cm³. The temperature of the hot side is 20°C, and the temperature of the cold side is -30°C.

- Sketch each process and state on a p - v diagram.
- Find the pressure at each state. [$p_1 = 965\text{kPa}$, $p_2 = 1286\text{kPa}$, $p_3 = 1067\text{kPa}$, $p_4 = 800\text{kPa}$]
- Create a table with the pressure, temperature, and specific volume at each state.
- Find the change of internal energy, work, and heat transfer for each process. Put this information in a table.
- Find the net work per cycle [-9.5 J], the heat transfer from the refrigerated side [46 J], and the coefficient of performance [4.86].

Question 3.21: A frictionless piston-cylinder device contains 0.2 kg of air at 100 kPa and 27°C. The air is now compressed adiabatically until it reaches a final temperature of 77°C.

- Sketch the P-V diagram of the process with respect to the relevant constant temperature lines, and indicate the work done on this diagram.
- Using the basic definition of boundary work done determine the boundary work done during the process [-7.18 kJ].
- Using the energy equation determine the heat transferred during the process [0 kJ], and verify that the process is in fact adiabatic.

Use values of specific heat capacity defined at 300K for the entire process.

Question 3.22: This is an extension of Example 3.5, in which we wish to use throughout all four processes the nominal standard specific heat capacity values for air at 300K. Using the values $c_v = 0.717$ kJ/kgK and $\gamma = 1.4$, determine:

- the temperature and pressure of the air at the end of each process [$p_2 = 1838$ kPa, $T_2 = 689\text{K}$, $T_3 = 1805\text{K}$, $p_3 = 4815$ kPa, $p_4 = 262$ kPa, $T_4 = 786\text{K}$]
- the net work output/cycle [451.5 kJ/kg], and
- the thermal efficiency of this engine cycle. [$\eta_{th} = 56\%$]

Question 3.23: Consider the expansion stroke of a typical Air Standard Diesel cycle engine which has a compression ratio of 20 and a cutoff ratio of 2. At the beginning of the process (fuel injection) the initial temperature is 627°C , and the air expands at a constant pressure of 6.2 MPa until cutoff. Subsequently the air expands adiabatically (no heat transfer) until it reaches the maximum volume.

- Sketch this process on a p - v diagram showing clearly all three states. Shade on the diagram the total work done during the entire expansion process.
- Determine the temperatures reached at the end of the constant pressure (fuel injection) process [1800K], as well as at the end of the expansion process [830K], and draw the three relevant constant temperature lines on the p - v diagram.
- Determine the total work done during the expansion stroke [1087 kJ/kg].
- Determine the total heat supplied to the air during the expansion stroke [1028 kJ/kg].
- Using Google Colab and CoolProp, plot the three states and two processes on a p - v diagram. Use the methodology in Section 3.9.3 to plot the curved expansion process.

Use the specific heat values defined at 1000K for the entire expansion process, obtained from the table of Specific Heat Capacities of Air.

Question 3.24: For an ideal Diesel engine with $r_c = 2$, what does the compression ratio need to be to match the efficiency of an ideal Otto engine with a compression ratio of $r = 8$?

Chapter 4

The First Law for Open Systems

4.1 Energy Conservation (First Law) for Open Systems

In this course we consider three types of **open systems** - steam power plants, refrigeration systems, and aircraft jet engines. These are complex systems, but we will break them down into components, then separately analyze each component. Figure 4.1 shows a representative component with mass flow, shaft work, and heat transfer.

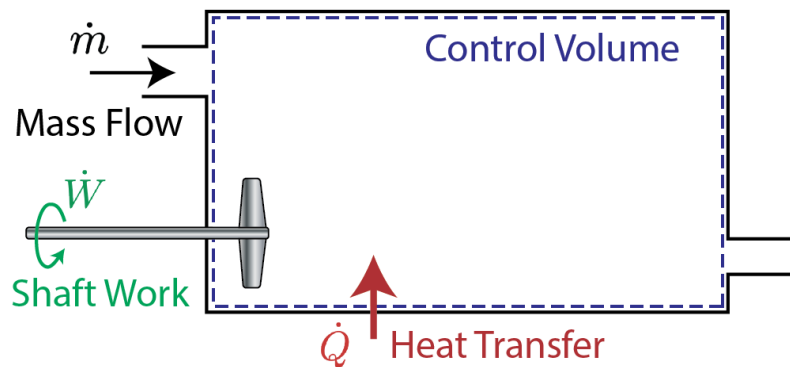


Figure 4.1: Diagram of a control volume (open system).

For every component we analyze in this class, we will assume that there will be no change over time inside the component. This is a very reasonable assumption as long as the component has been running for a while without any changes to its conditions. This is commonly known as the **steady-state** assumption.

For open systems, the consistent mass flow rate means that our components operate continuously, rather than through cycles. While we calculated the amount of energy in one cycle of operation for closed systems, we will calculate the amount of energy in one second of operation for open systems. Thus, it is typically more convenient to use power [kW] rather than energy [kJ]. The mass specific energy [kJ/kg] will still be useful in our analyses.

4.1.1 Mass Flow

Looking again at Figure 4.1, let's start by saying the change of mass inside the control volume, m_{cv} is equal to the mass entering, \dot{m}_{in} , minus the mass leaving, \dot{m}_{out} . The sums in Equation 4.1 are there to indicate that we may have multiple inlets and outlets in our problem.

$$\frac{\partial m_{cv}}{\partial t} = \sum \dot{m}_{in} - \sum \dot{m}_{out} \quad (4.1)$$

Assuming steady-state flow allows us to remove $\frac{\partial m_{cv}}{\partial t}$, as there should not be any change with respect to time:

$$\sum \dot{m}_{in} = \sum \dot{m}_{out} = \dot{m} \quad (4.2)$$

For a steady-state control volume with a single exit and outlet, we can simplify to the following:

$$\dot{m}_{in} = \dot{m}_{out} = \dot{m}$$

In other words, the mass flow must be a constant value through each component for a steady-state problem.

4.1.2 Flow Work

The fluid mass flows through the inlet and exit ports of the control volume. As it flows in, it performs pressure work on the fluid inside the control volume. This **flow work** is equivalent to the amount of work it would take to move a piston at the pressure of the incoming fluid.

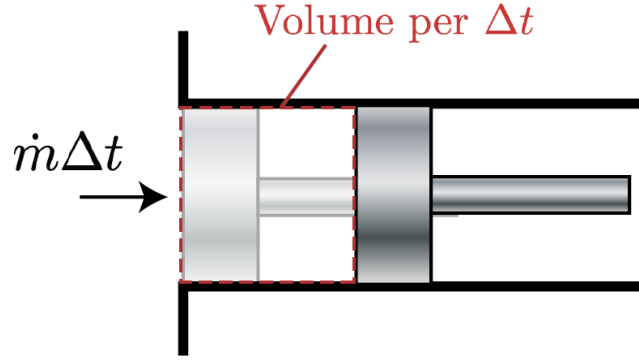


Figure 4.2: Fluid entering a control volume performs work just as if it were moving a piston.

Let us calculate the amount of work done by the flow coming in over a time Δt .

$$W_{flow} = \int F dx = \int_0^V p dV = pV = mpv \quad (4.3)$$

In other words, the amount of work done in Δt is equal to the pressure of the incoming flow times the amount of volume of fluid coming in over the same time period.

Considering the time that the operation takes, we rearrange Equation 4.3 as follows:

$$\frac{W_{flow}}{\Delta t} = \frac{mpv}{\Delta t} \quad \rightarrow \quad \dot{W}_{flow} = \dot{m}pv$$

The dotted values simply represent the quantity of work or mass **per unit time**. Thus, we can say that $m = \dot{m}\Delta t$.

We can define specific values in a similar manner to closed systems by using \dot{m} instead of m :

$$w_{flow} = \frac{\dot{W}_{flow}}{\dot{m}} = pv \quad (4.4)$$

This result will be useful as we develop the complete energy equation in the next section.

4.1.3 Conservation of Energy for an Open System

Let's look again at a sample control volume, but this time we will consider energy flow instead of mass flow.

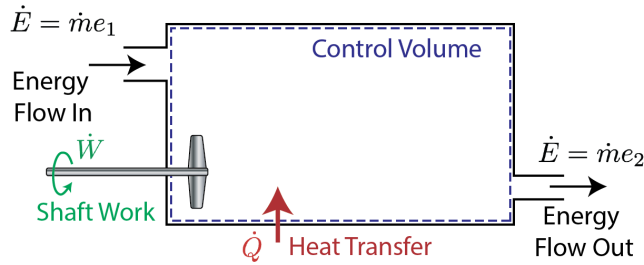


Figure 4.3: Diagram showing all sources of energy affecting an open system.

Figure 4.3 shows that energy enters the control volume through mass flow at 1 ($\dot{m}e_1$) and through heat transfer (\dot{Q}) and work (\dot{W}). Likewise, energy leaves the control volume through mass flow at 2 ($\dot{m}e_2$). As in closed systems, work is defined such that work performed on the system is negative.

$$\dot{m}e_1 + \dot{Q} - \dot{W} = \dot{m}e_2$$

At this point, we should talk about the energy, e . There are many different types of energy, and in the broadest sense, e incorporates all of them. We've discussed potential, kinetic, and internal energy back in Chapter 1. We could also include chemical energy (combustion, etc.), nuclear energy (for nuclear reactors), or other more exotic forms. For the sake of this discussion, we will limit ourselves to the three types discussed in Chapter 1.

$$e = u + \frac{1}{2}\vec{V}^2 + gz \quad (4.5)$$

The \dot{W} term will encompass three types of work. First, we will incorporate the flow work from Equation 4.3. Then, we will split our shaft work into work done to the system by pumps and compressors (\dot{W}_{pumps}) and work extracted from the system by turbines ($\dot{W}_{turbines}$).

$$\dot{W} = \dot{m}(p_2v_2 - p_1v_1) - \dot{W}_{pumps} + \dot{W}_{turbines} \quad (4.6)$$

Our energy equation now has the following form:

$$\dot{m} \left(u_1 + \frac{1}{2}\vec{V}_1^2 + gz_1 \right) + \dot{Q} - \left(\dot{m}(p_2v_2 - p_1v_1) - \dot{W}_{pumps} + \dot{W}_{turbines} \right) = \dot{m} \left(u_2 + \frac{1}{2}\vec{V}_2^2 + gz_2 \right)$$

We can distribute the flow work to group it with the other forms of energy, and move the \dot{W} terms to the right hand side:

$$\dot{m} \left(u_1 + p_1 v_1 + \frac{1}{2} \vec{V}_1^2 + gz_1 \right) + \dot{Q} = \dot{m} \left(u_2 + p_2 v_2 + \frac{1}{2} \vec{V}_2^2 + gz_2 \right) - \dot{W}_{pumps} + \dot{W}_{turbines}$$

Using the definition of enthalpy from Section 3.2, we recognize that $u + pv = h$.

$$\dot{m} \left(h_1 + \frac{1}{2} \vec{V}_1^2 + gz_1 \right) + \dot{Q} = \dot{m} \left(h_2 + \frac{1}{2} \vec{V}_2^2 + gz_2 \right) - \dot{W}_{pumps} + \dot{W}_{turbines} \quad (4.7)$$

For many problems, we neglect the effects of kinetic and potential energy. In those situations, we can further simplify:

$$\dot{m}h_1 + \dot{Q} = \dot{m}h_2 - \dot{W}_{pumps} + \dot{W}_{turbines}$$

If we define $\Delta h = h_2 - h_1$, we can get an equation very similar to Equation 3.1.

$$\dot{m}\Delta h = \dot{Q} + \dot{W}_{pumps} - \dot{W}_{turbines} \quad (4.8)$$

Finally, we can get the specific form of the equation by dividing through by \dot{m} .

$$\Delta h = q + w_{pumps} - w_{turbines} \quad (4.9)$$

Notice that enthalpy h is fundamental to the energy equation for open systems.

4.2 The p - h (Pressure-Enthalpy) Diagram

When dealing with closed systems we found that sketching T - v or p - v diagrams was a significant aid in describing and understanding the various processes. In steady flow systems we find that the pressure-enthalpy (p - h) diagrams serve a similar purpose, and we will use them extensively. In this course we consider two pure fluids - water and refrigerant R134a, and we have provided p - h diagrams for each in Appendices A.5 and B.4. We will illustrate their use in the following examples. The p - h diagram for water is shown below.

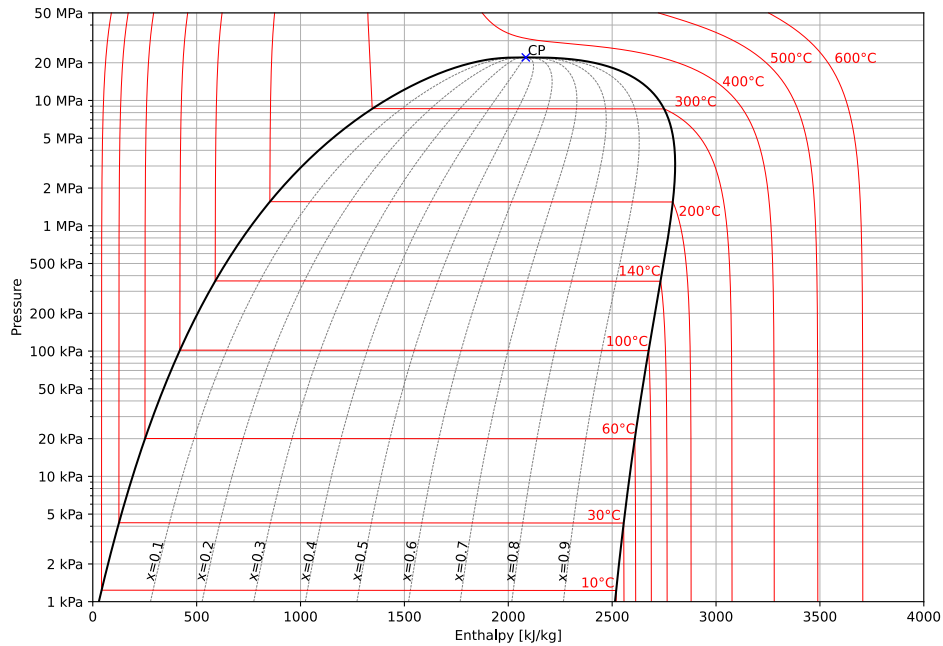


Figure 4.4: Pressure-enthalpy diagram for water. Note the vertical isothermal lines at high enthalpy and low pressure (ideal gas) and for supercooled water (incompressible), as well as the horizontal isothermal lines for the saturated mixture.

Importantly, in the ideal gas range, enthalpy is primarily determined by temperature ($\Delta h = c_p \Delta T$), leading to the vertical isothermal lines in that region. The same is true in the compressed liquid, as liquids are typically incompressible. For liquid-vapor mixtures, the enthalpy changes mostly with quality. Finally, in the supercritical region, both temperature and pressure will have an impact on enthalpy.

4.3 Components of Thermodynamic Systems

In open systems, fluids are transferred between multiple components, each of which has a specialized purpose. Most simply, a component will add or extract heat, perform work on the fluid as it passes through, or extract energy in the form of work from the fluid as it passes through. This is a different paradigm from the closed systems of Chapter 3, in which fluid was stuck in piston, heat was added, and work was extracted.

Cycles for open systems are created by chaining these components. For instance, in a Rankine Power Plant (see Section 4.4), water will move from a turbine to a condenser, then through a pump and a boiler, and finally back to the turbine. The states are defined between the various components, and each component is described as a thermodynamic process.

The rest of this section is devoted to describing the various components. Turbines, compressors, and pumps all deal with work. Boilers, condensers, and evaporators either add or remove heat. Finally, mixers, throttles, heat exchangers, and de-aerators all have other impacts on the fluid.

4.3.1 Turbines

Typically, anything that extracts work from a moving fluid (either liquid or gas) is known as a turbine. While there are many different varieties of turbines, all of them use high pressure within the fluid to provide torque to a rotating shaft, typically by way of specially designed blades. In the case of power plants, the rotating shaft is connected to an alternator, which turns the rotational mechanical energy into electrical energy.

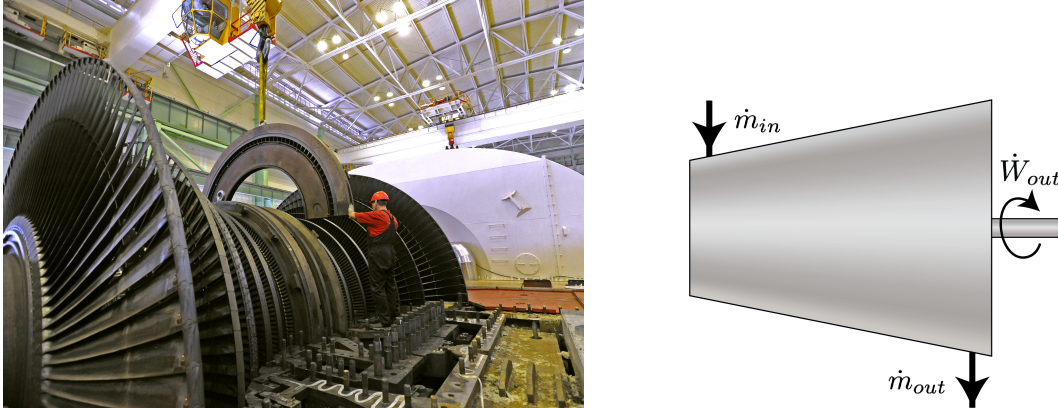


Figure 4.5: Left: From [Wikipedia](#). A large steam turbine without its cover. Right: A typical representation of a turbine for cycle diagrams.

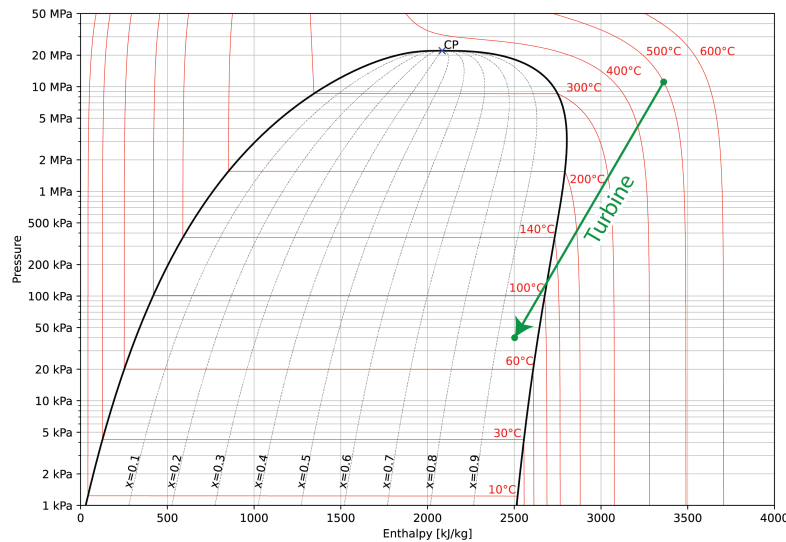


Figure 4.6: Typical turbine process plotted on a p - h diagram for steam.

Within a turbine, you should expect the pressure, enthalpy, and temperature to decrease while the specific volume increases. For a well designed steam turbine, the flow should start superheated or supercritical, and end either as saturated steam or as a gas-liquid mixture

with a quality of at least 90%. If the quality drops too low, the large number of water droplets will begin to erode the turbine blades.

Figure 4.6 shows a turbine process on a p - h diagram, which highlights the drop in pressure and enthalpy, as well as the final state of water being at a fairly high quality. In Example 4.2, we will see a Rankine cycle with two separate turbines. In modern power plants, the turbine is often subdivided even further to account for high temperature steam being bled off for various purposes. This can be seen in Example 4.3, in which some of the steam is used to preheat the water before boiling.

4.3.2 Compressors and Pumps

Compressors and pumps both act to add energy to a fluid through work. The primary difference is that pumps act on liquids, while compressors work on gases. Compressors and pumps come in many forms, varying from rotary compressors that look like the turbine in Figure 4.5 (though usually smaller) to reciprocating pumps that use pistons driven by a crankshaft, along with many other varieties.

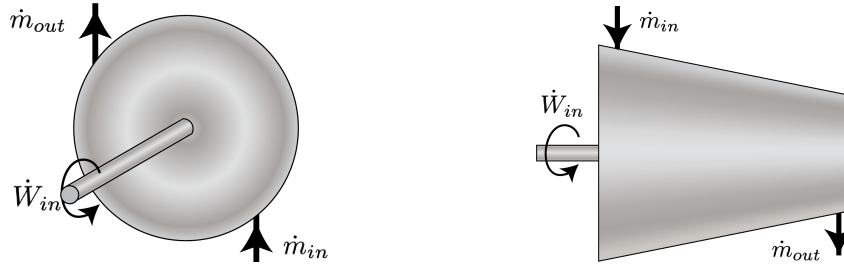


Figure 4.7: Left: Pump in for cycle diagrams. Right: Compressor in for cycle diagrams.

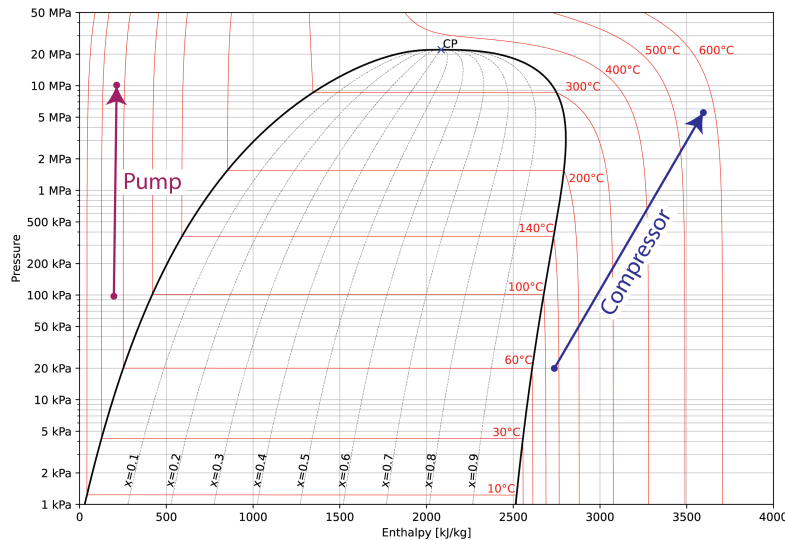


Figure 4.8: Typical compressor and pump processes plotted on a p - h diagram for steam.

Thermodynamically, compressors are nearly the opposite of turbines, as can be seen by comparing Figures 4.6 and 4.8. Pressure, enthalpy, and temperature all increase through the compressor, while specific volume decreases. One note is that compressors typically operate only in the superheated and supercritical regimes. In other words, droplets of liquid, as are present in the quality region, are typically worse for compressors than they are for turbines.

Pumps operate in liquids, which are usually modeled as incompressible. The result of this is that the amount of enthalpy change over a pump is very small compared to compressors. In fact, we can often neglect the energy requirements of the pump in our analysis, and expect less than 1% error in our final results. An example pump is also shown in Figure 4.8. If we do choose to calculate the enthalpy change in the pump, we can expand Δh as follows:

$$\begin{aligned}\Delta h &= \Delta u + \Delta(pv) \\ \Delta h &= c_p \Delta T + p \Delta v + v \Delta p \\ \Delta h &= v \Delta p\end{aligned}\tag{4.10}$$

The temperature change ΔT is assumed small, and the change in specific volume can be removed by assuming liquid water is incompressible. The only remaining term occurs due to the change in pressure with constant volume.

Like compressors, pumps should not operate in the quality region. Whereas droplets of water can cause erosion of the compressor machinery, bubbles of steam can occur as cavitation, which cause significantly more serious pressure spikes that can damage or destroy pump machinery.

4.3.3 Boilers, Evaporators, and Steam Generators

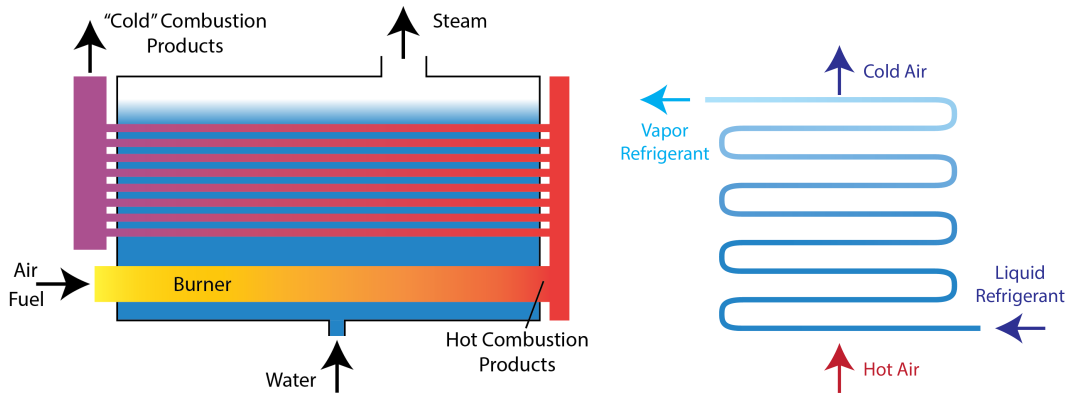


Figure 4.9: Left: Gas fired boiler, showing the inflow and burning of fuel and gas, the cooling of combustion products as the heat is transferred to the water, and the conversion of water into steam. Right: Evaporator, showing the flow of refrigerant as it evaporates inside the pipes, as well as the air flowing outside of the pipes.

Boilers, evaporators, and steam generators all add heat to a fluid, and the end result is a gas, as shown in Figure 4.10. Boilers and steam generators both add heat at elevated pressure. A boiler adds heat to a compressed liquid, changing its phase to steam through the vapor

envelope. Steam generators do not literally boil the liquid, as the starting point is water at supercritical pressure. As heat is added, the vapor envelope is avoided completely, and the end result is supercritical vapor.

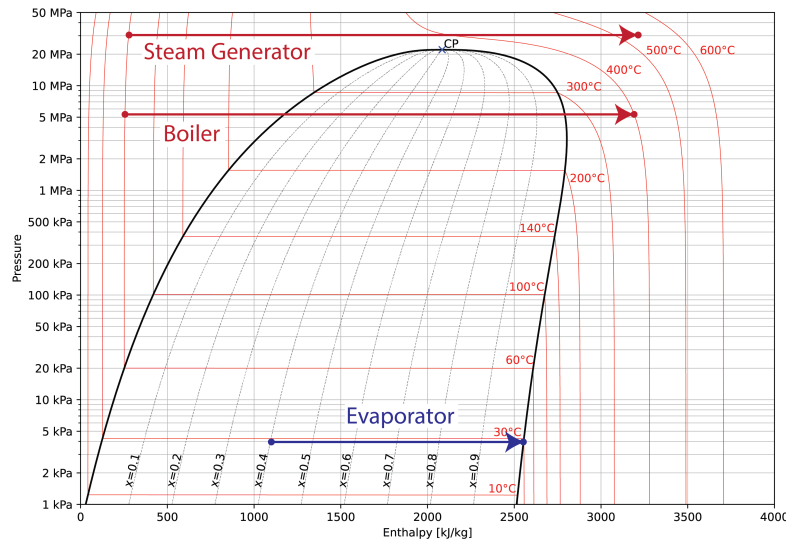


Figure 4.10: Typical boiler, evaporator, and steam generator processes plotted on a p - h diagram for steam.

Evaporators typically operate at lower pressure and temperature, and find use in HVAC (air conditioning) units. In this class, we will see evaporators using R134a, though they are used for water in the food industry (concentration of juices, for instance).

In all three cases, the pressure is assumed to remain constant while the enthalpy increases. In reality, there is typically a slight pressure decrease associated with friction in the fluid flow, but this is neglected in this class.

4.3.4 Condensers

A condenser is a type of heat exchanger in which heat is extracted from the steam or refrigerant in our system/cycle and sent into water or air outside of the system/cycle.

Like boilers and evaporators, the pressure is assumed to remain constant through a condenser, while the enthalpy drops drastically. For power plants, condensation occurs at relatively low temperatures, while for air conditioning units, condensation occurs at relatively high temperatures. Both are shown in Figure 4.11.

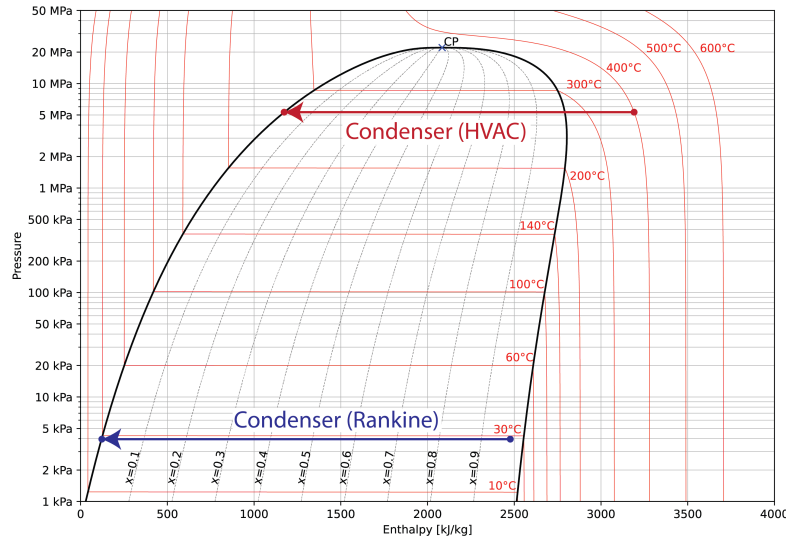


Figure 4.11: Typical condenser processes for power plants and air conditioning units plotted on a p - h diagram for steam.

Condensers in power plants typically consist of many small pipes containing cold water, with steam flowing over the pipes while cooling and condensing. The cold water heats up, and the steam condenses back to liquid water. Condensers therefore require a constant supply of cold water, typically from a nearby river or lake. The amount of water required can be reduced through the use of a cooling tower, which uses evaporative cooling to extract large amount of energy.

Condensers for HVAC applications are normally a single pipe containing the refrigerant, over which air is blown. Because of the high pressure of the refrigerant in HVAC systems, the air is normally significantly cooler, which allows the refrigerant to condense in the pipe. In essence, this works opposite from the evaporator shown in Figure 4.9.

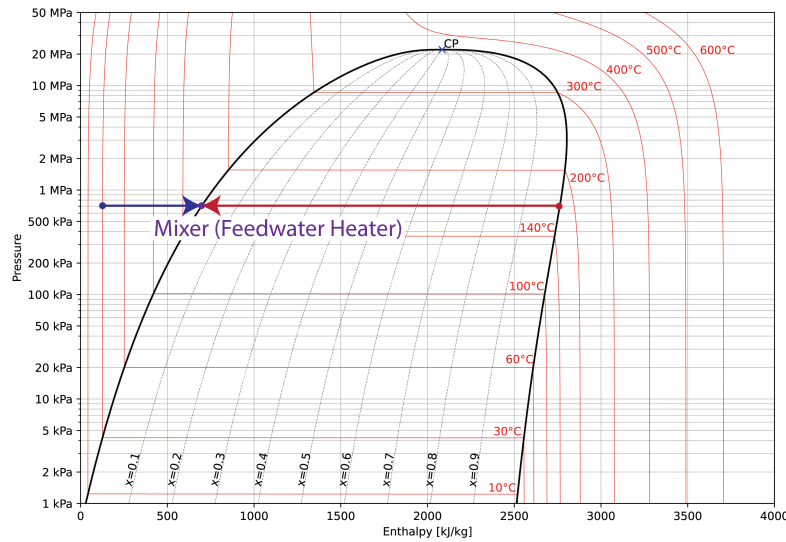
4.3.5 Mixers

Mixers allow multiple streams of fluid to reunite as a single stream. The output of a mixer is governed by the following equation:

$$\dot{m}_{out}h_{out} = \sum \dot{m}_{in}h_{in} \quad (4.11)$$

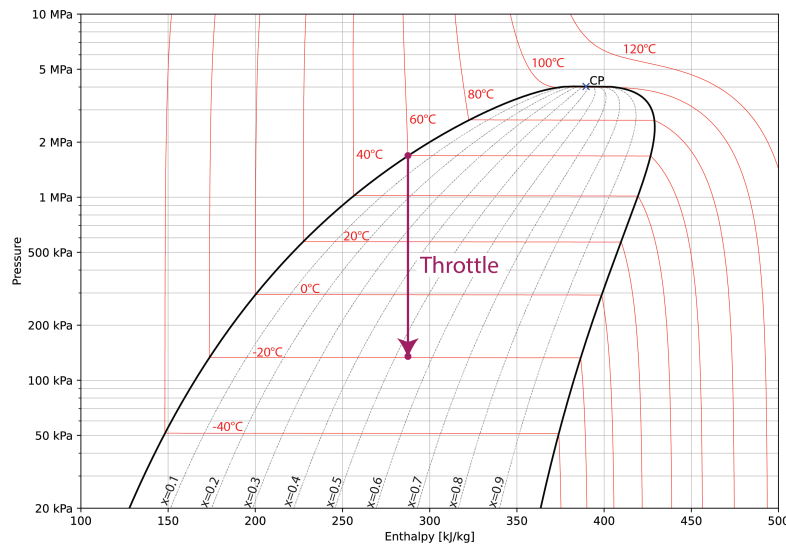
Equation 4.11 assumes that the mixing process requires no work and is adiabatic.

Open feedwater heaters are examples of mixers that are seen in more advanced power cycles. A portion of the hot steam is siphoned off after some work has been extracted, and that steam is mixed with water after condensation. The end result is that the water can be heated to near the boiling point, and less fuel is needed in the boiler.

Figure 4.12: Typical mixer process plotted on a $p-h$ diagram for steam.

4.3.6 Throttles

Throttles are components used primarily in HVAC systems. They are able to reduce the pressure by changing mechanical energy to thermal energy through friction. Essentially, they force the refrigerant to flow through very small holes or tubes to maximize the pressure loss. In the HVAC profession, the throttle will be called a metering device, which could be either a capillary tube (typical for refrigerators) or an expansion valve (for residential A/C or more complex units).

Figure 4.13: Typical throttle process plotted on a $p-h$ diagram for R-134a.

Throttles have no moving parts, and so cannot perform or extract any work. Additionally, it is typically assumed that throttles are adiabatic. Without any work or heat transfer, throttles cannot change the enthalpy in the flow. That is, the throttling process is isenthalpic, or constant enthalpy, resulting in a vertical process in a p - h diagram, as shown in Figure 4.13.

4.3.7 Heat Exchangers

Heat exchangers come in many shapes and sizes. The key concept is bringing a hot fluid in close thermal contact to a cold fluid, and letting the heat naturally move from hot to cold. As you can see in Figure 4.14, the hot side of the heat exchanger is losing enthalpy, while the cold side is gaining it.

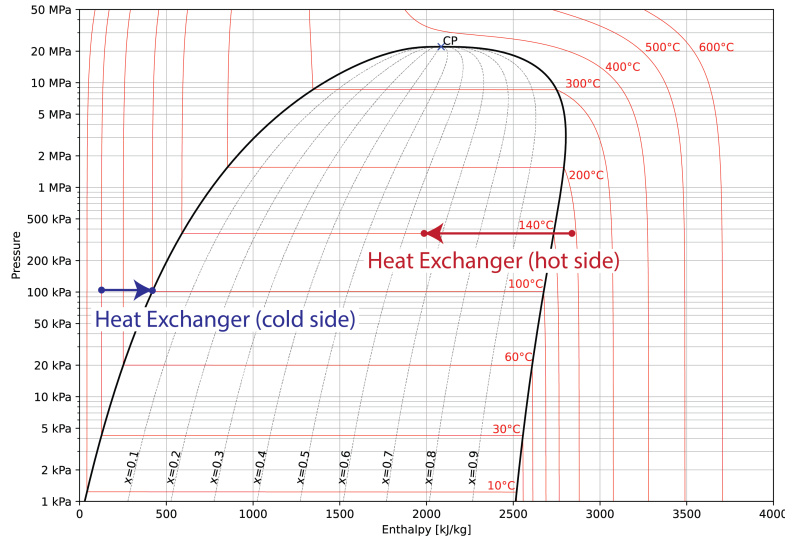


Figure 4.14: Both hot side and cold side of a heat exchanger plotted on a p - h diagram for steam.

The amount of enthalpy gained/lost must satisfy conservation of energy, as written in the Equation 4.12.

$$\dot{Q}_{hot} = \dot{m}_{hot} (h_{hot,in} - h_{hot,out}) = \dot{m}_{cold} (h_{cold,out} - h_{cold,in}) = \dot{Q}_{cold} \quad (4.12)$$

In reality, boilers, condensers, etc. all qualify as heat exchangers. The difference is purely in analysis. We typically consider both the hot and cold flows of heat exchangers, while we are only interested in the heat transfer in boilers, condensers, etc.

4.3.8 De-aerators

De-aerators are used to remove dissolved gases (mostly oxygen) from water in steam power plants, so that they cannot corrode the metal piping or components. They are therefore an important design consideration for power plants, even though they do not serve a thermodynamic purpose.

In other words, the level of analysis we will perform on these components is minimal. In Example 4.3, the de-aerator is combined with the open feedwater heater (a mixer), which is about as in-depth as we get. Because de-aerators don't significantly change the properties of the fluid, they typically show up as a single state on a p - h diagram, rather than a process.

The operation principle behind de-aerators is that the trapped gases have lower solubility (the ability to be dissolved) in hot water than in cold water. By heating the water up, the trapped gases are forced out of the water, and are then released through a pressure valve.

4.4 Basic Rankine Cycle

The most basic form of a steam power plant uses the Rankine cycle, shown in Figure 4.15. There are four components in the Rankine cycle: the turbine, boiler, condenser, and pump.

State 1 is traditionally located at the highest temperature/pressure combination, which occurs immediately after the boiler. The turbine removes work from the system in Process 1-2 (see Section 4.3.1). State 2 is usually either saturated steam or vapor-liquid mixture with a fairly high quality ($x_2 > 0.9$). This steam is piped into a condenser in Process 2-3, which removes enough heat to turn the steam into saturated water (see Section 4.3.4). Process 3-4 consists of a pump, which significantly increases the pressure of the water by adding work (see Section 4.3.2). Finally, the water is turned back into steam in the boiler of Process 4-1 (see Section 4.3.3). This cycle is depicted in Figure 4.15.

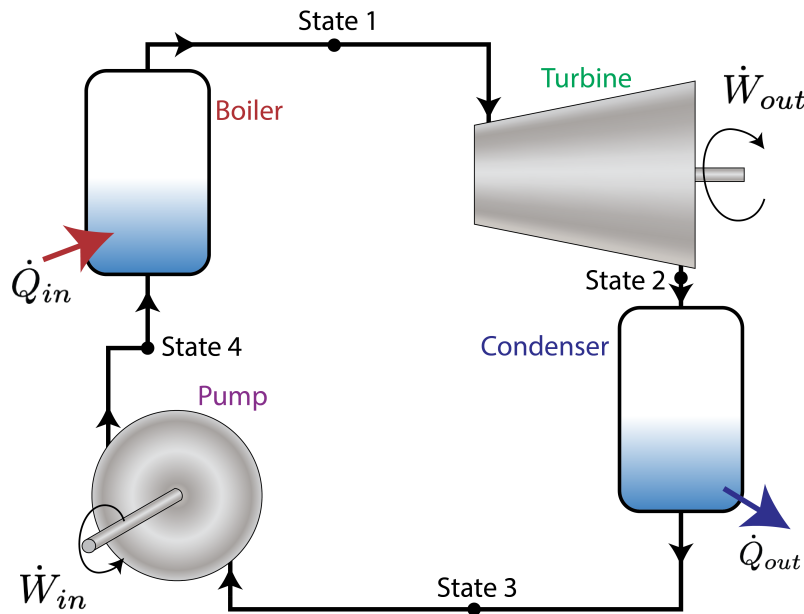


Figure 4.15: Diagram describing the Rankine Cycle for steam power plants.

Ideally, for each process, there is an assumption that only work or heat transfer occurs. In the case of pumps and turbines, this occurs through an adiabatic assumption, which is reasonable given sufficient insulation. The assumption of no work done occurs in the boiler and condenser because there are typically no moving parts in either of these components,

which is a necessary part of adding or extracting energy through work.

For the analysis of this system, we are typically interested in the thermal efficiency, which is defined almost identically to our original definition in Equation 3.20. Of course, we use \dot{Q} instead of Q and \dot{W} instead of W :

$$\eta_{th} = \frac{\dot{W}_{net}}{\dot{Q}_{in}} = \frac{w_{net}}{q_{in}}$$

The net work for the Rankine cycle can be defined as $w_{net} = w_{turb} - w_{pump}$. Often, we neglect the work from the pump, meaning that the net work can be assumed to be equal to the work from the turbine ($w_{net} \approx w_{turb}$). The heat transfer into the system occurs in the boiler, so $q_{in} = q_{boiler}$.

This gives us an efficiency for the Rankine cycle:

$$\eta_{th} = \frac{w_{turb} - w_{pump}}{q_{boiler}} \approx \frac{w_{turb}}{q_{boiler}} \quad (4.13)$$

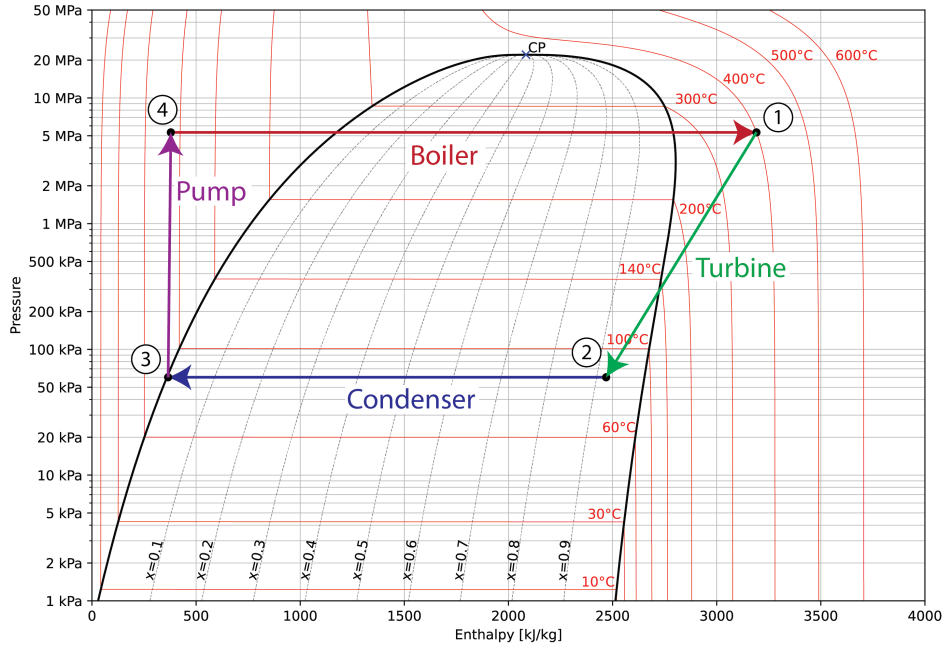


Figure 4.16: The four processes of the Rankine cycle plotted on a p - h diagram for steam.

Because no work occurs in the boiling process, and no heat transfer occurs in the turbine, Equation 4.8 can be used almost trivially for both the turbine and the boiler:

$$\begin{aligned} \dot{m}\Delta h_{1-2} &= \dot{Q}_{1-2} - \dot{W}_{1-2} \xrightarrow{0} w_{turb} = h_1 - h_2 \\ \dot{m}\Delta h_{4-1} &= \dot{Q}_{4-1} - \dot{W}_{1-2} \xrightarrow{0} q_{boil} = h_1 - h_4 \end{aligned}$$

Both of these differences can be seen visually in Figure 4.16 as the horizontal distance between states 1 and 2, and states 4 and 1. Clearly, $h_2 - h_1 < h_1 - h_4$, meaning that more heat is added to the boiler than is extracted from the turbine, giving us an efficiency of less than one, as expected.

We can rewrite the efficiency using only the state enthalpies as follows:

$$\eta_{th} \approx \frac{w_{turb}}{q_{boiler}} = \frac{h_1 - h_2}{h_1 - h_4} \quad (4.14)$$

Example 4.1 analyzes a specific Rankine cycle.

Example 4.1: Rankine Cycle

A steam power plant has been designed using the Rankine cycle. The turbine inlet temperature was chosen to be 500°C, with a pressure at that point of 10 MPa. After the turbine, the pressure will be 20 kPa, with a quality of 90%. The condenser will cool the water to 40°C, and a pump and boiler will complete the cycle. The flow rate of steam through the power plant will be 8 kg/s.

Using this information, do the following:

- draw each of the states and processes on the p - h diagram for water
- find the power of the turbine and feedwater pump
- find the heat transfer rates of the boiler and condenser
- calculate the thermal efficiency of the system

Solution Methodology

To start, we will create a table with the given information of each of the states, looking at Figure 4.15 to define each state. For instance, the “turbine inlet” refers to State 1, as it is the state immediately before the turbine.

After we find the relevant information for all of the states, we will create a p - h diagram for part a), then use that diagram to find the power and heat transfer rates for each of the components. The thermal efficiency will naturally follow once we know the specific work and heat transfer for each process.

State 1 and 2 Properties

It is possible to fill in the remaining information for States 1 and 2, as both have two properties defined. In this case, h_1 can be read directly from Appendix A.3, and h_2 can be found using Appendix A.2 and Equation 2.1:

$$h_2 = xh_g + (1 - x)h_f = 0.9 \left(2608.9 \frac{\text{kJ}}{\text{kg}} \right) + 0.1 \left(251.42 \frac{\text{kJ}}{\text{kg}} \right) = 2373.2 \frac{\text{kJ}}{\text{kg}}$$

	State 1	State 2	State 3	State 4
Pressure	10 MPa	20 kPa		
Temperature	500°C	60.06°C	40°C	
Enthalpy	3375.1 $\frac{\text{kJ}}{\text{kg}}$	2373.2 $\frac{\text{kJ}}{\text{kg}}$		
Quality	-	0.9		

State 3 Properties

To find the properties of State 3, we refer to Section 4.3.4. The key assumption for us here is that the pressure will remain constant across the condenser. This gives us a pressure of $p_3 = 20 \text{ kPa}$, which allows us to find the rest of the properties.

Unfortunately, we run into a small hitch in this process, which is that the most relevant chart, Appendix A.4 does not have data anywhere near 20 kPa. We therefore use the assumption that the enthalpy will be very near the enthalpy of saturated water at 40°C. This gives us an enthalpy of $h_3 = 167.53 \frac{\text{kJ}}{\text{kg}}$.

Alternatively, we can use CoolProp to determine the properties, including the effect of the slightly higher pressure. After our normal setup, we use the following command (converting to Pascals and Kelvin as needed):

```
print( CP.PropsSI('H', 'P', 20e3, 'T', 40 + 273.15, 'water') )
```

This command returns a value of 167544.2, which, according to Table 2.1, has units of J/kg. Converting to kJ/kg, we get that $h_3 = 167.54 \frac{\text{kJ}}{\text{kg}}$, which is functionally equivalent to the value from the tables.

State 4 Properties

To find the properties of State 4, we refer to Section 4.3.2 and Section 4.3.3. The assumption that boilers operate under constant pressure allows us to set $p_4 = p_1 = 10 \text{ MPa}$, while the assumption that pumps operate under constant temperature allows us to set $T_4 = T_3 = 40^\circ\text{C}$. The enthalpy after the pump can therefore be read directly from Appendix A.4 as $h_4 = 176.36 \frac{\text{kJ}}{\text{kg}}$.

Alternatively, if the tables are not so clean, we can follow the methodology from Equation 4.10:

$$\Delta h = v\Delta p \rightarrow h_4 - h_3 = v(p_4 - p_3)$$

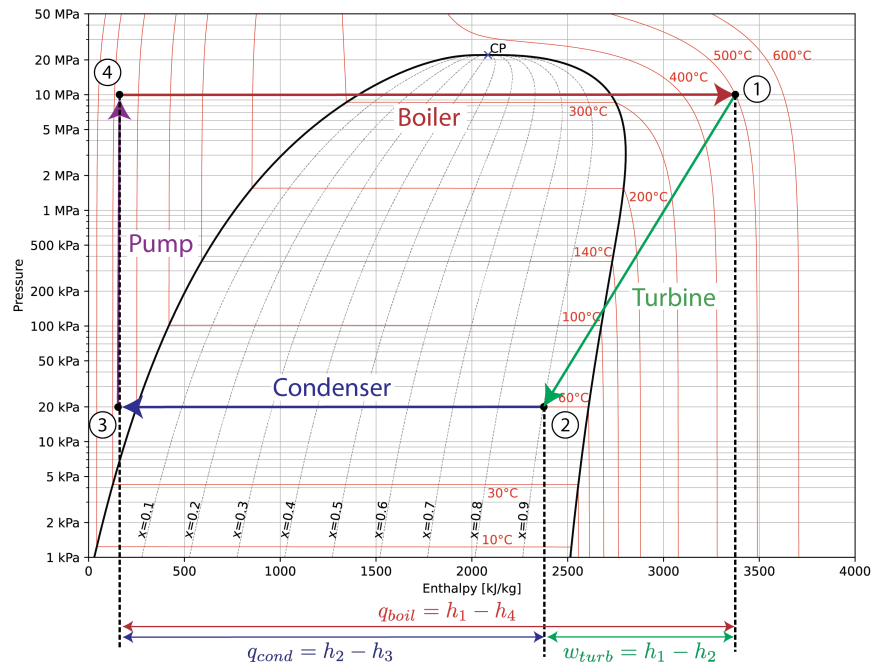
$$h_4 = 167.54 \frac{\text{kJ}}{\text{kg}} + 0.001 \frac{\text{m}^3}{\text{kg}} (10,000\text{kPa} - 20\text{kPa}) = 177.52 \frac{\text{kJ}}{\text{kg}}$$

While this is not perfect, it is a fair approximation. This allows us to complete our property table.

	State 1	State 2	State 3	State 4
Pressure	10 MPa	20 kPa	20 kPa	10 MPa
Temperature	500°C	60.06°C	40°C	40°C
Enthalpy	3375.1 $\frac{\text{kJ}}{\text{kg}}$	2373.2 $\frac{\text{kJ}}{\text{kg}}$	167.54 $\frac{\text{kJ}}{\text{kg}}$	176.36 $\frac{\text{kJ}}{\text{kg}}$
Quality	-	0.9	-	-

p - h Diagram

With the pressure and enthalpy defined for each state, we can now complete our p - h diagram. Δh for each process can be clearly seen by looking at the horizontal distance between the two states. Thus we see that the heat added in the boiler is the largest contribution, while the work done by the pump is negligible.



Additionally, plotting our cycle on the p - h diagram allows us to double check our enthalpies at each state. We can see that h_1 is about 3400 kJ/kg, h_2 is approximately 2300 kJ/kg, and both h_3 and h_4 are between 100 and 200 kJ/kg. These match our calculated values and let us know we are on the right track.

Process Work and Heat Transfer

With our enthalpies calculated, it is straightforward to calculate the work and heat transfer for each process. Both pumps and turbines are typically assumed to be

adiabatic, meaning that we can say $w_{turb} = h_1 - h_2 = 1002 \frac{\text{kJ}}{\text{kg}}$ and $w_{pump} = h_4 - h_3 = 11 \frac{\text{kJ}}{\text{kg}}$. Both of these are typically defined to be positive, even though the turbine results in positive work, while the pump is associated with negative work. Likewise, no work occurs in the condenser or boiler, meaning that $q_{boil} = h_1 - h_4 = 3199 \frac{\text{kJ}}{\text{kg}}$ and $q_{cond} = h_2 - h_3 = 2206 \frac{\text{kJ}}{\text{kg}}$. Again, both heat transfers are defined to be positive, even though the condenser removes heat from the system. The w and q values are all mass-specific. In order to find the actual values for the power plant, we need to multiply by the mass flow rate, \dot{m} .

$$\dot{W}_{turb} = \dot{m} w_{turb} = 8 \frac{\text{kg}}{\text{s}} \cdot 1002 \frac{\text{kJ}}{\text{kg}} = 8016 \text{ kW} = 8.0 \text{ MW}$$

$$\dot{W}_{pump} = \dot{m} w_{pump} = 8 \frac{\text{kg}}{\text{s}} \cdot 11 \frac{\text{kJ}}{\text{kg}} = 88 \text{ kW} = 0.088 \text{ MW}$$

$$\dot{Q}_{boil} = \dot{m} q_{boil} = 8 \frac{\text{kg}}{\text{s}} \cdot 3199 \frac{\text{kJ}}{\text{kg}} = 25,592 \text{ kW} = 25.6 \text{ MW}$$

$$\dot{Q}_{cond} = \dot{m} q_{cond} = 8 \frac{\text{kg}}{\text{s}} \cdot 2206 \frac{\text{kJ}}{\text{kg}} = 17,648 \text{ kW} = 17.6 \text{ MW}$$

Thermal Efficiency

Finally, we can calculate the efficiency of the process. The net work is simply $w_{net} = w_{turb} - w_{pump}$, and the heat transfer into the system occurs solely in the boiler: $q_{in} = q_{boil}$.

$$\eta_{th} = \frac{w_{net}}{q_{in}} = \frac{w_{turb} - w_{pump}}{q_{boil}} = \frac{1002 \frac{\text{kJ}}{\text{kg}} - 11 \frac{\text{kJ}}{\text{kg}}}{3199 \frac{\text{kJ}}{\text{kg}}} = 0.310$$

Thus, the thermal efficiency of the system above is 31.0%. Neglecting the work from the pump results in an efficiency of 31.3%. Clearly it is more accurate to include the work spent on running the pump, but there is a question whether the increased accuracy justifies the extra work in this case. Quite frankly, there are other losses and assumptions that are substantially more significant, and unless we address those, neglecting the pump is perfectly reasonable.

4.5 Modifications to the Rankine Cycle

The Rankine cycle is a good baseline to understand steam power plants. However, most power plants require additional components to create an accurate model. Most commonly, there is not a single turbine, but several, typically labeled as “high pressure” and “low pressure” turbines. Functionally, there is no difference between having a single turbine or multiple turbines, except that it provides an opportunity to do something with the steam in between.

Section 4.5.1 looks at the possibility of **re-heating** the steam in between the high pressure and low pressure turbines. The power plant in Section 4.5.2 extracts some of the steam

between the turbines and mixes it with cold water before the pump in a process called **regeneration**.

Other modifications to the cycle include getting heat from more than one source in a **hybrid** plant, or using excess heat for other purposes (known as **cogeneration**). These modifications are looked at in the homework problems for this chapter.

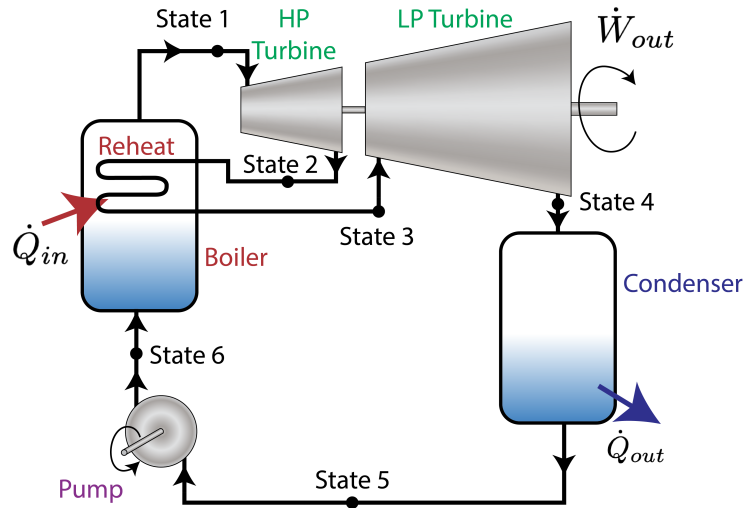
4.5.1 Rankine Cycle with Reheat

The Rankine Cycle with reheat includes a separation between the low pressure and high pressure turbines, which allows for steam to be sent back to the boiler. The steam is reheated either to around the same temperature as State 1, and is then sent to the low pressure turbine. As before, the boiler does not change the pressure of the steam in the reheat process.

Example 4.2: Nuclear Power Plant with Reheat

Modern nuclear power plants typically use pressurized water to transport heat outside of the core. Because of this, they tend to be rather limited in temperature (one of the benefits of molten salt nuclear reactors is the ability of the salt to transport heat at much higher temperatures).

For this example, we are looking at a steam turbine loosely based off of **General Electric's steam turbines for nuclear applications**. The high-pressure turbine inlet pressure and temperature will be 7.5 MPa and 300°C, respectively. After the high-pressure turbine, the steam goes back to the boiler and is again heated to 300°C. This lower pressure steam is sent to the low-pressure turbine, after which it finally goes through the rest of the cycle. The diagram for this is shown in the figure below.



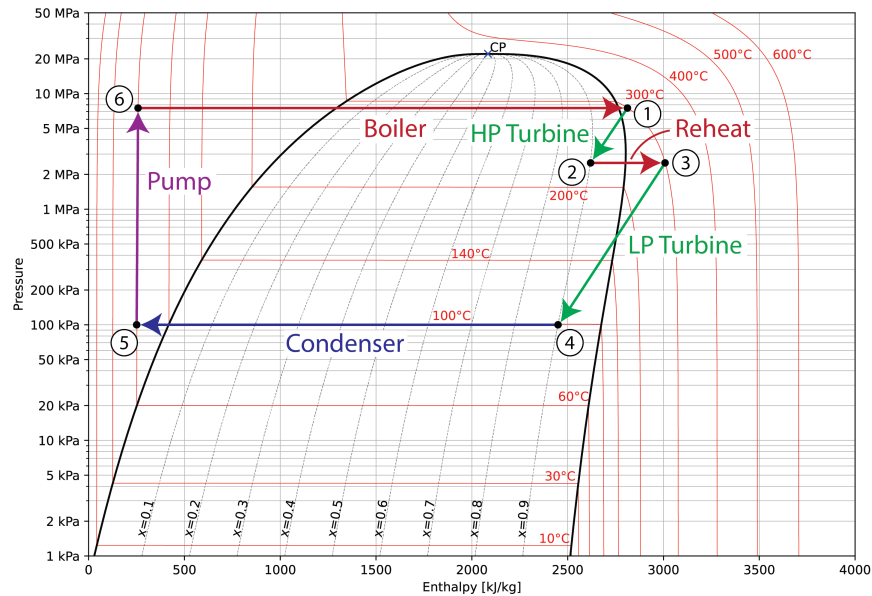
Solution Methodology

In this example, we will first build the p - h diagram, then fill out the table of properties afterwards. We know that State 1 is at 7.5 MPa and 300°C from GE's specifications. We will choose State 2 to be 2.5 MPa, with a quality of 0.9. State 3 will have the

same pressure as State 2, but with a temperature of 300°C again. Again, we choose a quality of 0.9 for State 4, this time with a pressure of 100 kPa. The condenser will cool the water at State 5 to 60°C, still with a pressure of 100 kPa. Finally, the pump will maintain a temperature of 60°C for State 6, but return the pressure to 7.5 MPa. With this information, we will be able to perform the following:

- Plot the cycle on a p - h diagram, using temperature, pressure, and quality lines to find the states.
- Complete a table of properties for each state.
- Determine the work extracted from the turbines per kg of steam.
- Determine the total energy needed from heat transfer in the boiler and reheat processes per kg of steam.
- Find the amount of heat rejected by the condenser per kg of steam.
- Determine the amount of work required to run the pump per kg of steam. Compare values from CoolProp and using the incompressible assumption.
- Find the mass flow rate of steam needed to generate 450 MW of power in the turbines.
- Find the efficiency of the power plant and the total heat transfer rate required from the nuclear reactor.

p - h Diagram



The beauty of the p - h diagram is that we can get a picture of the cycle without ever looking at tables.

- The relatively low temperature of State 1 means that the HP Turbine cannot extract much work without bringing the steam below a quality of 0.9.
- The heat absorbed in the reheat process is substantially less than the original boiling.
- The LP turbine is able to extract much more heat than the HP turbine, in conjunction with a much larger pressure drop.

Property Table

The properties of States 1, 3, and 6 are simple table look-ups (with interpolation required for States 1 and 6). States 2 and 4 require use of Equation 2.1, and State 5 can either be assumed to be equivalent to saturated water at 60°C or evaluated using CoolProp. The table below is populated purely from CoolProp, so small differences may be present compared to interpolation of tables. Quantities in black are taken from the problem statement, and quantities in red are calculated.

State	Pressure	Temperature	Enthalpy	Quality
1	7.5 MPa	300°C	2814.4 $\frac{\text{kJ}}{\text{kg}}$	-
2	2.5 MPa	224°C	2617.9 $\frac{\text{kJ}}{\text{kg}}$	0.9
3	2.5 MPa	300°C	3009.6 $\frac{\text{kJ}}{\text{kg}}$	-
4	100 kPa	100°C	2449.2 $\frac{\text{kJ}}{\text{kg}}$	0.9
5	100 kPa	60°C	251.3 $\frac{\text{kJ}}{\text{kg}}$	-
6	7.5 MPa	60°C	257.5 $\frac{\text{kJ}}{\text{kg}}$	-

Turbine/Pump Work, Boiler/Reheat/Condenser Heat Transfer

The specific work obtained from each turbine is exactly the difference in enthalpy between the states, due to the assumption of adiabatic turbines. Thus:

$$\begin{aligned}
 w_{HPT} &= h_1 - h_2 = 2814.4 \frac{\text{kJ}}{\text{kg}} - 2617.9 \frac{\text{kJ}}{\text{kg}} = 196.5 \frac{\text{kJ}}{\text{kg}} \\
 w_{LPT} &= h_3 - h_4 = 3009.6 \frac{\text{kJ}}{\text{kg}} - 2449.2 \frac{\text{kJ}}{\text{kg}} = 560.4 \frac{\text{kJ}}{\text{kg}} \\
 w_{turb} &= w_{HPT} + w_{LPT} = 756.9 \frac{\text{kJ}}{\text{kg}}
 \end{aligned}$$

The pump can be found the same way:

$$w_{pump} = h_6 - h_5 = 257.5 \frac{\text{kJ}}{\text{kg}} - 251.3 \frac{\text{kJ}}{\text{kg}} = 6.2 \frac{\text{kJ}}{\text{kg}}$$

Alternatively, we can use Equation 4.10:

$$w_{pump} = 0.001 \frac{\text{m}^3}{\text{kg}} (7500 \text{ kPa} - 100 \text{ kPa}) = 7.4 \frac{\text{kJ}}{\text{kg}}$$

We'll use the CoolProp value of $6.2 \frac{\text{kJ}}{\text{kg}}$.

Similarly, there is no work that occurs in the boiler or reheat cycle, meaning we can again use the difference in enthalpy:

$$\begin{aligned} q_{boiler} &= h_1 - h_6 = 2814.4 \frac{\text{kJ}}{\text{kg}} - 257.5 \frac{\text{kJ}}{\text{kg}} = 2556.9 \frac{\text{kJ}}{\text{kg}} \\ q_{reheat} &= h_3 - h_2 = 3009.6 \frac{\text{kJ}}{\text{kg}} - 2617.9 \frac{\text{kJ}}{\text{kg}} = 391.7 \frac{\text{kJ}}{\text{kg}} \\ q_{in} &= w_{boiler} + w_{reheat} = 2948.6 \frac{\text{kJ}}{\text{kg}} \end{aligned}$$

Finally, we can calculate the heat transfer in the condenser, again assuming no work occurs:

$$q_{cond} = h_4 - h_5 = 2449.2 \frac{\text{kJ}}{\text{kg}} - 251.3 \frac{\text{kJ}}{\text{kg}} = 2197.9 \frac{\text{kJ}}{\text{kg}}$$

Mass Flow Rate of Steam, Efficiency, and Total Heat Transfer

Using the requirement of 450 MW of net power, we can find the mass flow rate through the equation $\dot{W}_{net} = \dot{m}w_{net}$:

$$\dot{m} = \frac{\dot{W}_{net}}{w_{turb} - w_{pump}} = \frac{450 \text{ MW}}{756.9 \frac{\text{kJ}}{\text{kg}} - 6.2 \frac{\text{kJ}}{\text{kg}}} = 599.4 \frac{\text{kg}}{\text{s}}$$

The efficiency can be found with the total work from the turbines, the work from the pump, and the total heat in from the boiler and reheat processes:

$$\eta_{th} = \frac{w_{turb} - w_{pump}}{q_{in}} = \frac{756.9 \frac{\text{kJ}}{\text{kg}} - 6.2 \frac{\text{kJ}}{\text{kg}}}{2948.6 \frac{\text{kJ}}{\text{kg}}} = 25.5\%$$

This efficiency is not great, and a good part of the reason for that is the low maximum temperature of the cycle. However, the cost of nuclear fuel tends to be substantially cheaper than the cost of coal or natural gas, meaning that in kJ/\$, this power plant may be cheaper to operate than an similar sized coal-fired or gas-fired plant.

We can find the necessary amount of heat from the reactor one of two ways. First, we could use the efficiency, which works equally for extensive or intensive values. Alternatively, we could use the mass flow rate of steam to find the extensive equivalent of q_{in} . Both will give the same result, albeit with the possibility of rounding error.

$$\dot{Q}_{in} = \dot{m}q_{in} = \left(599.4 \frac{\text{kg}}{\text{s}} \right) \left(2948.6 \frac{\text{kJ}}{\text{kg}} \right) = 1.768 \text{ GW}$$

Thus, we need our nuclear reactor to generate about 1.77 GW of heat.

4.5.2 Rankine Cycle with Regeneration

There are a number of ways to increase the efficiency of steam power plants. In general, increasing the maximum temperature is the most straightforward method. However, higher temperatures also require more heat-resistant materials, which can drastically increase production costs or simply not be available.

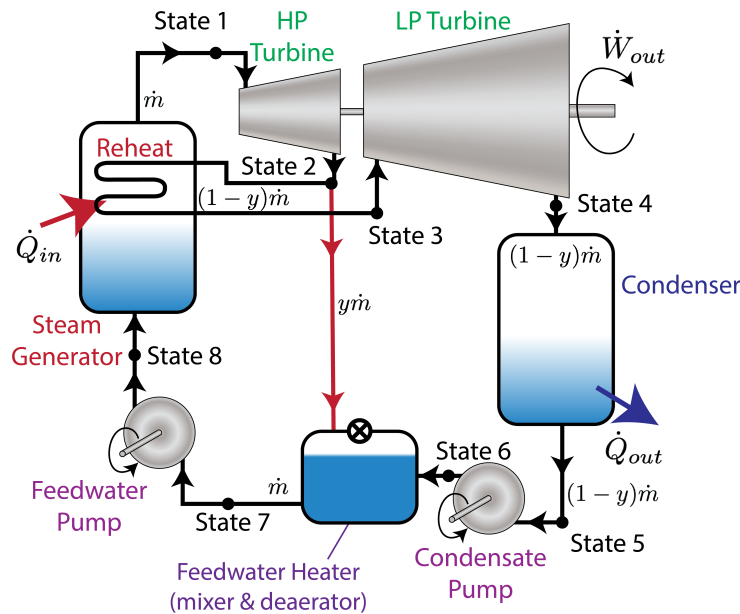
The other primary method is making use of the waste heat in one way or another. One way of doing this is through **cogeneration systems**. Cogeneration refers to the simultaneous generation of electric and thermal energy. While this is typically easier to design for industrial applications, where large sources of heat are useful, it has also been developed for residential power, where the waste heat can be used to heat water or for the heating side of air conditioning.

The following example will use waste heat (residual heat in the steam after work is extracted) to preheat the incoming cold feedwater, which allows us to use less fuel in the heating process.

Example 4.3: Supercritical Steam Power Plant with Feedwater Heater

In this example, we are again looking at a steam turbine from **General Electric**. This steam turbine is designed for “ultra-supercritical” steam, which is typically defined as steam with temperature in the range of 700°C and pressures up to 35 MPa. Specifically, our turbine will be designed for a main steam temperature of 650°C, a main steam pressure of 33 MPa, and a reheat steam temperature of 670°C. We will choose a power output of 1,000 MW for the plant.

Additionally, we will include consideration of a feedwater heater, which mixes steam from after the HP Turbine with cold water from the condenser. This mixing vessel will also serve as a de-aerator (see Section 4.3.8), though this does not significantly impact our thermodynamic analysis.

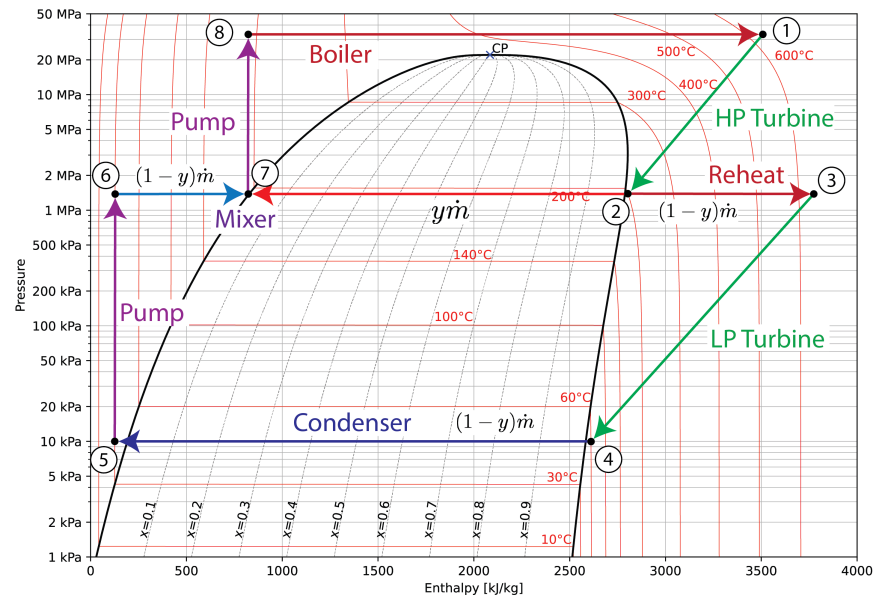


This system is referred to as a **regenerative reheat** cycle, and we will find that this simple extension of our previous system, along with the significantly higher maximum temperature, will result in an increase in thermal efficiency of the power plant. Note in particular the **mass fraction of steam**, y , that is taken directly from the outlet of the HP turbine to the feedwater heater. The remainder, $1 - y$, goes through reheat, the LP turbine, and the condenser. Also, we've changed the name of the boiler to a steam generator, since we are heating at supercritical temperatures, where boiling is not defined.

Solution Methodology

p - h Diagram

In order to build the p - h diagram, we need to define some additional parameters. First off, the outlet of the HP turbine will be at 1.5 MPa and 200°C. This also sets the pressure of the LP turbine inlet and the pressure of the feedwater heater. The outlet of the LP turbine will be at 10 kPa and 60°C. The condenser output water at 30°C, and the feedwater heater will output saturated water. This is sufficient to define all 8 states.



Note that only $(1 - y)\dot{m}$ of steam goes through States 3-6, while $y\dot{m}$ of steam goes directly from State 2 to 7. This results in less work output from the LP turbine, but also significantly less energy required from the boiler. The net effect is better thermal efficiency from the cycle as a whole. Otherwise, we will use our state definitions to find the enthalpy of each state, and analyze the cycle in a very similar way to Examples 4.1 and 4.2.

Property Table

As in Example 4.2, most properties are simple table look-ups, often with interpolation. States 5 and 6 require either the assumption of equivalence to saturated water at the same temperature, or evaluation with CoolProp. The property table below was generated solely with CoolProp. Quantities in black are taken from the problem statement, and quantities in red are calculated.

State	Pressure	Temperature	Enthalpy	Quality
1	33 MPa	650°C	3576.2 $\frac{\text{kJ}}{\text{kg}}$	-
2	1.5 MPa	200°C	2796.0 $\frac{\text{kJ}}{\text{kg}}$	-
3	1.5 MPa	670°C	3852.6 $\frac{\text{kJ}}{\text{kg}}$	-
4	10 kPa	60°C	2611.2 $\frac{\text{kJ}}{\text{kg}}$	-
5	10 kPa	30°C	125.7 $\frac{\text{kJ}}{\text{kg}}$	-
6	1.5 MPa	30°C	127.1 $\frac{\text{kJ}}{\text{kg}}$	-
7	1.5 MPa	198°C	844.6 $\frac{\text{kJ}}{\text{kg}}$	0.0
8	33 MPa	198°C	859.1 $\frac{\text{kJ}}{\text{kg}}$	-

We can check these values against the approximate locations of the States in the p - h diagram, and we see that each of the enthalpies is reasonable.

Finding y

The big unknown of the analysis at this point is the mass fraction of steam that goes directly from State 2 to State 7. In order to perform this analysis, we look at Equation 4.11:

$$\begin{aligned}
 \dot{m}_{out}h_{out} &= \sum \dot{m}_{in}h_{in} \\
 \dot{m}h_7 &= y\dot{m}h_2 + (1-y)\dot{m}h_6 \\
 h_7 - h_6 &= y(h_2 - h_6) \\
 y &= \frac{h_7 - h_6}{h_2 - h_6}
 \end{aligned}$$

Plugging in values here gives us the following:

$$y = \frac{844.6 \frac{\text{kJ}}{\text{kg}} - 127.1 \frac{\text{kJ}}{\text{kg}}}{2796.0 \frac{\text{kJ}}{\text{kg}} - 127.1 \frac{\text{kJ}}{\text{kg}}} = 0.269 = 26.9\%$$

Thus, 26.9% of the steam is diverted to the mixer, while 73.1% is reheated and sent to the LP turbine.

Turbine/Pump Work, Boiler/Reheat/Condenser Heat Transfer

The analysis here is nearly identical to that done in Example 4.2, with the exception of some change in State numbers for the pumps. Additionally, we will wait to lump together the turbines and heats until we can calculate the mass flow rates of steam through each component.

The turbines are found through the difference in enthalpies:

$$w_{HPT} = h_1 - h_2 = 3576.2 \frac{\text{kJ}}{\text{kg}} - 2796.0 \frac{\text{kJ}}{\text{kg}} = 780.2 \frac{\text{kJ}}{\text{kg}}$$

$$w_{LPT} = h_3 - h_4 = 3852.6 \frac{\text{kJ}}{\text{kg}} - 2611.2 \frac{\text{kJ}}{\text{kg}} = 1241.4 \frac{\text{kJ}}{\text{kg}}$$

The pumps can be found the same way:

$$w_{pump,5-6} = h_6 - h_5 = 127.1 \frac{\text{kJ}}{\text{kg}} - 125.7 \frac{\text{kJ}}{\text{kg}} = 1.4 \frac{\text{kJ}}{\text{kg}}$$

$$w_{pump,7-8} = h_8 - h_7 = 859.1 \frac{\text{kJ}}{\text{kg}} - 844.6 \frac{\text{kJ}}{\text{kg}} = 14.5 \frac{\text{kJ}}{\text{kg}}$$

Again, we could use Equation 4.10, though we choose to skip that analysis for this cycle.

There is no work that occurs in the boiler or reheat cycle, meaning we can again use the difference in enthalpy:

$$q_{boiler} = h_1 - h_8 = 3576.2 \frac{\text{kJ}}{\text{kg}} - 859.1 \frac{\text{kJ}}{\text{kg}} = 2717.1 \frac{\text{kJ}}{\text{kg}}$$

$$q_{reheat} = h_3 - h_2 = 3852.6 \frac{\text{kJ}}{\text{kg}} - 2796.0 \frac{\text{kJ}}{\text{kg}} = 1056.6 \frac{\text{kJ}}{\text{kg}}$$

Finally, we can calculate the heat transfer in the condenser, again assuming no work occurs:

$$q_{cond} = h_4 - h_5 = 2611.2 \frac{\text{kJ}}{\text{kg}} - 125.7 \frac{\text{kJ}}{\text{kg}} = 2485.5 \frac{\text{kJ}}{\text{kg}}$$

Mass Flow Rates of Steam, Efficiency, and Total Heat Transfer

The total mass flow rate, \dot{m} , must be sufficient to provide 1,000 MW of net power. We can calculate the net power, accounting for the mass fraction of steam that bypasses the LP Turbine, as follows:

$$\dot{W}_{net} = \dot{m} w_{HPT} + (1 - y)\dot{m} w_{LPT} - (1 - y)\dot{m} w_{pump,5-6} - \dot{m} w_{pump,7-8}$$

Thus, we are looking at the mass flow rate through each component to determine its contribution to the net work. Alternatively, we could define $\dot{W}_{net} = \dot{m} w_{net}$, and

calculate the following:

$$w_{net} = \frac{\dot{W}_{net}}{\dot{m}} = w_{HPT} + (1 - y)w_{LPT} - (1 - y)w_{pump,5-6} - w_{pump,7-8}$$

Plugging in numbers, we get the following:

$$w_{net} = 780.2 \frac{\text{kJ}}{\text{kg}} + 0.731 \left(1241.4 \frac{\text{kJ}}{\text{kg}} \right) - 0.731 \left(1.4 \frac{\text{kJ}}{\text{kg}} \right) - 14.5 \frac{\text{kJ}}{\text{kg}} = 1672.1 \frac{\text{kJ}}{\text{kg}}$$

Calculating \dot{m} is now trivial:

$$\dot{m} = \frac{\dot{W}_{net}}{w_{net}} = \frac{1,000 \text{ MW}}{1672.1 \frac{\text{kJ}}{\text{kg}}} = 598.1 \frac{\text{kg}}{\text{s}}$$

Defining $\dot{Q}_{in} = \dot{m} q_{in} = \dot{m} q_{boiler} + (1 - y)\dot{m} q_{reheat}$, we calculate the following for q_{in} :

$$q_{in} = q_{boiler} + (1 - y)q_{reheat} = 2717.1 \frac{\text{kJ}}{\text{kg}} + 0.731 \left(1056.6 \frac{\text{kJ}}{\text{kg}} \right) = 3489.5 \frac{\text{kJ}}{\text{kg}}$$

This allows us to calculate our efficiency:

$$\eta_{th} = \frac{w_{net}}{q_{in}} = \frac{1672.1 \frac{\text{kJ}}{\text{kg}}}{3489.5 \frac{\text{kJ}}{\text{kg}}} = 0.479 = 47.9\%$$

This is greatly improved from the efficiencies found in the previous examples of this chapter.

As a side-note, if we had chosen to forego the feedwater heater, we would have increased w_{net} to 2005.7 $\frac{\text{kJ}}{\text{kg}}$, due to the increased mass flow rate in the LP turbine.

However, we would have simultaneously increased q_{in} to 4491.2 $\frac{\text{kJ}}{\text{kg}}$, because of the increased heating load on the boiler, as well as the increased mass flow rate through the re-heater. Without the feedwater heater, the cycle would have had an efficiency of only 44.7% (which is still much better than the previous examples).

Finally, we will look at the total heat transfer required from the boiler and re-heater.

$$\dot{Q}_{in} = \dot{m} q_{in} = \left(598.1 \frac{\text{kg}}{\text{s}} \right) \left(3489.5 \frac{\text{kJ}}{\text{kg}} \right) = 2.087 \text{ GW}$$

Assuming perfect combustion, this requires 72 kg of high-quality (bituminous) coal a second, or 6200 tons of coal a day.

4.6 Refrigeration Cycle

4.6.1 Types of Refrigerants

Refrigeration cycles typically do not use air or steam. Instead, a great deal of chemical research is focused around finding substances that have p - h curves that are significantly more conducive to the refrigeration cycle.

In the early days of refrigeration, the two refrigerants in common use were ammonia and carbon dioxide. Both were problematic - ammonia is toxic and carbon dioxide requires extremely high pressures and pressures (roughly 100 atmospheres and 160°C).

In the last century or so, a large number of different refrigerants have been manufactured. The Heating, Refrigeration, and Air Conditioning Institute of Canada maintains a [list of refrigerants](#) with possible uses and dangers, including flammability, toxicity, and possible damage to the environment.

R12 (Phase-out started 1989) / R22 (Phase-out started 2010)

When the refrigerant R12 (dichloro-difluoro-methane) was discovered it took over as the refrigerant of choice. It is an extremely stable, non-toxic fluid, and it operates at pressures always somewhat higher than atmospheric, so that if any leakage occurred, air would not leak into the system. Thus, a technician can recharge a refrigeration system without having to apply vacuum.

Unfortunately when the refrigerant does ultimately leak, it will make its way up to the ozone layer. The ultraviolet radiation from the sun breaks up the molecule, it releases the highly active chlorine radicals, which deplete the ozone layer. Because of this environmental problem, R12 and other chloro-fluoro-carbons (CFCs) have since been banned from usage on a global scale.

For home air conditioning, most new units starting in the 1980s were using R22, which became the standard as Freon 12 was phased out. R22 has 5% of the ozone depletion potential of R12, but this is still considered too high for safe use. Currently, R22 cannot be produced or imported, and current units using R22 must rely on recycled or stockpiled materials.

R134a (Phase-out started 2021) / R410a (Phase-out starts 2025)

In order to replace R12, chlorine free R134a (tetrafluoro-ethane) was introduced for automotive air conditioners. R410a has been the most common replacement for R22 in residential HVAC systems in the last decade. Neither R134a nor R410a is known to cause ozone depletion.

Recently, however, the international scientific consensus is that Global Warming is caused by human energy related activity, and various man made substances are defined on the basis of Global Warming Potential (GWP) with reference to carbon dioxide (GWP=1).

R134a has been found to have a GWP of 1300 and, as of model year 2021, automobile air conditioning systems will be barred from using R134a as a refrigerant. Likewise, R410a has a GWP of 2088, and has a phasedown beginning in 2025.

Replacements for R134a

The option to replace R134a that has been adopted by most car manufacturers is the refrigerant R1234yf. R1234yf has a GWP of around 1 (roughly the same global warming potential as carbon dioxide), but is somewhat flammable. The risk of fire due to the refrigerant is negligible, but additional safety requirements have been set in place **by the EPA**.

Alternatively, a return to carbon dioxide (R744) is suggested. As mentioned before, this requires a much higher pressure than other refrigerants, and results in a much higher compressor temperature. There are no vehicles currently using R744, but several foreign automobile manufacturers are exploring its use.

Replacements for R410a

New residential air conditioning systems are changing from R410a to R32. Like R1234yf, R32 has a lower GWP (around 677). Ironically, R410a is a mixture that contains R32 and R125. R32 requires a slightly higher pressure than R410a, but significantly less refrigerant. While this is an improvement over R410a, it will likely be phased out in the next several decades as research into refrigerants with lower GWP continues.

4.6.2 Refrigeration

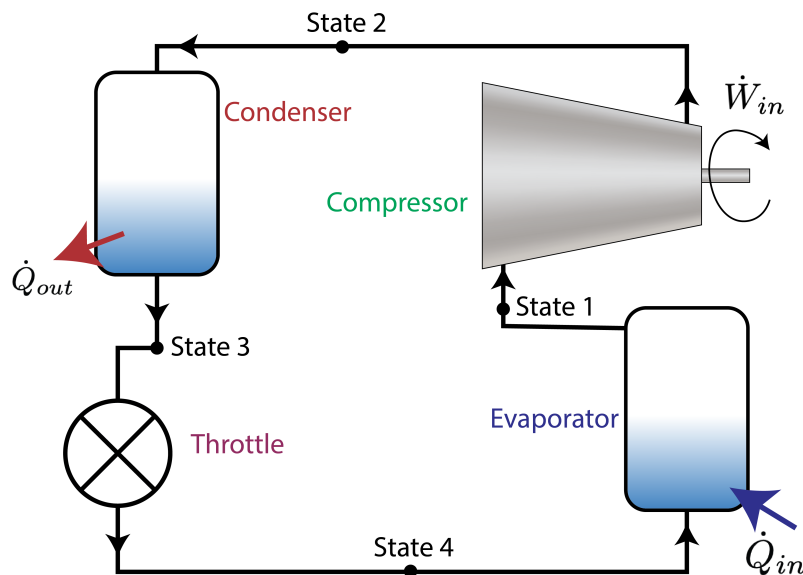


Figure 4.17: Diagram describing the refrigeration cycle.

Refrigeration cycles typically consist of a compressor, condenser, throttle, and evaporator. State 1 begins at low temperature, immediately before the compressor. Work is applied through the compressor in Process 1-2, which increases the pressure and temperature, and also provides the energy to move the fluid through the cycle (see Section 4.3.2). State 2 is the highest temperature point of the cycle. In Process 2-3, heat is extracted from the system (added to the environment) at high temperature from the condenser (see Section

4.3.4). State 3 is either saturated liquid or supercooled liquid. The throttle reduces the pressure across Process 3-4, similar to a turbine, but with no work extracted (see Section 4.3.6). State 4 is a mixture of liquid and vapor at low pressure and temperature. The evaporator in Process 4-1 adds heat to the system (extracts heat from the environment) at low temperature (see Section 4.3.3), and brings the system back to State 1.

As in the Rankine Cycle, only heat transfer occurs in the evaporator and condenser, as there are no moving parts in those components. The compressor is the only component that does work, and neither work nor heat transfer occurs in the throttle.

When considering efficiency, we look at the coefficient of performance, as originally defined in Section 3.5.

$$COP_R = \frac{\dot{Q}_{in}}{|\dot{W}_{net}|} = \frac{q_{in}}{|w_{net}|}$$

The net work for the refrigeration cycle comes solely from the compressor: $w_{net} = w_{comp}$. The heat transfer into the system (and out of the environment) comes from the evaporator, so $q_{in} = q_{evap}$.

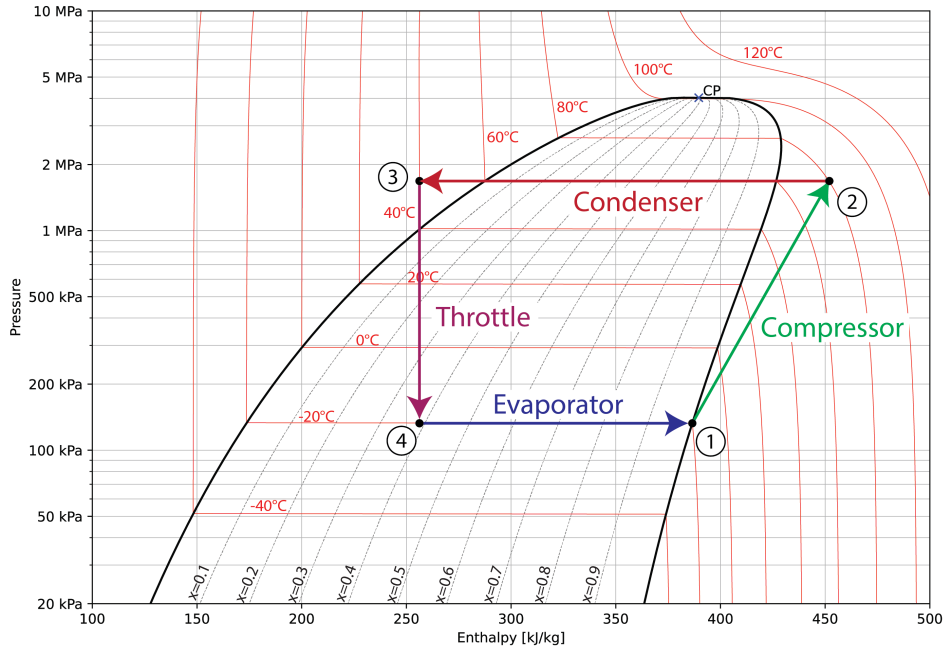


Figure 4.18: The states and processes of the refrigeration cycle shown on a p - h diagram.

Because no work occurs in the evaporation process, and no heat transfer occurs in the compressor, Equation 4.8 can again be used trivially for both components:

$$\begin{aligned} \dot{m}\Delta h_{1-2} &= \dot{Q}_{1-2} - \dot{W}_{1-2} \xrightarrow{0} w_{comp} = h_1 - h_2 \\ \dot{m}\Delta h_{4-1} &= \dot{Q}_{4-1} - \dot{W}_{1-2} \xrightarrow{0} q_{evap} = h_1 - h_4 \end{aligned}$$

Note that w_{comp} is actually negative, as h_2 will be larger than h_1 . This makes sense, as we are spending work to effect heat transfer from cold to hot.

We can rewrite the coefficient of performance using only the state enthalpies as follows:

$$COP_R = \frac{q_{evap}}{w_{comp}} = \frac{h_1 - h_4}{h_2 - h_1} \quad (4.15)$$

Note that it is possible to have a COP that is greater than 1. In fact, most residential HVAC units have COP values around 4, and the COP for AC units in cars is usually between 2 and 3.

The refrigeration cycle can also be used as a heat pump. The effect of a heat pump is still to move heat from cold to hot, but in this case our goal is to heat up an already warm environment, rather than cooling an already chilled environment. In this case, the “output” is the heat entering the environment from the condenser, q_{cond} .

In this case, we find that the coefficient of performance is as follows:

$$COP_{HP} = \frac{q_{cond}}{w_{comp}} = \frac{h_2 - h_3}{h_2 - h_1} \quad (4.16)$$

The coefficient of performance for a heat pump will always be exactly 1 greater than that for a refrigerator ($COP_{HP} = COP_R + 1$).

Example 4.4 contains an more detailed analysis of a simple refrigeration system.

Example 4.4: A Basic R134a Vapor-Compression Refrigeration System

A residential HVAC system uses R134a to transport heat out of the cooled living space and into the warmer outdoor environment. The refrigerant will be 60°C leaving the compressor (hot enough to lose heat to the environment on a hot summer day), and 5°C entering the evaporator (cold enough to take heat from the indoor space, but warm enough to avoid collecting freezing water vapor on the evaporator coils). The pressure in the condenser will be 1.2 MPa, the temperature leaving the condenser will be slightly supercooled at 40°C, and the temperature leaving the evaporator will be slightly superheated at 10°C. Using this information, do the following:

- draw each of the states and processes on the p - h diagram for R134a
- find the power of the compressor
- find the heat transfer rates of the evaporator and condenser
- calculate the coefficient of performance (COP) of the system, either as a refrigerator or as a heat pump

Solution Methodology

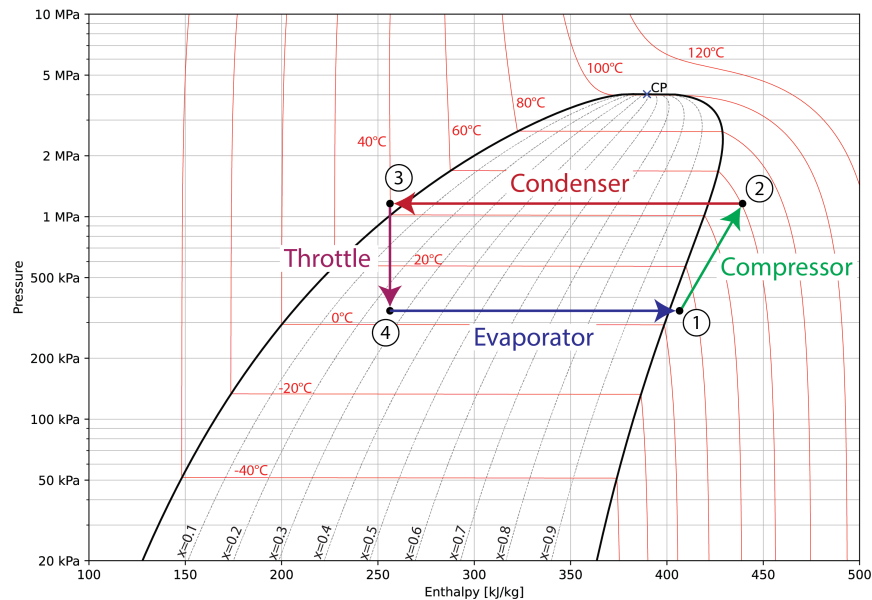
As before, we will use the information provided in order to plot the states and processes on the p - h diagram (this time for R134a). Once we have the diagram, we will find the enthalpy of each state using either CoolProp or interpolation from tables. Finally, we can find the various power and heat transfer rates, which will lead to the coefficient of performance.

p-h Diagram

In order to plot the *p-h* diagram, we do need some additional properties:

- The temperature of State 1 is defined as 10°C, but we need to find the pressure in the evaporator.
- State 2 is fully defined, with a pressure equal to the condenser pressure of 1.2 MPa, and the outlet temperature of the compressor defined as 60°C.
- State 3 is fully defined, again using the condenser pressure, but with a temperature of 40°C.
- State 4 again needs the pressure in the evaporator, though we know the temperature is 5°C.

The pressure inside the condenser is related to the temperature of state 4, as it is a saturated liquid-gas mixture. We can therefore find the pressure from Appendix B.1. Some quick interpolation gives us $p_4 = p_1 = 350$ kPa. With that, everything is defined, and we can create the diagram.



Property Table

We now know the pressure and temperature of each state, but this is actually not enough to define the enthalpy of State 4, as it is in the saturated region. Fortunately, because Process 3-4 is a throttle (which are isenthalpic), the enthalpy of State 4 will be equivalent to that of State 3. As before, CoolProp is used for all enthalpies, properties in black are taken from the problem statement, and quantities in red are calculated.

	State 1	State 2	State 3	State 4
Pressure	350 kPa	1.2 MPa	1.2 MPa	350 kPa
Temperature	10°C	60°C	40°C	5°C
Enthalpy	406.1 $\frac{\text{kJ}}{\text{kg}}$	437.8 $\frac{\text{kJ}}{\text{kg}}$	256.4 $\frac{\text{kJ}}{\text{kg}}$	256.4 $\frac{\text{kJ}}{\text{kg}}$
Quality	-	-	-	0.25

Once again, each of the values for enthalpy matches well with the locations of the States in the p - h diagram.

Compressor Work, Evaporator/Condenser Heat Transfer

Because the compressor is assumed to be adiabatic, and there are no moving parts in the evaporator or condenser, the specific work and heat transfer can be found directly through the difference in enthalpies.

The compressor sits between States 1 and 2:

$$w_{comp} = h_2 - h_1 = 437.8 \frac{\text{kJ}}{\text{kg}} - 406.1 \frac{\text{kJ}}{\text{kg}} = 31.7 \frac{\text{kJ}}{\text{kg}}$$

The condenser is between States 2 and 3, and the evaporator is between States 4 and 1:

$$q_{cond} = h_2 - h_3 = 437.8 \frac{\text{kJ}}{\text{kg}} - 256.4 \frac{\text{kJ}}{\text{kg}} = 181.4 \frac{\text{kJ}}{\text{kg}}$$

$$q_{evap} = h_1 - h_4 = 406.1 \frac{\text{kJ}}{\text{kg}} - 256.4 \frac{\text{kJ}}{\text{kg}} = 149.7 \frac{\text{kJ}}{\text{kg}}$$

There is no work or heat transfer associated with the throttle.

Coefficients of Performance

This system is designed for refrigeration of a living space, so we are primarily interested in the heat absorbed by the evaporator, q_{evap} . This means that the coefficient of performance can be calculated as follows:

$$COP_R = \frac{q_{evap}}{w_{comp}} = \frac{149.7 \frac{\text{kJ}}{\text{kg}}}{31.7 \frac{\text{kJ}}{\text{kg}}} = 4.72$$

This is significantly higher than typical coefficients of performance for residential AC, which usually run between 2 and 4. Likely, this means that our compressor is too efficient, and a real compressor for this system would require more energy in order to create the same increase in pressure.

If the goal of this refrigeration unit was instead heating up the inside of the home, and the evaporator and condenser temperatures were exactly the same (reasonable

for a chilly day, with an indoor temperature around 20°C), we would instead use the condenser heat transfer to calculate the coefficient of performance:

$$COP_{HP} = \frac{q_{evap}}{w_{comp}} = \frac{181.4 \frac{\text{kJ}}{\text{kg}}}{31.7 \frac{\text{kJ}}{\text{kg}}} = 5.72$$

This value is exactly 1 higher than COP_R , which matches our analysis of the Stirling Cooling cycle back in Section 3.5.

4.7 Summary

In this chapter, we analyzed several variations of the Rankine Cycle, including both reheat and regeneration. The Rankine Cycle is commonly used in steam power plants. Additionally, we looked at the vapor-compression refrigeration cycle, which is commonly used in air conditioning units and refrigerators.

Each of these cycles used the First Law of Thermodynamics, modified for open systems:

$$\Delta h = q - w$$

Because we used steam or refrigerant R134a for our examples, we were typically given enough information to find the enthalpies of each state, with only constant pressure processes being used to jump between states. Tables are provided to calculate these properties through linear interpolation, or CoolProp can be used to find properties directly.

Finally, we considered the efficiency of systems, which is defined as the thermal efficiency, η_{th} , for power plants, and as the coefficient of performance, COP_R or COP_{HP} for refrigerators or heat pumps:

$$\begin{aligned}\eta_{th} &= \frac{w_{net}}{q_{in}} \\ COP_R &= \frac{q_{evap}}{w_{comp}} \\ COP_{HP} &= \frac{q_{cond}}{w_{comp}}\end{aligned}$$

Chapter 4 Homework:

Question 4.1: Thinking about the energy equation, how do we get the enthalpy term? What concept gives us the additional term to change u into h ?

Question 4.2: What does the dot mean in \dot{m} , \dot{W} , and \dot{Q} ? What units do each of those terms have?

Question 4.3: For an adiabatic process, how would we represent work on a p - h diagram?

Question 4.4: Why are the constant temperature lines vertical to the left and right of the p - h diagram? Is enthalpy more dependent on temperature or pressure in those regions?

Question 4.5: Why is the constant temperature line horizontal in the central point of the diagram? Assuming constant temperature, what is the primary property that would determine enthalpy in that region?

Question 4.6: Which components add or remove energy as work from a system?

Question 4.7: Which components add or remove energy as heat from a system?

Question 4.8: Which components move heat within the system?

Question 4.9: Which components do not add or remove energy as either work or heat?

Question 4.10: What limitations are placed on pumps and turbines? (i.e. where in the p - h diagram are they allowed to operate?)

Question 4.11: What four components are present in the Rankine cycle?

Question 4.12: Why do we sometimes neglect the energy to run the pump?

Question 4.13: How do we define efficiency for a Rankine cycle? What are we “getting out” of the cycle? What are we “putting in”?

Question 4.14: What value did the reheat process add? Thinking about Example 4.2, what would have happened without re-heating? What would the quality have been at the outlet of the LP turbine? Is that acceptable for turbines?

Question 4.15: How did the efficiency equation change when we considered a reheat?

Question 4.16: What value did regeneration add in Example 4.3? Without regeneration, what happens to the heat added in the steam generator? What happens to the total work output? What happens to the efficiency?

Question 4.17: What has motivated the change in refrigerants over the years?

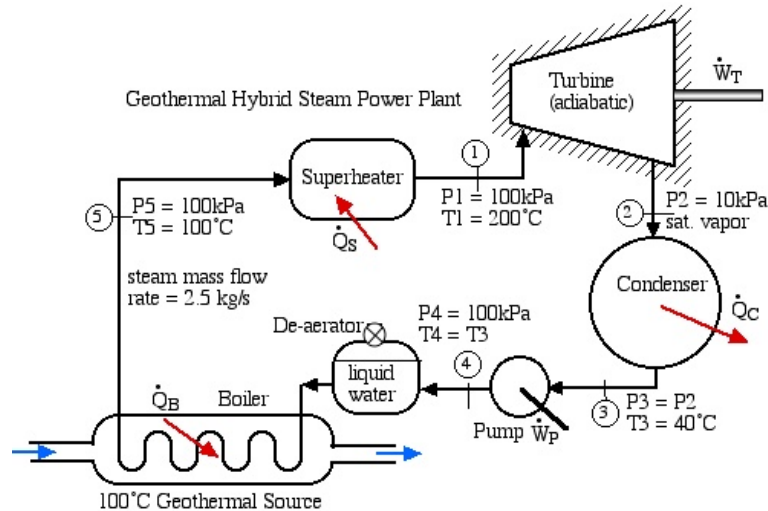
Question 4.18: What makes R1234yf preferable to R134a? What makes it worse?

Question 4.19: Why do we use a throttle instead of a turbine for a refrigeration cycle?

Question 4.20: A furnace directly burns fuel to heat homes. How does this differ from a heat pump? What about electric space heaters? For each, what are we paying for to get the end product of heat inside our house?

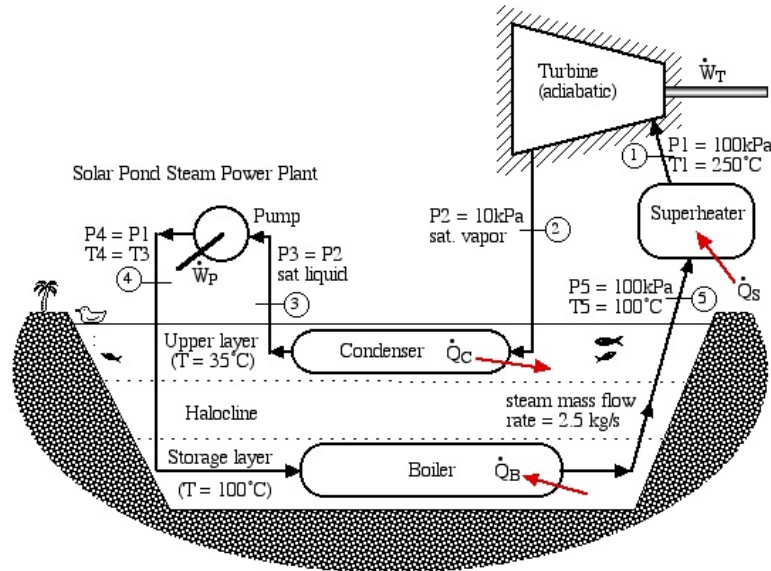
Question 4.21: Do refrigerators move heat from cold to hot? If so, how?

Question 4.22: A small community of about 500 households have discovered an underground geothermal brine source that can be used to boil water at 100°C and would like to use this to generate power. The following diagram shows the initial design of a low pressure geothermal plant in which the water is boiled by the geothermal source to 100°C and subsequently superheated to 200°C by a wood-fired superheater. Notice that the high pressure of the system is at 100kPa allowing a convenient de-aerator to be placed at the pump outlet.



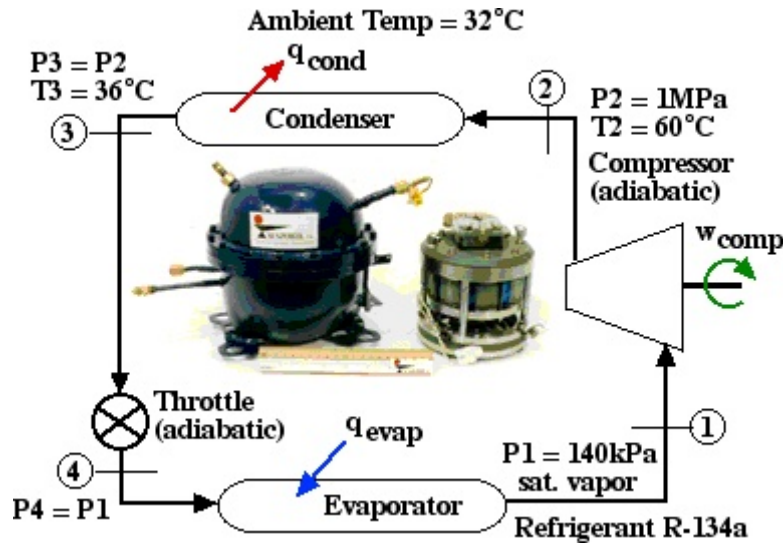
- Neatly sketch the complete cycle on the pressure-enthalpy p - h diagram for water, indicating clearly all 5 stations on the diagram.
- Assuming that the turbine is adiabatic, determine the power output of the turbine **[729kW]**.
- Assuming that the feedwater pump is adiabatic, and that the compressed liquid experiences no change in temperature while passing through the pump, determine the power required to drive the pump **[0.23kW]**.
- Using steam tables, determine the heat transferred to the boiler **[6271kW]** as well as the heat transferred to the superheater **[500kW]**.
- Determine the overall thermal efficiency η_{th} of this power plant **[11%]**. (Thermal efficiency is defined as the net work done by the system (turbine and feedwater pump) divided by the total heat supplied externally). Is this the best measure of efficiency for this power plant?
- Discuss the proposed system with respect to its environmental impact and feasibility. Is this a well designed system? What do you consider to be the major advantages and disadvantages of this system? Your discussion should include a comparison of the external fuel used and the turbine power.

Question 4.23: We wish to evaluate the proposed Solar-Pond Steam Power Plant shown in the following diagram. A solar pond is a large body of water having a varying salinity gradient (halocline) which traps the sun's energy such that the storage layer at the bottom of the pond can reach temperatures of greater than 100°C . The diagram following shows the initial design of a low pressure solar-pond steam power plant, using the storage layer as the boiler heat source, and the upper layer as the heat sink. Notice the wood-fired superheater in which the steam at the outlet of the boiler is heated from 100°C to 250°C .



- Neatly sketch the complete cycle on the pressure-enthalpy $p-h$ diagram for water, indicating clearly all 5 stations on the diagram.
- Assuming that the turbine is adiabatic, determine the power output of the turbine **[976kW]**.
- Assuming that the feedwater pump is adiabatic, and that the compressed liquid experiences no change in temperature while passing through the pump, determine the power required to drive the pump **[0.23kW]**.
- Using steam tables, determine the heat transferred to the boiler **[6210kW]** as well as the heat transferred to the superheater **[747kW]**.
- Determine the overall thermal efficiency η_{th} of this power plant **[14%]**. (Thermal efficiency is defined as the net work done by the system (turbine and feedwater pump) divided by the total heat supplied externally). Is this the best measure of efficiency for this power plant?
- Discuss the proposed system with respect to its environmental impact and feasibility. Is this a well designed system? What do you consider to be the major advantages and disadvantages of this system? Your discussion should include a comparison of the external fuel used and the turbine power.

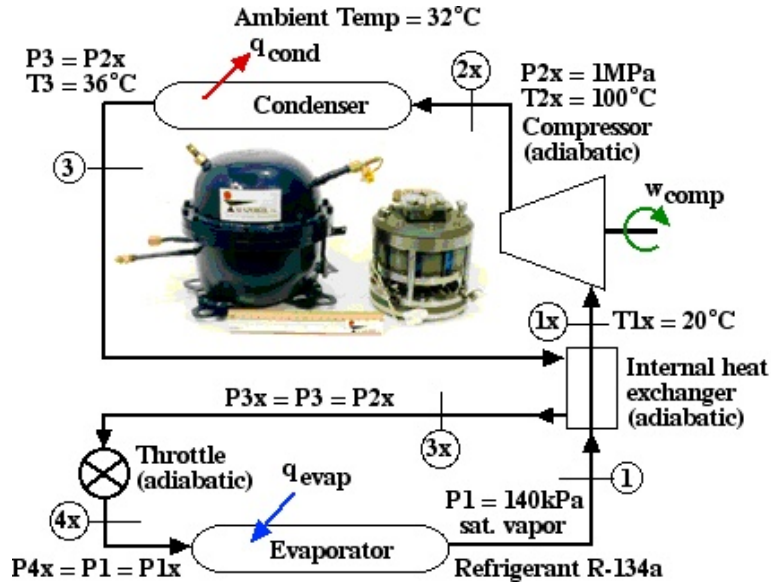
Question 4.25: We wish to do a preliminary thermodynamic evaluation of a refrigeration system designed for home usage which will use refrigerant R134a. Consider the following system flow diagram:



- Neatly sketch the complete cycle on the pressure-enthalpy p - h diagram for R134a, indicating clearly all 4 stations on the diagram.
- Determine the work done by the compressor [54 kJ/kg].
- Determine the heat absorbed by the evaporator [137 kJ/kg], and that rejected by the condenser [191 kJ/kg].
- Determine the Coefficient of Performance of the refrigerator (COP_R) (defined as the heat absorbed in the evaporator divided by the work done on the compressor - always presented as a positive value even though the work done w_c is negative) [$COP_R = 2.53$].

Question 4.26: It is common practice in the refrigeration industry to use an internal heat exchanger to subcool the refrigerant at the outlet of the condenser by means of the refrigerant exiting the evaporator. This practice obtains a much larger refrigeration capacity using the same components.

Continuing the previous problem, we will add a heat exchanger as indicated.



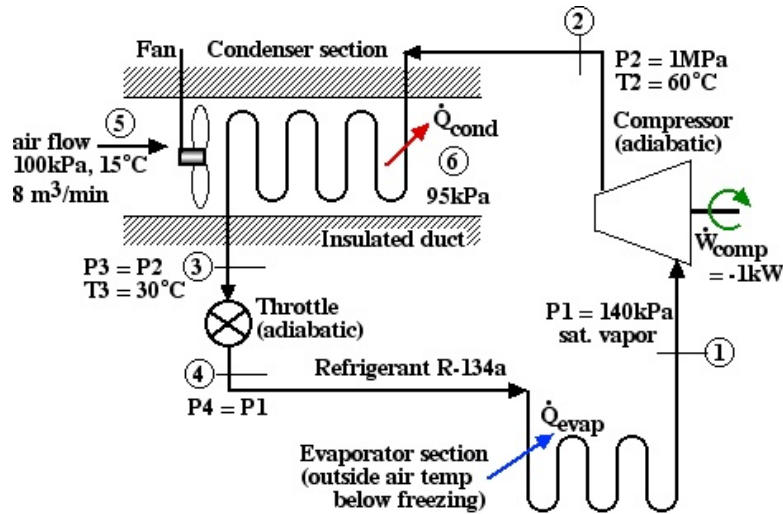
Notice that we have included an internal heat exchanger that heats the refrigerant exiting the evaporator (as a saturated vapor at 140kPa) to 20°C . We have chosen a state numbering system (1x, 2x, and so on) so as to allow the new system to be plotted on the same p - h diagram as above, and thus to be able to qualitatively compare the increase and improvement of performance provided by adding the internal heat exchanger.

- Add the 4 states (1x, 2x, 3x, 4x) to the pressure-enthalpy p - h diagram from the previous problem and sketch the new cycle.
- Determine the heat transferred in the internal heat exchanger, assuming it to be externally adiabatic [32.2 kJ/kg], and the temperature of the subcooled liquid entering the throttle (3x) [13.4°C]. The heat transferred can be calculated through the equation:

$$q = h_{1x} - h_1 = h_3 - h_{3x}$$

- Determine the heat absorbed by the evaporator [169 kJ/kg].
- Determine the heat rejected by the condenser [235 kJ/kg].
- Determine the Coefficient of Performance of the refrigerator (COP_R) (defined as the heat absorbed in the evaporator divided by the work done on the compressor) [$COP_R = 2.65$].

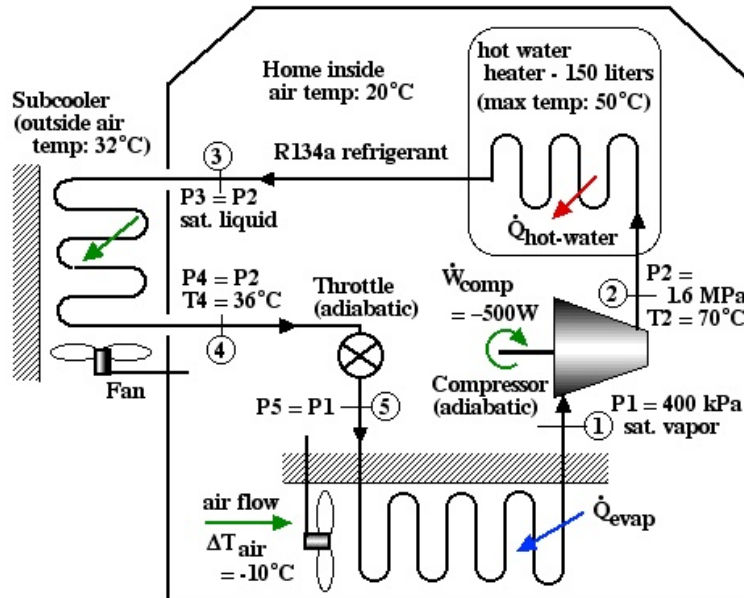
Question 4.27: We wish to do a preliminary thermodynamic evaluation of a 1kW input power home heat pump system for space heating using refrigerant R134a. Consider the following system flow diagram:



Thus the heat pump system absorbs heat from the evaporator placed outside in order to pump heat into the air flowing through the insulated duct over the condenser section. The fan provides an air flow of $8 \text{ m}^3/\text{min}$, which is enough to cool the refrigerant in the condenser to 30°C . In this analysis we will neglect the power provided to the fan. We also assume that the duct is adiabatic, and that all the heat rejected by the condenser is absorbed by the air flowing in the duct.

- Neatly sketch the complete cycle on the pressure-enthalpy p - h diagram for R134a, indicating clearly all 4 stations on the diagram (stations 5 and 6 are air, and are therefore not on the same p - h diagram).
- Determine the mass flow rate of the refrigerant R134a $[0.0185\text{kg/s}]$.
- Determine the mass flow rate of the air flowing in the insulated duct $[0.161\text{kg/s}]$.
- Determine the heat rejected by the condenser $[3.7\text{kW}]$. Assuming that all this heat is absorbed by the air, determine the exit temperature of the air at station (6) $[37.9^\circ\text{C}]$. Is this value reasonable? Why?
- Determine the heat absorbed by the evaporator $[2.7\text{kW}]$.
- Determine the Coefficient of Performance of the heat pump (COP_{HP}) (defined as the heat rejected by the condenser divided by the work done on the compressor) $[3.7]$.

Question 4.28: We wish to do a preliminary thermodynamic evaluation of a 500W input power home heat pump system as applied to summertime use for both hot water heating to 50°C, and space cooling (air conditioning), and thus maintain the inside home temperature at a comfortable 20°C.



This unique combined air conditioning / hot water heating system is designed to absorb heat from the air flowing through the insulated duct in order to pump heat into the hot water heating tank. The fan provides enough air flow over the evaporator to cool the air by 10°C as it passes through the duct, and the hot water is heated to a maximum of 50°C. In this analysis we neglect the power provided to the fan. We also assume that both the duct and the hot water tank are externally adiabatic.

- Neatly sketch the complete cycle on the pressure-enthalpy p - h diagram for R134a, indicating clearly all 5 stations on the diagram.
- Determine the enthalpy values at all five stations [kJ/kg], and indicate these values on the p - h diagram.
- Determine the mass flow rate of the refrigerant R134a [0.0133 kg/s].
- Determine the heat rejected by the condenser [2.09 kW]. Assuming that all this heat is absorbed by the water in the hot water tank, determine the time taken for 150 liters of water at 30°C to reach the required temperature of 50°C [1 hr 40 min].
- Determine the heat power absorbed by the refrigerant in the evaporator [2.04 kW]. Assuming that all this heat is absorbed from the air in the duct and neglecting the fan power, determine the required mass flow rate of the in order reduce the air temperature by 10°C while passing through the duct [0.204 kg/s].

- f) Determine the Coefficient of Performance of the hot water heater (COP_{HW}) (defined as the heat rejected by the condenser divided by the work done on the compressor) [$COP_{HW} = 4.17$].
- g) Determine the Coefficient of Performance of the air conditioner (COP_{AC}) (defined as the heat absorbed by the evaporator divided by the work done on the compressor) [$COP_{AC} = 4.07$].
- h) If we bypass the outside subcooler (State (4) becomes saturated liquid as in State (3)) determine the change in Coefficient of Performance of the air conditioner evaluated above. Indicate this change on the $p-h$ diagram and discuss the relevance of the outside subcooling section in this system. [COP_{AC} reduced to 3.17].

Chapter 5

The Second Law and Entropy

In this chapter, we discuss the Second Law of Thermodynamics, and how it ties to entropy, which is the final property we will discuss in this book. The Second Law helps us determine which processes are possible, and does so through entropy balance. We will start with a brief discussion of entropy.

5.1 Entropy

Entropy is a property that is typically associated with the disorder of a system. In other words, when a system is perfectly ordered (crystalline structure, with no atoms out of place), entropy will be at a minimum. Disorder for a system means that it is hard to tell where an atom or molecule is at any particular point in time. Figure 5.1 shows the relative level of disorder for solids, liquids, and gases.

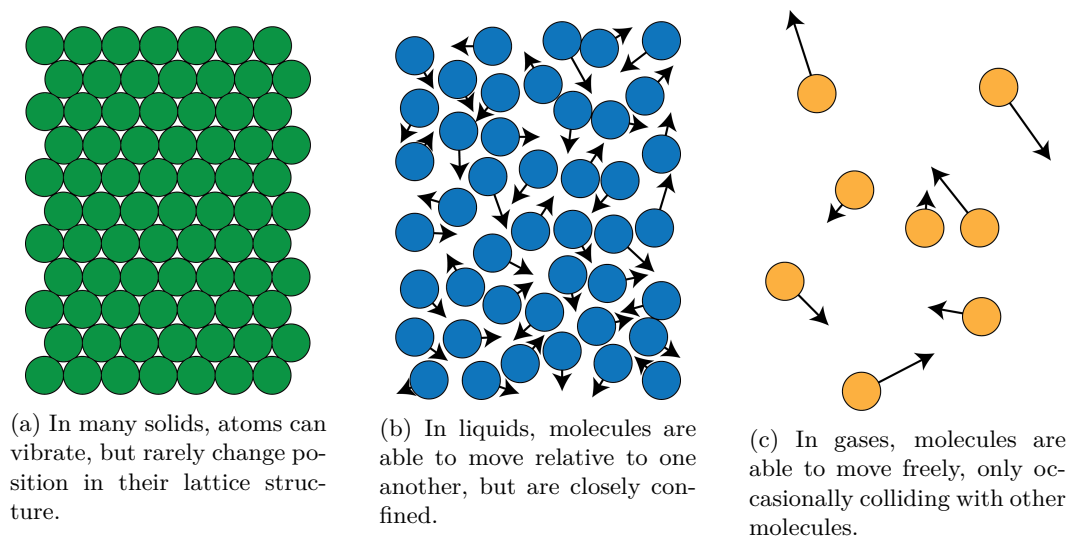


Figure 5.1: Entropy is related to the amount of disorder in a system, which varies significantly between solids, liquids, and gases.

For solids, atoms are mostly constrained, though there will be some vibration as temperature increases. This means that there are relatively few options for an atoms location, and thus, minimal disorder. For liquids, molecules are able to move somewhat freely, but are generally packed closely together with other molecules, which leads to significantly more entropy than for solids. For gases, molecules are even more free to move than in liquids, with only occasional collisions with other molecules. Thus, we assume that gases will have more disorder (and more entropy) than either liquids or solids.

5.1.1 Entropy as a Property

In addition to the difference between phases, we also know that temperature leads to an increase of the average movement of molecules within a substance. An increase in temperature, therefore, can be assumed to increase the level of entropy.

Likewise, more volume available for a substance to spread out in leads to more locations for a given molecule to be at a given time. So larger specific volume is also associated with higher entropy. Figure 5.2 is a contour plot of entropy on the T - v diagram for water.

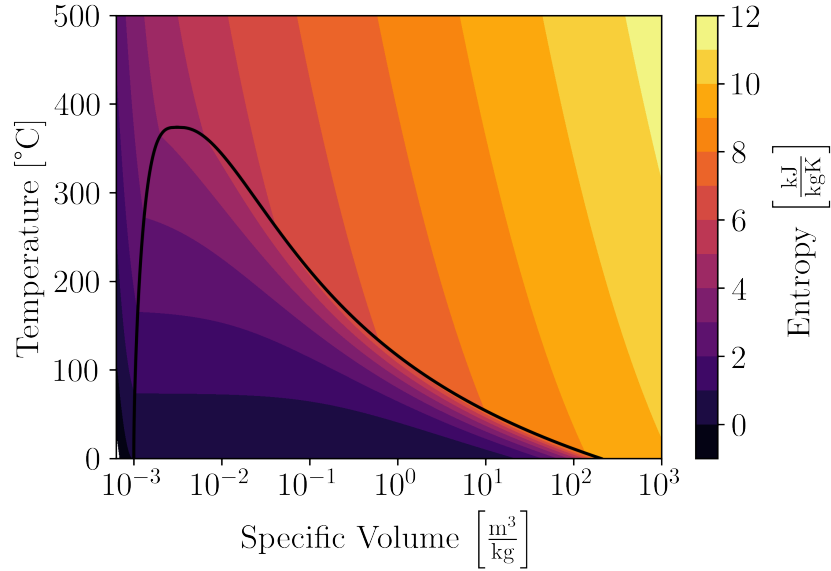


Figure 5.2: A contour plot of entropy overlaid on the vapor curve in the T - v diagram.

Finally, thinking of the ideal gas law, higher pressure will lead to lower specific volume for a given temperature, so we typically associate higher pressure with lower entropy.

5.1.2 Entropy in Statistical Mechanics

There is a way to determine the entropy by considering the number of **microstates** that could make up a given state. A microstate is a specific arrangement of molecules, including their position, velocity, vibration, and rotation. States with more entropy could be generated through a greater number of microstates. Entropy is defined as follows:

$$S = k_B \log \Omega \quad (5.1)$$

In Equation 5.1, S is the entropy of the system in $\frac{\text{kJ}}{\text{K}}$, k_B is the Boltzmann constant: $1.381 \times 10^{-26} \frac{\text{kJ}}{\text{K}}$, and Ω is the number of microstates for a given volume, number of atoms, and internal energy of the state.

Note that if only a single arrangement of molecules could define a state, the entropy would be exactly zero. This occurs for molecules in crystalline solids (restricting their position) with zero energy (restricting their movement). However, for common use, a reference value is typically chosen to be a zero point. In this book, for instance, the triple point of water is given both an internal energy and entropy of zero, even though it could be cooled below that point. In other words, if you see a negative entropy value, it means that it is less than a reference point, not negative in a true sense.

5.2 Reversibility

Entropy plays a central role as we consider **ideal** processes. Specifically, entropy is **generated** any time we have a non-ideal process. Non-ideal processes occur primarily because of friction, but can also occur when we violate the **quasi-equilibrium** assumption that we discussed back in Section 1.8.

5.2.1 Friction

Consider a block being pushed up a hill. In the idealized case (no friction between the block and the hill), all of the work done is transferred to potential energy. We could theoretically return down the hill, letting the block push us instead, and the net amount of work would be exactly zero. This situation, where we can return to an initial state with no net loss is a **reversible** process.

In a more real scenario, we could consider friction between the block and the hill. In this situation, some fraction of our work will be used to overcome friction. The potential energy at the top of the hill will be less than the amount of work we used to push the block. If we tried to extract our work this time, we would start at a deficit, and then lose even more as friction stole energy on the way down! This more realistic situation is an example of an **irreversible** process.

Thinking about the First Law of Thermodynamics, it may seem like the irreversible process above violates the conservation of energy. However, friction takes mechanical energy and transforms it into thermal energy. After pushing the block and returning to the bottom, we have lost some amount of work, but the block has gained an equivalent amount of heat.

5.2.2 Non-equilibrium Processes

Another possibility of wasting work can occur when we perform a process too quickly. Consider the piston undergoing compression shown in Figure 5.3. Pressure does not change instantaneously. In fact, any movement of the piston will create a wave in the volume which moves at the speed of sound.

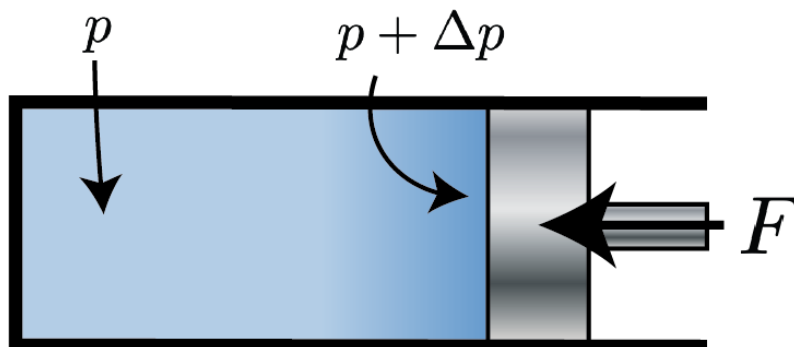


Figure 5.3: A piston undergoing high-speed compression experiences higher pressure close to the piston.

If the piston continues to compress before the pressure wave travels through the volume, it will push on air that has a higher pressure than the average within the volume. Higher pressure means more work, but eventually, the pressure will even out. In the end, there will be no additional compression because of the higher work. Instead, that additional energy will be transformed into heat.

Likewise, in an expansion process, if the pressure is not given a chance to equilibrate, it will be lower near the piston, meaning that less work will be extracted than expected. Once again, the energy that was not extracted will be converted into heat.

Thus, in both types of irreversibility, we see that the final effect is to lose some amount of mechanical energy (work), with a commensurate increase in the internal energy (heat) of our system.

5.2.3 Reversibility and Entropy

We stated above that entropy was generated for non-ideal, or irreversible processes. Based off the generation of heat for both types of irreversible processes above, we can guess that there will also be a change of entropy whenever heat is transferred in or out of a system.

However, when a process is reversible and no heat is transferred, there will be no change in entropy. Recall our discussion of adiabatic processes in Section 3.7, in which we teased the term **isentropic relations**. In fact, there is an additional requirement to use those equations: the processes must be reversible. Any friction or non-equilibrium effects will lead to conversion of work energy to heat energy, breaking the validity of the equations.

For isentropic (adiabatic and reversible) processes in steam or R134a, we can instead use the steam tables or CoolProp, assuming that $\Delta s = 0$.

For adiabatic and irreversible processes, we know that entropy will be generated. This means that $\Delta s > 0$. There cannot be an adiabatic process in which entropy is reduced. In order to reduce entropy within a system, we must transfer it out of the system. That will be the subject of Section 5.3.

One final note: there is no real process that is also reversible. In fact, we often use the word **ideal** to mean reversible and **real** to mean irreversible.

5.3 Entropy Transfer

We saw in Section 5.2 that entropy was related to heat energy. In fact, heat transfer is the only way to transfer entropy into or out of a system.

5.3.1 Entropy Transfer in Statistical Mechanics

From a molecular point of view, the only “true” properties are the internal energy (how fast are the molecules moving, vibrating, or spinning on average), entropy (how many microstates are possible for a given macrostate), and specific volume (how much volume do these molecules inhabit). With this context, we can return to an analysis similar to Section 3.3:

$$\begin{aligned} u &= u(s, v) \\ du &= \frac{\partial u}{\partial s} ds + \frac{\partial u}{\partial v} dv \\ du &= \left[\frac{du}{ds} \right]_v ds + \left[\frac{du}{dv} \right]_s dv \end{aligned}$$

Let us first think of an isentropic situation, and set $ds = 0$. In this case, we get:

$$du = \left[\frac{du}{dv} \right]_s dv$$

Now we are left to ask what sorts of situations can change the internal energy without changing entropy. Because entropy is so closely related to heat transfer, we will assume that any heat transfer will cause a change in entropy. Energy transfer through work is the only other option. Knowing that $\delta w = p dv$ and $du = \delta q - \delta w$ leads us to the following realization:

$$\left[\frac{du}{dv} \right]_s = -p \quad (5.2)$$

Pressure emerges as a property that defines how the internal energy changes as volume changes.

Now, for constant volume, we know that heat transfer is the only way to affect the internal energy.

$$du = \delta q = \left[\frac{du}{ds} \right]_v ds$$

Although we will not prove it here, this derivative is actually temperature:

$$\left[\frac{du}{ds} \right]_v = T \quad (5.3)$$

Like pressure, temperature emerges out of this molecular understanding of matter. The end result of this exercise is a new form of the First Law of Thermodynamics, based solely on properties and their changes:

$$du = T ds - p dv \quad (5.4)$$

Finally, let's define the mechanism of entropy transfer:

$$ds = \frac{\partial q}{T} \quad (5.5)$$

Equation 5.5 states that heat transfer naturally yields a change in entropy, with heat transfer into a system causing an increase of entropy. Additionally, more heat transfer is required at higher temperature to cause a change in entropy. At low temperatures, less heat transfer is required to result in a large change in entropy.

5.3.2 Defining Entropy for an Ideal Gas

For ideal gases, we can go further in our analysis. Taking the ideal gas law ($pv = RT$) in conjunction with Equation 5.4 allow us to write:

$$ds = \frac{du}{T} + \frac{pdv}{T} \rightarrow ds = c_v \frac{dT}{T} + R \frac{dv}{v} \quad (5.6)$$

Integrating (and assuming that c_v is constant) results in:

$$\begin{aligned} \int ds &= c_v \int \frac{dT}{T} + R \int \frac{dv}{v} \\ s_2 - s_1 &= c_v \ln \left(\frac{T_2}{T_1} \right) + R \ln \left(\frac{v_2}{v_1} \right) \end{aligned} \quad (5.7)$$

Alternatively, we can look back at the definition of enthalpy in Equation 3.9 and replace du in Equation 5.4 as follows:

$$dh - p dv - v dp = T ds - p dv$$

Following the same steps that led to Equation 5.7:

$$\begin{aligned} ds &= \frac{dh}{T} - \frac{vdp}{T} \rightarrow ds = c_p \frac{dT}{T} + R \frac{dv}{v} \\ \int ds &= c_p \int \frac{dT}{T} - R \int \frac{dp}{p} \\ s_2 - s_1 &= c_p \ln \left(\frac{T_2}{T_1} \right) - R \ln \left(\frac{p_2}{p_1} \right) \end{aligned} \quad (5.8)$$

Equations 5.7 and 5.8 describe the change of entropy in an ideal gas.

Change of Entropy in an Ideal Gas

$$s_2 - s_1 = c_v \ln \left(\frac{T_2}{T_1} \right) + R \ln \left(\frac{v_2}{v_1} \right) = c_p \ln \left(\frac{T_2}{T_1} \right) - R \ln \left(\frac{p_2}{p_1} \right)$$

These equations can also be used to produce the adiabatic relations we derived in Section 3.7 by setting the change in entropy ($s_2 - s_1$) equal to zero.

5.3.3 Entropy for Liquids and Solids

For liquids and solids, we can again use Equation 5.4, recognizing that dv is zero for incompressible materials:

$$du = T ds - p \overset{0}{dv} \rightarrow ds = \frac{du}{T}$$

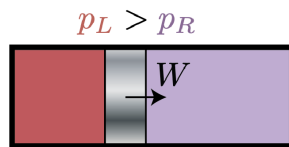
Assuming that the specific heat is a constant c , we can integrate as follows:

$$\int ds = c \int \frac{dT}{T} \rightarrow s_2 - s_1 = c \ln \left(\frac{T_2}{T_1} \right) \quad (5.9)$$

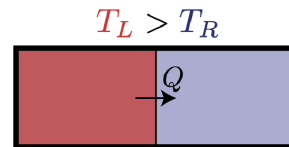
Therefore, entropy is defined purely by the temperature change for solids and liquids.

5.3.4 Additional Thoughts on Pressure and Temperature

Consider a piston between two cavities, as shown in Figure 5.4a. Each cavity is at a different pressure, the net effect of which is to cause a force on the piston. Assuming a frictionless environment, the piston will move until the pressure is balanced on both sides.



(a) The difference in pressure will cause work on the piston until pressure is balanced between the two cavities.



(b) The difference in temperature will cause heat transfer until temperature is balanced between the two cavities.

Figure 5.4: Work and heat transfer cause pressure and temperature to tend toward balance.

Likewise, looking at Figure 5.4b, the difference in temperature across the two cavities leads to heat transfer across the boundary. Heat will continue to flow until the temperature is balanced on both sides. This process will transfer entropy from left to right, and also increase the total amount of entropy. This process is investigated in Example 5.1.

Example 5.1: Maximizing Entropy Through Heat Transfer

A thin barrier separates equal masses of water such that each side has a volume of 10 L and a mass of 0.1 kg. The walls are insulated, meaning that no heat can escape. On the left, the steam is at 300°C, and on the right, the steam is at 200°C. Find the resting state of the steam, and evaluate the amount of entropy change in the system.

- Find the entropy on each side prior to heat transfer, as well as the total amount of entropy in the system.
- Find the temperature of the final system.
- Find the entropy on each side after heat transfer, as well as the total amount of entropy in the system.
- How much did entropy change on each side? How much did the total entropy change?

Entropy Before Heat Transfer

The entropy prior to heat transfer can be determined from the steam tables. Each side has the same specific volume, calculated as follows:

$$v = \frac{V}{m} = \frac{0.01 \text{ m}^3}{0.1 \text{ kg}} = 0.1 \frac{\text{m}^3}{\text{kg}}$$

Thus, the left side is superheated steam, with $s_{left} = 6.651 \frac{\text{kJ}}{\text{kg K}}$. The right side is a mixture with a quality of 0.784, leading to an entropy of $s_{left} = 5.545 \frac{\text{kJ}}{\text{kg K}}$.

We can find the total amount of entropy as follows:

$$S_i = \sum_j m_j s_j = 0.1 \text{ kg} \left(6.651 \frac{\text{kJ}}{\text{kg K}} \right) + 0.1 \text{ kg} \left(5.545 \frac{\text{kJ}}{\text{kg K}} \right) = 1.220 \frac{\text{kJ}}{\text{K}}$$

Final Temperature

At the end of heat transfer, both states should have the same temperature. Since both also have the same specific volume, they will share the same state exactly.

This means that we can calculate the total internal energy of the system, divide it up between the two sides, and we will have our final state.

$$U = \sum_j m_j u_j = 0.1 \text{ kg} \left(2762.8 \frac{\text{kJ}}{\text{kg}} \right) + 0.1 \text{ kg} \left(2217.8 \frac{\text{kJ}}{\text{kg K}} \right) = 498.06 \text{ kJ}$$

The final internal energy will be $u_f = U / \sum_i m_i$, or $498.06 \text{ kJ} / 0.2 \text{ kg}$. Thus, $u_f = 2490.3 \frac{\text{kJ}}{\text{kg}}$.

We again go to the tables (or CoolProp) with u and v to find temperature:

$T_f = 208.86^\circ\text{C}$. This occurs as a mixture with a quality of 0.937.

Final Entropy

We head to the tables one final time to find the final entropy: $s_f = 6.116 \frac{\text{kJ}}{\text{kg K}}$.

Since this is the same for both sides, we can find the total entropy through $S_f =$

$0.2 \text{ kg} s_f$. Thus, $S_f = 1.223 \frac{\text{kJ}}{\text{K}}$.

Change in Entropy

We can see that the specific entropy on the left side decreased from $6.651 \frac{\text{kJ}}{\text{kg K}}$ to $6.116 \frac{\text{kJ}}{\text{kg K}}$, while the specific entropy on the right side increased from $5.545 \frac{\text{kJ}}{\text{kg K}}$ to

$$6.116 \frac{\text{kJ}}{\text{K}}.$$

$$\Delta s_L = 6.116 \frac{\text{kJ}}{\text{kg K}} - 6.651 \frac{\text{kJ}}{\text{kg K}} = -0.535 \frac{\text{kJ}}{\text{kg K}}$$

$$\Delta s_L = 6.116 \frac{\text{kJ}}{\text{kg K}} - 5.545 \frac{\text{kJ}}{\text{kg K}} = 0.571 \frac{\text{kJ}}{\text{kg K}}$$

$$\Delta S = 1.223 \frac{\text{kJ}}{\text{K}} - 1.220 \frac{\text{kJ}}{\text{K}} = 0.003 \frac{\text{kJ}}{\text{K}}$$

Working through the example more accurately gives a change of entropy of $3.474 \frac{\text{J}}{\text{K}}$.

Discussion

First off, through heat transfer, we were able to transfer entropy as well. The entropy of the left side decreased, while the entropy of the right side increased. However, the act of transferring heat also increased the total entropy of the system. Thus, unlike energy, it is possible to **generate** entropy. Additionally, the entropy of an isolated system (like the one in the example) will always be maximized when all components are at the same temperature.

Example 5.1 showed that heat transfer naturally led to entropy transfer between two systems. Section 5.4 investigates the byproduct of that heat transfer: the entropy generated by an irreversible process.

5.4 Defining the Second Law

The simplest form of the Second Law is as follows:

The Second Law

There can be no process or cycle in which entropy decreases globally.

We saw the possibility of an isentropic process in Section 5.2.3. However, for irreversible processes without heat transfer, entropy increased. In Section 5.3, we saw the possibility of **locally** decreasing entropy, but always at the cost of increasing it elsewhere. For any realistic heat transfer, there will always be a **global** increase of entropy in the process.

Global, in this context, means considering the system that includes the surroundings. For a simplified problem, that often just means including the hot and cold **reservoirs** attached to the system. Reservoirs are regions that have enough heat capacity to maintain a constant temperature even with the heat transfer associated with our system.

In addition to this simple form, we will investigate two other statements of the Second Law: the Clausius Statement and the Kelvin-Planck Statement.

5.4.1 The Clausius Statement of the Second Law

The Clausius Statement

It is impossible to construct a device which operates on a cycle and produces no other effect than the transfer of heat from a cooler body to a hotter body.

To unpack that, let's first consider the “transfer of heat” part. We saw in Example 5.1 that the natural flow of hot to cold increases global entropy. It stands to reason that heat flow from cold to hot would instead decrease entropy, which violates the simple form of the First Law.

It may be obvious that heat transfer will not naturally happen in this manner, but the Clausius Statement forbids any device from completing this feat. This situation is pictured on the left of Figure 5.5.

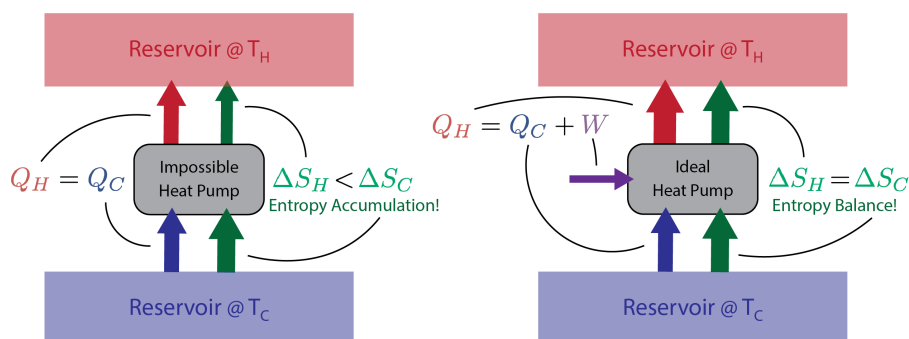


Figure 5.5: A visual depiction of a device outlawed by the Clausius statement of the Second Law beside an ideal device.

It can be seen that in the case where $Q_C = Q_H$, there is more entropy entering the heat pump than is leaving. This is a direct result of Equation 5.5: $ds = \frac{\delta q}{T}$. In order to balance the entropy transferred through Q_C and Q_H , more heat transfer is needed at T_H than at T_C . To get that extra energy without changing the inflow of entropy, we must add work to the heat pump. The corrected situation is shown on the right side of Figure 5.5.

There is an exception implicitly mentioned in the Clausius Statement: devices that do not “operate on a cycle”. It is technically possible to remove heat from a cold body and add heat to a hot body without violating the Second Law. However, the only way this is possible is by increasing the overall entropy inside the device. If the entropy inside the device increased, there is no way for it to return to its original state **without some other effect**, and therefore it cannot form a cycle.

Bottom Line: a cycle requires an influx of energy as work to move heat from cold to hot.

5.4.2 Kelvin-Planck Statement of the Second Law

There is an alternative statement of the Second Law, known as the **Kelvin-Planck Statement**:

The Kelvin-Planck Statement

It is impossible to construct a device which operates on a cycle and produces no other effect than the transfer of heat from a single body in order to produce work.

Thinking only about the First Law of Thermodynamics (energy conservation), there is nothing in the Kelvin-Planck Statement that seems impossible. Energy enters a system as heat and leaves as work, thus maintaining a constant amount of energy in the system.

Unfortunately, when we consider entropy flow, we run into trouble. Heat transfer into the system will always bring entropy with it. Work leaving the system will either generate entropy (irreversible processes) or maintain a constant amount of it (reversible processes). Combining the two, we know that there will be a net gain of entropy in the system. Once again, with a net gain of entropy, we cannot form a cycle.

The left side of Figure 5.6 shows a device breaking the Kelvin-Planck statement by absorbing heat from a heat reservoir at T_H and transforming that heat into work. Note that entropy is also absorbed with the heat, but there is nowhere that entropy is rejected.

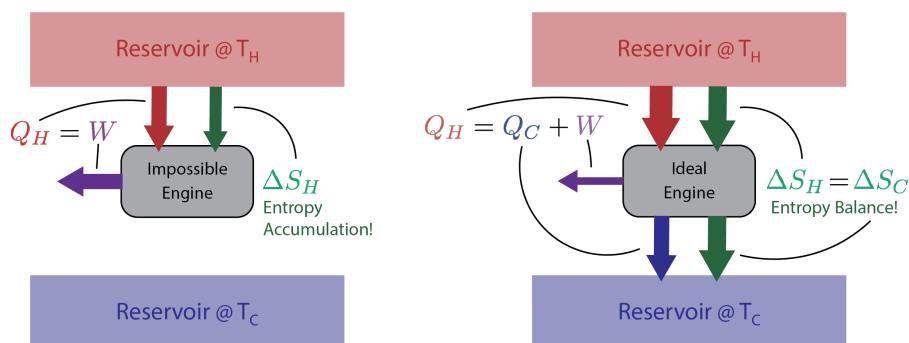


Figure 5.6: A visual depiction of a scenario outlawed by the Kelvin-Planck statement of the Second Law beside an ideal device.

The right side of Figure 5.6 give the corrected picture. In order to rid itself of excess entropy, the device must reject heat.

Bottom line: for a cycle to produce work, it must both accept and reject heat.

5.4.3 Equivalence of Kelvin-Planck and Clausius Statements

We showed above that the simple form of the Second Law implied both the Kelvin-Planck and Clausius Statements. It should be no surprise, then, that we are able to show that both statements are equivalent.

First, let us consider the combination of cycles shown on the left side of 5.7. First, we have a heat pump that violates the Clausius Statement, which we have labeled an “impossible heat pump”. Simultaneously, we run a perfectly legitimate engine. The important thing to note here is that the heat transfer through the impossible heat pump perfectly matches the heat flow through the engine (the part not turned into work).

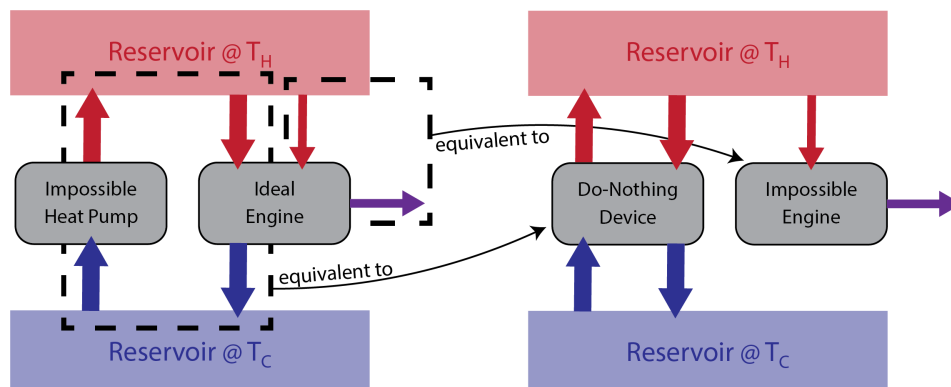


Figure 5.7: A heat pump in violation of the Clausius Statement paired with a perfectly legal heat engine.

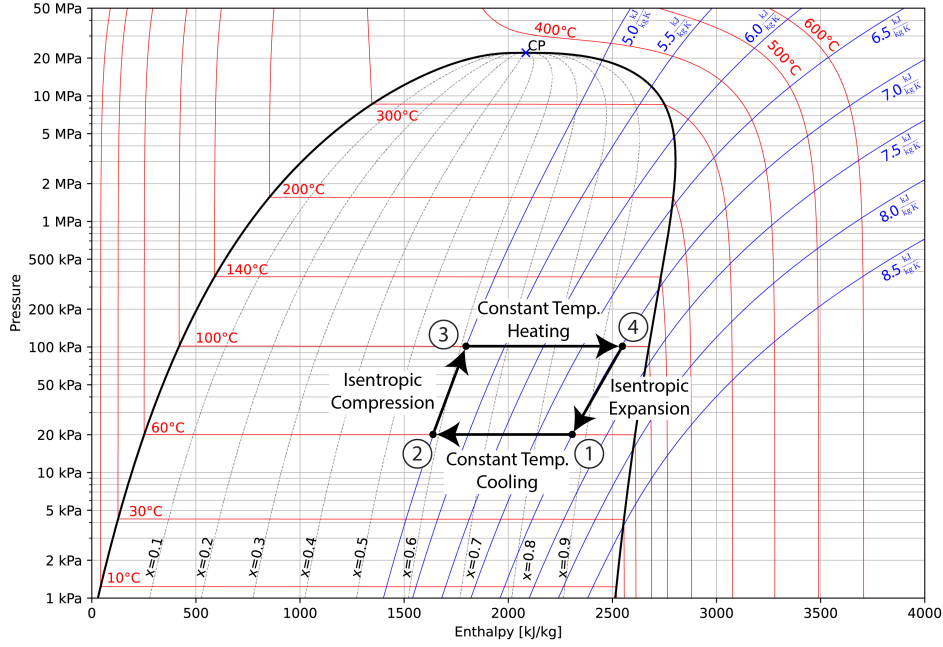
A different grouping of the heat flows allows us to see the violation of the Kelvin-Planck Statement. Combining the illegal heat flow with the heat flow through the legitimate engine creates a device in which all the heat flows cancel: a “do-nothing device”. What remains is an “impossible engine”, which converts heat directly into work.

Thus, each of our statements of the Second Law are functionally equivalent.

5.5 Carnot Cycles and Ideal Efficiency

The Carnot cycle is a primarily theoretical cycle that is intended to provide the maximum efficiency between two temperature reservoirs.

Heat transfer into the system occurs isothermally at the highest temperature possible, to minimize the entropy transferred into the system. Likewise, heat transfer out of the system occurs at the lowest temperature possible, to maximize the entropy transferred out of the system. Temperature change occurs adiabatically and isentropically. Figure 5.8 shows the cycle drawn on a p - h diagram.

Figure 5.8: A Carnot Cycle plotted on a p - h diagram for steam.

Note the frequent switching between constant temperature environments (necessitating ample heat flow) and adiabatic environments (requiring zero heat flow). This change of environment is the primary problem with the Carnot cycle for practical purposes.

As you can see in Figure 5.8, the difference in entropy between states 1 and 2 perfectly matches the difference between states 3 and 4. In other words, the entropy entering the system through Q_H exits the system through Q_C .

5.5.1 Carnot Cycle Analysis

By setting the inflow of entropy equal to the outflow, we can determine the maximum possible difference between Q_H and Q_C , which in turn will maximize work (recall that $W = Q_H - Q_C$).

$$\begin{aligned} ds_{in} &= \frac{\delta q}{T_H} & ds_{out} &= \frac{\delta q}{T_C} \\ \Delta s_{in} &= \int \frac{\delta q}{T_H} = \frac{Q_H}{T_H} & \Delta s_{out} &= \int \frac{\delta q}{T_C} = \frac{Q_C}{T_C} \end{aligned}$$

This leads to a relationship between the heat transfers and the temperatures:

$$\Delta s_{in} = \Delta s_{out} \quad \rightarrow \quad \frac{Q_H}{T_H} = \frac{Q_C}{T_C}$$

We can rearrange this to build an equivalence between the heat transfer ratio and the

temperature ratio for the Carnot cycle.

$$\frac{Q_C}{Q_H} = \frac{T_C}{T_H} \quad (5.10)$$

Finally, let's revisit the equation for efficiency of an engine, letting $Q_{in} = Q_H$.

$$\eta_{th} = \frac{W_{net}}{Q_H} = \frac{Q_H - Q_C}{Q_H} = 1 - \frac{Q_C}{Q_H}$$

Using Equation 5.10, we can find the efficiency for a Carnot cycle based purely on the temperature of the reservoirs:

$$\eta_{th,max} = 1 - \frac{T_C}{T_H} \quad (5.11)$$

With a bit of work, we can also come up with the maximum theoretical coefficient of power for the refrigeration and heat pump cycles.

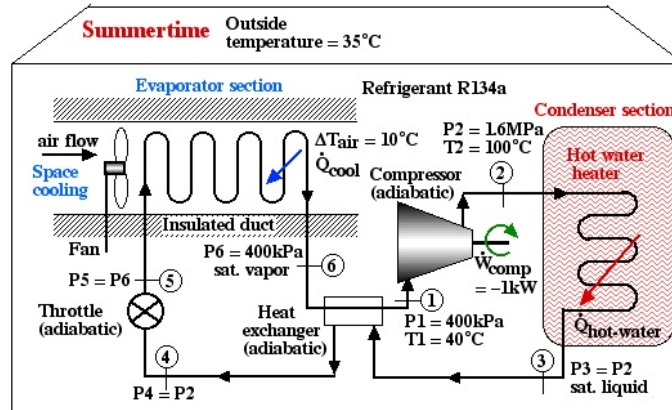
$$COP_{R,max} = \frac{T_C}{T_H - T_C} \quad (5.12)$$

$$COP_{HP,max} = \frac{T_H}{T_H - T_C} \quad (5.13)$$

As always, the coefficient of performance for a heat pump will be exactly 1 larger than that for a refrigerator.

Example 5.2: Reversible Home Air Conditioner and Hot Water Heater

We wish to do a preliminary thermodynamic evaluation of the following proposed heat pump system designed for summertime hot water heating and space cooling.

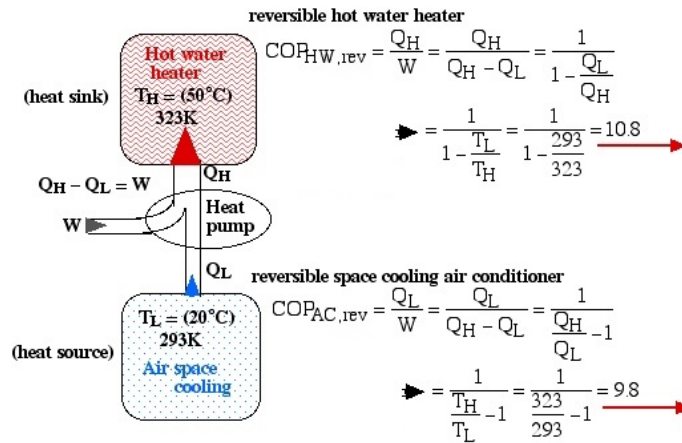


Assuming that this system is required to maintain the home air at 20°C and the hot water at 50°C , we wish to determine the maximum possible theoretical Coefficients of Performance that could be obtained under these conditions.

- Draw a diagram representing the heat pump system showing the flow of energy and source and sink temperatures.

- b) Determine the maximum possible Coefficient of Performance of a hot water heater (COP_{HW}) that could be obtained by a reversible heat pump.
- c) Determine the maximum possible Coefficient of Performance of a space cooling air conditioner (COP_{AC}) that could be obtained by a reversible heat pump.
- d) The actual Coefficients of Performance for the system shown above were found to be $COP_{HW} = 4.41$ and $COP_{AC} = 3.41$. Compare those values to those of the reversible heat pump determine if the actual heat pump shown above is feasible. State the reasons for your conclusion.

Solution: For a), b), and c) we need to reduce the system complexity shown above to an energy flow diagram showing only the basic requirements - cool the air to 20°C and heat the water to 50°C. Given the temperature of the heat source (20°C) and the heat sink (50°C) we can evaluate the respective reversible Coefficients of Performance.



For question d) we use the accumulated knowledge that the conditions for thermal and mechanical reversibility are so difficult to realize that no practical heat pump or heat engine can attain more than 50% to 60% of the equivalent reversible (Carnot) performance. Thus comparing the actual to the reversible Coefficients of Performance we obtain:

$$\frac{COP_{HW,act}}{COP_{HW,rev}} = \frac{4.4}{10.8} = 41\% \quad (< 50\% \rightarrow \text{feasible!})$$

$$\frac{COP_{AC,act}}{COP_{AC,rev}} = \frac{3.4}{9.8} = 35\% \quad (< 50\% \rightarrow \text{feasible!})$$

Both coefficients of performance were less than 50% of their reversible values, meaning that they are indeed feasible.

Chapter 5 Homework:

Question 5.1: A heat pump is used to meet the heating requirements of a house and maintain it at 20°C . On a day when the outdoor air temperature drops to -10°C it is estimated that the house loses heat at the rate of 10 kW. Under these conditions the actual Coefficient of Performance (COP_{HP}) of the heat pump is 2.5.

- Draw a diagram representing the heat pump system showing the flow of energy and the temperatures.
- Determine the actual power consumed by the heat pump [4 kW].
- Determine the power that would be consumed by a reversible heat pump under these conditions [1.02 kW].
- Determine the power that would be consumed by an electric resistance heater under these conditions [10 kW].
- Determine if the performance of the actual heat pump is feasible.

Question 5.2: During an experiment conducted at 25°C , a student measured that a Stirling cycle refrigerator that draws 250W of power has removed 1000kJ of heat from the refrigerated space maintained at -30°C . The running time of the refrigerator during the experiment was 20 min.

- Draw a diagram representing the refrigerator system showing the flow of energy and the temperatures.
- Determine the actual and reversible coefficients of performance [$COP_R = 3.33$, $COP_{R,rev} = 4.42$].
- Determine if these measurements are reasonable. [$COP_R/COP_{R,rev} = 75\% > 60\%$ - not feasible]. State the reasons for your conclusions.

Chapter 6

Isentropic Efficiency and the Brayton Cycle

In this chapter, we take another look at a number of components in thermodynamic systems, using entropy to compare their actual performance to their real performance. After looking specifically at compressors and turbines, we will be able to revisit the Rankine and Refrigeration cycles, as well as explore the Brayton cycle, which is used in jet engines and gas generators.

6.1 Isentropic Efficiency

One of the important applications of isentropic processes is in defining the ideal performance of various adiabatic components. These include turbines, compressors, pumps, and aircraft jet nozzles. Before, we have made the statement that steam turbines are designed to be adiabatic, and that any heat loss from the turbine will result in a reduction in output power. Now, however, we can make the statement that the ideal turbine is isentropic. This enables us to evaluate the **isentropic efficiency** of these components.

There are two property diagrams involving entropy in common usage, the temperature-entropy (T - s) and enthalpy-entropy (h - s) "Mollier" diagrams. We will find that the h - s diagram is extremely useful for evaluating adiabatic turbines and compressors, and complements the p - h diagram which we used in Chapter 4 to evaluate entire steam power plants

or refrigerator systems. The h - s diagram for steam is presented below:

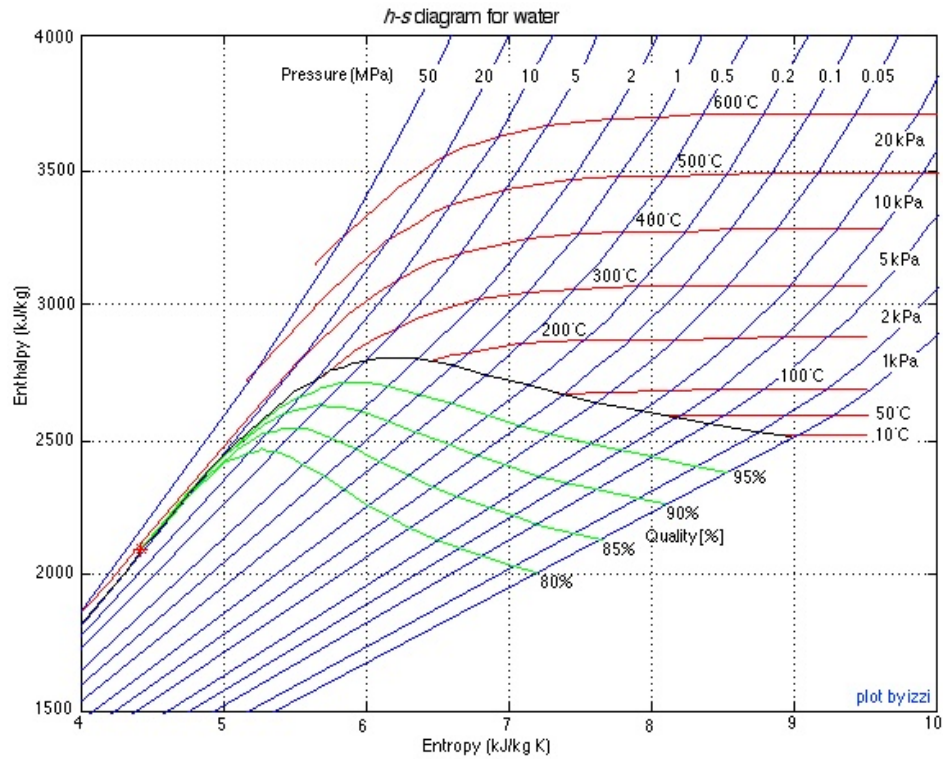


Figure 6.1: The Mollier, or h - s diagram, for water

6.1.1 Isentropic Efficiency of Turbines

The important characteristic of the h - s diagram is that the ideal adiabatic turbine can be conveniently plotted as a vertical line, allowing an intuitive visual appreciation of the turbine

performance. We define the turbine adiabatic efficiency as follows:

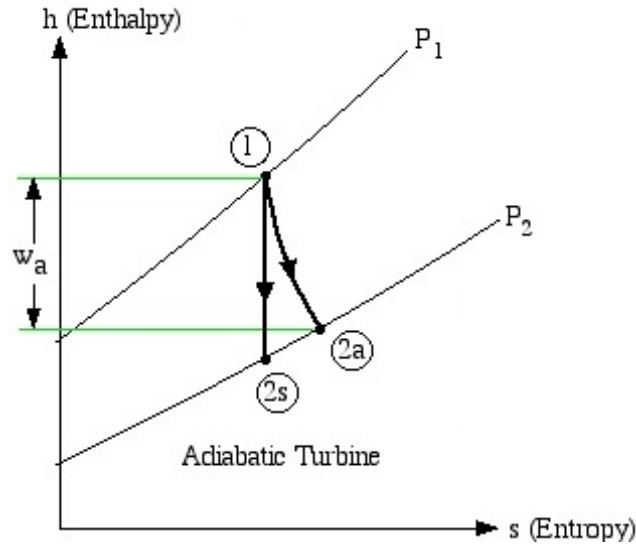


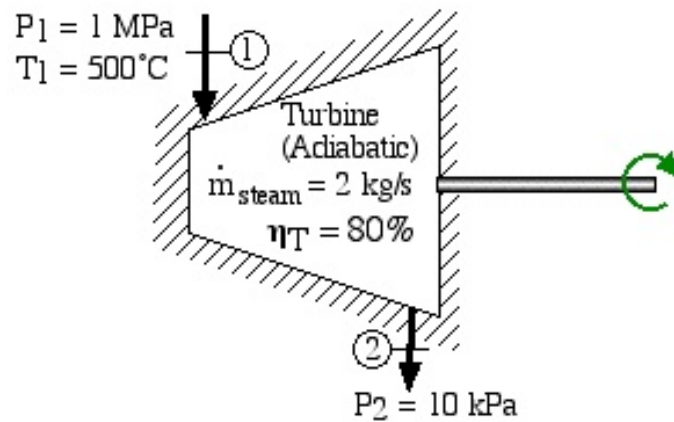
Figure 6.2: An isentropic turbine process (1-2s) shown next to an actual turbine process (1-2a).

$$\eta_T = \frac{\text{actual work}}{\text{isentropic work}} = \frac{w_a}{w_s} = \frac{h_1 - h_{2a}}{h_1 - h_{2s}} \quad (6.1)$$

Notice that for the actual turbine there will always be an increase in entropy, which means that the turbine adiabatic efficiency will always be less than 100%.

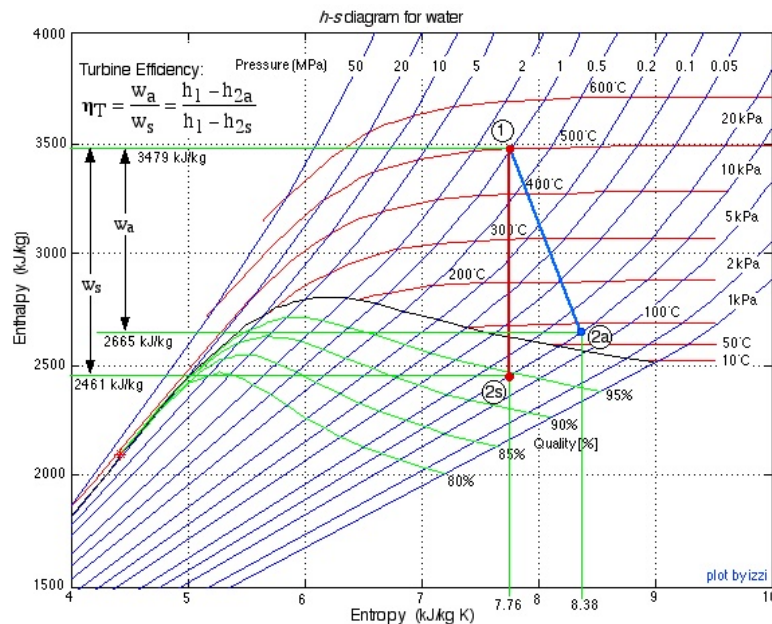
Example 6.1: Adiabatic Steam Turbine

Consider an adiabatic steam turbine having a turbine adiabatic efficiency $\eta_T = 80\%$, operating under the conditions shown in the following diagram:



- Using steam tables, determine the enthalpy and entropy values at station (1) and station (2s) assuming that the turbine is isentropic. [$h_1 = 3479$ kJ/kg, $s_1 = 7.764$ kJ/kg.K; $h_{2s} = 2461$ kJ/kg, $s_{2s} = s_1$]
- From the definition of turbine adiabatic efficiency, and given that $\eta_T = 80\%$, determine the actual enthalpy and entropy values as well as the temperature at station (2a). [$h_{2a} = 2665$ kJ/kg, $s_{2a} = 8.38$ kJ/kg.K, $T_{2a} = 88^\circ\text{C}$]
- Plot the actual and isentropic turbine processes (Stations (1)-(2a) and (1)-(2s)) on the enthalpy-entropy h - s "Mollier" diagram, and indicate the actual turbine specific work (w_a) as well as the isentropic turbine specific work (w_s) on the diagram.
- Determine the actual power output of the turbine (kW). [1629 kW]

The h - s diagram plot follows. Notice that we have indicated all the enthalpy and entropy values (which we determined from the steam tables) on the plot. This allows a check on the feasibility of our results.



6.1.2 Isentropic Efficiency of Compressors

One of the interesting aspects of compressors is that they can be made more efficient by cooling. The reason why we still consider the adiabatic efficiency of compressors that are normally found in refrigeration, air-condition and heat pump systems is that it is considered to be impractical to cool them. Thus the ideal compressor (absorbing a minimum of power)

is considered to be isentropic, and we define compressor adiabatic efficiency as follows:

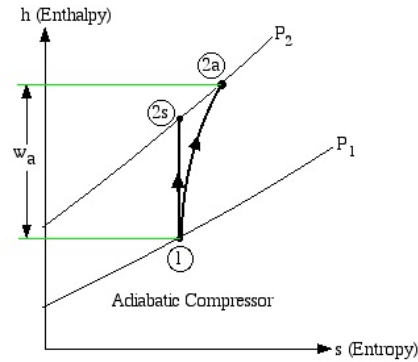


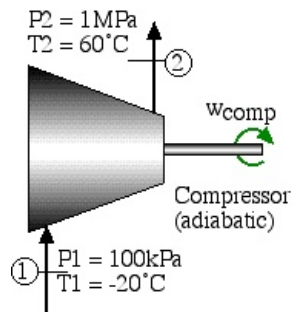
Figure 6.3: An isentropic compressor process (1-2s) shown next to an actual turbine process (1-2a).

$$\eta_C = \frac{\text{isentropic work}}{\text{actual work}} = \frac{w_s}{w_a} = \frac{h_{2s} - h_1}{h_{2a} - h_1} \quad (6.2)$$

Notice that for the actual compressor there will always be an increase in entropy, leading to a compressor adiabatic efficiency which is less than 100%.

Example 6.2: Adiabatic Compressor

Consider the R134a refrigerator compressor shown below.

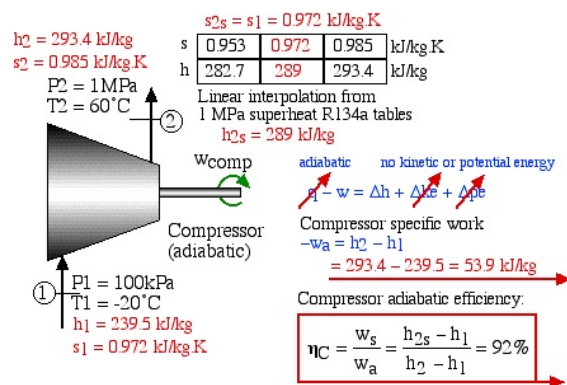
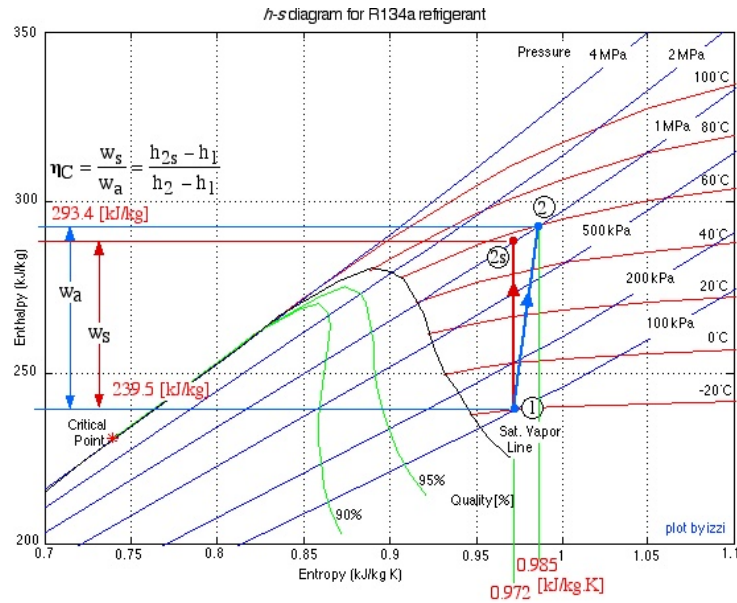


- Carefully plot the actual and isentropic compression processes on the h - s diagram and indicate the actual and isentropic specific work done to drive the compressor on the plot.
- Using R134a refrigerant tables determine the specific work required to drive the compressor [53.9 kJ/kg].
- Using R134a refrigerant tables determine the adiabatic efficiency of the compressor [$\eta_C = 92\%$].
- Discuss these results and determine if this is a feasible compressor design.

Solution:

Filling in information from tables for states 1 and 2:

	State 1	State 2	State 2s	
p	100	1000	1000	kPa
T	-20	60	?	°C
s	0.972	0.985	0.972	kJ/kgK
h	239.5	293.4	?	kJ/kg



6.2 Aircraft Engines and the Brayton Cycle

There are many different forms and modifications of aircraft gas turbine engines, and in this course we discuss two variants - the ideal turbojet engine, and the gas turbine engine for usage in helicopters.

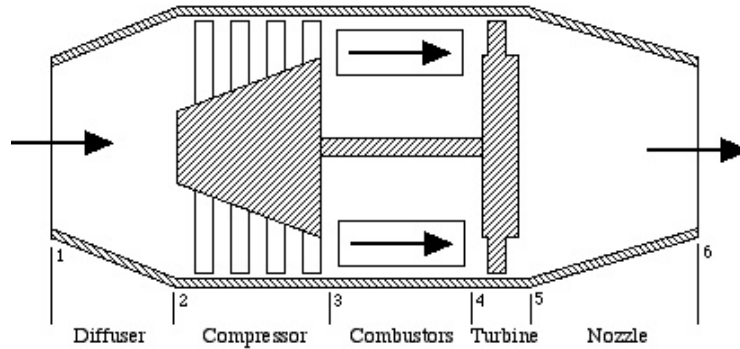


Figure 6.4: A simple diagram showing the five processes that make up a turbojet aircraft engine.

The ideal turbojet engine shown schematically in Figure 6.4 is comprised of the series connection of five components - diffuser, compressor, combustor, turbine, and nozzle. The analysis of the complete system, is best done in terms of the h - s (enthalpy-entropy) diagram, which we will develop in class. Throughout the system we assume that the fluid is pure air, and the combustors are considered to be constant-pressure heat-addition devices. Notice that the sole purpose of the turbine is to drive the compressor, the nozzle providing the final kinetic energy increase to drive the aircraft.

The gas turbine engine for usage in helicopters is shown below:

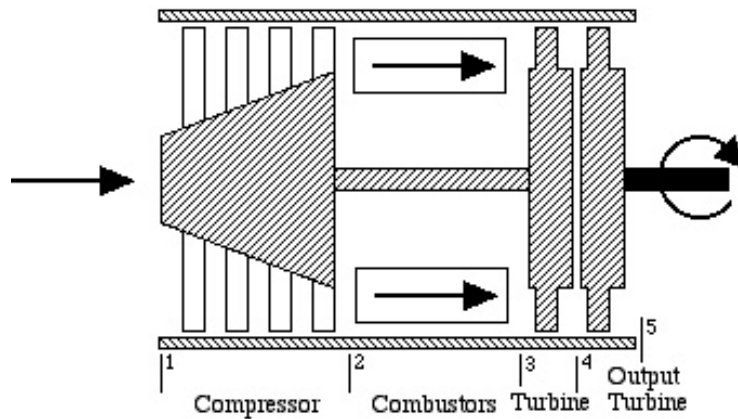
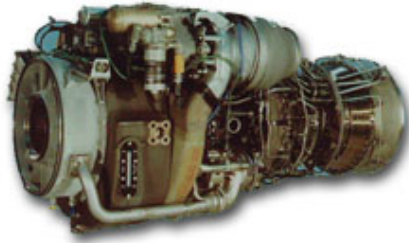


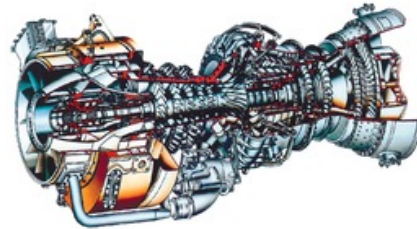
Figure 6.5: A simple diagram showing the five processes that make up a gas turbine helicopter engine.

In this case we see that there is no diffuser or nozzle, and that the turbine section has

been replaced by two independent turbines - a "gas generator", or "gassifier" turbine to drive the compressor, and an output turbine to drive the helicopter blades. A typical gas turbine engine of this type is the General Electric T700 engine shown below, which is used in the Army Black Hawk helicopter.



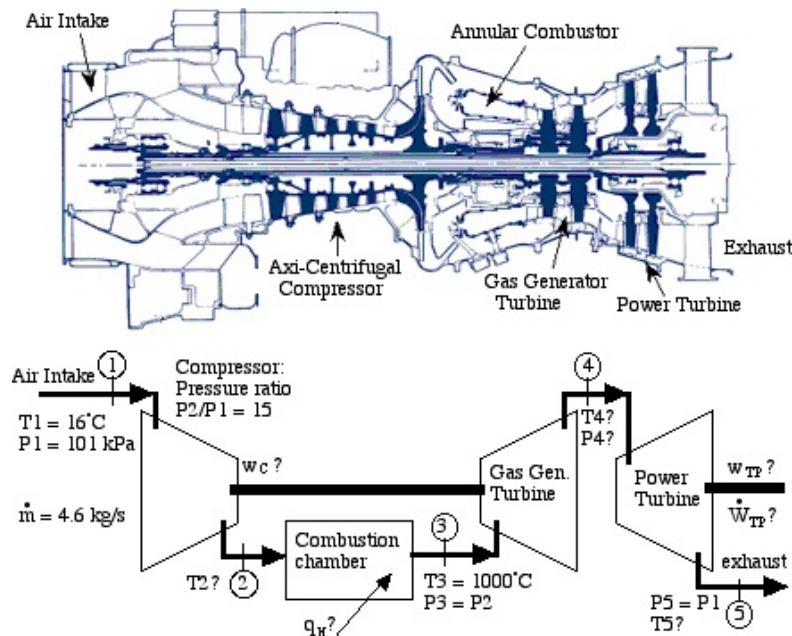
(a) Photograph of the GE T700 engine.



(b) Section view schematic of GE T700 engine.

Example 6.3: GE T700 Gas Turbine Engine

We wish to do an ideal thermodynamic analysis of the General Electric T700 gas turbine engine.



Notice that there are two turbines operating on independent output shafts. The High Pressure (first) turbine, named the Gas Generator Turbine, is directly connected by a shaft to the compressor. Its sole purpose is to drive the the axial/centrifugal compressor, thus the energy output of this turbine must equal the energy consumed by the compressor. The Low Pressure (second) turbine, named the Power Turbine, is connected via gearing to the helicopter rotor.

Note: Because of the large temperature variation throughout this problem we will

need to consider the temperature dependence of the specific heat capacities of air. In the above schematic diagram we see that the temperature extremes of the system are 16°C - 1000°C (289 K - 1273 K), giving an average temperature of 781 K. From the table of Specific Heat Capacities of Air we see that at 800 K, $c_p = 1.099$ [kJ/kgK] and the ratio of specific heat capacities $\gamma = 1.354$, thus we use those values throughout this problem.

Assume that the compressor and both turbines are isentropic, and that the combustion process occurs at constant pressure (isobaric). Using the information shown on the schematic diagram above, do the following:

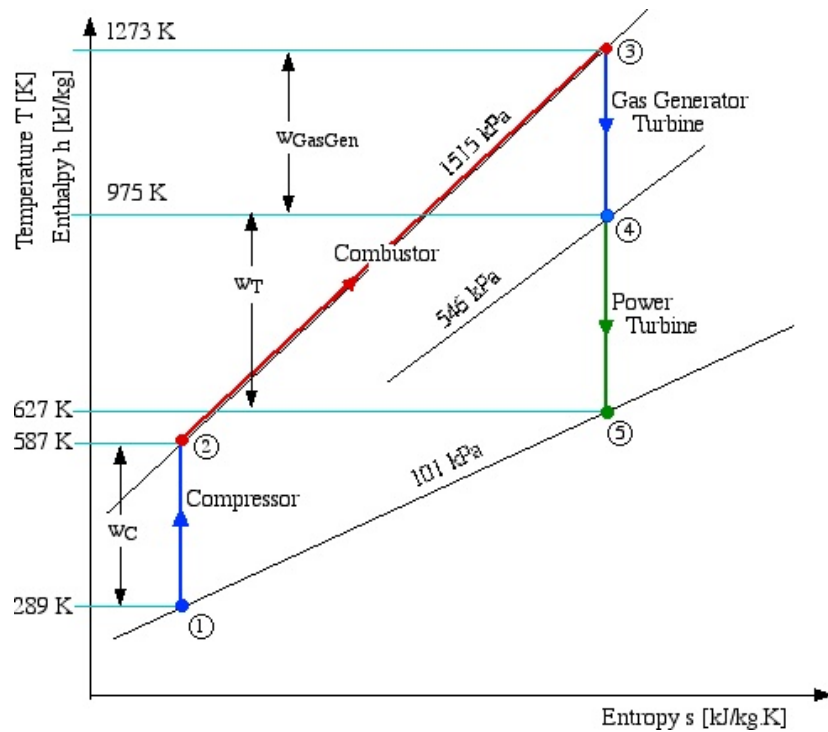
- Sketch the entire process on an h - s diagram, clearly showing the 5 stations on the diagram and the relevant isentropic and constant pressure lines.
- Determine the energy consumed by the compressor [$w_C = -328$ kJ/kg], and the temperature at the outlet of the compressor [$T_2 = 587$ K].
- Determine the heat energy absorbed by the working gas in the combustion chamber [$q_H = 754$ kJ/kg].
- Determine the temperature [$T_4 = 975$ K] and the pressure [$p_4 = 546$ kPa] at the outlet of the gas generator turbine.
- Determine the temperature [$T_5 = 627$ K] and energy output of the power turbine [$w_{PT} = 382.5$ kJ/kg].
- Given that the mass flow rate of the working gas through the system is 4.6 kg/s, determine the power output of the power turbine [1.76 MW].

Solution: information shown on the schematic diagram above, do the following:

- Sketch the entire process on an h - s diagram, clearly showing the 5 stations on the diagram and the relevant isentropic and constant pressure lines.

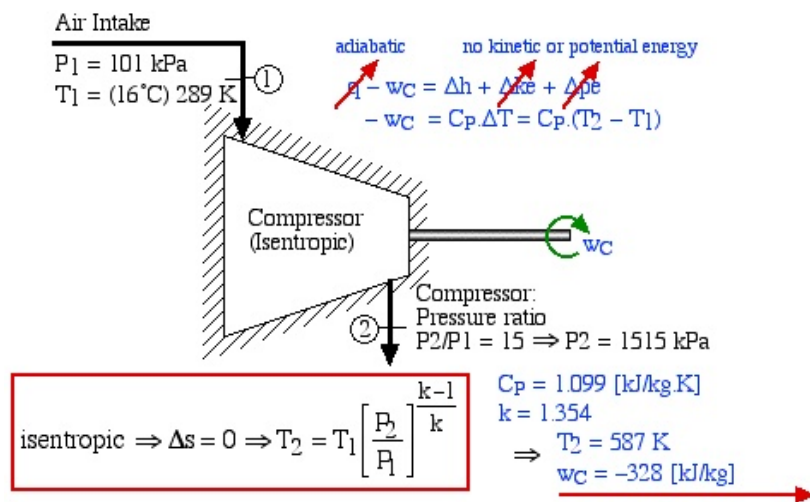
Unlike the case with a pure fluid such as steam, the h - s diagram is not drawn to scale. Instead, is sketched in order to provide an intuitive graphical understanding of the problem.

Furthermore, for an ideal gas the enthalpy is proportional to the temperature, hence the y-axis can be considered either an enthalpy or temperature axis. The various temperatures and pressures shown on this diagram are evaluated and plotted as we progress with the solution.

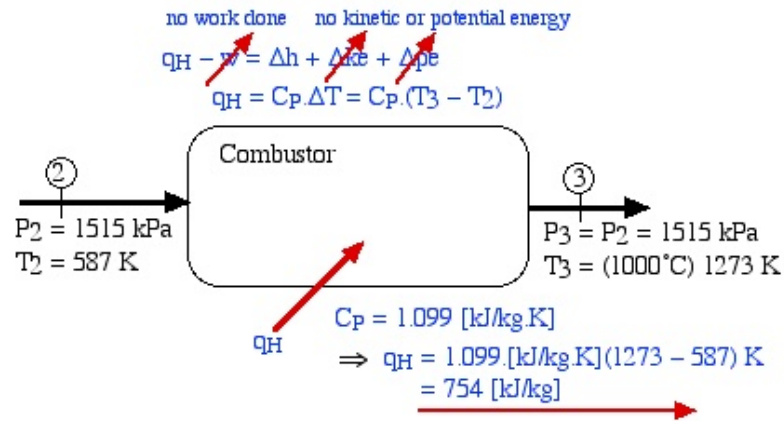


- b) Determine the energy consumed by the compressor, and the temperature at the outlet of the compressor.

Ideally both the compressor and the turbine are isentropic devices, thus given the pressure ratio, in order to determine the temperature we consider the isentropic relations developed for an ideal gas.

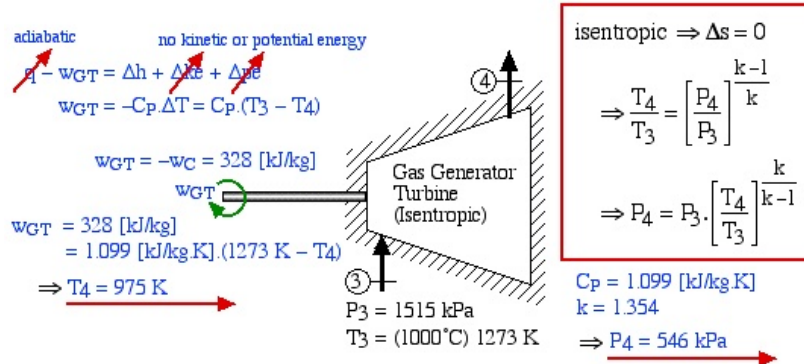


- c) Determine the heat energy absorbed by the working gas in the combustion chamber.

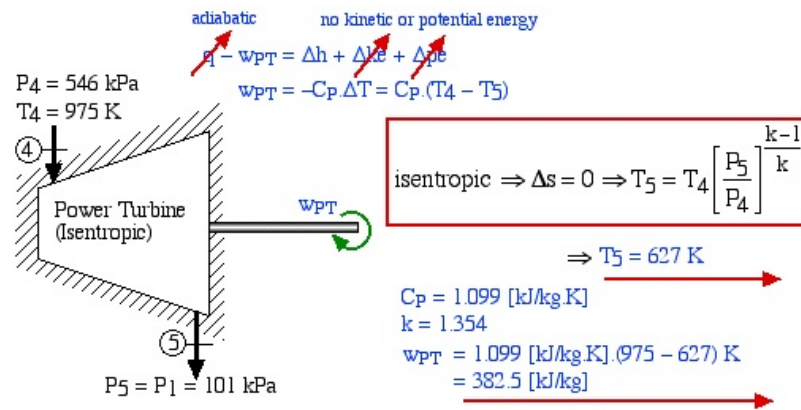


- d) Determine the temperature and the pressure at the outlet of the gas generator turbine.

Once more, since both turbines are isentropic, we use the pressure temperature relations developed for an isentropic process of an ideal gas.



- e) Determine the temperature and energy output of the power turbine.



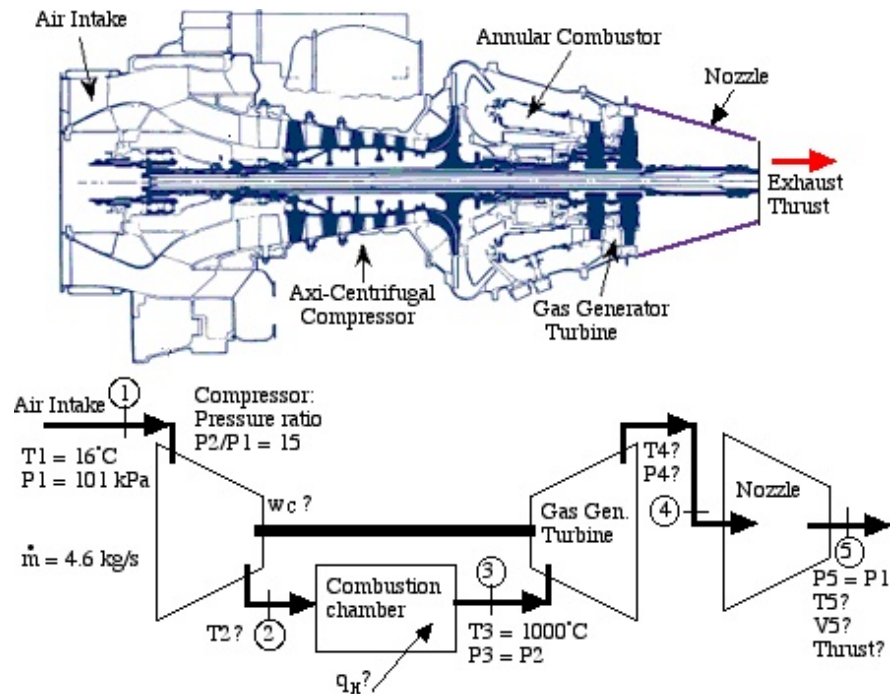
- f) Given that the mass flow rate of the working gas through the system is 4.6 kg/s, determine the power output of the power turbine.

$$\dot{W}_{PT} = \dot{m} w_{PT} = 4.6 \frac{\text{kg}}{\text{s}} \cdot 382.5 \frac{\text{kJ}}{\text{kg}} = 1.76 \text{ MW}$$

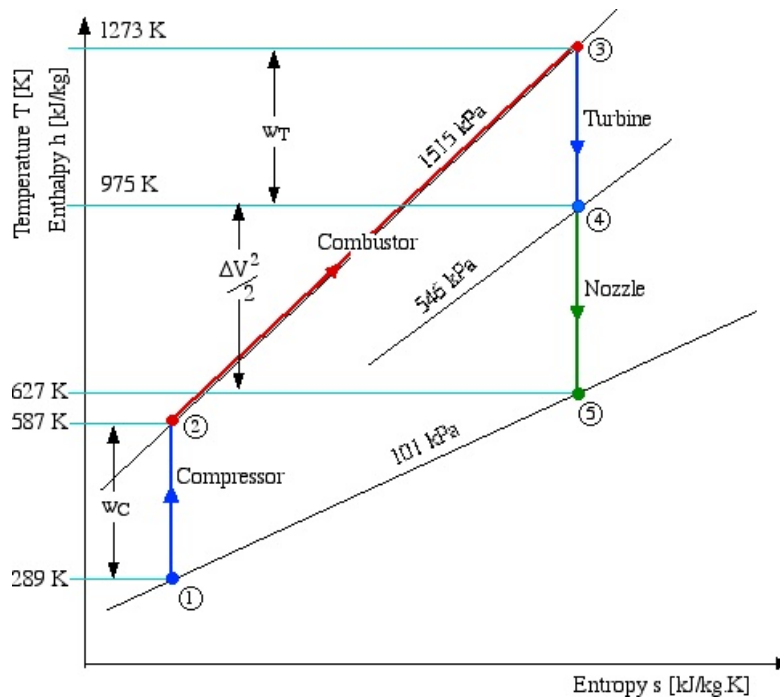
Note that the actual power output of the T700 engine is around 1800 hp, which is significantly less than the above value (≈ 2360 hp). This is because we have assumed that the compressor and both turbines are isentropic, which will never occur in practice. In a later problem, we will extend this exercise to consider a non-isentropic compressor and turbines.

Example 6.4: T700 Proposed Turbojet Conversion

In Example 6.3, we did an ideal thermodynamic analysis of the General Electric T700 helicopter gas turbine engine. In theory, one could unclamp and remove the power turbine at the rear and replace it with a nozzle to form a turbojet engine. Consider the schematic diagram of this conversion shown in the figure below:



Notice that the h - s diagram has retained the same shape and characteristics as that in Example 6.3, with the only difference being that the change in enthalpy on the nozzle gives rise to a kinetic energy increase, rather than work output.



On replacing the power turbine with an isentropic jet nozzle, determine the temperature (T_5) and exhaust velocity (V_5 - m/s). Given that the mass flow rate of the working gas through the system is 4.6 kg/s, determine the thrust force at the outlet of the nozzle (Thrust5 - lbf). [Note: 1 lbf (pound force) = 4.46 N]

adiabatic no work done no potential energy

$$q_1 - w = \Delta h + \frac{\Delta V^2}{2} + \Delta p_e$$

$P_4 = 546 \text{ kPa}$ ④ $P_5 = P_1 = 101 \text{ kPa}$ ⑤

$T_4 = 975 \text{ K}$

$\frac{V_5^2 - V_4^2}{2} = C_p (T_4 - T_5)$

isentropic $\Rightarrow \Delta s = 0 \Rightarrow T_5 = T_4 \left[\frac{P_5}{P_4} \right]^{\frac{k-1}{k}}$

$C_p = 1.099 \text{ [kJ/kg K]}$
 $k = 1.354$
 $\Rightarrow T_5 = 627 \text{ K}$

Thrust = $\dot{m} \cdot V_5$

$\frac{V_5^2}{2} \frac{[\text{m}^2/\text{s}^2]}{1000} = 1.099 \text{ [kJ/kg K]} \cdot (975 - 627) \text{ K}$
 $\Rightarrow V_5 = 874 \text{ [m/s]}$

$\dot{m}_{\text{air}} = 4.6 \text{ kg/s}$
 $\Rightarrow \text{Thrust} = 4.6 \text{ [kg/s]} (874 \text{ [m/s]}) = 4020 \text{ N (approx 900 lbf)}$

Example 6.5: Flying at altitude

A theoretical jet engine is used to power a plane's flight at an altitude of 20 km and a speed of 200 m/s.

All that is known about the engine is that the compressor's pressure ratio is 18, and the turbine inlet temperature is 1250 K.

Determine the thrust of the engine if the mass flow of air is 15 kg/s. Find the thermal efficiency.

Solution:

There are 6 total states that we need to determine [Answers using CoolProp]:

- Atmosphere (a) - pressure and temperature at 20 km, with a velocity of 200 m/s [5474 Pa, 216.65 K]
- Post-diffuser (1) - the diffuser will convert all of the kinetic energy of the air into enthalpy. We will assume this is isentropic. Kinetic energy is typically $KE = 1/2mV^2$, but we tend to work in mass-specific units, so instead, we will find $ke = 1/2V^2$. This is in units of J/kg, so remember to convert to kJ/kg! [$h_1 = 362.9$ kJ/kg, $p_1 = 7446$ Pa]
- Post-compressor, ideal (2s) - the compressor has a pressure ratio of 18. We will assume isentropic for this state. [$h_{2s} = 668$ kJ/kg]
- Post-compressor, actual (2) - we will use the isentropic efficiency of a compressor (around 80-85 percent, we will use 82%) to find the new enthalpy of this state, while assuming that the pressure remains the same. [$h_2 = 735$ kJ/kg]
- Post-combustion (3) - we know the turbine inlet temperature, and will assume that the pressure remains the same from state 2. [$h_3 = 1463$ kJ/kg]
- Post-turbine (4, 4s) - the total amount of work from the turbine ($h_4 - h_3$) should be equal to the amount of work from the compressor ($h_2 - h_1$). This is not enough to determine a state by itself. We assume that the pressure is constant between states 4 and 4s, and use the isentropic efficiency of a turbine (around 85-90 percent, we will use 91%) to determine the enthalpy at state 4s. The entropy is known there (same as s_3), so we can determine the pressure at state 4s, at which point all information is known for state 4. [$h_4 = 1091$ kJ/kg, $h_{4s} = 1055$ kJ/kg]
- Post-nozzle (5) - we set the pressure back to atmospheric and assume that the nozzle operates isentropically. The difference in energy ($h_5 - h_4$) will be converted back to kinetic energy. Remember to convert from kJ/kg to J/kg! [$h_5 = 699$ kJ/kg]

Once we have all the states, we need to determine the thrust and the thermal efficiency:

- Thrust - $T = \dot{m}(V_e - V_i)$

Once we have both the inlet and outlet velocity, this is relatively trivial. [Thrust=10.3 kN]

- Thermal efficiency - $\eta_{th} = (ke_{\text{nozzle}} - ke_{\text{diffuser}}) / q_{\text{combustor}}$

We care about the thrust, which is in units of Newtons. We are putting in energy as heat into the combustor, which is in units of kJ/kg. However, to find efficiency, we need to have consistent units between the two values. The easiest thing to do is to find the kinetic energy converted within the nozzle and diffuser (since we already have the enthalpy values).

We can rewrite this as:

$$\eta_{th} = \frac{(h_1 - h_a) - (h_4 - h_5)}{(h_3 - h_2)} \quad (6.3)$$

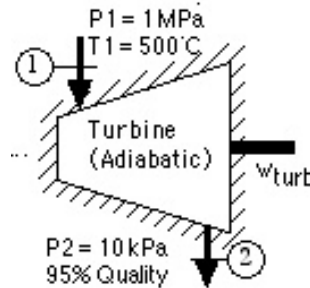
$$[\eta_{th} = 0.51]$$

Important notes:

- We should find the enthalpy at each state at a minimum. If we are using CoolProp, we need the enthalpy and entropy at each state, at a minimum. We can avoid finding the actual entropy if we are using the ideal gas assumption, unless we want to plot the cycle on the h - s diagram.
- Whichever path we choose (ideal gas vs. CoolProp), we can write a code in Colab to find our states and efficiency. If you use the ideal gas law, you can actually use Excel instead. Be sure to choose good values for air properties!
- Pressure ratio is a ratio of pressures (unlike the compression ratio from Chapter 3, which was a ratio of volumes).

Chapter 6 Homework:

Question 6.1: Consider the adiabatic steam turbine shown below.

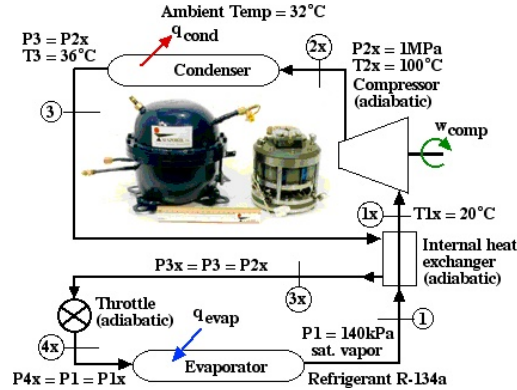


- Carefully plot the turbine process on the h - s diagram and indicate the turbine work done on the plot.
- Using steam tables determine the specific work output of the turbine [1015 kJ/kg].
- Using steam tables determine the entropy generated by this process. [0.01 kJ/kg.K].
- Discuss these results and determine if this is a feasible turbine design.

Question 6.2: Adiabatic Evaluation of a R134a Compressor A young engineer was assigned to evaluate the compressor of a proposed R-134a refrigeration system shown below, and assumed it to be adiabatic.

- Under this assumption, and ignoring kinetic and potential energy effects, determine the specific work input to the compressor [48.3 kJ/kg].
- His supervisor checked the results and told him that his assumption was incorrect and not physically feasible. On the h - s diagram for R134a, draw the actual compression process (as assumed) as well as the equivalent isentropic compression process.
- Using the h - s diagram, as well as relevant data evaluated at the inlet and outlet of the compressor, explain how she arrived at that conclusion.

Question 6.3: Previously, we provided a problem concerning a home refrigerator, and examining it's performance before and after adding an internal heat exchanger.

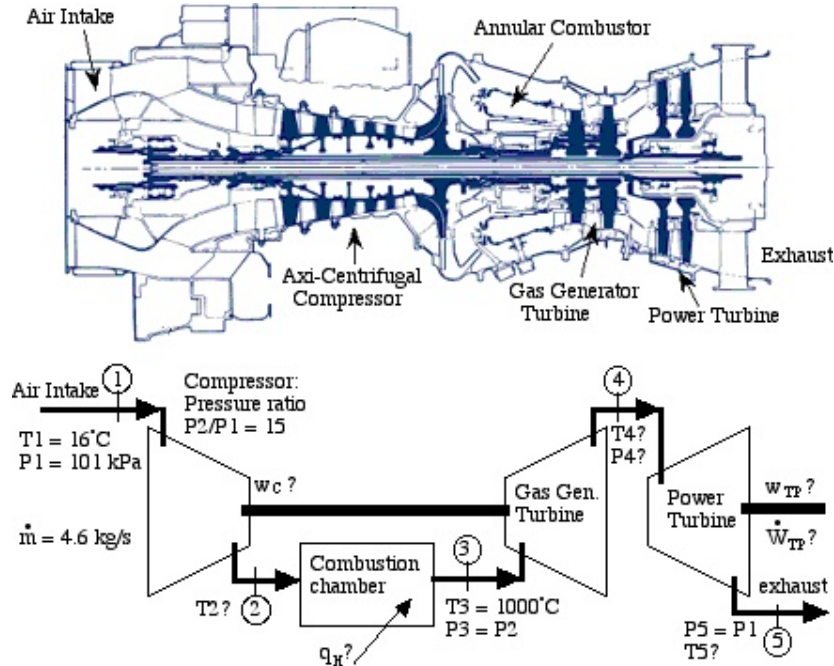


- Plot the actual and the isentropic compressor processes on the enthalpy-entropy (h - s) diagram for both cases: with and without the internal heat exchanger.
- Using the R134a tables determine the actual compressor adiabatic efficiency (η_C) for both cases [75%, 76%]

Question 6.4: A compressor is used to drive R134a through a heat pump. R134a enters the compressor at a pressure $p_1 = 400$ kPa and a temperature $T_1 = 40^\circ\text{C}$ and leaves at a pressure $p_2 = 1.6$ MPa. For the actual compressor, the exit temperature is $T_2 = 100^\circ\text{C}$.

- Carefully plot the actual and isentropic compression processes on the h - s diagram and indicate the actual and isentropic specific work done to drive the compressor on the plot.
- Using R134a refrigerant tables, determine the specific work required to drive the compressor [43.5 kJ/kg].
- Using R134a refrigerant tables, determine the adiabatic efficiency of the compressor [$\eta_C = 78\%$].
- Discuss these results and determine if this is a feasible compressor design.

Question 6.5: In Example 6.3 we did an ideal thermodynamic analysis of the General Electric T700 helicopter gas turbine engine, shown in the following schematic diagram:



Notice again that there are two turbines operating on independent output shafts. The High Pressure (first) turbine, named the Gas Generator turbine, is directly connected by a shaft to the compressor. Its sole purpose is to drive the the compressor, thus the energy output of this turbine must equal the energy consumed by the compressor. The Low Pressure (second) turbine, named the Power turbine, is connected via gearing to the helicopter rotor.

In Example 6.3 we assumed that the compressor and both turbines were isentropic. In this exercise we wish to extend the analysis to non-isentropic compressor and turbines.

Assume that the compressor adiabatic efficiency $\eta_C = 88\%$, and that each turbine has an adiabatic efficiency $\eta_T = 86\%$.

- Sketch the entire process on an h - s diagram, clearly showing the 5 stations on the diagram and the relevant isentropic and constant pressure lines. Indicate the relevant actual and isentropic work values on the sketch.
- Determine the actual energy consumed by the compressor [$w_{C,act} = 373 \text{ kJ/kg}$], and the actual temperature at the outlet of the compressor [$T_{2a} = 628\text{K}$].
- Determine the heat energy absorbed by the working gas in the combustion chamber [$q_H = 709 \text{ kJ/kg}$].
- Determine the actual temperature [$T_{4a} = 934\text{K}$] and the pressure [$P_4 = 366 \text{ kPa}$] at the outlet of the gas generator turbine.
- Determine the actual temperature [T_{5a}] and energy output of the power turbine [$w_{PT,act} = 252 \text{ kJ/kg}$].

- f) Given that the mass flow rate of the working gas through the system is 4.6 kg/s, determine the actual power output of the power turbine [1161 kW].
- g) Determine the thermal efficiency (η_{th}) of the T700 gas turbine. Compare this value to the equivalent reversible thermal efficiency and discuss your results.

Appendices

Appendix A

Steam Tables

Unless otherwise indicated, all data was sourced from the [NIST Chemistry WebBook](#), Accessed Dec. 2021.

A.1 Saturation Properties for Steam - Temperature

Temp.	Pressure	Spec. Volume		Int. Energy		Enthalpy		Entropy	
T	p	v_f	v_g	u_f	u_g	h_f	h_g	s_f	s_g
°C	MPa	m ³ /kg	m ³ /kg	kJ/kg	kJ/kg	kJ/kg	kJ/kg	kJ/kgK	kJ/kgK
0.01	0.00061	0.00100	205.99	0	2374.9	0	2500.9	0.00	9.1555
5	0.00087	0.00100	147.01	21.019	2381.8	21.019	2510.1	0.0763	9.0248
10	0.00123	0.00100	106.3	42.020	2388.6	42.020	2519.2	0.1511	8.8998
15	0.00171	0.00100	77.875	62.980	2395.5	62.980	2528.3	0.2245	8.7803
20	0.00234	0.00100	57.757	83.912	2402.3	83.912	2537.4	0.2965	8.6660
25	0.00317	0.00100	43.337	104.83	2409.1	104.83	2546.5	0.3672	8.5566
30	0.00425	0.00100	32.878	125.73	2415.9	125.73	2555.5	0.4368	8.4520
35	0.00563	0.00101	25.205	146.63	2422.7	146.63	2564.5	0.5051	8.3517
40	0.00738	0.00101	19.515	167.53	2429.4	167.53	2573.5	0.5724	8.2555
45	0.00960	0.00101	15.252	188.43	2436.1	188.43	2582.4	0.6386	8.1633
50	0.01235	0.00101	12.027	209.33	2442.7	209.33	2591.3	0.7038	8.0748
55	0.01576	0.00101	9.5643	230.24	2449.3	230.24	2600.1	0.7680	7.9898
60	0.01995	0.00102	7.6672	251.16	2455.9	251.16	2608.8	0.8313	7.9081
65	0.02504	0.00102	6.1935	272.09	2462.4	272.09	2617.5	0.8937	7.8296
70	0.03120	0.00102	5.0395	293.03	2468.9	293.03	2626.1	0.9551	7.7540
75	0.03860	0.00103	4.1289	313.99	2475.2	313.99	2634.6	1.0158	7.6812
80	0.04741	0.00103	3.4052	334.96	2481.6	334.96	2643.0	1.0756	7.6111
85	0.05787	0.00103	2.8258	355.95	2487.8	355.95	2651.3	1.1346	7.5434
90	0.07018	0.00104	2.3591	376.97	2494.0	376.97	2659.5	1.1929	7.4781
95	0.08461	0.00104	1.9806	398.00	2500.0	398.00	2667.6	1.2504	7.4151
100	0.10142	0.00104	1.6718	419.06	2506.0	419.06	2675.6	1.3072	7.3541
105	0.12090	0.00105	1.4184	440.15	2511.9	440.15	2683.4	1.3633	7.2952
110	0.14338	0.00105	1.2093	461.26	2517.7	461.26	2691.1	1.4188	7.2381
115	0.16918	0.00106	1.0358	482.41	2523.3	482.41	2698.6	1.4737	7.1828
120	0.19867	0.00106	0.8912	503.60	2528.9	503.60	2705.9	1.5279	7.1291
125	0.23224	0.00106	0.7700	524.83	2534.3	524.83	2713.1	1.5816	7.0770
130	0.27028	0.00107	0.6680	546.09	2539.5	546.09	2720.1	1.6346	7.0264
135	0.31323	0.00107	0.5817	567.41	2544.7	567.41	2726.9	1.6872	6.9772
140	0.36154	0.00108	0.5085	588.77	2549.6	588.77	2733.4	1.7392	6.9293
145	0.41568	0.00109	0.4460	610.19	2554.4	610.19	2739.8	1.7907	6.8826
150	0.47616	0.00109	0.3925	631.66	2559.1	631.66	2745.9	1.8418	6.8371

Temp.	Pressure	Spec. Volume		Int. Energy		Enthalpy		Entropy	
T	p	v_f	v_g	u_f	u_g	h_f	h_g	s_f	s_g
°C	MPa	m ³ /kg	m ³ /kg	kJ/kg	kJ/kg	kJ/kg	kJ/kg	kJ/kgK	kJ/kgK
155	0.54350	0.00110	0.3465	653.19	2563.5	653.19	2751.8	1.8924	6.7926
160	0.61823	0.00110	0.3068	674.79	2567.8	674.79	2757.4	1.9426	6.7491
165	0.70093	0.00111	0.2724	696.46	2571.9	696.46	2762.8	1.9923	6.7066
170	0.79219	0.00111	0.2426	718.20	2575.7	718.20	2767.9	2.0417	6.6650
175	0.8926	0.00112	0.2166	740.02	2579.4	740.02	2772.7	2.0906	6.6241
180	1.0028	0.00113	0.1938	761.92	2582.8	761.92	2777.2	2.1392	6.5840
185	1.1235	0.00113	0.1739	783.91	2586.0	783.91	2781.4	2.1875	6.5447
190	1.2552	0.00114	0.1564	806.00	2589.0	806.00	2785.3	2.2355	6.5059
195	1.3988	0.00115	0.1409	828.18	2591.7	828.18	2788.8	2.2832	6.4678
200	1.5549	0.00116	0.1272	850.47	2594.2	850.47	2792.0	2.3305	6.4302
205	1.7243	0.00116	0.1151	872.87	2596.4	872.87	2794.8	2.3777	6.3930
210	1.9077	0.00117	0.1043	895.39	2598.3	895.39	2797.3	2.4245	6.3563
215	2.1058	0.00118	0.0947	918.04	2599.9	918.04	2799.3	2.4712	6.3200
220	2.3196	0.00119	0.0861	940.82	2601.2	940.82	2800.9	2.5177	6.2840
225	2.5497	0.00120	0.0784	963.74	2602.2	963.74	2802.1	2.5640	6.2483
230	2.7971	0.00121	0.0715	986.81	2602.9	986.81	2802.9	2.6101	6.2128
235	3.0625	0.00122	0.0653	1010.0	2603.2	1010.0	2803.2	2.6561	6.1775
240	3.3469	0.00123	0.0597	1033.4	2603.1	1033.4	2803.0	2.7020	6.1423
245	3.6512	0.00124	0.0547	1057.0	2602.7	1057.0	2802.2	2.7478	6.1072
250	3.9762	0.00125	0.0501	1080.8	2601.8	1080.8	2800.9	2.7935	6.0721
255	4.3229	0.00126	0.0459	1104.8	2600.5	1104.8	2799.1	2.8392	6.0369
260	4.6923	0.00128	0.0422	1129.0	2598.7	1129.0	2796.6	2.8849	6.0016
265	5.0853	0.00129	0.0387	1153.4	2596.5	1153.4	2793.5	2.9307	5.9661
270	5.5030	0.00130	0.0356	1178.1	2593.7	1178.1	2789.7	2.9765	5.9304
275	5.9464	0.00132	0.0328	1203.1	2590.3	1203.1	2785.2	3.0224	5.8944
280	6.4166	0.00133	0.0302	1228.3	2586.4	1228.3	2779.9	3.0685	5.8579
285	6.9147	0.00135	0.0278	1253.9	2581.8	1253.9	2773.7	3.1147	5.8209
290	7.4418	0.00137	0.0256	1279.9	2576.5	1279.9	2766.7	3.1612	5.7834
295	7.9991	0.00138	0.0235	1306.2	2570.5	1306.2	2758.7	3.2080	5.7451
300	8.5879	0.00140	0.0217	1332.9	2563.6	1332.9	2749.6	3.2552	5.7059
305	9.2094	0.00143	0.0199	1360.2	2555.9	1360.2	2739.4	3.3028	5.6657
310	9.8651	0.00145	0.0183	1387.9	2547.1	1387.9	2727.9	3.3510	5.6244
315	10.556	0.00147	0.0169	1416.3	2537.2	1416.3	2715.1	3.3998	5.5816
320	11.284	0.00150	0.0155	1445.3	2526.0	1445.3	2700.6	3.4494	5.5372
325	12.051	0.00153	0.0142	1475.1	2513.4	1475.1	2684.3	3.5000	5.4908
330	12.858	0.00156	0.0130	1505.8	2499.2	1505.8	2666.0	3.5518	5.4422
335	13.707	0.00160	0.0118	1537.6	2483.0	1537.6	2645.4	3.6050	5.3906
340	14.601	0.00164	0.0108	1570.6	2464.4	1570.6	2621.8	3.6601	5.3356
345	15.541	0.00168	0.0098	1605.3	2443.1	1605.3	2594.9	3.7176	5.2762
350	16.529	0.00174	0.0088	1642.1	2418.1	1642.1	2563.6	3.7784	5.2110
355	17.570	0.00181	0.0079	1682.0	2388.4	1682.0	2526.6	3.8439	5.1380
360	18.666	0.00190	0.0069	1726.3	2351.8	1726.3	2481.5	3.9167	5.0536
365	19.821	0.00202	0.0060	1777.8	2303.8	1777.8	2422.9	4.0014	4.9497
370	21.044	0.00222	0.0050	1844.1	2230.3	1844.1	2334.5	4.1112	4.8012
373.9	22.052	0.00286	0.0034	1977.9	2058.6	1977.9	2133.5	4.3402	4.4831

A.2 Saturation Properties for Steam - Pressure

Pressure	Temp.	Spec. Volume		Int. Energy		Enthalpy		Entropy	
p MPa	T °C	v_f m ³ /kg	v_g m ³ /kg	u_f kJ/kg	u_g kJ/kg	h_f kJ/kg	h_g kJ/kg	s_f kJ/kgK	s_g kJ/kgK
0.0010	6.970	0.00100	129.18	29.298	2384.5	29.299	2513.7	0.1059	8.9749
0.0012	9.654	0.00100	108.67	40.568	2388.2	40.569	2518.6	0.1460	8.9082
0.0014	11.969	0.00100	93.899	50.279	2391.3	50.280	2522.8	0.1802	8.8521
0.0016	14.010	0.00100	82.743	58.831	2394.1	58.833	2526.5	0.2100	8.8035
0.002	17.495	0.00100	66.987	73.426	2398.9	73.428	2532.9	0.2606	8.7226
0.003	24.079	0.00100	45.653	100.97	2407.9	100.98	2544.8	0.3543	8.5764
0.004	28.960	0.00100	34.791	121.38	2414.5	121.39	2553.7	0.4224	8.4734
0.005	32.874	0.00101	28.185	137.74	2419.8	137.75	2560.7	0.4762	8.3938
0.006	36.159	0.00101	23.733	151.47	2424.2	151.48	2566.6	0.5208	8.3290
0.008	41.509	0.00101	18.099	173.83	2431.4	173.84	2576.2	0.5925	8.2273
0.010	45.806	0.00101	14.670	191.80	2437.2	191.81	2583.9	0.6492	8.1488
0.012	49.419	0.00101	12.358	206.90	2442.0	206.91	2590.3	0.6963	8.0849
0.014	52.547	0.00101	10.691	219.98	2446.1	219.99	2595.8	0.7366	8.0311
0.016	55.313	0.00101	9.4306	231.55	2449.8	231.57	2600.6	0.7720	7.9846
0.02	60.058	0.00102	7.6480	251.40	2456.0	251.42	2608.9	0.8320	7.9072
0.03	69.095	0.00102	5.2284	289.24	2467.7	289.27	2624.5	0.9441	7.7675
0.04	75.857	0.00103	3.9930	317.58	2476.3	317.62	2636.1	1.0261	7.6690
0.05	81.317	0.00103	3.2400	340.49	2483.2	340.54	2645.2	1.0912	7.5930
0.06	85.926	0.00103	2.7317	359.84	2489.0	359.91	2652.9	1.1454	7.5311
0.08	93.486	0.00104	2.0871	391.63	2498.2	391.71	2665.2	1.2330	7.4339
0.10	99.606	0.00104	1.6939	417.40	2505.6	417.50	2674.9	1.3028	7.3588
0.12	104.78	0.00105	1.4284	439.23	2511.7	439.36	2683.1	1.3609	7.2977
0.14	109.29	0.00105	1.2366	458.27	2516.9	458.42	2690.0	1.4110	7.2461
0.16	113.30	0.00105	1.0914	475.21	2521.4	475.38	2696.0	1.4551	7.2014
0.2	120.21	0.00106	0.88568	504.49	2529.1	504.70	2706.2	1.5302	7.1269
0.3	133.52	0.00107	0.60576	561.10	2543.2	561.43	2724.9	1.6717	6.9916
0.4	143.61	0.00108	0.46238	604.22	2553.1	604.65	2738.1	1.7765	6.8955
0.5	151.83	0.00109	0.37481	639.54	2560.7	640.09	2748.1	1.8604	6.8207
0.6	158.83	0.00110	0.31558	669.72	2566.8	670.38	2756.1	1.9308	6.7592
0.8	170.41	0.00111	0.24034	719.97	2576.0	720.86	2768.3	2.0457	6.6616
1.0	179.88	0.00113	0.19436	761.39	2582.7	762.52	2777.1	2.1381	6.5850
1.2	187.96	0.00114	0.16326	796.96	2587.8	798.33	2783.7	2.2159	6.5217
1.4	195.04	0.00115	0.14078	828.36	2591.8	829.97	2788.8	2.2835	6.4675
1.6	201.37	0.00116	0.12374	856.60	2594.8	858.46	2792.8	2.3435	6.4199
2	212.38	0.00118	0.09959	906.14	2599.1	908.50	2798.3	2.4468	6.3390
3	233.85	0.00122	0.06666	1004.7	2603.2	1008.3	2803.2	2.6455	6.1856
4	250.35	0.00125	0.04978	1082.5	2601.7	1087.5	2800.8	2.7968	6.0696
5	263.94	0.00129	0.03945	1148.2	2597.0	1154.6	2794.2	2.9210	5.9737
6	275.58	0.00132	0.03245	1206.0	2589.9	1213.9	2784.6	3.0278	5.8901
8	295.01	0.00138	0.02353	1306.2	2570.5	1317.3	2758.7	3.2081	5.7450
10	311.00	0.00145	0.01803	1393.5	2545.2	1408.1	2725.5	3.3606	5.6160
12	324.68	0.00153	0.01426	1473.1	2514.3	1491.5	2685.4	3.4967	5.4939
14	336.67	0.00161	0.01149	1548.4	2477.1	1571.0	2637.9	3.6232	5.3727
16	347.35	0.00171	0.00931	1622.3	2431.8	1649.7	2580.8	3.7457	5.2463
20	365.75	0.00204	0.00587	1786.4	2295.0	1827.2	2412.3	4.0156	4.9314
22	373.71	0.00270	0.00365	1951.8	2092.8	2011.3	2173.1	4.2945	4.5446
22.064	373.95	0.00311	0.00311	2015.7	2015.7	2084.3	2084.3	4.4070	4.4070

A.3 Superheated Vapor Properties for Steam

$p = 0.010 \text{ MPa } (T_{sat} = 45.8^\circ\text{C})$					$p = 0.050 \text{ MPa } (T_{sat} = 81.3^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	14.670	2437.2	2583.9	8.1488	3.2400	2483.2	2645.2	7.5930
50	14.867	2443.3	2592.0	8.1741				
100	17.196	2515.5	2687.5	8.4489	3.4187	2511.5	2682.4	7.6953
150	19.513	2587.9	2783.0	8.6892	3.8897	2585.7	2780.2	7.9413
200	21.826	2661.3	2879.6	8.9049	4.3562	2660.0	2877.8	8.1592
250	24.136	2736.1	2977.4	9.1015	4.8206	2735.1	2976.1	8.3568
300	26.446	2812.3	3076.7	9.2827	5.2840	2811.6	3075.8	8.5386
350	28.755	2890.0	3177.5	9.4513	5.7469	2889.4	3176.8	8.7076
400	31.063	2969.3	3279.9	9.6094	6.2094	2968.9	3279.3	8.8659
450	33.371	3050.3	3384.0	9.7584	6.6717	3049.9	3383.5	9.0151
500	35.680	3132.9	3489.7	9.8998	7.1338	3132.6	3489.3	9.1566
550	37.988	3217.2	3597.1	10.034	7.5957	3217.0	3596.8	9.2913
600	40.296	3303.3	3706.3	10.163	8.0576	3303.1	3706.0	9.4201
650	42.603	3391.2	3817.2	10.287	8.5195	3391.0	3816.9	9.5436
700	44.911	3480.8	3929.9	10.406	8.9812	3480.6	3929.7	9.6625
750	47.219	3572.2	4044.4	10.520	9.4430	3572.0	4044.2	9.7773
800	49.527	3665.3	4160.6	10.631	9.9047	3665.2	4160.4	9.8882
850	51.835	3760.3	4278.6	10.739	10.366	3760.1	4278.5	9.9957
900	54.142	3856.9	4398.3	10.843	10.828	3856.8	4398.2	10.100
950	56.450	3955.2	4519.7	10.944	11.290	3955.1	4519.6	10.201
1000	58.758	4055.2	4642.8	11.043	11.751	4055.1	4642.7	10.300

$p = 0.10 \text{ MPa } (T_{sat} = 99.6^\circ\text{C})$					$p = 0.20 \text{ MPa } (T_{sat} = 120.2^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	1.6939	2505.6	2674.9	7.3588	0.8857	2529.1	2706.2	7.1269
100	1.6959	2506.2	2675.8	7.3610				
150	1.9367	2582.9	2776.6	7.6148	0.9599	2577.1	2769.1	7.2810
200	2.1724	2658.2	2875.5	7.8356	1.0805	2654.6	2870.7	7.5081
250	2.4062	2733.9	2974.5	8.0346	1.1989	2731.4	2971.2	7.7100
300	2.6388	2810.6	3074.5	8.2172	1.3162	2808.8	3072.1	7.8941
350	2.8710	2888.7	3175.8	8.3866	1.4330	2887.3	3173.9	8.0644
400	3.1027	2968.3	3278.6	8.5452	1.5493	2967.1	3277.0	8.2236
450	3.3342	3049.4	3382.8	8.6946	1.6655	3048.5	3381.6	8.3734
500	3.5655	3132.2	3488.7	8.8361	1.7814	3131.4	3487.7	8.5152
550	3.7968	3216.6	3596.3	8.9709	1.8973	3215.9	3595.4	8.6502
600	4.0279	3302.8	3705.6	9.0998	2.0130	3302.2	3704.8	8.7792
650	4.2590	3390.7	3816.6	9.2234	2.1287	3390.2	3815.9	8.9030
700	4.4900	3480.4	3929.4	9.3424	2.2443	3479.9	3928.8	9.0220
750	4.7209	3571.8	4043.9	9.4572	2.3599	3571.4	4043.4	9.1369
800	4.9519	3665.0	4160.2	9.5681	2.4755	3664.7	4159.8	9.2479
850	5.1828	3760.0	4278.2	9.6757	2.5910	3759.6	4277.8	9.3555
900	5.4137	3856.6	4398.0	9.7800	2.7066	3856.3	4397.6	9.4598
950	5.6446	3955.0	4519.5	9.8813	2.8221	3954.7	4519.1	9.5612
1000	5.8754	4055.0	4642.6	9.9800	2.9375	4054.8	4642.3	9.6599

$p = 0.30 \text{ MPa } (T_{sat} = 133.5^\circ\text{C})$					$p = 0.40 \text{ MPa } (T_{sat} = 143.6^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.6058	2543.2	2724.9	6.9916	0.4624	2553.1	2738.1	6.8955
150	0.6340	2571.0	2761.2	7.0791	0.4709	2564.4	2752.8	6.9306
200	0.7164	2651.0	2865.9	7.3131	0.5343	2647.2	2860.9	7.1723
250	0.7964	2728.9	2967.9	7.5180	0.5952	2726.4	2964.5	7.3804
300	0.8753	2807.0	3069.6	7.7037	0.6549	2805.1	3067.1	7.5677
350	0.9536	2885.9	3172.0	7.8750	0.7140	2884.4	3170.0	7.7399
400	1.0315	2966.0	3275.5	8.0347	0.7726	2964.9	3273.9	7.9002
450	1.1092	3047.5	3380.3	8.1849	0.8311	3046.6	3379.0	8.0508
500	1.1867	3130.6	3486.6	8.3271	0.8894	3129.8	3485.5	8.1933
550	1.2641	3215.3	3594.5	8.4623	0.9475	3214.6	3593.6	8.3287
600	1.3414	3301.6	3704.0	8.5914	1.0056	3301.0	3703.2	8.4580
650	1.4186	3389.7	3815.3	8.7153	1.0636	3389.1	3814.6	8.5820
700	1.4958	3479.5	3928.2	8.8344	1.1215	3479.0	3927.6	8.7012
750	1.5729	3571.0	4042.9	8.9494	1.1794	3570.6	4042.4	8.8162
800	1.6500	3664.3	4159.3	9.0604	1.2373	3663.9	4158.8	8.9273
850	1.7271	3759.3	4277.4	9.1680	1.2951	3759.0	4277.0	9.0350
900	1.8042	3856.0	4397.3	9.2724	1.3530	3855.7	4396.9	9.1394
950	1.8812	3954.4	4518.8	9.3739	1.4108	3954.2	4518.5	9.2409
1000	1.9582	4054.5	4642.0	9.4726	1.4686	4054.3	4641.7	9.3396

$p = 0.50 \text{ MPa } (T_{sat} = 151.8^\circ\text{C})$					$p = 0.60 \text{ MPa } (T_{sat} = 158.8^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.3748	2560.7	2748.1	6.8207	0.3156	2566.8	2756.1	6.7592
200	0.4250	2643.3	2855.8	7.0610	0.3521	2639.3	2850.6	6.9683
250	0.4744	2723.8	2961.0	7.2724	0.3939	2721.2	2957.6	7.1832
300	0.5226	2803.2	3064.6	7.4614	0.4344	2801.4	3062.0	7.3740
350	0.5702	2883.0	3168.1	7.6346	0.4743	2881.6	3166.1	7.5481
400	0.6173	2963.7	3272.3	7.7955	0.5137	2962.5	3270.8	7.7097
450	0.6642	3045.6	3377.7	7.9465	0.5530	3044.7	3376.5	7.8611
500	0.7109	3129.0	3484.5	8.0892	0.5920	3128.2	3483.4	8.0041
550	0.7576	3213.9	3592.7	8.2249	0.6309	3213.2	3591.8	8.1399
600	0.8041	3300.4	3702.5	8.3543	0.6698	3299.8	3701.7	8.2695
650	0.8506	3388.6	3813.9	8.4784	0.7085	3388.1	3813.2	8.3937
700	0.8970	3478.5	3927.0	8.5977	0.7473	3478.1	3926.4	8.5131
750	0.9433	3570.2	4041.8	8.7128	0.7859	3569.8	4041.3	8.6283
800	0.9897	3663.6	4158.4	8.8240	0.8246	3663.2	4157.9	8.7395
850	1.0360	3758.6	4276.6	8.9317	0.8632	3758.3	4276.2	8.8472
900	1.0823	3855.4	4396.6	9.0362	0.9018	3855.1	4396.2	8.9518
950	1.1285	3953.9	4518.2	9.1377	0.9404	3953.6	4517.8	9.0533
1000	1.1748	4054.0	4641.4	9.2364	0.9789	4053.7	4641.1	9.1521

$p = 0.80 \text{ MPa } (T_{sat} = 170.4^\circ\text{C})$					$p = 1.0 \text{ MPa } (T_{sat} = 179.9^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.2403	2576.0	2768.3	6.6616	0.1944	2582.7	2777.1	6.5850
200	0.2609	2631.0	2839.7	6.8176	0.2060	2622.2	2828.3	6.6955
250	0.2932	2715.9	2950.4	7.0401	0.2328	2710.4	2943.1	6.9265
300	0.3242	2797.5	3056.9	7.2345	0.2580	2793.6	3051.6	7.1246
350	0.3544	2878.6	3162.2	7.4106	0.2825	2875.7	3158.2	7.3029
400	0.3843	2960.2	3267.6	7.5734	0.3066	2957.9	3264.5	7.4669
450	0.4139	3042.8	3373.9	7.7257	0.3305	3040.9	3371.3	7.6200
500	0.4433	3126.6	3481.3	7.8692	0.3541	3125.0	3479.1	7.7641
550	0.4726	3211.9	3590.0	8.0054	0.3777	3210.5	3588.1	7.9008
600	0.5019	3298.7	3700.1	8.1354	0.4011	3297.5	3698.6	8.0310
650	0.5310	3387.1	3811.9	8.2598	0.4245	3386.0	3810.5	8.1557
700	0.5601	3477.2	3925.3	8.3794	0.4478	3476.2	3924.1	8.2755
750	0.5892	3569.0	4040.3	8.4947	0.4711	3568.1	4039.3	8.3909
800	0.6182	3662.4	4157.0	8.6061	0.4944	3661.7	4156.1	8.5024
850	0.6472	3757.6	4275.4	8.7139	0.5176	3757.0	4274.6	8.6103
900	0.6762	3854.5	4395.5	8.8185	0.5408	3853.9	4394.8	8.7150
950	0.7052	3953.1	4517.2	8.9201	0.5640	3952.5	4516.5	8.8166
1000	0.7341	4053.2	4640.5	9.0189	0.5872	4052.7	4639.9	8.9155

$p = 1.2 \text{ MPa } (T_{sat} = 188.0^\circ\text{C})$					$p = 1.4 \text{ MPa } (T_{sat} = 195.0^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.1633	2587.8	2783.7	6.5217	0.1408	2591.8	2788.8	6.4675
200	0.1693	2612.9	2816.1	6.5909	0.1430	2602.7	2803.0	6.4975
250	0.1924	2704.7	2935.6	6.8313	0.1636	2698.9	2927.9	6.7488
300	0.2139	2789.7	3046.3	7.0335	0.1823	2785.7	3040.9	6.9552
350	0.2346	2872.7	3154.2	7.2139	0.2003	2869.7	3150.1	7.1379
400	0.2548	2955.5	3261.3	7.3793	0.2178	2953.1	3258.1	7.3046
450	0.2748	3038.9	3368.7	7.5332	0.2351	3037.0	3366.1	7.4594
500	0.2946	3123.4	3476.9	7.6779	0.2522	3121.8	3474.8	7.6047
550	0.3143	3209.1	3586.3	7.8150	0.2691	3207.7	3584.5	7.7422
600	0.3339	3296.3	3697.0	7.9455	0.2860	3295.1	3695.4	7.8730
650	0.3535	3385.0	3809.2	8.0704	0.3028	3384.0	3807.8	7.9982
700	0.3730	3475.3	3922.9	8.1904	0.3195	3474.4	3921.7	8.1183
750	0.3924	3567.3	4038.2	8.3060	0.3362	3566.5	4037.2	8.2340
800	0.4118	3661.0	4155.2	8.4176	0.3529	3660.2	4154.3	8.3457
850	0.4312	3756.3	4273.8	8.5256	0.3695	3755.6	4273.0	8.4538
900	0.4506	3853.3	4394.0	8.6303	0.3861	3852.7	4393.3	8.5587
950	0.4699	3952.0	4515.9	8.7320	0.4027	3951.4	4515.2	8.6604
1000	0.4893	4052.2	4639.4	8.8310	0.4193	4051.7	4638.8	8.7594

$p = 1.6 \text{ MPa } (T_{sat} = 201.4^\circ\text{C})$					$p = 1.8 \text{ MPa } (T_{sat} = 207.1^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.1237	2594.8	2792.8	6.4199	0.1104	2597.2	2795.9	6.3775
250	0.1419	2692.9	2919.9	6.6753	0.1250	2686.7	2911.7	6.6087
300	0.1587	2781.6	3035.4	6.8863	0.1403	2777.4	3029.9	6.8246
350	0.1746	2866.6	3146.0	7.0713	0.1546	2863.6	3141.8	7.0120
400	0.1901	2950.7	3254.9	7.2394	0.1685	2948.3	3251.6	7.1814
450	0.2053	3035.0	3363.5	7.3950	0.1821	3033.1	3360.9	7.3380
500	0.2203	3120.1	3472.6	7.5409	0.1955	3118.5	3470.4	7.4845
550	0.2352	3206.3	3582.6	7.6788	0.2088	3205.0	3580.8	7.6228
600	0.2500	3293.9	3693.9	7.8100	0.2220	3292.7	3692.3	7.7543
650	0.2647	3382.9	3806.5	7.9354	0.2351	3381.9	3805.1	7.8799
700	0.2794	3473.5	3920.5	8.0557	0.2482	3472.6	3919.4	8.0004
750	0.2940	3565.7	4036.1	8.1716	0.2613	3564.9	4035.1	8.1164
800	0.3087	3659.5	4153.3	8.2834	0.2743	3658.8	4152.4	8.2284
850	0.3232	3755.0	4272.2	8.3916	0.2872	3754.3	4271.3	8.3367
900	0.3378	3852.1	4392.6	8.4965	0.3002	3851.5	4391.9	8.4416
950	0.3523	3950.9	4514.6	8.5984	0.3131	3950.3	4514.0	8.5435
1000	0.3669	4051.2	4638.2	8.6974	0.3261	4050.7	4637.6	8.6426

$p = 2.0 \text{ MPa } (T_{sat} = 212.4^\circ\text{C})$					$p = 2.5 \text{ MPa } (T_{sat} = 224.0^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.0996	2599.1	2798.3	6.3390	0.0799	2602.1	2801.9	6.2558
250	0.1115	2680.2	2903.2	6.5475	0.0871	2663.3	2880.9	6.4107
300	0.1255	2773.2	3024.2	6.7684	0.0989	2762.2	3009.6	6.6459
350	0.1386	2860.5	3137.7	6.9583	0.1098	2852.5	3127.0	6.8424
400	0.1512	2945.9	3248.3	7.1292	0.1201	2939.8	3240.1	7.0170
450	0.1635	3031.1	3358.2	7.2866	0.1302	3026.2	3351.6	7.1767
500	0.1757	3116.9	3468.2	7.4337	0.1400	3112.8	3462.7	7.3254
550	0.1877	3203.6	3579.0	7.5725	0.1497	3200.1	3574.3	7.4653
600	0.1996	3291.5	3690.7	7.7043	0.1593	3288.5	3686.8	7.5979
650	0.2115	3380.8	3803.8	7.8302	0.1689	3378.2	3800.4	7.7243
700	0.2233	3471.6	3918.2	7.9509	0.1784	3469.3	3915.2	7.8455
750	0.2350	3564.0	4034.1	8.0670	0.1878	3562.0	4031.5	7.9620
800	0.2467	3658.0	4151.5	8.1790	0.1972	3656.2	4149.2	8.0743
850	0.2584	3753.6	4270.5	8.2874	0.2066	3752.0	4268.5	8.1830
900	0.2701	3850.9	4391.1	8.3925	0.2160	3849.4	4389.3	8.2882
950	0.2818	3949.8	4513.3	8.4945	0.2253	3948.4	4511.7	8.3904
1000	0.2934	4050.2	4637.0	8.5936	0.2347	4048.9	4635.6	8.4896

$p = 3.0 \text{ MPa } (T_{sat} = 233.9^\circ\text{C})$					$p = 3.5 \text{ MPa } (T_{sat} = 242.6^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.0667	2603.2	2803.2	6.1856	0.0571	2602.9	2802.6	6.1243
250	0.0706	2644.7	2856.5	6.2893	0.0588	2624.0	2829.7	6.1764
300	0.0812	2750.8	2994.3	6.5412	0.0685	2738.8	2978.4	6.4484
350	0.0906	2844.4	3116.1	6.7449	0.0768	2836.0	3104.8	6.6601
400	0.0994	2933.5	3231.7	6.9234	0.0846	2927.2	3223.2	6.8427
450	0.1079	3021.2	3344.8	7.0856	0.0920	3016.1	3338.0	7.0074
500	0.1162	3108.6	3457.2	7.2359	0.0992	3104.5	3451.6	7.1593
550	0.1244	3196.6	3569.7	7.3768	0.1063	3193.1	3565.0	7.3014
600	0.1325	3285.5	3682.8	7.5103	0.1133	3282.5	3678.9	7.4356
650	0.1405	3375.6	3796.9	7.6373	0.1202	3372.9	3793.5	7.5633
700	0.1484	3467.0	3912.2	7.7590	0.1270	3464.7	3909.3	7.6854
750	0.1563	3559.9	4028.9	7.8758	0.1338	3557.8	4026.3	7.8027
800	0.1642	3654.3	4146.9	7.9885	0.1406	3652.5	4144.6	7.9156
850	0.1721	3750.3	4266.5	8.0973	0.1474	3748.6	4264.4	8.0247
900	0.1799	3847.9	4387.5	8.2028	0.1541	3846.4	4385.7	8.1303
950	0.1877	3947.0	4510.1	8.3051	0.1608	3945.6	4508.4	8.2328
1000	0.1955	4047.7	4634.1	8.4045	0.1675	4046.4	4632.7	8.3324

$p = 4.0 \text{ MPa } (T_{sat} = 250.4^\circ\text{C})$					$p = 4.5 \text{ MPa } (T_{sat} = 257.4^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.0498	2601.7	2800.8	6.0696	0.0441	2599.7	2797.9	6.0197
300	0.0589	2726.2	2961.7	6.3639	0.0514	2713.0	2944.2	6.2854
350	0.0665	2827.4	3093.3	6.5843	0.0584	2818.6	3081.5	6.5153
400	0.0734	2920.7	3214.5	6.7714	0.0648	2914.2	3205.6	6.7070
450	0.0800	3011.0	3331.2	6.9386	0.0708	3005.8	3324.2	6.8770
500	0.0864	3100.3	3446.0	7.0922	0.0765	3096.0	3440.4	7.0323
550	0.0927	3189.5	3560.3	7.2355	0.0821	3186.0	3555.6	7.1767
600	0.0989	3279.4	3674.9	7.3705	0.0877	3276.4	3670.9	7.3127
650	0.1049	3370.3	3790.1	7.4988	0.0931	3367.7	3786.6	7.4416
700	0.1110	3462.4	3906.3	7.6214	0.0985	3460.0	3903.3	7.5646
750	0.1170	3555.8	4023.6	7.7390	0.1039	3553.7	4021.0	7.6826
800	0.1229	3650.6	4142.3	7.8523	0.1092	3648.8	4140.0	7.7962
850	0.1289	3747.0	4262.4	7.9616	0.1145	3745.3	4260.3	7.9057
900	0.1348	3844.8	4383.9	8.0674	0.1197	3843.3	4382.1	8.0118
950	0.1407	3944.2	4506.8	8.1701	0.1250	3942.8	4505.2	8.1146
1000	0.1465	4045.1	4631.2	8.2697	0.1302	4043.9	4629.8	8.2144

$p = 5.0 \text{ MPa } (T_{sat} = 263.9^\circ\text{C})$					$p = 6.0 \text{ MPa } (T_{sat} = 275.6^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.0394	2597.0	2794.2	5.9737	0.0324	2589.9	2784.6	5.8901
300	0.0453	2699.0	2925.7	6.2110	0.0362	2668.4	2885.5	6.0703
350	0.0520	2809.5	3069.3	6.4516	0.0423	2790.4	3043.9	6.3357
400	0.0578	2907.5	3196.7	6.6483	0.0474	2893.7	3178.2	6.5432
450	0.0633	3000.6	3317.2	6.8210	0.0522	2989.9	3302.9	6.7219
500	0.0686	3091.7	3434.7	6.9781	0.0567	3083.1	3423.1	6.8826
550	0.0737	3182.4	3550.9	7.1237	0.0610	3175.2	3541.3	7.0307
600	0.0787	3273.3	3666.8	7.2605	0.0653	3267.2	3658.7	7.1693
650	0.0836	3365.0	3783.2	7.3901	0.0694	3359.6	3776.2	7.3001
700	0.0885	3457.7	3900.3	7.5136	0.0735	3453.0	3894.3	7.4246
750	0.0934	3551.6	4018.4	7.6320	0.0776	3547.5	4013.2	7.5438
800	0.0982	3646.9	4137.7	7.7458	0.0816	3643.2	4133.1	7.6582
850	0.1029	3743.6	4258.3	7.8556	0.0857	3740.3	4254.2	7.7685
900	0.1077	3841.8	4380.2	7.9618	0.0896	3838.8	4376.6	7.8751
950	0.1124	3941.5	4503.6	8.0648	0.0936	3938.7	4500.3	7.9784
1000	0.1172	4042.6	4628.3	8.1648	0.0976	4040.1	4625.4	8.0786

$p = 7.0 \text{ MPa } (T_{sat} = 285.8^\circ\text{C})$					$p = 8.0 \text{ MPa } (T_{sat} = 295.0^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.0274	2581.0	2772.6	5.8148	0.0235	2570.5	2758.7	5.7450
300	0.0295	2633.5	2839.9	5.9337	0.0243	2592.3	2786.5	5.7937
350	0.0353	2770.1	3016.9	6.2304	0.0300	2748.3	2988.1	6.1321
400	0.0400	2879.5	3159.2	6.4502	0.0343	2864.6	3139.4	6.3658
450	0.0442	2979.0	3288.3	6.6353	0.0382	2967.8	3273.3	6.5579
500	0.0482	3074.3	3411.4	6.8000	0.0418	3065.4	3399.5	6.7266
550	0.0520	3167.9	3531.6	6.9506	0.0452	3160.5	3521.8	6.8799
600	0.0557	3260.9	3650.6	7.0910	0.0485	3254.7	3642.4	7.0221
650	0.0593	3354.3	3769.3	7.2231	0.0517	3348.9	3762.3	7.1556
700	0.0629	3448.3	3888.2	7.3486	0.0548	3443.6	3882.2	7.2821
750	0.0664	3543.3	4007.9	7.4685	0.0579	3539.1	4002.6	7.4028
800	0.0699	3639.5	4128.4	7.5836	0.0610	3635.7	4123.8	7.5184
850	0.0733	3736.9	4250.1	7.6944	0.0641	3733.5	4246.0	7.6297
900	0.0768	3835.7	4373.0	7.8014	0.0671	3832.6	4369.3	7.7371
950	0.0802	3935.9	4497.1	7.9050	0.0701	3933.1	4493.8	7.8411
1000	0.0836	4037.5	4622.5	8.0055	0.0731	4035.0	4619.6	7.9419

$p = 9.0 \text{ MPa } (T_{sat} = 303.3^\circ\text{C})$					$p = 10.0 \text{ MPa } (T_{sat} = 311.0^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.0205	2558.5	2742.9	5.6791	0.0180	2545.2	2725.5	5.6160
350	0.0258	2724.9	2957.3	6.0380	0.0224	2699.6	2924.0	5.9459
400	0.0300	2849.2	3118.8	6.2876	0.0264	2833.1	3097.4	6.2141
450	0.0335	2956.3	3258.0	6.4872	0.0298	2944.5	3242.3	6.4219
500	0.0368	3056.3	3387.4	6.6603	0.0328	3047.0	3375.1	6.5995
550	0.0399	3153.0	3512.0	6.8164	0.0357	3145.4	3502.0	6.7585
600	0.0429	3248.4	3634.1	6.9605	0.0384	3242.0	3625.8	6.9045
650	0.0458	3343.4	3755.2	7.0953	0.0410	3337.9	3748.1	7.0408
700	0.0486	3438.8	3876.1	7.2229	0.0436	3434.0	3870.0	7.1693
750	0.0514	3534.9	3997.3	7.3443	0.0461	3530.7	3992.0	7.2916
800	0.0541	3632.0	4119.1	7.4606	0.0486	3628.2	4114.5	7.4085
850	0.0569	3730.2	4241.9	7.5724	0.0511	3726.8	4237.8	7.5207
900	0.0596	3829.6	4365.7	7.6802	0.0535	3826.5	4362.0	7.6290
950	0.0622	3930.3	4490.6	7.7844	0.0560	3927.5	4487.3	7.7335
1000	0.0649	4032.4	4616.7	7.8855	0.0584	4029.9	4613.8	7.8349

$p = 12.5 \text{ MPa } (T_{sat} = 327.8^\circ\text{C})$					$p = 15.0 \text{ MPa } (T_{sat} = 342.16^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.0135	2505.6	2674.3	5.4638	0.0103	2455.6	2610.7	5.3106
350	0.0161	2624.8	2826.6	5.7130	0.0115	2520.9	2693.1	5.4437
400	0.0200	2789.6	3040.0	6.0433	0.0157	2740.6	2975.7	5.8819
450	0.0230	2913.7	3201.4	6.2749	0.0185	2880.7	3157.9	6.1434
500	0.0256	3023.2	3343.6	6.4650	0.0208	2998.4	3310.8	6.3480
550	0.0280	3126.1	3476.5	6.6317	0.0229	3106.2	3450.4	6.5230
600	0.0303	3225.8	3604.6	6.7828	0.0249	3209.3	3583.1	6.6796
650	0.0325	3324.1	3730.2	6.9227	0.0268	3310.1	3712.1	6.8233
700	0.0346	3422.0	3854.6	7.0539	0.0286	3409.8	3839.1	6.9572
750	0.0367	3520.1	3978.6	7.1782	0.0304	3509.4	3965.2	7.0836
800	0.0387	3618.7	4102.8	7.2967	0.0321	3609.2	4091.1	7.2037
850	0.0407	3718.3	4227.5	7.4102	0.0338	3709.8	4217.1	7.3185
900	0.0427	3818.9	4352.9	7.5194	0.0355	3811.2	4343.7	7.4288
950	0.0447	3920.6	4479.2	7.6249	0.0372	3913.6	4471.0	7.5350
1000	0.0466	4023.5	4606.5	7.7269	0.0388	4017.1	4599.2	7.6378

$p = 17.5 \text{ MPa } (T_{sat} = 354.7^\circ\text{C})$					$p = 20.0 \text{ MPa } (T_{sat} = 365.8^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	0.0079	2390.5	2529.3	5.1431	0.0059	2295.0	2412.3	4.9314
400	0.0125	2684.3	2902.4	5.7211	0.0100	2617.9	2816.9	5.5525
450	0.0152	2845.4	3111.4	6.0212	0.0127	2807.2	3061.7	5.9043
500	0.0174	2972.4	3276.7	6.2424	0.0148	2945.3	3241.2	6.1446
550	0.0193	3085.8	3423.6	6.4266	0.0166	3064.7	3396.1	6.3389
600	0.0211	3192.5	3561.3	6.5890	0.0182	3175.3	3539.0	6.5075
650	0.0227	3295.8	3693.8	6.7366	0.0197	3281.4	3675.3	6.6593
700	0.0243	3397.5	3823.5	6.8734	0.0211	3385.1	3807.8	6.7990
750	0.0259	3498.6	3951.7	7.0019	0.0225	3487.7	3938.1	6.9297
800	0.0274	3599.7	4079.3	7.1236	0.0239	3590.1	4067.5	7.0531
850	0.0289	3701.2	4206.8	7.2398	0.0252	3692.6	4196.4	7.1705
900	0.0303	3803.4	4334.5	7.3511	0.0265	3795.7	4325.4	7.2829
950	0.0318	3906.6	4462.9	7.4582	0.0278	3899.5	4454.7	7.3909
1000	0.0332	4010.7	4592.0	7.5616	0.0290	4004.3	4584.7	7.4950

$p = 25.0 \text{ MPa (supercritical)}$					$p = 30.0 \text{ MPa (supercritical)}$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
400	0.00600	2428.5	2578.6	5.1400	0.00280	2068.9	2152.8	4.4757
450	0.00918	2721.2	2950.6	5.6759	0.00674	2618.9	2821.0	5.4421
500	0.01114	2887.3	3165.9	5.9642	0.00869	2824.0	3084.7	5.7956
550	0.01274	3020.8	3339.2	6.1816	0.01018	2974.5	3279.7	6.0402
600	0.01414	3140.0	3493.5	6.3637	0.01145	3103.4	3446.7	6.2373
650	0.01543	3251.9	3637.7	6.5242	0.01259	3221.7	3599.4	6.4074
700	0.01664	3359.9	3776.0	6.6702	0.01365	3334.3	3743.9	6.5598
750	0.01780	3465.8	3910.9	6.8054	0.01466	3443.6	3883.4	6.6997
800	0.01892	3570.7	4043.8	6.9322	0.01563	3551.2	4020.0	6.8300
850	0.02001	3675.4	4175.6	7.0523	0.01656	3658.0	4154.9	6.9529
900	0.02108	3780.2	4307.1	7.1668	0.01747	3764.6	4288.8	7.0695
1000	0.02315	3991.5	4570.2	7.3820	0.01924	3978.6	4555.8	7.2880

$p = 35.0 \text{ MPa (supercritical)}$					$p = 40.0 \text{ MPa (supercritical)}$			
Temp. T $^\circ\text{C}$	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Volume v m^3/kg	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
400	0.00211	1914.9	1988.6	4.2143	0.00191	1854.9	1931.4	4.1145
450	0.00496	2497.5	2671.0	5.1945	0.00369	2364.2	2511.8	4.9448
500	0.00693	2755.3	2997.9	5.6331	0.00562	2681.6	2906.5	5.4744
550	0.00835	2925.8	3218.0	5.9092	0.00698	2875.0	3154.4	5.7857
600	0.00952	3065.6	3398.9	6.1228	0.00809	3026.8	3350.4	6.0170
650	0.01057	3190.9	3560.7	6.3030	0.00905	3159.5	3521.6	6.2078
700	0.01152	3308.3	3711.6	6.4622	0.00993	3282.0	3679.1	6.3740
750	0.01242	3421.2	3855.9	6.6069	0.01075	3398.6	3828.4	6.5236
800	0.01328	3531.5	3996.3	6.7409	0.01152	3511.8	3972.6	6.6612
850	0.01410	3640.5	4134.2	6.8665	0.01226	3623.1	4113.6	6.7896
900	0.01490	3748.9	4270.6	6.9853	0.01298	3733.3	4252.5	6.9106
1000	0.01645	3965.8	4541.5	7.2069	0.01436	3952.9	4527.3	7.1355

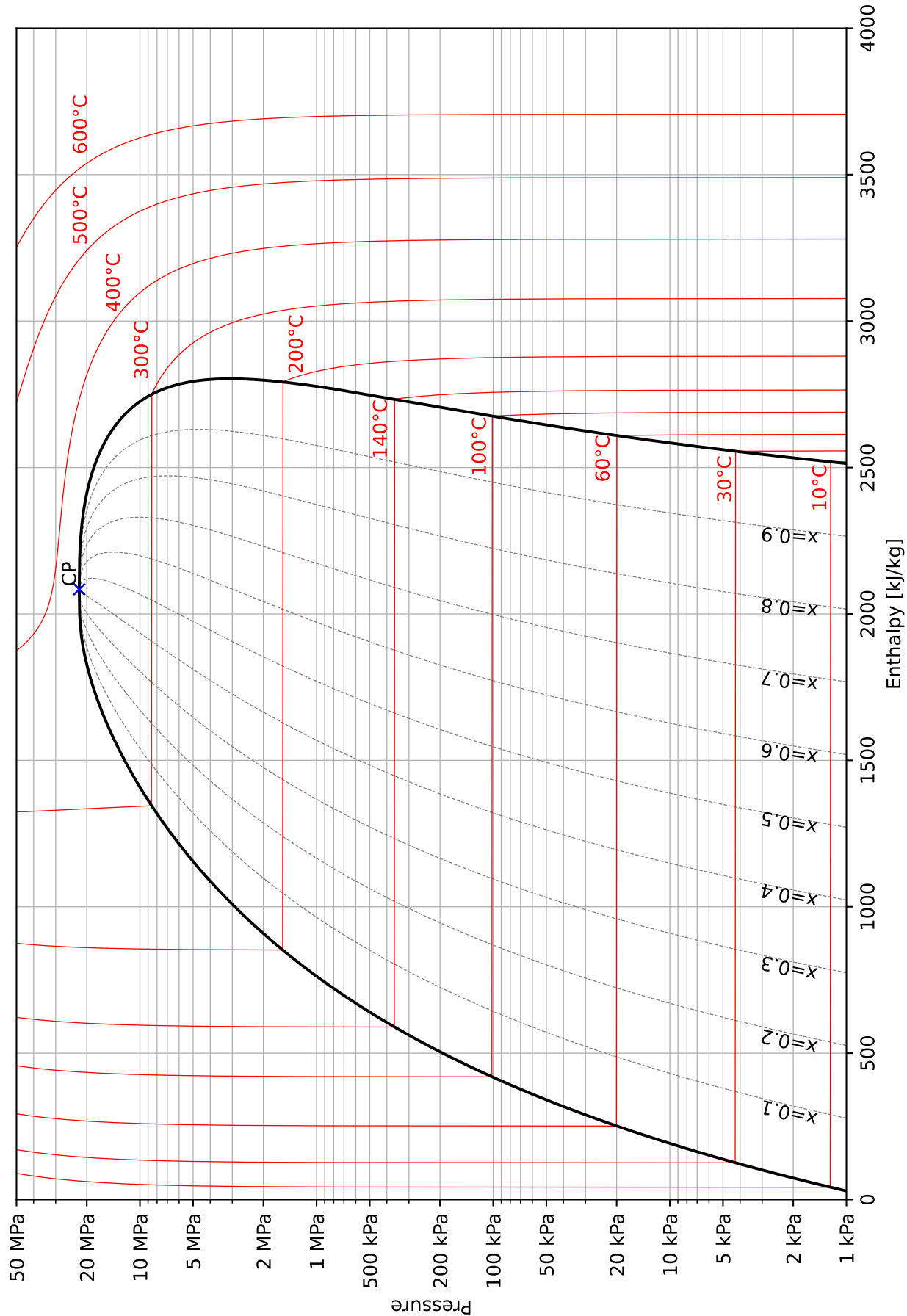
A.4 Compressed Liquid Properties for Water

$p = 5.0 \text{ MPa } (T_{sat} = 263.9^\circ\text{C})$					$p = 10 \text{ MPa } (T_{sat} = 311.0^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Density ρ kg/m^3	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Density ρ kg/m^3	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
20	1000.4	83.609	88.607	0.29543	1002.7	83.308	93.281	0.29435
40	994.36	166.92	171.95	0.57046	996.52	166.33	176.36	0.56851
60	985.33	250.29	255.36	0.82865	987.48	249.42	259.55	0.8260
80	973.97	333.82	338.95	1.0723	976.17	332.69	342.94	1.0691
100	960.63	417.64	422.85	1.3034	962.93	416.23	426.62	1.2996
120	945.49	501.90	507.19	1.5236	947.94	500.18	510.73	1.5191
140	928.63	586.79	592.18	1.7344	931.28	584.71	595.45	1.7293
160	910.05	672.55	678.04	1.9374	912.95	670.06	681.01	1.9315
180	889.65	759.46	765.08	2.1338	892.88	756.48	767.68	2.1271
200	867.26	847.91	853.68	2.3251	870.94	844.31	855.80	2.3174
220	842.58	938.39	944.32	2.5127	846.84	934.00	945.81	2.5037
240	815.10	1031.6	1037.7	2.6983	820.18	1026.1	1038.3	2.6876
260	784.03	1128.5	1134.9	2.8841	790.30	1121.6	1134.3	2.8710
280					756.07	1221.8	1235.0	3.0565
300					715.29	1329.4	1343.3	3.2488
Sat.	777.37	1148.2	1154.6	2.9210	688.42	1393.5	1408.1	3.3606

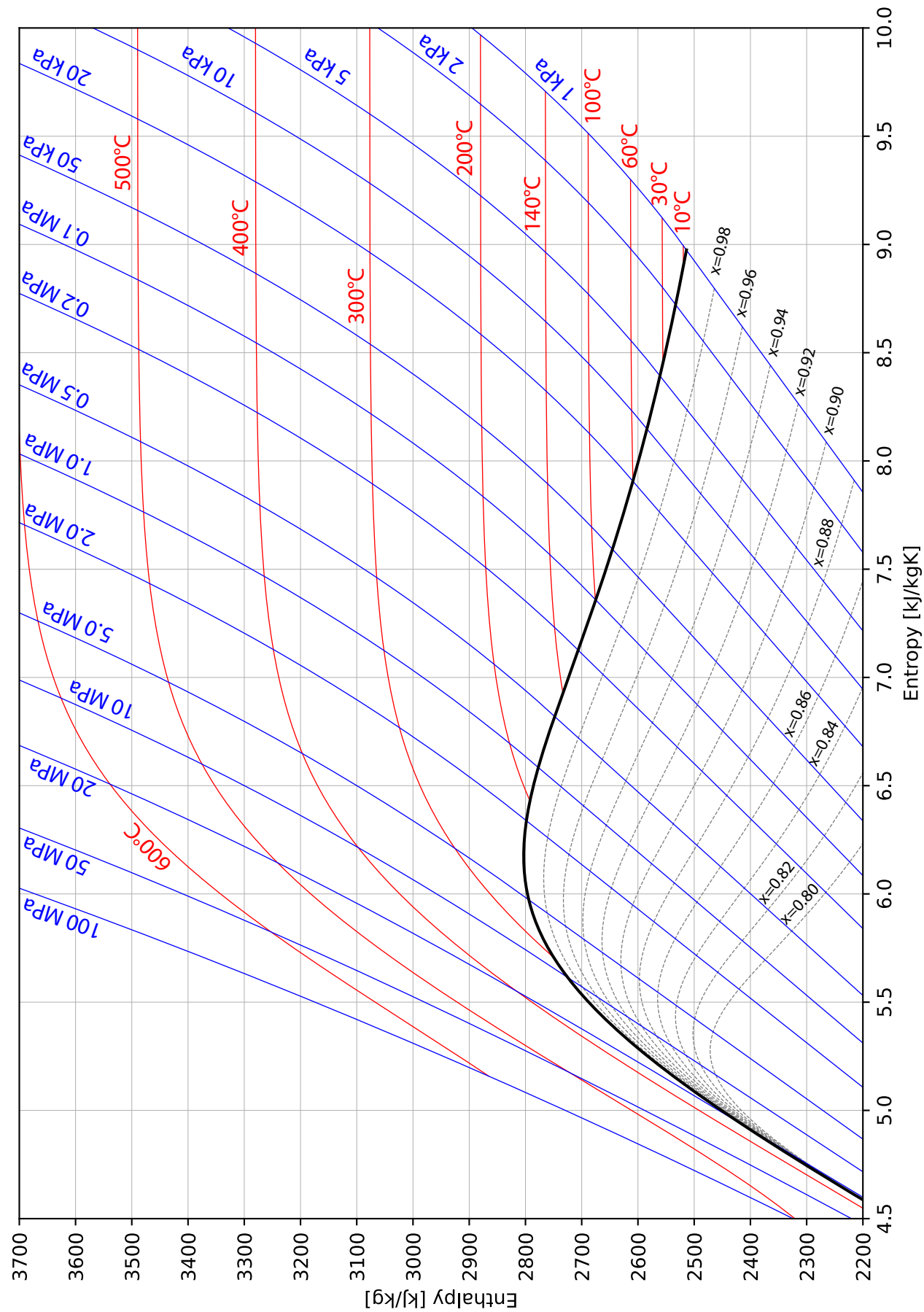
$p = 15.0 \text{ MPa } (T_{sat} = 342.2^\circ\text{C})$					$p = 20.0 \text{ MPa } (T_{sat} = 365.8^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Density ρ kg/m^3	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Density ρ kg/m^3	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
20	1004.9	83.007	97.934	0.29323	1007.1	82.708	102.57	0.29207
40	998.65	165.75	180.77	0.56656	1000.8	165.17	185.16	0.56461
60	989.60	248.58	263.74	0.8234	991.71	247.75	267.92	0.8208
80	978.35	331.59	346.92	1.0659	980.49	330.50	350.90	1.0627
100	965.20	414.85	430.39	1.2958	967.44	413.50	434.17	1.2920
120	950.35	498.49	514.28	1.5148	952.72	496.85	517.84	1.5105
140	933.87	582.69	598.75	1.7243	936.42	580.71	602.07	1.7194
160	915.79	667.63	684.01	1.9259	918.57	665.27	687.05	1.9203
180	896.04	753.58	770.32	2.1206	899.12	750.77	773.02	2.1143
200	874.50	840.84	857.99	2.3100	877.97	837.49	860.27	2.3027
220	850.95	929.80	947.43	2.4951	854.91	925.77	949.16	2.4867
240	825.03	1021.0	1039.2	2.6774	829.67	1016.1	1040.2	2.6676
260	796.20	1115.1	1134.0	2.8586	801.78	1109.0	1134.0	2.8469
280	763.58	1213.4	1233.0	3.0409	770.52	1205.5	1231.5	3.0265
300	725.55	1317.6	1338.3	3.2279	734.71	1307.1	1334.4	3.2091
320	678.77	1431.9	1454.0	3.4263	692.06	1416.6	1445.5	3.3996
340	613.07	1567.9	1592.4	3.6555	637.23	1540.2	1571.6	3.6086
360					548.01	1703.6	1740.1	3.8787
Sat.	603.52	1585.3	1610.2	3.6846	490.19	1786.4	1827.2	4.0156

Temp. T °C	$p = 25.0$ MPa (supercritical)				$p = 30.0$ MPa (supercritical)			
	Density	Energy	Enthalpy	Entropy	Density	Energy	Enthalpy	Entropy
	ρ kg/m ³	u kJ/kg	h kJ/kg	s kJ/kgK	ρ kg/m ³	u kJ/kg	h kJ/kg	s kJ/kgK
20	1009.3	82.409	107.18	0.29089	1011.5	82.112	111.77	0.28968
40	1002.9	164.61	189.53	0.56265	1004.9	164.05	193.90	0.56069
60	993.79	246.94	272.09	0.81821	995.84	246.14	276.26	0.81564
80	982.61	329.44	354.88	1.0595	984.71	328.40	358.86	1.0564
100	969.65	412.17	437.95	1.2883	971.82	410.87	441.74	1.2847
120	955.05	495.24	521.41	1.5062	957.35	493.66	525.00	1.5020
140	938.93	578.78	605.41	1.7146	941.39	576.89	608.76	1.7098
160	921.30	662.98	690.11	1.9148	923.97	660.74	693.21	1.9094
180	902.13	748.05	775.76	2.1081	905.07	745.40	778.54	2.1020
200	881.33	834.24	862.61	2.2956	884.62	831.10	865.02	2.2888
220	858.75	921.88	951.00	2.4786	862.46	918.14	952.93	2.4707
240	834.12	1011.4	1041.3	2.6582	838.40	1006.9	1042.7	2.6491
260	807.06	1103.2	1134.2	2.8357	812.10	1097.8	1134.7	2.8250
280	777.01	1198.3	1230.5	3.0129	783.10	1191.5	1229.8	3.0001
300	743.02	1297.6	1331.3	3.1919	750.66	1288.9	1328.9	3.1760
320	703.49	1403.4	1438.9	3.3764	713.58	1391.6	1433.7	3.3557
340	655.13	1519.4	1557.5	3.5731	669.70	1502.3	1547.1	3.5438
360	589.31	1656.2	1698.6	3.7993	614.39	1626.7	1675.6	3.7498

A.5 Steam Pressure-Enthalpy (p - h) Diagram



A.6 Steam Enthalpy-Entropy (*h-s*) Diagram



Appendix B

R134a (Tetraflouroethane) Tables

Unless otherwise indicated, all data was sourced from the [NIST Chemistry WebBook](#), Accessed Dec. 2021.

B.1 Saturation Properties for R134a - Temperature

Temp.	Pressure	Density		Int. Energy		Enthalpy		Entropy	
T	p	ρ_f	ρ_g	u_f	u_g	h_f	h_g	s_f	s_g
°C	MPa	kg/m ³	kg/m ³	kJ/kg	kJ/kg	kJ/kg	kJ/kg	kJ/kgK	kJ/kgK
-40	0.05121	1417.7	2.7695	148.11	355.51	148.14	374.00	0.7956	1.7643
-38	0.05682	1411.9	3.0529	150.62	356.66	150.66	375.27	0.8063	1.7615
-36	0.06291	1406.1	3.3590	153.14	357.81	153.18	376.54	0.8170	1.7588
-34	0.06951	1400.2	3.6890	155.66	358.96	155.71	377.80	0.8276	1.7563
-32	0.07666	1394.3	4.0441	158.19	360.11	158.25	379.06	0.8381	1.7538
-30	0.08438	1388.4	4.4259	160.73	361.25	160.79	380.32	0.8486	1.7515
-28	0.09270	1382.4	4.8356	163.28	362.40	163.34	381.57	0.8591	1.7492
-26	0.10167	1376.5	5.2748	165.83	363.55	165.90	382.82	0.8694	1.7471
-24	0.11130	1370.4	5.7450	168.39	364.70	168.47	384.07	0.8798	1.7451
-22	0.12165	1364.4	6.2477	170.96	365.84	171.05	385.32	0.8900	1.7432
-20	0.13273	1358.3	6.7845	173.54	366.99	173.64	386.55	0.9003	1.7413
-18	0.14460	1352.1	7.3571	176.12	368.13	176.23	387.79	0.9104	1.7396
-16	0.15728	1345.9	7.9673	178.72	369.28	178.83	389.02	0.9205	1.7379
-14	0.17082	1339.7	8.6168	181.32	370.41	181.44	390.24	0.9306	1.7363
-12	0.18524	1333.4	9.3074	183.93	371.55	184.07	391.46	0.9407	1.7348
-10	0.20060	1327.1	10.041	186.55	372.69	186.70	392.66	0.9507	1.7334
-8	0.21693	1320.8	10.820	189.17	373.82	189.34	393.87	0.9606	1.7320
-6	0.23428	1314.3	11.646	191.81	374.95	191.99	395.06	0.9705	1.7307
-4	0.25268	1307.9	12.521	194.45	376.07	194.65	396.25	0.9804	1.7294
-2	0.27217	1301.4	13.448	197.11	377.19	197.32	397.43	0.9902	1.7282
0	0.29280	1294.8	14.428	199.77	378.31	200.00	398.60	1.0000	1.7271
2	0.31462	1288.1	15.465	202.45	379.42	202.69	399.77	1.0098	1.7260
4	0.33766	1281.4	16.560	205.13	380.53	205.40	400.92	1.0195	1.7250
6	0.36198	1274.7	17.717	207.83	381.63	208.11	402.06	1.0292	1.7240
8	0.38761	1267.9	18.938	210.53	382.73	210.84	403.20	1.0388	1.7230
10	0.41461	1261.0	20.226	213.25	383.82	213.58	404.32	1.0485	1.7221
12	0.44301	1254.0	21.584	215.98	384.90	216.33	405.43	1.0581	1.7212
14	0.47288	1246.9	23.015	218.71	385.98	219.09	406.53	1.0677	1.7204
16	0.50425	1239.8	24.522	221.46	387.05	221.87	407.61	1.0772	1.7196
18	0.53718	1232.6	26.109	224.23	388.11	224.66	408.69	1.0867	1.7188
20	0.57171	1225.3	27.780	227.00	389.17	227.47	409.75	1.0962	1.7180

Temp.	Pressure	Density		Int. Energy		Enthalpy		Entropy	
T	p	ρ_f	ρ_g	u_f	u_g	h_f	h_g	s_f	s_g
°C	MPa	kg/m ³	kg/m ³	kJ/kg	kJ/kg	kJ/kg	kJ/kg	kJ/kgK	kJ/kgK
22	0.60789	1218.0	29.539	229.79	390.21	230.29	410.79	1.1057	1.7173
24	0.64578	1210.5	31.389	232.59	391.25	233.12	411.82	1.1152	1.7166
26	0.68543	1202.9	33.335	235.40	392.28	235.97	412.84	1.1246	1.7159
28	0.72688	1195.2	35.382	238.23	393.29	238.84	413.84	1.1341	1.7152
30	0.77020	1187.5	37.535	241.07	394.30	241.72	414.82	1.1435	1.7145
32	0.81543	1179.6	39.799	243.93	395.29	244.62	415.78	1.1529	1.7138
34	0.86263	1171.6	42.180	246.80	396.27	247.54	416.72	1.1623	1.7131
36	0.91185	1163.4	44.683	249.69	397.24	250.48	417.65	1.1717	1.7124
38	0.96315	1155.1	47.316	252.60	398.19	253.43	418.55	1.1811	1.7118
40	1.0166	1146.7	50.085	255.52	399.13	256.41	419.43	1.1905	1.7111
42	1.0722	1138.2	52.998	258.46	400.05	259.41	420.28	1.1999	1.7103
44	1.1301	1129.5	56.064	261.42	400.96	262.43	421.11	1.2092	1.7096
46	1.1903	1120.6	59.292	264.40	401.84	265.47	421.92	1.2186	1.7089
48	1.2529	1111.5	62.690	267.41	402.71	268.53	422.69	1.2280	1.7081
50	1.3179	1102.3	66.272	270.43	403.55	271.62	423.44	1.2375	1.7072
52	1.3854	1092.9	70.047	273.47	404.37	274.74	424.15	1.2469	1.7064
54	1.4555	1083.2	74.030	276.54	405.17	277.89	424.83	1.2563	1.7055
56	1.5282	1073.4	78.235	279.64	405.94	281.06	425.47	1.2658	1.7045
58	1.6036	1063.2	82.679	282.76	406.67	284.27	426.07	1.2753	1.7035
60	1.6818	1052.9	87.379	285.91	407.38	287.50	426.63	1.2848	1.7024
62	1.7628	1042.2	92.358	289.09	408.06	290.78	427.14	1.2944	1.7013
64	1.8467	1031.2	97.637	292.30	408.69	294.09	427.61	1.3040	1.7000
66	1.9337	1020.0	103.24	295.55	409.29	297.44	428.02	1.3137	1.6987
68	2.0237	1008.3	109.21	298.83	409.84	300.84	428.36	1.3234	1.6972
70	2.1168	996.25	115.57	302.16	410.33	304.28	428.65	1.3332	1.6956
72	2.2132	983.76	122.37	305.53	410.78	307.78	428.86	1.3430	1.6939
74	2.3130	970.78	129.65	308.95	411.16	311.33	429.00	1.3530	1.6920
76	2.4161	957.25	137.48	312.42	411.47	314.94	429.04	1.3631	1.6899
78	2.5228	943.10	145.93	315.95	411.70	318.63	428.98	1.3733	1.6876
80	2.6332	928.24	155.08	319.55	411.83	322.39	428.81	1.3836	1.6850
82	2.7473	912.56	165.05	323.23	411.87	326.24	428.51	1.3942	1.6821
84	2.8653	895.91	175.97	327.00	411.77	330.20	428.05	1.4049	1.6789
86	2.9874	878.10	188.05	330.88	411.53	334.28	427.42	1.4159	1.6752
88	3.1136	858.86	201.52	334.89	411.10	338.51	426.55	1.4273	1.6710
90	3.2442	837.83	216.76	339.06	410.45	342.93	425.42	1.4390	1.6662
92	3.3793	814.43	234.31	343.44	409.49	347.59	423.92	1.4514	1.6604
94	3.5193	787.75	255.08	348.11	408.13	352.58	421.92	1.4645	1.6534
96	3.6645	756.09	280.73	353.23	406.13	358.07	419.18	1.4789	1.6445
98	3.8152	715.51	315.13	359.14	403.03	364.47	415.14	1.4957	1.6322
100	3.9724	651.18	373.01	367.20	397.03	373.30	407.68	1.5188	1.6109
101.06	4.0591	511.90	511.90	381.71	381.71	389.64	389.64	1.5621	1.5621

B.2 Saturation Properties for R134a - Pressure

Pressure p MPa	Temp. T °C	Density		Int. Energy		Enthalpy		Entropy	
		ρ_f kg/m ³	ρ_g kg/m ³	u_f kJ/kg	u_g kJ/kg	h_f kJ/kg	h_g kJ/kg	s_f kJ/kgK	s_g kJ/kgK
0.02	-56.41	1464.3	1.1484	127.75	346.17	127.77	363.58	0.7051	1.7931
0.04	-44.60	1430.9	2.1973	142.36	352.88	142.39	371.09	0.7707	1.7713
0.06	-36.94	1408.8	3.2131	151.96	357.27	152.00	375.94	0.8120	1.7601
0.08	-31.12	1391.7	4.2096	159.31	360.61	159.37	379.62	0.8428	1.7528
0.10	-26.36	1377.5	5.1932	165.37	363.34	165.44	382.60	0.8676	1.7475
0.12	-22.31	1365.3	6.1677	170.56	365.67	170.65	385.12	0.8884	1.7435
0.14	-18.76	1354.5	7.1353	175.14	367.70	175.24	387.32	0.9066	1.7402
0.16	-15.59	1344.7	8.0979	179.25	369.51	179.37	389.27	0.9226	1.7376
0.18	-12.71	1335.7	9.0566	183.00	371.15	183.13	391.02	0.9371	1.7353
0.20	-10.08	1327.4	10.012	186.45	372.64	186.60	392.62	0.9503	1.7334
0.22	-7.64	1319.6	10.966	189.65	374.02	189.82	394.09	0.9624	1.7317
0.24	-5.37	1312.3	11.918	192.65	375.30	192.83	395.44	0.9736	1.7303
0.26	-3.24	1305.4	12.869	195.47	376.50	195.67	396.70	0.9841	1.7290
0.28	-1.23	1298.8	13.820	198.14	377.62	198.35	397.89	0.9940	1.7278
0.30	0.67	1292.6	14.770	200.67	378.68	200.90	399.00	1.0033	1.7267
0.32	2.48	1286.5	15.721	203.09	379.69	203.34	400.04	1.0121	1.7257
0.34	4.20	1280.8	16.671	205.40	380.64	205.66	401.03	1.0204	1.7249
0.36	5.84	1275.2	17.623	207.61	381.54	207.90	401.97	1.0284	1.7240
0.38	7.42	1269.9	18.575	209.74	382.41	210.04	402.87	1.0360	1.7233
0.40	8.93	1264.7	19.529	211.79	383.24	212.11	403.72	1.0433	1.7226
0.45	12.48	1252.3	21.918	216.63	385.16	216.99	405.69	1.0604	1.7210
0.50	15.74	1240.8	24.317	221.10	386.91	221.50	407.47	1.0759	1.7197
0.55	18.75	1229.9	26.729	225.27	388.51	225.72	409.09	1.0903	1.7185
0.60	21.57	1219.5	29.155	229.19	389.99	229.68	410.57	1.1037	1.7175
0.65	24.22	1209.7	31.596	232.89	391.36	233.43	411.94	1.1162	1.7165
0.70	26.71	1200.2	34.054	236.41	392.64	236.99	413.20	1.1280	1.7156
0.80	31.33	1182.2	39.025	242.97	394.96	243.65	415.46	1.1497	1.7140
0.90	35.53	1165.4	44.078	249.01	397.01	249.78	417.43	1.1695	1.7126
1.0	39.39	1149.3	49.222	254.63	398.85	255.50	419.16	1.1876	1.7113
1.1	42.97	1134.0	54.465	259.90	400.49	260.87	420.69	1.2044	1.7100
1.2	46.32	1119.2	59.815	264.88	401.98	265.95	422.04	1.2201	1.7087
1.3	49.46	1104.8	65.280	269.60	403.32	270.78	423.24	1.2349	1.7075
1.4	52.42	1090.9	70.870	274.12	404.54	275.40	424.30	1.2489	1.7062
1.5	55.23	1077.2	76.595	278.45	405.64	279.84	425.23	1.2622	1.7049
1.6	57.91	1063.7	82.464	282.61	406.64	284.11	426.04	1.2748	1.7036
1.8	62.90	1037.3	94.682	290.52	408.35	292.26	427.36	1.2987	1.7007
2.0	67.48	1011.4	107.63	297.98	409.70	299.95	428.28	1.3209	1.6976
2.2	71.73	985.47	121.42	305.07	410.72	307.30	428.84	1.3417	1.6941
2.4	75.69	959.38	136.24	311.88	411.42	314.38	429.04	1.3615	1.6902
2.6	79.41	932.74	152.28	318.47	411.80	321.26	428.88	1.3805	1.6858
2.8	82.90	905.19	169.84	324.92	411.84	328.01	428.33	1.3990	1.6807
3.0	86.20	876.21	189.35	331.28	411.49	334.70	427.34	1.4171	1.6748
3.2	89.33	845.09	211.44	337.64	410.70	341.43	425.83	1.4350	1.6679
3.4	92.30	810.67	237.19	344.12	409.32	348.31	423.65	1.4533	1.6595
3.6	95.12	770.79	268.69	350.91	407.11	355.58	420.50	1.4724	1.6487
3.8	97.80	720.15	311.11	358.50	403.41	363.77	415.63	1.4939	1.6336
4.0	100.34	632.87	390.23	369.25	395.13	375.57	405.38	1.5247	1.6046
4.0591	101.06	511.90	511.90	381.71	381.71	389.64	389.64	1.5621	1.5621

B.3 Superheated Vapor Properties for Refrigerant R134a

$p = 0.06 \text{ MPa } (T_{sat} = -36.9^\circ\text{C})$					$p = 0.10 \text{ MPa } (T_{sat} = -26.4^\circ\text{C})$			
Temp.	Density	Energy	Enthalpy	Entropy	Density	Energy	Enthalpy	Entropy
T	ρ	u	h	s	ρ	u	h	s
$^\circ\text{C}$	kg/m^3	kJ/kg	kJ/kg	kJ/kgK	kg/m^3	kJ/kg	kJ/kg	kJ/kgK
Sat.	3.2131	357.27	375.94	1.7601	5.1933	363.34	382.60	1.7475
-30	3.1105	361.93	381.22	1.7821				
-20	2.9755	368.74	388.91	1.8131	5.0401	367.81	387.65	1.7677
-10	2.8532	375.70	396.73	1.8433	4.8208	374.89	395.64	1.7986
0	2.7415	382.80	404.69	1.8730	4.6232	382.10	403.73	1.8288
10	2.6390	390.07	412.81	1.9022	4.4433	389.45	411.95	1.8584
20	2.5444	397.50	421.08	1.9310	4.2784	396.94	420.31	1.8874
30	2.4567	405.10	429.52	1.9593	4.1266	404.59	428.82	1.9159
40	2.3752	412.86	438.12	1.9872	3.9860	412.40	437.49	1.9441
50	2.2991	420.79	446.88	2.0147	3.8554	420.37	446.30	1.9718
60	2.2280	428.88	455.81	2.0419	3.7336	428.49	455.28	1.9991
70	2.1613	437.14	464.90	2.0688	3.6198	436.78	464.41	2.0261
80	2.0986	445.56	474.15	2.0954	3.5130	445.23	473.70	2.0528
90	2.0395	454.14	483.56	2.1216	3.4126	453.84	483.14	2.0792
100	1.9837	462.89	493.14	2.1477	3.3181	462.61	492.74	2.1053

$p = 0.14 \text{ MPa } (T_{sat} = -18.8^\circ\text{C})$					$p = 0.18 \text{ MPa } (T_{sat} = -12.7^\circ\text{C})$			
Temp.	Density	Energy	Enthalpy	Entropy	Density	Energy	Enthalpy	Entropy
T	ρ	u	h	s	ρ	u	h	s
$^\circ\text{C}$	kg/m^3	kJ/kg	kJ/kg	kJ/kgK	kg/m^3	kJ/kg	kJ/kg	kJ/kgK
Sat.	7.1353	367.70	387.32	1.7402	9.0566	371.15	391.02	1.7353
-10	6.8467	374.05	394.50	1.7680	8.9369	373.17	393.31	1.7441
0	6.5517	381.37	402.74	1.7987	8.5310	380.62	401.72	1.7754
10	6.2861	388.81	411.08	1.8287	8.1699	388.15	410.18	1.8059
20	6.0446	396.37	419.53	1.8580	7.8445	395.79	418.73	1.8355
30	5.8234	404.08	428.12	1.8868	7.5484	403.55	427.40	1.8646
40	5.6197	411.93	436.84	1.9151	7.2772	411.46	436.19	1.8931
50	5.4312	419.94	445.72	1.9430	7.0273	419.51	445.12	1.9212
60	5.2561	428.10	454.74	1.9705	6.7959	427.71	454.20	1.9489
70	5.0928	436.42	463.91	1.9977	6.5808	436.06	463.41	1.9761
80	4.9401	444.90	473.24	2.0244	6.3802	444.56	472.78	2.0030
90	4.7968	453.53	482.72	2.0509	6.1923	453.22	482.29	2.0296
100	4.6621	462.32	492.35	2.0771	6.0160	462.03	491.95	2.0558

$p = 0.20 \text{ MPa } (T_{sat} = -10.1^\circ\text{C})$					$p = 0.24 \text{ MPa } (T_{sat} = -5.4^\circ\text{C})$			
Temp.	Density	Energy	Enthalpy	Entropy	Density	Energy	Enthalpy	Entropy
T	ρ	u	h	s	ρ	u	h	s
$^\circ\text{C}$	kg/m^3	kJ/kg	kJ/kg	kJ/kgK	kg/m^3	kJ/kg	kJ/kg	kJ/kgK
Sat.	10.012	372.64	392.62	1.7334	11.918	375.30	395.44	1.7303
-10	10.009	372.70	392.68	1.7337				
0	9.5410	380.23	401.20	1.7654	11.605	379.43	400.11	1.7475
10	9.1281	387.81	409.73	1.7961	11.079	387.13	408.79	1.7787
20	8.7577	395.49	418.33	1.8259	10.612	394.89	417.51	1.8090
30	8.4220	403.29	427.04	1.8551	10.192	402.75	426.30	1.8385
40	8.1152	411.22	435.87	1.8838	9.8102	410.74	435.20	1.8674

$p = 0.20 \text{ MPa } (T_{sat} = -10.1^\circ\text{C})$					$p = 0.24 \text{ MPa } (T_{sat} = -5.4^\circ\text{C})$			
Temp.	Density	Energy	Enthalpy	Entropy	Density	Energy	Enthalpy	Entropy
T	ρ	u	h	s	ρ	u	h	s
$^\circ\text{C}$	kg/m^3	kJ/kg	kJ/kg	kJ/kgK	kg/m^3	kJ/kg	kJ/kg	kJ/kgK
50	7.8332	419.29	444.83	1.9120	9.4609	418.86	444.22	1.8957
60	7.5725	427.51	453.92	1.9397	9.1393	427.11	453.37	1.9236
70	7.3306	435.88	463.16	1.9670	8.8417	435.51	462.65	1.9511
80	7.1051	444.39	472.54	1.9939	8.5652	444.06	472.08	1.9781
90	6.8944	453.06	482.07	2.0206	8.3072	452.75	481.64	2.0048
100	6.6967	461.88	491.75	2.0468	8.0658	461.59	491.35	2.0312

$p = 0.28 \text{ MPa } (T_{sat} = -1.2^\circ\text{C})$					$p = 0.32 \text{ MPa } (T_{sat} = 2.5^\circ\text{C})$			
Temp.	Density	Energy	Enthalpy	Entropy	Density	Energy	Enthalpy	Entropy
T	ρ	u	h	s	ρ	u	h	s
$^\circ\text{C}$	kg/m^3	kJ/kg	kJ/kg	kJ/kgK	kg/m^3	kJ/kg	kJ/kg	kJ/kgK
Sat.	13.820	377.62	397.89	1.7278	15.721	379.69	400.04	1.7257
0	13.733	378.59	398.98	1.7318				
10	13.079	386.42	407.83	1.7636	15.131	385.68	406.83	1.7501
20	12.505	394.27	416.66	1.7943	14.440	393.64	415.80	1.7812
30	11.994	402.21	425.55	1.8241	13.829	401.65	424.79	1.8113
40	11.532	410.25	434.53	1.8532	13.280	409.75	433.85	1.8407
50	11.111	418.41	443.61	1.8818	12.783	417.96	443.00	1.8695
60	10.725	426.71	452.81	1.9098	12.330	426.30	452.25	1.8977
70	10.369	435.14	462.14	1.9374	11.912	434.77	461.63	1.9254
80	10.039	443.71	471.61	1.9646	11.527	443.37	471.13	1.9527
90	9.7319	452.43	481.20	1.9914	11.169	452.11	480.77	1.9796
100	9.4452	461.30	490.94	2.0178	10.835	461.00	490.54	2.0062
110	9.1765	470.31	500.82	2.0440	10.523	470.03	500.44	2.0324
120	8.9241	479.47	510.85	2.0698	10.230	479.21	510.49	2.0583

$p = 0.40 \text{ MPa } (T_{sat} = 8.9^\circ\text{C})$					$p = 0.50 \text{ MPa } (T_{sat} = 15.7^\circ\text{C})$			
Temp.	Density	Energy	Enthalpy	Entropy	Density	Energy	Enthalpy	Entropy
T	ρ	u	h	s	ρ	u	h	s
$^\circ\text{C}$	kg/m^3	kJ/kg	kJ/kg	kJ/kgK	kg/m^3	kJ/kg	kJ/kg	kJ/kgK
Sat.	19.529	383.24	403.72	1.7226	24.317	386.91	407.47	1.7197
10	19.415	384.12	404.72	1.7261				
20	18.446	392.32	414.01	1.7584	23.744	390.55	411.61	1.7339
30	17.607	400.50	423.22	1.7893	22.554	398.99	421.16	1.7659
40	16.865	408.73	432.45	1.8192	21.526	407.40	430.63	1.7967
50	16.201	417.05	441.74	1.8484	20.619	415.86	440.11	1.8265
60	15.600	425.47	451.11	1.8770	19.808	424.40	449.64	1.8555
70	15.050	434.01	460.58	1.9050	19.074	433.04	459.25	1.8839
80	14.546	442.67	470.17	1.9325	18.406	441.78	468.95	1.9118
90	14.080	451.47	479.88	1.9596	17.792	450.65	478.75	1.9392
100	13.647	460.41	489.72	1.9864	17.225	459.65	488.67	1.9661
110	13.244	469.48	499.68	2.0127	16.700	468.78	498.72	1.9927
120	12.867	478.69	509.78	2.0387	16.211	478.04	508.88	2.0189
130	12.513	488.05	520.02	2.0645	15.753	487.44	519.18	2.0447
140	12.181	497.55	530.39	2.0899	15.324	496.98	529.60	2.0703

$p = 0.60 \text{ MPa } (T_{sat} = 21.6^\circ\text{C})$					$p = 0.70 \text{ MPa } (T_{sat} = 26.7^\circ\text{C})$			
Temp.	Density	Energy	Enthalpy	Entropy	Density	Energy	Enthalpy	Entropy
T	ρ	u	h	s	ρ	u	h	s
$^\circ\text{C}$	kg/m^3	kJ/kg	kJ/kg	kJ/kgK	kg/m^3	kJ/kg	kJ/kg	kJ/kgK
Sat.	29.155	389.99	410.57	1.7175	34.054	392.64	413.20	1.7156
30	27.790	397.37	418.96	1.7455	33.371	395.62	416.60	1.7269
40	26.409	406.01	428.73	1.7772	31.549	404.53	426.72	1.7598
50	25.215	414.63	438.43	1.8077	30.010	413.35	436.67	1.7910
60	24.161	423.30	448.13	1.8373	28.674	422.16	446.57	1.8212
70	23.218	432.04	457.88	1.8661	27.493	431.01	456.47	1.8505
80	22.366	440.87	467.70	1.8943	26.435	439.94	466.42	1.8791
90	21.590	449.82	477.61	1.9220	25.478	448.97	476.44	1.9070
100	20.877	458.88	487.62	1.9492	24.605	458.09	486.54	1.9345
110	20.219	468.06	497.74	1.9759	23.804	467.34	496.74	1.9615
120	19.609	477.37	507.97	2.0023	23.064	476.70	507.05	1.9880
130	19.040	486.82	518.33	2.0283	22.377	486.19	517.47	2.0142
140	18.509	496.39	528.81	2.0540	21.737	495.80	528.01	2.0400

$p = 0.80 \text{ MPa } (T_{sat} = 31.3^\circ\text{C})$					$p = 0.90 \text{ MPa } (T_{sat} = 35.5^\circ\text{C})$			
Temp.	Density	Energy	Enthalpy	Entropy	Density	Energy	Enthalpy	Entropy
T	ρ	u	h	s	ρ	u	h	s
$^\circ\text{C}$	kg/m^3	kJ/kg	kJ/kg	kJ/kgK	kg/m^3	kJ/kg	kJ/kg	kJ/kgK
Sat.	39.025	394.96	415.46	1.7140	44.078	397.01	417.43	1.7126
40	36.988	402.96	424.59	1.7436	42.781	401.28	422.32	1.7283
50	35.030	412.00	434.84	1.7758	40.307	410.59	432.92	1.7616
60	33.363	420.97	444.95	1.8067	38.247	419.74	443.28	1.7932
70	31.908	429.96	455.03	1.8364	36.478	428.87	453.54	1.8236
80	30.619	438.99	465.12	1.8654	34.928	438.01	463.78	1.8530
90	29.462	448.10	475.25	1.8937	33.550	447.21	474.03	1.8816
100	28.414	457.30	485.45	1.9214	32.309	456.48	484.34	1.9096
110	27.457	466.60	495.73	1.9486	31.183	465.85	494.71	1.9370
120	26.578	476.01	506.12	1.9754	30.153	475.32	505.17	1.9640
130	25.764	485.55	516.60	2.0017	29.205	484.90	515.72	1.9905
140	25.009	495.21	527.20	2.0277	28.327	494.61	526.38	2.0166
150	24.305	504.99	537.91	2.0533	27.511	504.43	537.14	2.0423
160	23.646	514.91	548.74	2.0786	26.750	514.38	548.02	2.0678

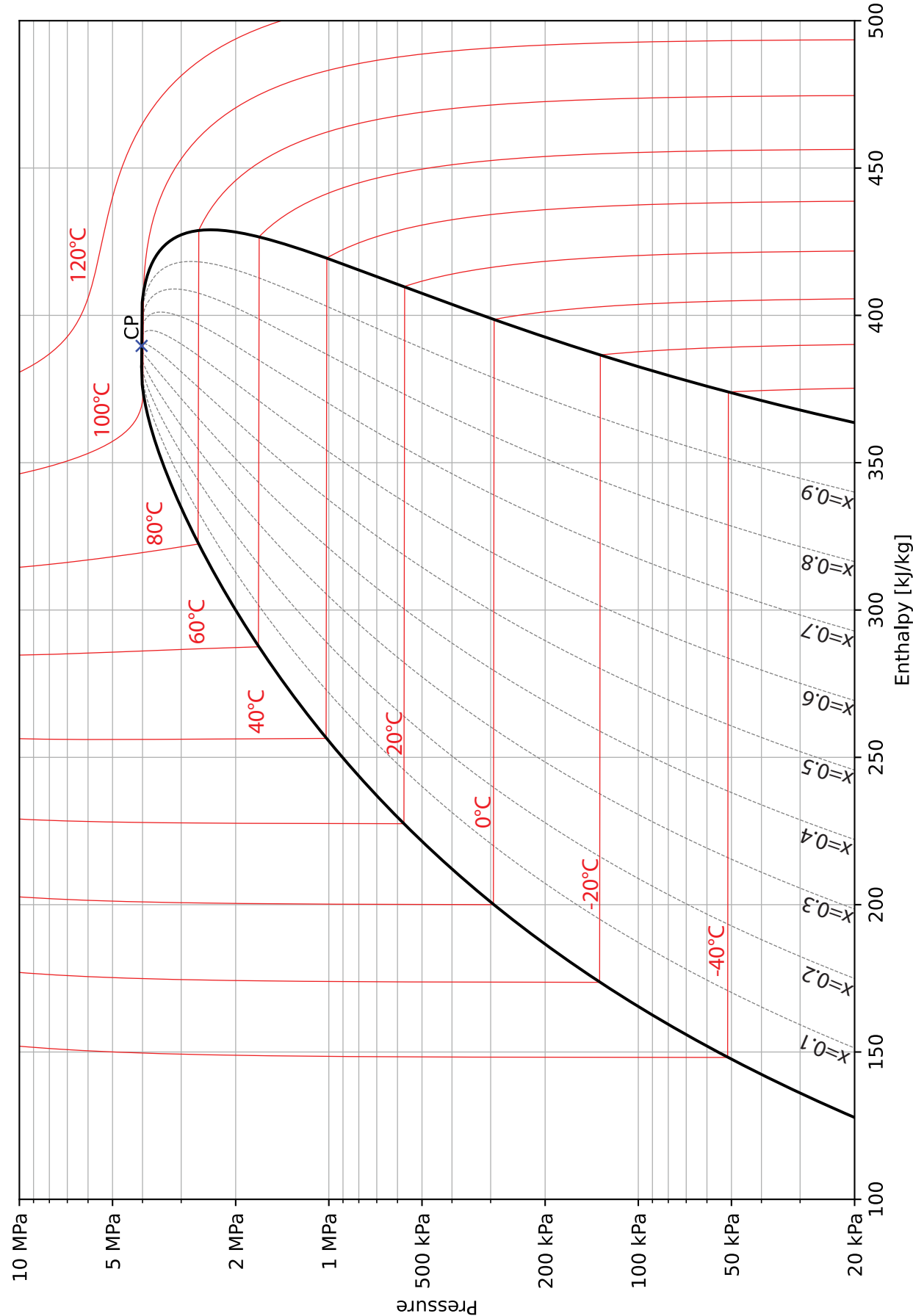
$p = 1.0 \text{ MPa } (T_{sat} = 39.4^\circ\text{C})$					$p = 1.2 \text{ MPa } (T_{sat} = 46.3^\circ\text{C})$			
Temp.	Density	Energy	Enthalpy	Entropy	Density	Energy	Enthalpy	Entropy
T	ρ	u	h	s	ρ	u	h	s
$^\circ\text{C}$	kg/m^3	kJ/kg	kJ/kg	kJ/kgK	kg/m^3	kJ/kg	kJ/kg	kJ/kgK
Sat.	49.222	398.85	419.16	1.7113	59.815	401.98	422.04	1.7087
40	49.004	399.45	419.86	1.7135				
50	45.880	409.09	430.88	1.7482	58.136	405.77	426.41	1.7223
60	43.350	418.46	441.53	1.7806	54.335	415.70	437.79	1.7570
70	41.218	427.74	452.00	1.8116	51.277	425.36	448.76	1.7895
80	39.372	437.00	462.40	1.8414	48.710	434.90	459.53	1.8204
90	37.746	446.30	472.79	1.8705	46.499	444.41	470.22	1.8502
100	36.295	455.65	483.20	1.8988	44.558	453.94	480.87	1.8792
110	34.985	465.09	493.67	1.9264	42.830	463.53	491.54	1.9074
120	33.792	474.62	504.21	1.9536	41.275	473.18	502.25	1.9350
130	32.700	484.25	514.83	1.9803	39.862	482.92	513.02	1.9621

$p = 1.0 \text{ MPa } (T_{sat} = 39.4^\circ\text{C})$					$p = 1.2 \text{ MPa } (T_{sat} = 46.3^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Density ρ kg/m^3	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Density ρ kg/m^3	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
140	31.692	494.00	525.55	2.0065	38.569	492.76	523.87	1.9886
150	30.758	503.86	536.37	2.0324	37.379	502.70	534.81	2.0148
160	29.889	513.84	547.30	2.0579	36.277	512.76	545.84	2.0405
170	29.076	523.95	558.34	2.0831	35.252	522.93	556.97	2.0660
180	28.314	534.19	569.51	2.1080	34.295	533.23	568.22	2.0910

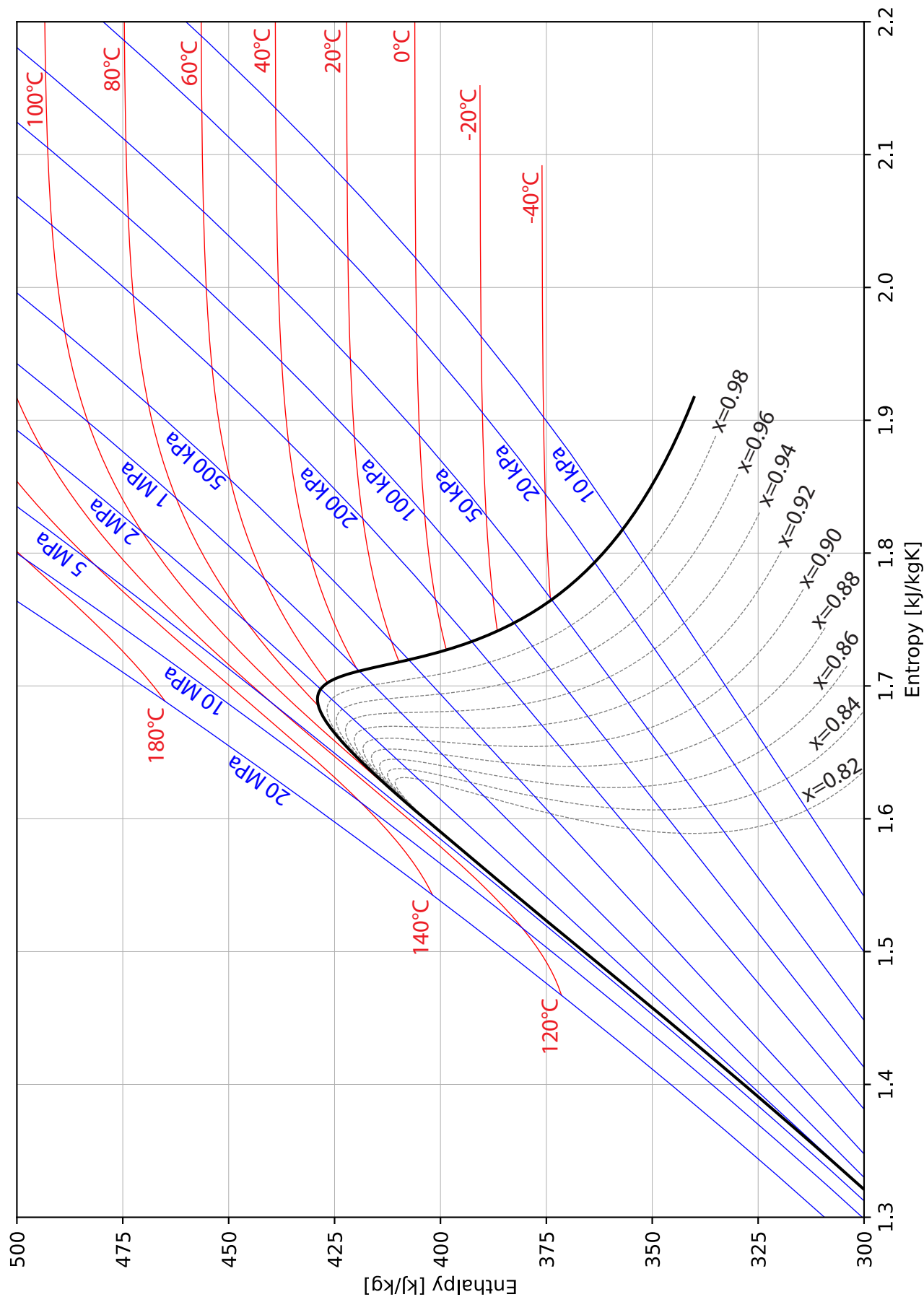
$p = 1.4 \text{ MPa } (T_{sat} = 52.4^\circ\text{C})$					$p = 1.6 \text{ MPa } (T_{sat} = 57.9^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Density ρ kg/m^3	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Density ρ kg/m^3	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	70.871	404.54	424.30	1.7062	82.464	406.64	426.04	1.7036
60	66.643	412.61	433.62	1.7345	80.824	409.04	428.84	1.7120
70	62.267	422.77	445.25	1.7689	74.458	419.91	441.40	1.7491
80	58.745	432.65	456.48	1.8012	69.628	430.24	453.22	1.7831
90	55.794	442.42	467.51	1.8320	65.722	440.32	464.66	1.8150
100	53.255	452.16	478.45	1.8617	62.443	450.29	475.91	1.8456
110	51.028	461.90	489.34	1.8905	59.620	460.22	487.06	1.8751
120	49.049	471.70	500.24	1.9186	57.143	470.16	498.16	1.9037
130	47.269	481.55	511.17	1.9460	54.941	480.15	509.27	1.9316
140	45.654	491.49	522.16	1.9730	52.960	490.19	520.40	1.9589
150	44.177	501.52	533.21	1.9994	51.163	500.32	531.59	1.9856
160	42.818	511.65	544.35	2.0254	49.519	510.53	542.84	2.0119
170	41.560	521.89	555.58	2.0510	48.007	520.84	554.17	2.0378
180	40.392	532.25	566.91	2.0763	46.608	531.26	565.59	2.0632

$p = 1.8 \text{ MPa } (T_{sat} = 62.9^\circ\text{C})$					$p = 2.0 \text{ MPa } (T_{sat} = 67.5^\circ\text{C})$			
Temp. T $^\circ\text{C}$	Density ρ kg/m^3	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK	Density ρ kg/m^3	Energy u kJ/kg	Enthalpy h kJ/kg	Entropy s kJ/kgK
Sat.	94.682	408.35	427.36	1.7007	107.63	409.70	428.28	1.6976
70	88.275	416.68	437.07	1.7293	104.46	412.91	432.06	1.7086
80	81.570	427.61	449.67	1.7655	94.886	424.70	445.77	1.7480
90	76.405	438.07	461.63	1.7989	88.003	435.66	458.38	1.7833
100	72.198	448.32	473.25	1.8305	82.612	446.24	470.45	1.8160
110	68.652	458.46	484.68	1.8607	78.182	456.63	482.21	1.8471
120	65.590	468.58	496.02	1.8899	74.426	466.93	493.80	1.8770
130	62.900	478.70	507.32	1.9183	71.172	477.21	505.31	1.9059
140	60.504	488.86	518.61	1.9460	68.304	487.49	516.77	1.9340
150	58.348	499.08	529.93	1.9731	65.745	497.82	528.24	1.9614
160	56.389	509.38	541.30	1.9996	63.438	508.21	539.74	1.9883
170	54.597	519.77	552.74	2.0257	61.339	518.68	551.28	2.0146
180	52.948	530.25	564.25	2.0514	59.417	529.23	562.89	2.0405

B.4 R134a Pressure-Enthalpy (*p-h*) Diagram



B.5 R134a Enthalpy-Entropy (*h-s*) Diagram



Appendix C

Properties of Various Ideal Gases

C.1 Properties of Select Ideal Gases at 300 K

Gas Name	Formula	Gas Constant		Specific Heats		Critical Properties			
		M kg/kmol	R_{gas} kJ/kgK	c_v kJ/kgK	c_p kJ/kgK	T_c K	p_c MPa	ρ_c kg/m ³	Z_c -
Air ¹	-	28.970	0.2870	0.7177	1.0048	132.53	3.786	342.7	0.2905
Argon	Ar	39.948	0.2081	0.3122	0.5203	150.69	4.863	535.6	0.2895
Butane	C ₄ H ₁₀	58.122	0.1431	1.5594	1.7024	425.13	3.796	228.0	0.2738
Carbon Dioxide	CO ₂	44.010	0.1889	0.6569	0.8459	304.13	7.377	467.6	0.2746
Carbon Monoxide	CO	28.010	0.2968	0.7435	1.0404	132.86	3.494	303.9	0.2915
Ethane	C ₂ H ₆	30.069	0.2765	1.4760	1.7526	305.32	4.872	206.2	0.2799
Ethylene	C ₂ H ₄	28.053	0.2964	1.2393	1.5357	282.50	5.060	214.1	0.2822
Helium	He	4.0026	2.0773	3.1161	5.1932	5.1953	0.228	69.58	0.3041
Hydrogen	H ₂	2.0159	4.1245	10.186	14.310	33.145	1.296	31.26	0.3033
Methane	CH ₄	16.043	0.5183	1.7119	2.2301	190.56	4.599	162.7	0.2863
Neon	Ne	20.180	0.4120	0.6181	1.0303	44.400	2.662	486.0	0.2994
Nitrogen	N ₂	28.013	0.2968	0.7429	1.0397	126.19	3.396	313.3	0.2894
Octane	C ₈ H ₁₈	114.23	0.0728	1.5901	1.0458	568.74	2.484	232.0	0.2586
Oxygen	O ₂	31.999	0.2598	0.6585	0.9183	154.58	5.043	436.1	0.2879
Propane	C ₃ H ₈	44.096	0.1886	1.4828	1.6713	369.89	4.251	220.0	0.2771
Steam	H ₂ O	18.015	0.4615	1.4033	1.8649	647.10	22.064	322.0	0.2294

C.2 Ideal Gas Specific Heats of Air¹

Temp. K	h kJ/kg	c_v kJ/kgK	c_p kJ/kgK	γ -	Temp. K	h kJ/kg	c_v kJ/kgK	c_p kJ/kgK	γ -
200	326.20	0.7154	1.0024	1.4013	1100	1287.46	0.8717	1.1587	1.3293
300	426.53	0.7177	1.0048	1.3999	1150	1345.60	0.8798	1.1668	1.3263
400	527.37	0.7263	1.0133	1.3952	1200	1404.14	0.8874	1.1744	1.3235
500	629.45	0.7423	1.0294	1.3867	1250	1463.04	0.8945	1.1815	1.3209
600	733.43	0.7638	1.0509	1.3758	1300	1522.28	0.9011	1.1882	1.3185
700	839.70	0.7877	1.0747	1.3644	1350	1581.85	0.9074	1.1944	1.3163
800	948.37	0.8115	1.0985	1.3537	1400	1641.72	0.9132	1.2003	1.3143
900	1059.36	0.8337	1.1208	1.3443	1450	1701.87	0.9188	1.2058	1.3124
1000	1172.46	0.8539	1.1409	1.3362	1500	1762.29	0.9239	1.2110	1.3107

¹Data from CoolProp

C.3 Compressibility Charts

